

Synthesis, Characterization and Biological Activities of Binuclear Metal Complexes of 2-Benzoylpyridine and Phenyl(Pyridin-2-yl)Methanediol Derived from 1-Phenyl-1-(Pyridin-2-yl)-N-(Pyrimidin-2-yl)Methanimine Dihydrate Schiff Base

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Abstract

The Schiff base, 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L_1) has been synthesized by the condensation reaction between 2-aminopyrimidine and 2-benzoylpyridine and characterized using ^{13}C -NMR, ^1H -NMR, microanalysis, FT-IR, DEPT-135, HSQC, HMBC, COZY, NOESY. The reaction of 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L_1) with salts of V(IV), Co(II) and Cu(II), however, resulted in the hydrolysis of L_1 to give binuclear metal complexes of 2-benzoylpyridine (L_2) and phenyl(pyridin-2-yl)methanediol (L_3) which were characterized using UV-visible spectroscopy, and TGA. The single crystal x-ray structure determined for the Copper(II) complex revealed that we obtained a compound previously obtained using a different method of synthesis. The Schiff base ligand (L_1) is soluble in methanol, ethanol, DMSO, acetone and DMF. Microanalysis and Spectroscopic studies indicated that binuclear metal complexes were obtained by the coordination of metal ion to 2-benzoylpyridine (L_2) and phenyl(pyridin-2-yl)methanediol (L_3) from the hydrolysis of L_1 . Spectroscopic and elemental analysis reveal the formation of square pyramidal complexes of Co(II) and Cu(II) and a square planar complex of V(IV). *In vitro* antibacterial and antifungal activity against three bacterial strains (*Escherichia coli*, *Klebsiella pneumoniae* and *Staphylococcus aureus*) and three fungal stains (*Candida albicans*, *Candida glabrata* and *Candida parapsilosis*) showed moderate biological activity. Antioxidant studies reveal that only the binuclear

Co(II) and oxovanadium(II) complexes are potent to eliminate free radicals.

Keywords

Binuclear Complexes, Schiff Base, 2-Benzoylpyridine, Crystal Structure, Antimicrobial Activity, Antioxidant Activity

1. Introduction

Heterocyclic Schiff base ligands containing O- or N-donors and their metal complexes have been shown to exhibit interesting properties [1] [2] [3] and those containing pyrimidine and pyridine rings represent a promising class of compounds due to their inherent biological and pharmaceutical properties [2] [3] [4] [5]. Pyrimidine and pyridine based compounds exhibit a broad spectrum of biological activities [6] [7]. Pyrimidine with its interesting heterocyclic structure has extensively been used in drug design due to its potential antimicrobial, antifungal, antiviral and antitumor properties. Many ligands containing the pyridine ring and their transition metal complexes have been synthesized and tested for biological and therapeutic properties where they sometimes act as models for the active sites of biomolecules [4] [8] [9]. The importance of 2-benzoylpyridine metal complexes, for example, is due to their ability to facilitate substrate binding thus stabilizing the reactive intermediates of transition metals [4] [10].

Copper complexes have recently been the subject of intense research because of their potentials as radiopharmaceuticals for targeting hypoxic tissues [6] and as effective drugs for the treatment of refractory neuroblastoma in children [7]. Copper is an important trace element for plants and animals and is involved in mixed ligand complex formation in a number of biological processes [11]. Many cobalt(II) complexes of Schiff base ligands containing 2-benzoylpyridine behaved as a growth inhibitor for microorganisms [9]. Binuclear copper(II) complexes with monoatomic bridges, such as halides, have attracted recent attention due to their structural diversity and interesting catalytic, magnetic and biological properties [8] [12]. Depending on the nature of the co-ligands, chloride or oxygen bridged compounds, having square planar $\text{Cu}(\text{O})_2\text{Cu}$ bridging loops show different molecular structures [13] [14]. The crystalline architectures of these compounds are found to be interesting due to variations in the nature of intervening intermolecular forces.

The presence of the primary amine functionality in heterocyclic compounds containing rings such as pyridine and pyrimidine has prompted researchers to embark on the synthesis of Schiff bases of pyridine derivatives and their metal complexes. Also, Schiff bases easily undergo hydrolysis which is considered to be a very significant reaction in many pharmacological and biological processes resulting in the formation of new species [15] [16]. Recently our group embarks on studies on the biological activities of complexes of heterocyclic Schiff base ligands

[17]. In this paper we report on the synthesis, characterization and evaluation of the biological activities of binuclear metal complexes of 2-benzoylpyridine (L_2) and phenyl(pyridin-2-yl)methanediol (L_3) obtained from the hydrolysis of 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L_1) in the presence of salts of V(IV), Co(II) and Cu(II).

2. Experimental

2.1. Materials and Methods

All reagents and solvents were obtained from commercial sources and used as received without any further purification. Microanalysis (C, H and N) data were obtained using a Perkin-Elmer automated model 2400 series II CHNS/O analyzer. Melting points were determined using an analogy-SMP11 melting point measuring instrument of the Stuart Scientific type. The 1D and 2D spectra were recorded at room temperature in DMSO using a Bruker ARX 500 NMR Spectrometer (proton at 600 Mhz and carbon ^{13}C at 150 MHz). Infrared spectra were recorded using single reflectance ATR spectrometers on KBr disc. Electronic spectra were obtained using a HACH DR 3900 spectrophotometer. Thermal analyses were carried out using a Perkin-Elmer Pyris 6 TGA type device up to 600°C. Conductivity measurements were carried out using a Labtech® Digital Conductivity Meter type AVI-846 at room temperature and a cell constant of 1.056. Single crystal X-ray structure determination was carried out using a Rigaku Super Nova diffractometer.

2.2. Synthesis

The Schiff base ligand and its complexes were synthesized according to reported general synthetic procedures with some slight modifications [18] [19].

2.2.1. Synthesis of the Schiff Base (L_1)

To a solution of 10 mL of methanol containing 0.95 g (10 mmol) of 2-aminopyrimidine, 1.83 g (10 mmol) of 2-benzoylpyridine dissolved in 10 mL of methanol was added. The reaction mixture was maintained under reflux for 4 hours under continuous stirring. The mixture was allowed to cool at room temperature to obtain a pale yellow precipitate which was filtered, washed several times with ethanol and dried in a desiccator.

Yield: 1.60 g (54.05%), m.p. 64°C - 66°C. Analysis for $\text{C}_{16}\text{H}_{12}\text{N}_4 \cdot 2\text{H}_2\text{O}$ (%)
Found (calc): C, 64.92 (64.85); H, 5.00 (5.44); N, 23.00 (18.91).

2.2.2. Synthesis of Metal (II) Complexes

All complexes were synthesized by adding 20 mL of an ethanolic solution of either VOSO_4 (0.326 g, 2 mmol), $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ (0.582 g, 2 mmol) or $\text{Cu}(\text{NO}_3)_2 \cdot 5\text{H}_2\text{O}$ (0.554 g, 2 mmol) to a 30 mL ethanol-acetone mixture of (1.35 g, 4 mmol) Schiff base (L_1). The final mixtures were stirred at room temperature for about 24 hours to obtain precipitates which were filtered, washed with methanol and dried in a desiccator over calcium chloride.

2.3. Crystal Structure Determination

A suitable single crystal of $C_{51}H_{40}Cu_2N_7O_{12}$ was mounted on a Rigaku Supernova diffractometer and data were collected using Mo $K\alpha$ ($\lambda = 0.71073 \text{ \AA}$) radiation at a temperature of 95.01 K. The structure was solved using olex2 [20] with the ShelXT [21] structure solution program using Intrinsic Phasing and refined with the olex2 refine refinement package [20], using Gauss-Newton minimisation.

2.4. Antimicrobial Study

The synthesized compounds were used for antimicrobial screening by testing against three bacterial strains (*Escherichia coli*, *Klebsiella pneumonia* and *Staphylococcus aureus*) and three fungal strains, (*Candida albicans*, *Candida glabrata* and *Candida parapsilosis*), Ceftriaxone, Rifampicin and Fluconazole were used as antibacterial and antifungal references respectively.

Antimicrobial studies were carried out using the disc diffusion method based on the determination of the zone of inhibition around the paper discs [22]. In fact, 40 mg of each compound was dissolved in 1 mL of 10% DMSO making a final concentration of 40 mg/mL. Discs of wattman N°3 filter paper of 5 mm in diameter were deposited on the surface of the seeded medium. 10 μL of products were then deposited on each corresponding disk and left for pre-diffusion for 15 minutes at room temperature before being incubated at 37°C for 24 hours for bacteria and 48 hours for yeasts. The antimicrobial activities were assessed by measuring the diameter of the zone of inhibition of the bacterial and fungal growth around every Disc with a ruler following two axes. For every product tested, three determinations were made.

2.5. In Vitro Antioxidant Activities

The synthesized compounds were tested for *in vitro* antioxidant activities at 37°C using free both the radical scavenging assay (DPPH) method and Ferric Ion Reducing Antioxidant Power Assay (FRAP) method.

2.5.1. DPPH Free Radical Trapping Assay Method

In this method, the antiradical activity of the ligand L_1 and binuclear metal complexes were tested against the stable free radical of 1, 1-diphenyl-2-picrylhydrazyl (DPPH), using the free radical scavenging assay with some modifications [23]. Solutions of compounds at different concentrations (1000, 500, 250, 125, 62.5, 31.25, 15.625 and 7.8125 $\mu\text{g/mL}$) were prepared in methanol as solvent. 25 μL of each sample at different concentrations, and 75 μL of DPPH (0.02%) solution were measured into different test tubes and the mixture shaken vigorously for about 30 minutes. The contents of the test tubes were then incubated in the dark for 30 minutes at room temperature. A blank DPPH solution without the sample used for the baseline correction gave a strong absorption maximum at 517 nm (purple color with $\epsilon = 8.32 \times 10^3 \text{ M}^{-1}\cdot\text{cm}^{-1}$). After incubation, the absorbance was determined at wavelength between 510 and 520 nm for each sample using a UV-visible spectrometer. The test was performed in triplicate using ascorbic acid

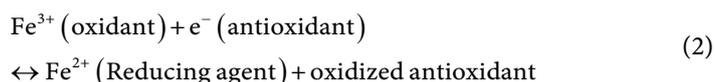
(vitamin C) as positive control. The observed decrease in absorbance values indicate that the compounds show scavenging activity and the relative free radical scavenging activity was calculated using formula 1 below:

$$\%RSA = \frac{A_o - A_s}{A_o} \times 100 \quad (1)$$

where **RSA** represents the radical Scavenging Activity; **A_o** is the absorbance of white (DPPH + methanol) and **A_s** is the absorbance of the radical DPPH + compounds.

2.5.2. Ferric Ion Reducing Antioxidant Power Assay (FRAP)

In this method, the antiradical power is based on the reduction of Fe³⁺ to Fe²⁺ by compounds which, in the presence of 1, 10-phenanthroline, form a brown or orange-red colored complex [24]. The complex absorbs at 505 nm and the intensity of the coloration is proportional to the amount of Fe³⁺ converted to Fe²⁺ by the compound as shown in **Scheme 2**.

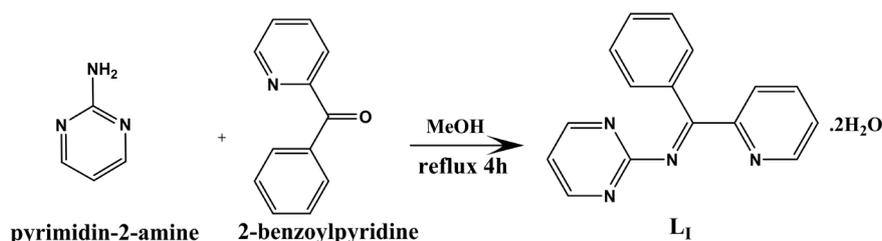


Solutions of each compound were prepared in methanol. 25 μL of each solution were added to a new microplate and 25 μL of Fe³⁺ solution (1.2 mg/mL) added. The plates were pre-incubated for 15 minutes at room temperature. After this period, 50 μL of ortho-phenanthroline (0.2%) was added to obtain final concentrations of 1000, 500, 250, 125, 62.5, 31.25, 15.625, 7.8125, 3.90625 and 1.95325 μg/mL. The reaction mixtures were then incubated for 15 minutes at room temperature and the absorbance measured at 505 nm using a spectrophotometer. Ascorbic acid was used as a positive control and was treated in the same manner as for the compounds. The test was performed in triplicate. A graph of optical density versus concentration was plotted to the degree of correlation between the concentration of the compounds and their reducing power.

3. Results and Discussion

3.1. Synthesis and Characterization of the Schiff Base L₁

The Schiff base L₁ was prepared by the condensation of equimolar amounts of 2-aminopyrimidine with 2-benzoylpyridine in absolute methanol under reflux at 65 °C as shown in **Scheme 1**.



Scheme 1. Synthesis of 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L₁).

The Schiff base is brownish yellow, stable at room temperature, soluble in organic solvents such as methanol, ethanol, DMSO, acetone and DMF. The 1D and 2D NMR sequences were recorded for the characterization of the molecule, in order to provide the structure of the ligand synthesized.

3.2. Infrared Study of the Schiff Base L₁

The infrared spectrum of the Schiff base is presented in **Figure 1**. The Infrared spectrum exhibits an intense absorption band at 1643 cm⁻¹ attributed to azomethine group $\nu_{(C=N)}$. The very strong absorption band at 1556 cm⁻¹ is attributed to $\nu_{(C=N)}$ group of the ring. The band observed at 3326 cm⁻¹ is attributed to (O-H) vibration mode of water molecules [25]. The loss of the absorption band around 3317-3385 cm⁻¹ attributed to $\nu_{(N-H)}$ stretch of 2-aminopyrimidine and at 1706 cm⁻¹ for $\nu_{(C=O)}$ of 2-benzoylpyridine and the appearance of a band at 1643 cm⁻¹ attributed to azomethine group $\nu_{(C=N)}$ in the ligand, are indicative of the formation of the Schiff base L₁.

3.3. ¹H NMR Spectral Analysis of the Schiff Base L₁

The ¹H-NMR spectrum of the Schiff base L₁ is presented in **Figure 2** and the ¹H-NMR data of the compounds and the chemical shifts of the different types of protons are listed in **Table 1**.

Table 1. (¹H) and (¹³C) NMR chemical shift data observed for the Schiff base (L₁).

N° C	¹³ C(δ, ppm)	DEPT 135	¹ H*(δ, ppm)	HMBC
1	110.63	CH	6.55 (<i>dd</i> , 1H, Pyrim.)	—
2, 6	158.42	CH	8.23 (<i>d</i> , 2H, Pyrim.)	—
4	164.10	—	—	2.4
8	164.10	—	—	—
9	154.96	—	—	20.9
10	136.48	—	—	—
11	131.07		7.97 (<i>m</i> , 1H, Harom.)	
12	124.63		8.01 (<i>m</i> , 1H, Harom.)	
13	128.70	CH	7.55 (<i>m</i> , 1H, Harom.)	—
14	124.63		8.01 (<i>m</i> , 1H, Harom.)	
15	131.07		7.97 (<i>m</i> , 1H, Harom.)	
17	149.05		8.74 (<i>d</i> , 1H, Pyr.)	
18	127.22		7.68 (<i>t</i> , 1H, Pyr.)	
19	133.47	CH	7.68 (<i>td</i> , 1H, Pyr.)	—
20	138.47		8.08 (<i>d</i> , 1H, Pyr.)	
—	—	H ₂ O	3.36 (<i>s</i> , 2H ₂ O)	

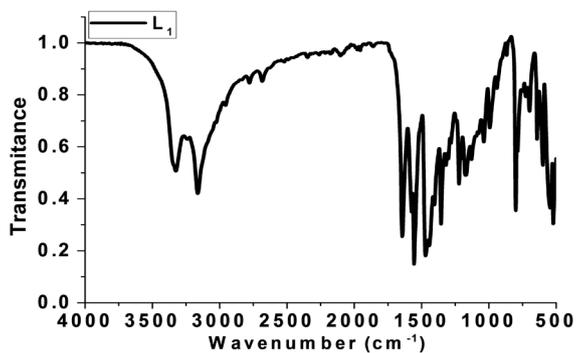


Figure 1. Infrared spectrum of the Schiff base ligand L_1 .

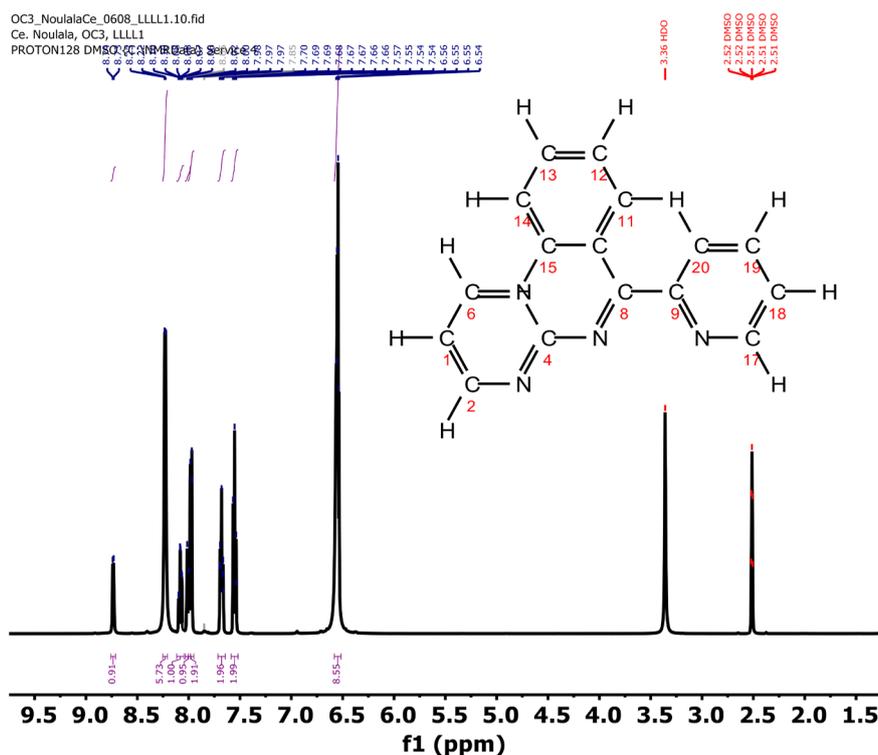


Figure 2. ^1H -NMR spectrum of the Schiff base ligand L_1 .

Examination of the proton NMR spectrum shows the presence of three doublets. The first at $\delta = 8.08$ ppm is attributed to the proton H-20; the second at $\delta = 8.23$ ppm is attributed to the protons of the H-2, H-6 group of the pyrimidine ring and the third at $\delta = 7.74$ ppm is attributed to the H-17 proton of the pyridine ring. This spectrum also contains a signal in the form of doublet of a doublet at $\delta = 6.55$ ppm attributable to the proton at position C-1. The spectrum also shows doublet of a triplet corresponding to the proton H-19 at $\delta = 7.68$. The protons of the aromatic nucleus (H-11, H-12, H-13, H-14 and H-15) are represented by the presence of multiplets between $\delta = 7.97 - 8.01$ ppm.

3.4. ^{13}C NMR Spectral Analysis of the Schiff Base L_1

Analysis of the ^{13}C -NMR spectra (**Figure 3**) shows twelve distinguished carbon

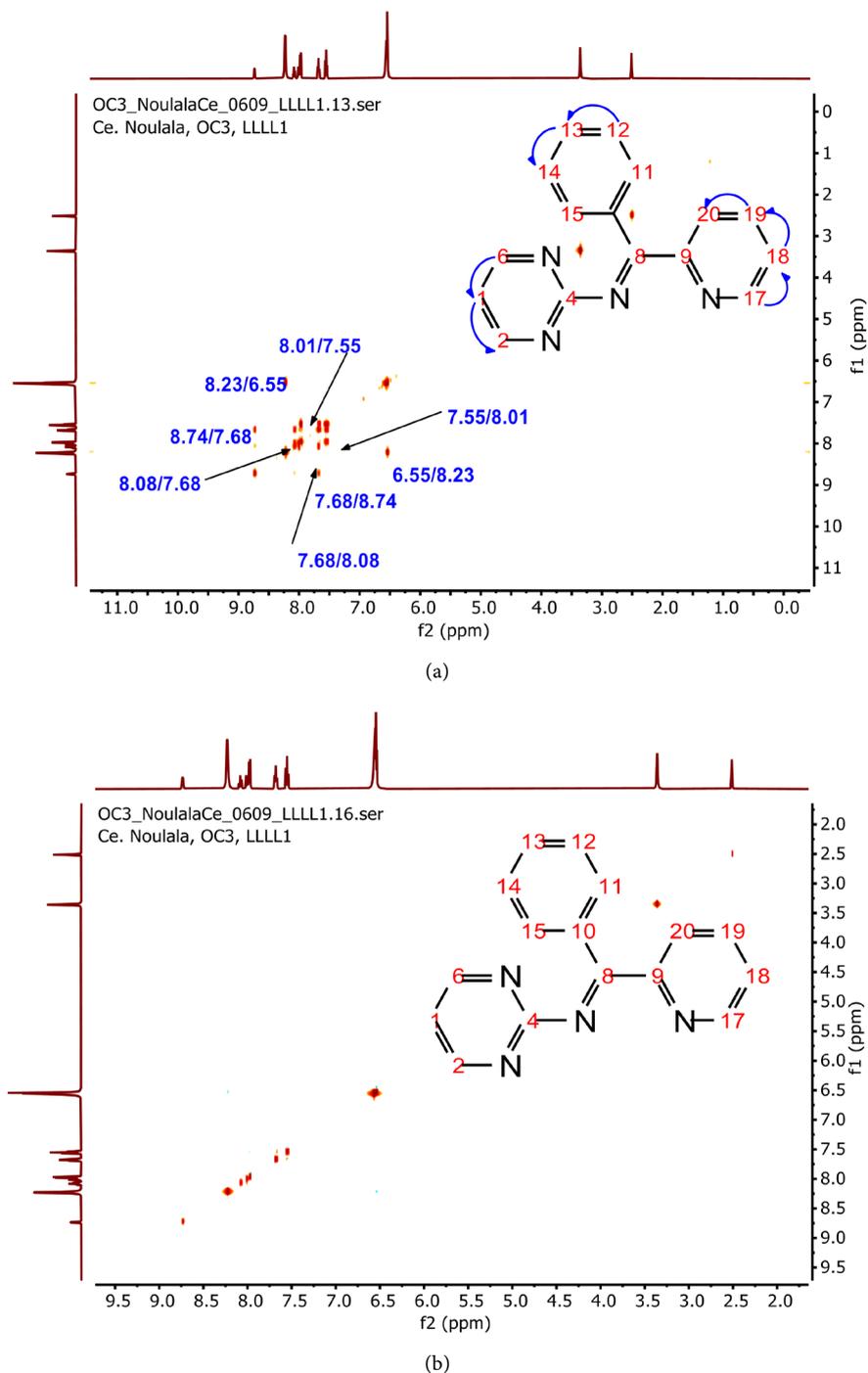


Figure 4. (a) COSY and (b) NOESY correlations of Schiff base L_1 .

An examination of the COSY spectrum (**Figure 4(a)**) shows a correlation between the H-1 nucleus ($\delta = 6.55$ ppm) and the H-2, H-6 proton ($\delta = 8.23$ ppm) of the pyrimidine ring. On the same spectrum, a correlation is observed between the proton H-12 ($\delta = 8.01$ ppm) and its neighbour H-13 ($\delta = 7.55$ ppm). This spectrum also shows correlation between the H-17, H-20 nuclei of pyridine with the protons H-18, H-19.

All protons and carbons were assigned after analysis of the HSQC and HMBC (Figure 5) spectra and the data results are summarizing in Table 1.

HSQC spectrum (Figure 5(a)) makes it possible to locate C-1 and C-2, C-6 at $\delta = 110.63$ ppm and $\delta = 158.42$ ppm; C-18, C-19 at $\delta = 127.22$ and $\delta = 133.47$ ppm; C-12 at $\delta = 124.63$ ppm; C-13 at $\delta = 128.07$ ppm; C-20 at $\delta = 138.47$ ppm; C-17 at $\delta = 149.05$ ppm; C-11, C-15 at $\delta = 131.07$ ppm.

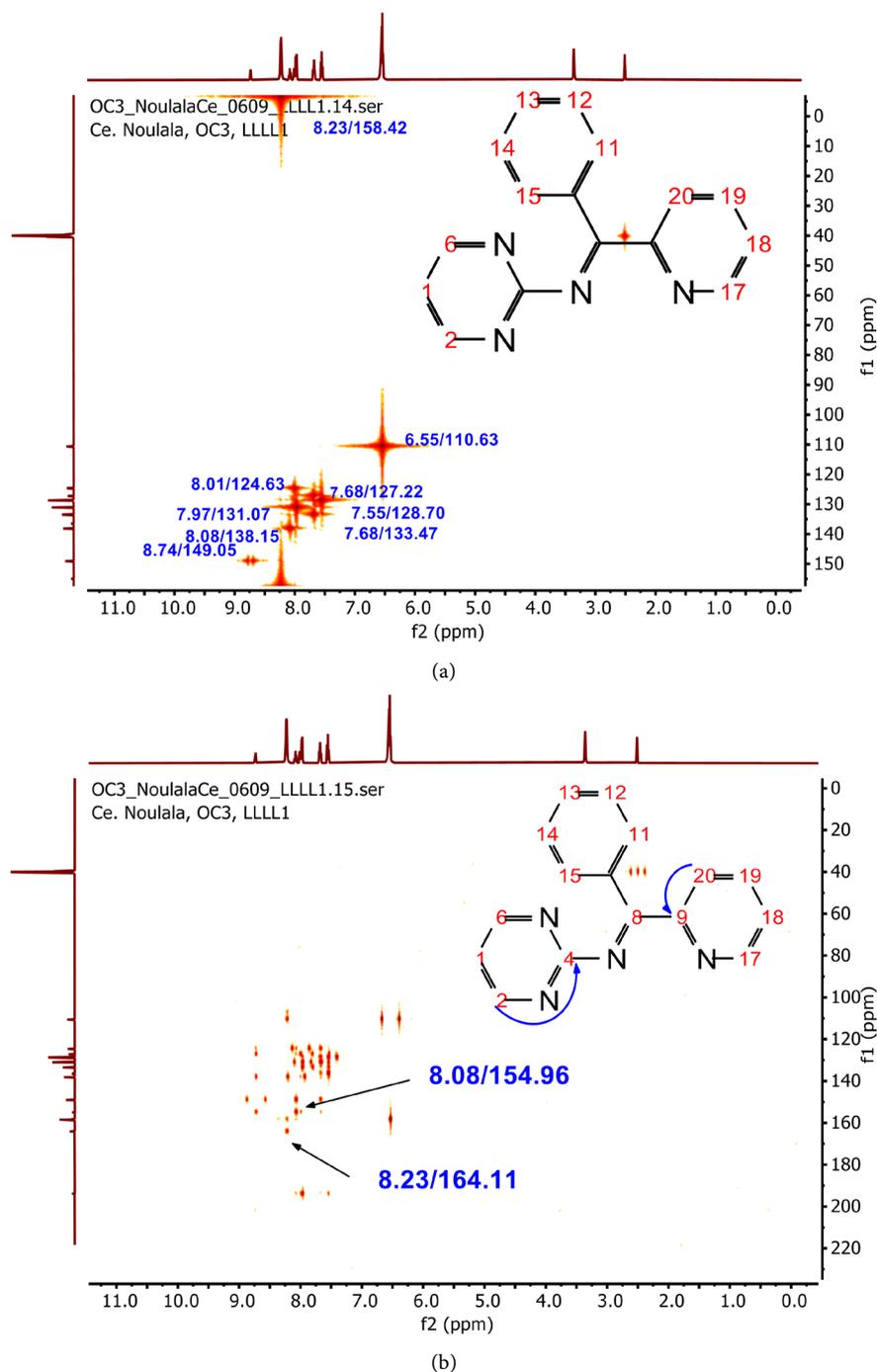


Figure 5. (a) HSQC correlations of the Schiff base (L₁); (b) HMBC correlations of the Schiff base (L₁) showing the different of HMBC correlations.

HMBC spectrum (**Figure 5(b)**) shows correlation of the H-1 proton with the quaternary carbon of the imine function at $\delta = 164.11$ ppm, which allows the attribution of the latter to the C-7 carbon. The same spectrum also shows a correlation of proton H-11 with carbon at $\delta = 154.96$ ppm which can only be C-9 and another correlation of proton H-15 with oxygenated carbon at $\delta = 193.92$ ppm attributed to C-5. Proton H-15 with the oxygenated carbon at $\delta = 193.92$ ppm attributed to C-5 [26].

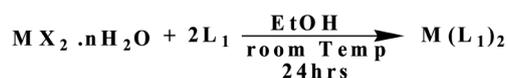
The analysis of all the IR and NMR data confirm the proposed structure of the Schiff base ligand (L_1) as shown in **Scheme 3**.

4. Synthesis and Characterization of Metal Complexes of the Schiff Base L_1

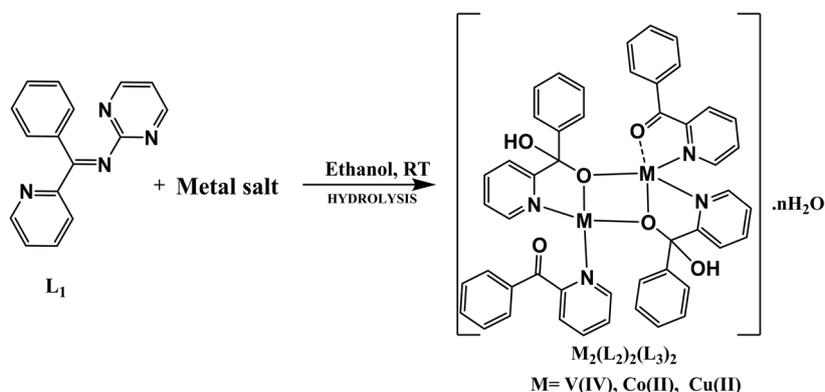
The Schiff base (L_1) was reacted with metal salts in a 1:2 molar ratio in an attempt to synthesize vanadium(IV), cobalt(II) and copper(II) complexes of the form, $[M(L_1)_2]$ following **Scheme 2**.

The reaction of 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L_1) with salts of V(IV), and Co(II) Cu(II), however resulted in the hydrolysis of L_1 to give binuclear metal complexes of 2-benzoylpyridine (L_2) and phenyl(pyridin-2-yl)methanediol (L_3) as revealed by the various analysis, suggesting metal assisted hydrolysis of the Schiff base ligand, L_1 during complexation as illustrated in **Scheme 3**.

We obtained suitable crystals for $[Cu_2(L_2)_2(L_3)_2]$ for single crystal x-ray structure determination. The structure of this [27] and similar compounds [28], obtained by different synthetic procedures (the reactions of 2-benzoylpyridine with copper(II) ion [27] and metal assisted hydrolysis of N' -[(*E*)-phenyl(pyridin-2-yl)methylidene]furan-2-carbohydrazide [28]) have however been reported elsewhere. The analysis and physical properties of the prepared compounds are given in **Table 2**.



Scheme 2. Equation reaction of the synthesis of $[M(L_1)_2]$ complexes.



Scheme 3. Reaction equation for the synthesis of binuclear $M(II)$ complexes.

Table 2. Analysis and physical properties of Schiff base and binuclear Cu(II) complex.

Compounds	Empirical formula	Mol.Wt (g/mol)	Aspect	M.p (°C)	Yield (%)	Elemental analysis (%)		
						Calc (Exp)		
						%C	%H	%N
L ₁	C ₁₆ H ₁₂ N ₄ ·2H ₂ O	296.33	Brownish Yellow	64	54.48	64.85 (64.92)	5.44 (5.00)	18.91 (23.00)
[Cu ₂ (L ₂) ₂ (L ₃) ₂]	C ₅₁ H ₄₀ Cu ₂ N ₇ O ₁₂	1081.03	Greenish crystal	124	55.16	55.55 (53.42)	4.47 (3.83)	7.77 (8.44)
[VO ₂ (L ₂) ₂]	VO ₂ (C ₁₂ H ₉ NO) ₂ (SO ₄) ₂ ·(H ₂ O) ₃	746.46	Yellowish green powder	≥360	39.10	38.62 (38.43)	3.24 (3.01)	3.75 (3.12)
[Co ₂ (L ₂)(L ₃) ₂]	[Co ₂ (C ₁₂ H ₉ NO) (C ₁₂ H ₁₀ NO ₂) ₂ (H ₂ O)](NO ₃) ₂	905.54	Pink powder	202	50.33	47.75 (48.54)	4.45 (3.88)	9.28 (9.71)

Elemental analysis values for the Schiff base and Metal(II) complexes (**Table 2**) tallied with the expected values, thus confirming the purity and proposed formulation of the Schiff base ligand and its complexes.

4.1. Infrared Study of the Binuclear Metal(II) Complexes

The IR spectrum of the ligand when compared to that of M(II) complexes reveals the binding of the ligands, L₂ and L₃ to M(II) ions. In the spectra of the complexes (**Figure 6**) and IR data (**Table 3**), the disappearance of the strong band observed for the free ligand at 1643 cm⁻¹ and attributed to the imine group $\nu_{(C=N)}$, and the appearance of a new band around 1768 - 1668 cm⁻¹ attributed to the carbonyl group, $\nu_{(C=O)}$ are indicative of the hydrolysis of the Schiff base L₁ and the formation of 2-benzoylpyridine L₂ [29] [30]. The strong absorption band at 1556 cm⁻¹ attributed to $\nu_{(C=N)}$ group of the pyridyl ring is shifted to the higher values around 1630 - 1598 cm⁻¹ in the spectra of complexes, suggesting the coordination of the nitrogen of the pyridine ring to the metal centre. The band at 1329 cm⁻¹ in the spectra of the Cu(II) and Co(II) complexes is attributed to $\nu_{(C-O)}$ of phenyl(pyridin-2-yl)methanediol (L₃). This is confirmed by the presence of a broad band around 3350 - 3292 cm⁻¹ attributed to hydroxyl group $\nu_{(O-H)}$ of phenyl(pyridin-2-yl)methanediol (L₃) and crystal water molecules [31]. The bands appearing at lower frequencies around 594 - 454 and 595 - 548 cm⁻¹ have been attributed to $\nu_{(M-N)}$ and $\nu_{(M-O)}$ respectively [32]. Important IR vibrations and corresponding assignments of the L₁ ligand and M(II) complexes have been presented in **Table 3**.

4.2. Conductivity Measurement

The molar conductivities of the complexes were determined in methanol for [(VO)₂(L₂)₂] and [Co₂(L₂)(L₃)₂] and in ethanol for [Cu₂(L₂)₂(L₃)₂] and the obtained values are summarized in **Table 4**. The low value of the molar conductivity (18.25 S·mol⁻¹·cm²) for [(VO)₂(L₂)₂], indicated its molecular nature. While the high values of the molar conductivities for [Cu₂(L₂)₂(L₃)₂] (204.34 S·mol⁻¹·cm²) and [Co₂(L₂)(L₃)₂] (295.45 S·mol⁻¹·cm²) indicate that they are 1: 2 and 1:3 electrolytes in solution respectively [33] [34].

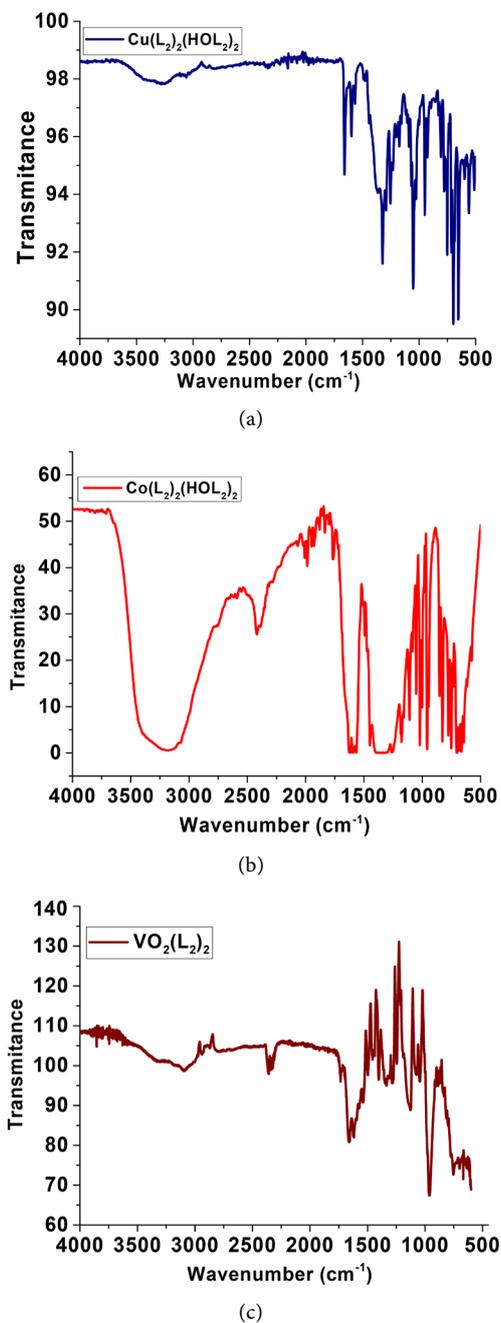


Figure 6. Infrared spectra of the (a) $[\text{Cu}_2(\text{L}_2)_2(\text{L}_3)_2]$, (b) $[\text{Co}_2(\text{L}_2)(\text{L}_3)_2]$ and (c) $[(\text{VO}_2)(\text{L}_2)_2]$ binuclear complexes.

Table 3. Infrared data of the Schiff base (L_1) and M(II) complex.

Compound	$\nu_{(\text{C}=\text{N})}$	$\nu_{(\text{C}=\text{Npyr})}$	$\nu_{(\text{NH}_2)}$	$\nu_{\text{OH}(\text{H}_2\text{O})}$	$\nu_{(\text{C}=\text{O})}$	$\nu_{(\text{C}-\text{O})}$	$\nu_{(\text{NO}_3)}$	$\nu_{(\text{M}-\text{N})}$	$\nu_{(\text{M}-\text{O})}$
L_1	1643	1560	3326	3330 - 3163	—	—	—	—	—
$[\text{Cu}_2(\text{L}_2)_2(\text{L}_3)_2]$	/	1598	/	3292 - 3292	1668	1329	1254	454	548
$[\text{Co}_2(\text{L}_2)_2(\text{L}_3)_2]$	/	1590	/	3292 - 3350	1768	1355	1254	451	510
$[(\text{VO})_2(\text{L}_2)_2]$	/	1642	/	3321 - 3274	1728	1322	—	594	595

Table 4. Assignments of the electronic spectra of L₁ ligand and its complexes.

Compound	Bande Positions (ν en cm ⁻¹)	Assignments	Geometries	Molar conductivity S·mol ⁻¹ ·cm ²
L ₁	21740	π → π*	—	—
[(VO) ₂ (L ₂) ₂]	28248	MLCT	Square pyramidal	18.25
[Co ₂ (L ₂)(L ₃) ₂ (H ₂ O)](NO ₃) ₃	21362 28011	MLCT	Square pyramidal	295.45
[Cu ₂ (L ₂) ₂ (L ₃) ₂](NO ₃) ₂ ·H ₂ O	13869	² B _{1g} → ² A _{1g}	Distorted square planar	204.34

4.3. UV-Visible Spectroscopy

The electronic spectral data for the compounds are summarized in **Table 4**. High intensity bands appeared in the ultraviolet and near-visible regions at 382 and 460 nm in the spectrum of ligand L₁ and was assigned to the *n* → π* transition.

In the spectrum of the [(VO)₂(L₂)₂] complex, the band appearing in the ultraviolet region at 354 nm is attributed to the MLCT, which suggests a square-based pyramidal geometry of the complex [34] [35] [36]. The UV-visible spectrum of the Cobalt complex show two bands at 357 and 468 nm attributed to the MLCT with square pyramidal geometry around cobalt ion [35]. The spectrum of the [Cu₂(L₂)₂(L₃)₂] complex (**Figure 7**) shows a band at 721 nm attributed to ligand-metal charge transfer [37].

4.4. Thermogravimetric Analysis of the Binuclear Copper (II) Complex

The differential and thermogravimetric analyses of the [Cu₂(L₂)₂(L₃)₂] complex was determined under an inert atmosphere in the range 0°C and 600°C. The thermogram for the [Cu₂(L₂)₂(L₃)₂] is represented in **Figure 8**. The thermal decomposition curves of Cu(II) complex showed three steps decomposition. The first step between 30°C and 151°C corresponds to the loss of a water molecule and an ethanolate ion (Found: 5.36%; Calc: 5.65%). The second step between 151°C and 180°C (found: 6.19%; calc: 6.48%) corresponding to the loss of a nitrate ion. While the third decomposition step between 180°C and 310°C corresponds to the loss of the organic moieties (2C₁₂H₉NO and C₁₂H₁₀NO₂) (found: 62.28%; calc: 63.75%) leaving a residue which corresponds to a mixture of copper oxides, CuO·Cu₂O (found: 22.5%; calc. 20.6%).

4.5. Crystal Structure of the Binuclear Copper Complex

Suitable crystals of Cu(II) complex were obtained for single crystal x-ray structure determination. **Table 5** presents the crystal data and refinement parameters. **Table 6** lists the selected bond lengths and bond angles. The binuclear complexes are presumed to occur as a result of bridging between the two Metal(II) centers through the deprotonated *gem-diol* oxygen atoms. The structure of the Cu(II) complex has however been reported elsewhere [27] [28], even though obtained

by different synthetic procedures involving the reactions of 2-benzoylpyridine with copper(II) ion [27] and metal assisted hydrolysis of N'-[(E)-phenyl(pyridine-2-yl)methylidene]furan-2-carbohydrazide [28].

Table 5. Crystal Data and structure refinement for $[\text{Cu}_2(\text{L}_2)_2(\text{L}_3)_2](\text{NO}_3)_2 \cdot \text{H}_2\text{O} \cdot \text{CH}_3\text{CH}_2\text{O}$.

Empirical formula	$\text{C}_{51}\text{H}_{40}\text{Cu}_2\text{N}_7\text{O}_{12}$
Formula weight	1081.04
Temperature/K	95.01 (10)
Crystal system	monoclinic
Space group	$\text{P}2_1/\text{n}$
a/Å	18.3200 (4)
b/Å	13.9885 (4)
c/Å	18.6759 (5)
$\alpha/^\circ$	90
$\beta/^\circ$	101.765 (2)
$\gamma/^\circ$	90
Volume/Å ³	4685.5 (2)
Z	4
$\rho_{\text{calc}}/\text{cm}^3$	1.5324
F(000)	2231.8
Crystal size/mm ³	N/A × N/A × N/A
Radiation/Å	Mo K α ($\lambda = 0.71073$)
2 θ range for data collection/ $^\circ$	3.5 to 64.22
Reflections collected	30,842
Independent reflections	14,639 [$R_{\text{int}} = 0.0337$, $R_{\text{sigma}} = 0.0595$]
Reflections with $I \geq 2\sigma(I)$	10,704
Goodness-of-fit on F^2	1.041
Final R indexes [$I \geq 2\sigma(I)$]	$R_1 = 0.0558$, $wR_2 = 0.1313$
Final R indexes [all data]	$R_1 = 0.0822$, $wR_2 = 0.1500$

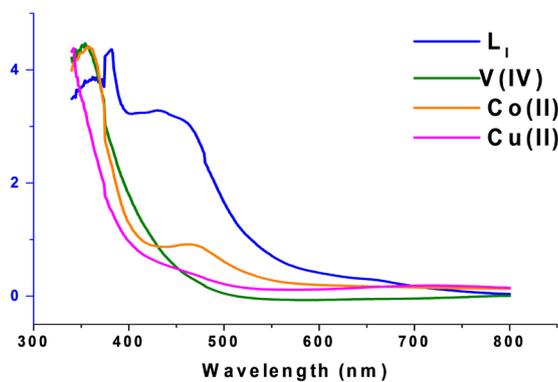


Figure 7. UV-visible spectra of Schiff base and VO(II), Co(II) and Cu(II) binuclear complexes.

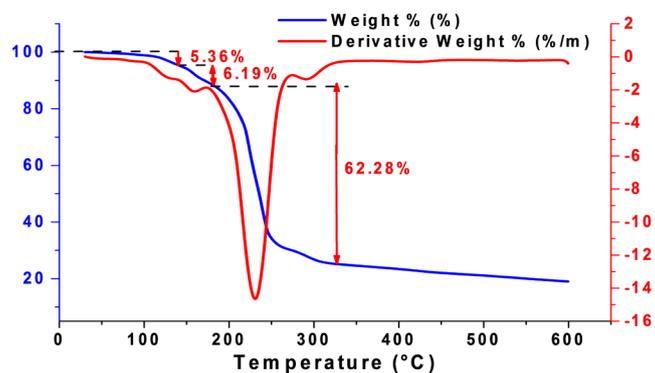


Figure 8. Thermogram of the $[\text{Cu}_2(\text{L}_2)_2(\text{L}_3)_2]$ binuclear complex.

Table 6. Selected bond Lengths and bond Angles for $\text{C}_{51}\text{H}_{40}\text{Cu}_2\text{N}_7\text{O}_{12}$.

Atom	Atom	Length/Å	
Cu1	Cu2	3.0301 (4)	
Cu1	O11	1.9280 (18)	
Cu1	O26	1.9600 (17)	
Cu1	O61	2.475 (2)	
Cu1	N12	1.986 (2)	
Cu1	N46	2.007 (2)	
Cu2	O11	1.9321 (17)	
Cu2	O26	1.9097 (18)	
Cu2	O54	2.420 (2)	
Cu2	N32	1.983 (2)	
Cu2	N47	1.981 (2)	
O11	C9	1.399 (3)	
O40	C39	1.226 (3)	
O25	C24	1.403 (3)	
N12	C13	1.341 (3)	
Atom	Atom	Atom	Angle/°
O11	Cu1	Cu2	38.33 (5)
O26	Cu1	Cu2	37.87 (5)
O26	Cu1	O11	76.15 (7)
O61	Cu1	Cu2	88.88 (5)
O61	Cu1	O11	89.34 (8)
O61	Cu1	O26	86.35 (8)
N12	Cu1	Cu2	119.91 (7)
N12	Cu1	O11	81.61 (8)
N12	Cu1	O26	157.75 (8)
N12	Cu1	O61	93.16 (8)
N46	Cu1	Cu2	138.11 (6)
N46	Cu1	O11	174.14 (9)
N46	Cu1	O26	100.71 (8)

The molecular structure of the binuclear Cu(II) complex is given in **Figure 9** and the crystallographic data summarized in **Table 5**. The complex crystallizes in a monoclinic crystal system with two ligands in the unit cell just like the reported structure [27]. The selected bond lengths and angles are shown in Table 6. The molecular structure of the complex consists of a non-centro symmetric $[\text{Cu}_2(\text{L}_1)_2(\text{L}_3)_2]^{2+}$ cation, a water molecule, an ethanolate ion and two nitrate ions. The two copper atoms are separated by a distance of 3.023 Å comparable to that reported for bridged alkoxo complexes. On the other hand, the bridge angles Cu(1)-O(11)-Cu(2) and Cu(1)-O(26)-Cu(2) of $38.33^\circ(5)$ and $37.87^\circ(5)$ are different from those observed for other alkoxo complexes [27]. The bond lengths (Cu(1)-N(py)) are respectively [Cu(1)-N(12) = 1.986 (2) Å and Cu(1)-N(46) = 2.007(2) Å] and compare well with those found in the reported binuclear copper(II) complexes [27] [28].

5. Biological Activity of the Synthesized Compounds

5.1. Antimicrobial Activity of the Schiff Base and Binuclear Metal Complexes

The Schiff base and complexes were tested against three bacterial strains *Staphylococcus aureus*, *Klebsiella pneumoniae* and *Escherichia coli* and three fungal strains *Candida albicans*, *Candida glabrata* and *Candida parapsilosis*. Rifampicin and Fluconazole were used as the standards for bacterial and fungal studies respectively. The diameters of the zone of inhibition of the ligand and binuclear complexes were determined using disc diffusion method [23]. The diameters of zone of inhibition of the bacteria are summarized in **Table 7**. The results reveal that the Schiff base and the binuclear complexes do not exhibit antifungal activities but for the Cobalt complex which exhibited moderate activity against the *C. albicans* species. The antifungal activity of Co(II) complex on *C. albicans* is comparable to the activity of Rifampicin used as standard. $[(\text{VO})_2(\text{L}_2)_2]$ complex exhibited high antibacterial activity on all the tested species, when compared to Rifampicin used as standard.

Table 7. Diameters of inhibition zones at the end of the screening.

Compounds	Bacteria				Fungi	
	<i>Escherichia coli</i>	<i>Klebsiella pneumoniae</i>	<i>Staphylococcus aureus</i>	<i>Candida albicans</i>	<i>Candida glabrata</i>	<i>Candida parapsilosis</i>
L_1	—	—	—	—	—	—
$[(\text{VO})_2(\text{L}_2)_2]$	15 ± 0.8	13 ± 1.01	15 ± 0.23	—	—	—
$[\text{Co}_2(\text{L}_2)(\text{L}_3)_2]$	—	—	—	8 ± 2	—	—
$[\text{Cu}_2(\text{L}_2)_2(\text{L}_3)_2]$	—	—	—	—	—	—
Control	Fluconazole	/	/	8 ± 3	20 ± 2.3	25 ± 1.08
	Rifampicin	15 ± 1.3	18 ± 1.6	22 ± 1.05	/	/

Compounds are considered active when their diameters of the zone of inhibition are greater than 6 mm (DZI > 6 mm) and (—) means no activities observed.

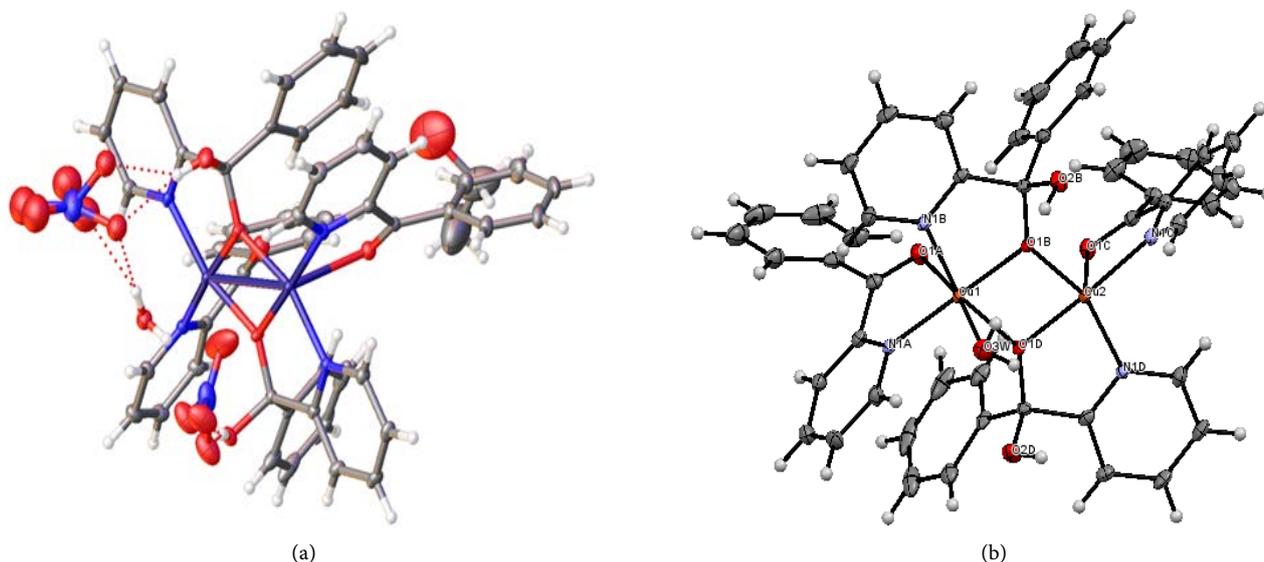


Figure 9. Molecular structure of complex (a) $[\text{Cu}_2(\text{L}_1)_2(\text{L}_2)_2](\text{NO}_3)_2 \cdot \text{H}_2\text{O} \cdot \text{CH}_3\text{CH}_2\text{O}$ (this work); (b) $[\text{Cu}_2(\text{L}_1)_2(\text{L}_2)_2](\text{NO}_3)_2 \cdot \text{H}_2\text{O}$ [27].

5.2. Anti-Radical Activities of Synthetic Products

5.2.1. DPPH Free Radical Scavenging Test

The synthesized compounds were screened for their free radical scavenging activities by DPPH method using Vitamin C as a standard. Antioxidant activities of these compounds were investigated by determining the concentration of substance necessary to reduce 50% of the DPPH[•] radical (EC_{50} values) of each compound. The results of the determination of EC_{50} values of the compounds are shown in **Table 8**. It is evident from these results that, all the compounds do not exhibited antiradical activity even at different concentrations. This observation could be due to the absence of proton in the Schiff base structure and the deprotonation of *gem-diol* proton in the complexes during chelation.

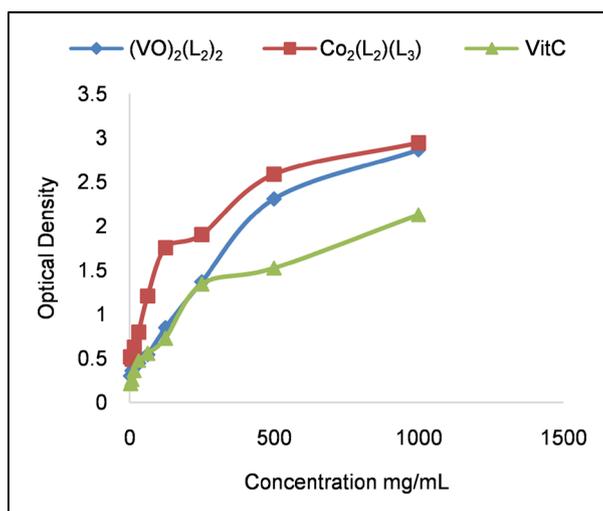
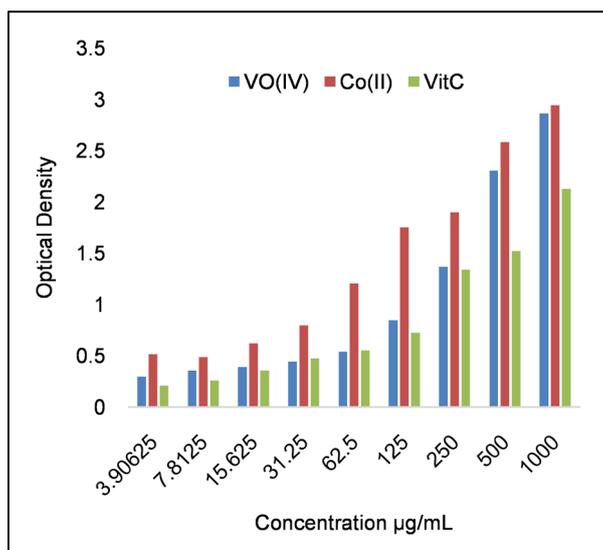
5.2.2. Ferric Ion Reducing Antioxidant Power Assay (FRAP) Test

The reducing powers of the compounds are associated with their antiradical power. The reducing powers of the synthesized compounds were determined using the FRAP method. This technique determines the ability of the tested compounds to reduce ferric iron (Fe^{3+}) present in $\text{K}_3\text{Fe}(\text{CN})_6$ complex to ferrous iron (Fe^{2+}) [38]. Only $[(\text{VO})_2(\text{L}_2)_2]$ and $[\text{Co}_2(\text{L}_2)(\text{L}_3)_2]$ showed reduction capacity which are presented in **Figure 10** and **Figure 11**. It is evident from these results that, the reducing ability of the compounds are concentration dependent [39].

From the result, we find that the reduction of iron by the FRAP method is more pronounced by $[(\text{VO})_2(\text{L}_2)_2]$ and $[\text{Co}_2(\text{L}_2)(\text{L}_3)_2]$ complexes with maxima at Optical Density, OD = 2.87 and 2.94. We can deduce that $[(\text{VO})_2(\text{L}_2)_2]$ and $[\text{Co}_2(\text{L}_2)(\text{L}_3)_2]$ complexes have the capacity to reduce iron. This ability to reduce iron is greater than that of ascorbic acid (OD = 2.12), employed as standard. Thus, the antioxidant activity of these complexes can be attributed to the oxidation of Co(II) to Co(III) and V(IV) to V(V) during the reduction of Iron(III) to Iron(II).

Table 8. EC₅₀ data of Schiff base and binuclear complexes.

Samples	EC ₅₀ (µg/mL)	EC ₅₀ × 10 ³ (µg/mL)
L ₁	/	–
[(VO) ₂ (L ₂) ₂]	/	–
[Co ₂ (L ₂)(L ₃) ₂]	/	–
[Cu ₂ (L ₂) ₂ (L ₃) ₂]	/	–
VitC	9.777 ± 0.520 ^a	0.652 ± 0.029 ^a

**Figure 10.** Ferric ion reducing power of the complexes [(VO)₂(L₁)₂], [Co₂(L₁)(L₂)₂] and Vitamin C at different concentrations.**Figure 11.** Histogram of the OD of VO(IV) and Co(II) complexes and Vitamin C studied by FRAP at different concentrations.

6. Conclusion

We have synthesized a new Schiff base (L_1) derived from 2-aminopyrimidine and 2-benzoylpyridine. The reaction of 1-phenyl-1-(pyridin-2-yl)-N-(pyrimidin-2-yl)methanimine dihydrate (L_1) with salts of V(IV), Co(II) and Cu(II) in a molar ratio of 2:1, resulted in the hydrolysis of L_1 to give binuclear metal complexes of 2-benzoylpyridine (L_2) and phenyl(pyridin-2-yl)methanediol (L_3) and to give alkoxo bridged binuclear complexes. The results of elemental analysis, molar conductivity measurement and spectral analyzes (FT-IR, Uv-Vis) support the suggested structures of metal complexes. The Schiff base and some of the complexes showed no antimicrobial activity but for the $[(VO)_2(L_2)_2]$ complex which was found to be active against *Eschericia coli*, *Klebsiella pneumoniae*, *Staphylococcus aureus* strains. Antioxidant test results revealed that $[(VO)_2(L_2)_2]$ and $[Co_2(L_2)(L_3)_2]$ complexes had higher activity than the reference antioxidant (vitamin C).

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Conflicts of Interest

The authors declare no conflicts of interest regarding the publication of this paper.

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