

Synthesis and Characterization of Lactone Functional Macromonomers by End Group Deactivation and Their Use in Miktoarm Star Polymer

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Received February 29, 2012; revised April 19, 2012; accepted April 28, 2012

ABSTRACT

Newly designed miktoarm star-shaped copolymers made of poly[(benzyl methacrylate(BMA)-co-(ϵ -caprolactone)(CL)] and poly[(BMA-*b*-MMA-*b*-BMA)-co- ϵ -caprolactone)(CL)] were synthesized by combining ring-opening polymerization (ROP) of ϵ -caprolactone (CL) and poly(BMA) five membered lactone functionalized prepared via atom transfer radical polymerization (ATRP) of BMA, and ϵ -CL and P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer, in the presence of tin(II) bis(2-ethylhexanoate) (Sn(Oct)₂). Although lactone ended poly(benzyl methacrylate) with ϵ -caprolactone monomer gave ring open polymerization by Sn(Oct)₂, the macromonomer itself did not give any polymerization. The macromonomers, and the miktoarm star-shaped copolymers were analyzed by FT-IR and ¹H-NMR spectroscopies and GPC (gel permeation chromatograph), Differential scanning calorimetry (DSC-50) and thermogravimetric analysis (TGA-50). These copolymers exhibited the expected structure. The crystallization of star-shaped copolymers was studied by DSC. The results show that when the content of the BMA block increased, the T_m of the star-shaped block copolymer increased.

Keywords: Miktoarm Star Polymer; Lactone; Ring Open Polymerization

1. Introduction

Increasing attention is paid nowadays to the synthesis and characterization of miktoarm star-shaped copolymers that consist of three different polymer chains emanating from a central junction point [1-6]. Generally, star-shaped polymers can be prepared by two different routes: the “arm-first” [7] strategy and the “core-first” [8,9] method. In the arm-first approach, the linear arms of the star-shaped polymer are synthesized first followed by binding of the arms to form the core. The binding of the arms is achieved by using either a difunctional monomer or a multifunctional terminating agent. The core-first method is based on a multifunctional core used as initiator to initiate the polymerization of monomer to form a multi-arm star-shaped polymer [10]. Between the two methods, the core-first approach with multifunctional initiator has received greater attention, since it is easy to control the structure of a star polymer. The “core-first” method has been used in this work to prepare a miktoarm star-shaped copolymer via phloroglucin. Ring-opening polymerization (ROP) [11-13] have been applied to prepare such star-shaped polymers, and the ROP technique has great

advantages in preparing well-defined polymers. PCL and its star-shaped polymers are biodegradable. Therefore, it may have potential applications as biomaterials. Polylactones have potential utility for such usage as a result of their permeability, biocompatibility and biodegradability [14]. One of the convenient strategies in synthesizing these polymers is the ring opening polymerization of the corresponding cyclic lactone monomers or functionally related compounds [15,16]. In recent years, attention on synthesis of star-shaped PCL and star-shaped block copolymers with PCL as their arms has been increasing. For example, Rieger *et al.* [6] reported controlled synthesis of an ABC Miktoarm Star-Shaped Copolymer by sequential ring-opening polymerization of ethylene oxide, benzyl β -Malolactonate, and ϵ -Caprolactone. Thus, they have designed as a macroinitiator for the sequential synthesis of two biocompatible/biodegradable polyesters, *i.e.*, poly(benzyl β -malolactonate) (PMLABz) and poly(ϵ -caprolactone) (PCL) consisting of biocompatible/bioresorbable arms. Lele *et al.* [17] have reported the synthesis of star-shaped poly(ϵ -caprolactone)-*b*-poly(N-(2-hydroxypropyl)methacrylamide) by combination of ring-opening and chain transfer polymerization. Also, Chen *et*

al. [18] have reported newly designed star-shaped block copolymers made of poly(ϵ -caprolactone) (ϵ -PCL) and polystyrene (PS). Hedrick *et al.* [19] have also contributed to this field with their dendrimer-like block copolymer which was obtained upon combining ring-opening polymerization of ϵ -caprolactone for the inner part and ATRP of methyl methacrylate for the external branches. The polymerization of ϵ -caprolactone is commonly initiated by alcohols in the presence of a tin catalysts [20,21].

In this paper, we focused on the synthesis of miktoarm star copolymers using “core-first” method by ring-opening (ROP) process using appropriate (co)polymers with poly(benzyl methacrylate) lacton functionalized and P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer as macromonomer. Thus, the first aim of our study was to synthesize suitable macroinitiators. Low-molecular-weight poly(benzyl methacrylate) one and P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer prepared by atom transfer radical polymerization (ATRP) of benzyl methacrylate at 110°C was used as starting materials for ring open polymerization. Then, the ROP of ϵ -caprolactone with one and dual functionalized macromonomers was carried out in the presence of tin(II)bis(2-ethylhexanoate) (Sn(Oct)₂) at 130°C.

2. Experimental Section

2.1. Materials

MMA was vacuum-distilled after washing with a 5% NaOH aqueous solution just before homopolymerization. Cuprous(I) bromide, 2,2'-bipyridyne (bpy), and ethyl 2-bromoacetate (analytical reagent) were used as received. 2-Chloroacetyl chloride (Aldrich, 99%) was used as received. The ϵ -caprolactone (CL, Aldrich) was stirred with CaH₂ overnight at 80°C and then distilled under vacuum. Phloroglucin dihydrate (Aldrich) was dried to remove dihydrat *in vacuo* (at 2 mmHg) at 120°C. Benzyl methacrylate (BMA) was prepared by aid of literature [22]. So, BMA was synthesized by the reaction of benzyl alcohol and methacryloyl chloride at 60°C - 70°C by using potassium carbonate. It was distilled *in vacuo* (b.p.: 103°C at 2 mmHg).

2.2. Synthesis of 7-Hydroxy-4-chloromethyl Coumarin (CIMOHC)

Synthesis of 7-hydroxy-4-chloromethylcoumarin was prepared as follows [23].

To resorcinol (1.1 g, 10 mmol) stirring in concentrated H₂SO₄ (20 ml) and ethyl acetoacetate (1.3 g, 10 mmol) was added. The mixture was stirred for 10 h at room temperature, then poured into water (100 ml) and left to stand overnight. The corresponding 7-hydroxy-4-chloromethyl coumarin was precipitated and was recovered by

filtration and then purified using 9:1 n-hexane/ethylacetate as eluent.

¹H-NMR (δ , ppm): 4.89 (s, ²H, Cl-CH₂-), 6.40 (s, ¹H, olefinic proton on coumarin), 6.79 and 6.91 (protons on 5 and 6 carbons in aromatic ring), 7.71 (s, ¹H, between two rings coumarin group).

¹³C-NMR (δ , ppm): 41.3 (Cl-CH₂-), 126.2, 112.9, 111.6, 110.1, 102.8 (carbons with hydrogen on coumarin)

2.3. Synthesis of 4-(Chloromethyl)-2-oxo-2H-chromen-7-yl Chloroacetate (CMOCA)

4-(Chloromethyl)-2-oxo-2H-chromen-7-yl chloroacetate was prepared as follows [24,25]. 7-Hydroxy-4-chloromethyl-lcoumarin (2.1 g, 0.01 mol), triethylamine (1.01 g, 0.01 mol), and tetrahydrofuran (20 mL) were added into a 100 mL-round-bottom flask equipped with a magnetic stirrer. After cooling to 0°C with an ice bath, a solution of a chloroacetylchloride (1.13 g, 0.01 mol) in 5 mL of tetrahydrofurane dried was added dropwise. The reaction mixture was then stirred at room temperature for additional 24 h, and the salt formed in the reaction was filtered off. The filtrate was washed with NaHCO₃ solution and following by water until neutralization, and the solution was dried over MgSO₄. After filtration, the solvent was removed by rotary evaporator, and then purified using 9:1 hexane/ethylacetate as eluent.

¹H NMR of CMOCA (δ , ppm): 4.34 (s, ²H, -CH₂-Cl), 4.65 (s, ²H, -C(=O)CH₂-Cl), 6.55 (s, ¹H 8 on aromatic ring, 7.15 (6 on aromatic ring) 7.69 (5 on aromatic ring).

2.4. Characterization

¹H NMR spectra were recorded on a Bruker AVANCE-400 nuclear magnetic resonance (NMR) instrument with CDCl₃ as solvent and tetramethylsilane (TMS) as internal standard. Infrared spectra were obtained on a Perkin Elmer Spectrum one FT-IR spectrometer.

Thermogravimetric analysis (TGA) measurements were carried out under a nitrogen flow with a TGA-50 thermobalance at a heating rate of 10°C min⁻¹. Differential scanning calorimetry (DSC) thermograms of polymers were measured by DSC-50. The molecular weight and molecular weight distribution of the polymers were measured at 30°C on a Agilent 1100 gel permeation chromatography (GPC) equipped with microstyragel columns and detector. Molecular weights were calibrated against narrowed polymethylmethacrylate standards. THF was used as eluent at a flow rate of 1.0 mL/min.

2.5. Synthesis of Lacton Functionalized Macromonomer by ATRP Method

The ATRP reaction of lacton functionalized macromo-

monomer was carried out according to procedure recorded in literature [22].

2.6. Synthesis of Poly(Methyl Methacrylate) Ended 7-Hydroxy Coumarin by ATRP Method

The dried polymerization tube was charged with CuBr (14.4 mg, 0.1 mmol), 2,2'-bipyridine (31.2 mg, 0.2 mmol), 7-hydroxy-(4-chloromethyl)coumarin (21 mg, 0.1 mmol) and passed from Ar gas. Then, the methyl methacrylate (1 gr, 10 mmol) was added into polymerization tube. The tube was sealed under vacuum after three freeze-pump-thaw cycles, and the sealed tube was immersed into an oil bath at 110°C. After 12 h, the polymerization mixture was cooled to ambient room and precipitated three times in ethyl alcohol containing 1% HCl solution. The resultant product was dried in a vacuum oven at 40°C for 24 h.

2.7. Synthesis of Dual Lacton Functionalized Macromonomer

The 4-[(PMMA)methyl]-2-oxo-2H-chromen-7-yl chloroacetate (CAPMMA) was synthesized by reaction of 7-hydroxy-[(4-PMMA)methyl]coumarin (HAPMMA) and chloroacetyl chloride by using CIMOHC as initiator in ATRP in presence of triethylamine at 0°C - 5°C. The CAPMMA was used in synthesis of macromonomer dual lacton functionalized. For this purpose, the dried polymerization tube was charged with CuBr (7.2 mg, 0.05 mmol), 2,2'-bipyridine (315.6 g, 0.1 mmol), ($M_n = 4000$ gr/mol; 315 mg, 0.05 mmol) and passed from Ar gas. Then, benzyl methacrylate (0.88 gr, 5 mmol) was added into the polymerization tube. The tube was sealed under vacuum after three freeze-pump-thaw cycles, and the sealed tube was immersed into an oil bath at 110°C. After 72 h, the polymerization mixture was cooled to ambient room and precipitated three times in cooled ethylalcohol. The resultant product was dried in a vacuum oven for 24 h at 40°C.

2.8. Synthesis of Three Armed Miktoarm Star Shaped Copolymer from Macromonomer with ϵ -Caprolacton

Phloroglucin (0.001 mol), monomer (ϵ -CL) (0.06 mol), macromonomer (poly(BMA) ended lacton) and stannous octoate (7.5×10^{-6} mol) were added under nitrogen in previously flamed and argon-purged schlenk tube equipped with magnetic stirrer. The ring opened polymerization of ϵ -CL was carried out in bulk at 130°C. After 48 h, the polymerization was terminated by cooling the tube to the room temperature, then diluted with CH_2Cl_2 and poured into n-hexane. The resulting polymer

was dried at 30°C in a vacuum for 2 days.

3. Results and Discussion

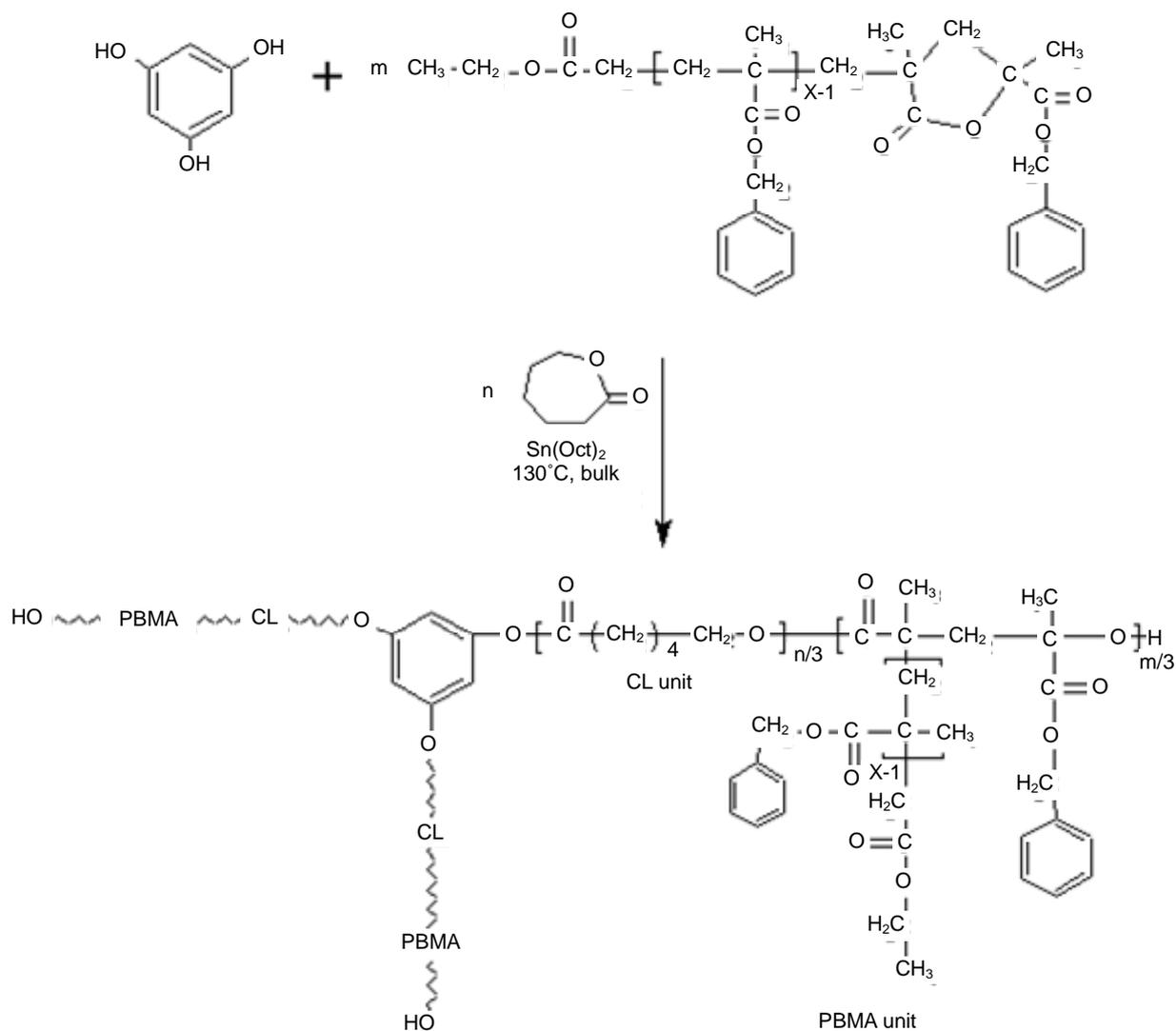
Our strategy for the synthesis of lacton functional macromonomers by end group deactivation during ATRP of benzyl methacrylate was to design miktoarm star copolymers with ϵ -caprolacton (**Scheme 1**).

The linear poly(benzyl methacrylate) lacton ended was synthesized by ATRP method at 110°C with CuBr (0.1 mmol), 2,2'-bipyridine (0.2 mmol) and ethyl 2-bromoacetate as initiator. The FT-IR and $^1\text{H-NMR}$ spectra of poly(BMA) obtained via ATRP are shown in **Figures 1(a)** and **2(a)**, respectively. The strong band at 1730 cm^{-1} is the stretching vibration of the carbonyl group in ester, and the bands at 1162 cm^{-1} and 1253 cm^{-1} are symmetric stretching and asymmetric stretching of the $-\text{C}(=\text{O})-\text{O}-\text{C}$ group, respectively. The intensity band at 1779 cm^{-1} showed a lactone ring with five members. It may be suggested that a lactone ring forms by means of removing benzyl bromide at the chain end, as suggested in **Scheme 2**. This band is not observed in that of poly(BMA) synthesized under free-radical polymerization conditions [22]. Designed miktoarm star shaped copolymers made of poly[(ϵ -caprolacton)(ϵ -PCL)-co-benzyl methacrylate(BMA)] were synthesized by combining ring-opening polymerization (ROP), as was illustrated in **Scheme 3**.

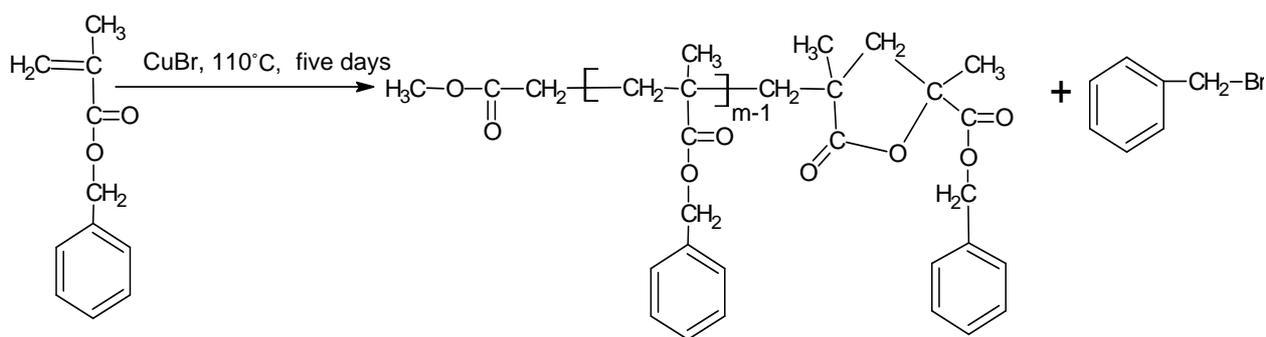
The structure of copolymer synthesized was characterized by FT-IR (**Figure 1(b)**) and $^1\text{H-NMR}$ (**Figure 2(b)**) spectra. The FT-IR spectrum of miktoarm star shaped poly[(ϵ -caprolacton)(PCL)-co-benzyl methacrylate (BMA) prepared from ring opened not showed the band at 1779 cm^{-1} , which characterizes lacton ring. Disappearance of this band is an important evidence of ring open for poly(BMA) lacton ended as macromonomer at 130°C. The $^1\text{H NMR}$ spectrum of poly(BMA) by ATRP are illustrated in **Figure 2(a)**, and the formation reaction of miktoarm star copolymer is showed in **Scheme 1**. The peaks at 7.16 - 7.32 ppm reveal the phenyl ring and the signal at 4.92 ppm assigned to methylene is adjacent to oxygen in the BMA unit, the signal at 5.18 ppm was assigned to methylene protons of benzyl group on lacton ring. This demonstrates that the PBMA ended lacton group has completely been copolymerized with ϵ -CL. But this band did not disappear in only ROP of poly(BMA) lacton ended. So, although poly(benzyl methacrylate) ended lactone with ϵ -caprolacton shows ring open polymerization by $\text{Sn}(\text{Oct})_2$, poly(benzyl methacrylate) lactone ended did not show ring opening reaction. This is probably due to that unsubstituted cyclic esters activated by $\text{Sn}(\text{Oct})_2$ are attacked by nucleophiles like hydroxyl groups, but because of steric hinder of substituents on lacton ring is not activated by $\text{Sn}(\text{Oct})_2$.

The most reliable method to determine the actual molar mass or composition of such copolymers is by $^1\text{H-NMR}$. The copolymer composition of the miktoarm star shaped P(BMA-co- ϵ -CL) was determined based on the integral values of the separate signals, which are chara-

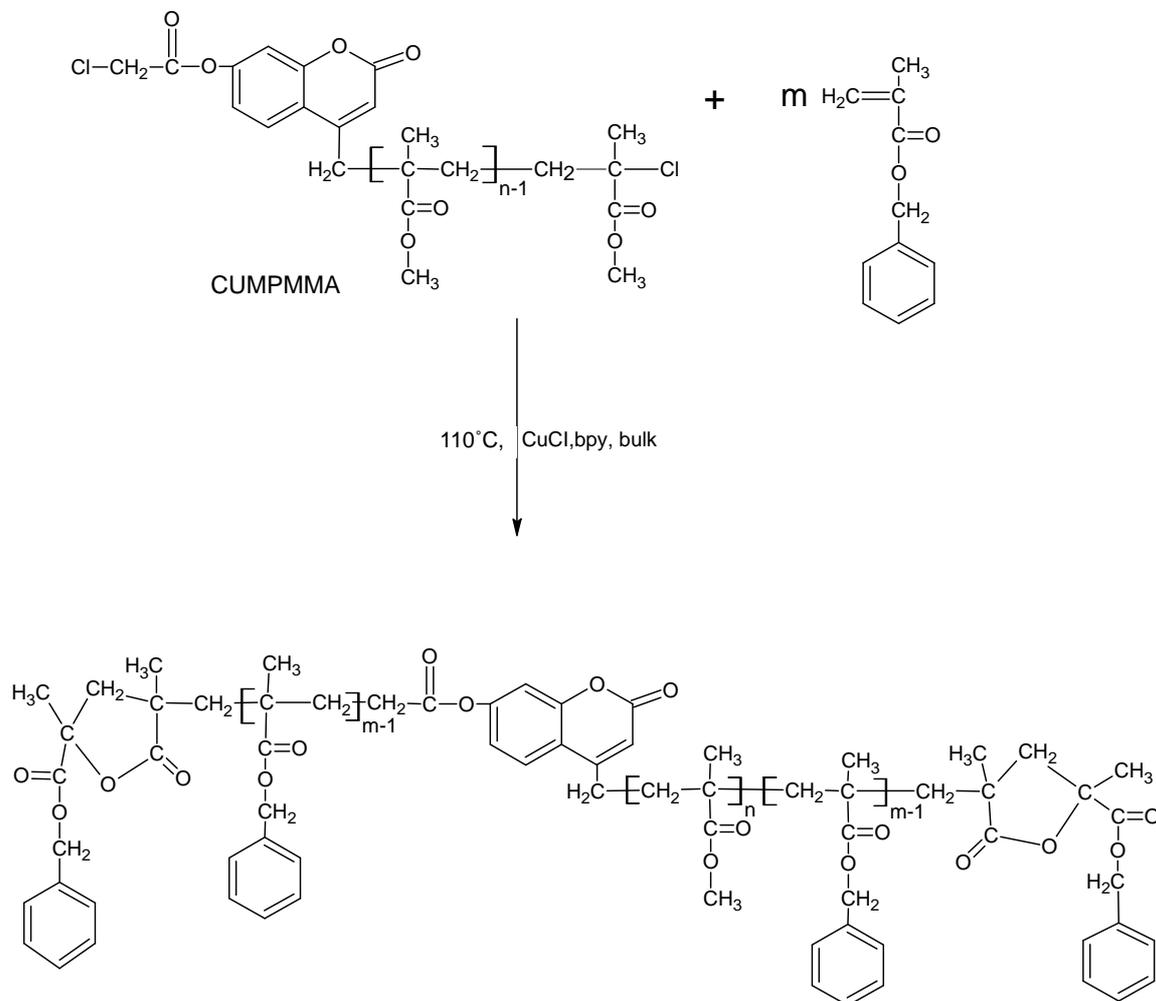
cteristic for BMA and ϵ -CL moieties at 4.86 ppm (20 mm) and 4.05 ppm (2.1 mm), respectively. The copolymer compositions by using the integration values of the separate signals, can be calculated according to following equation:



Scheme 1. The structure of miktoarm star shaped poly(ϵ -CL-co-BMA).



Scheme 2. The structure of lacton functionalized poly(BMA) [22].



Scheme 3. The reaction of P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer.

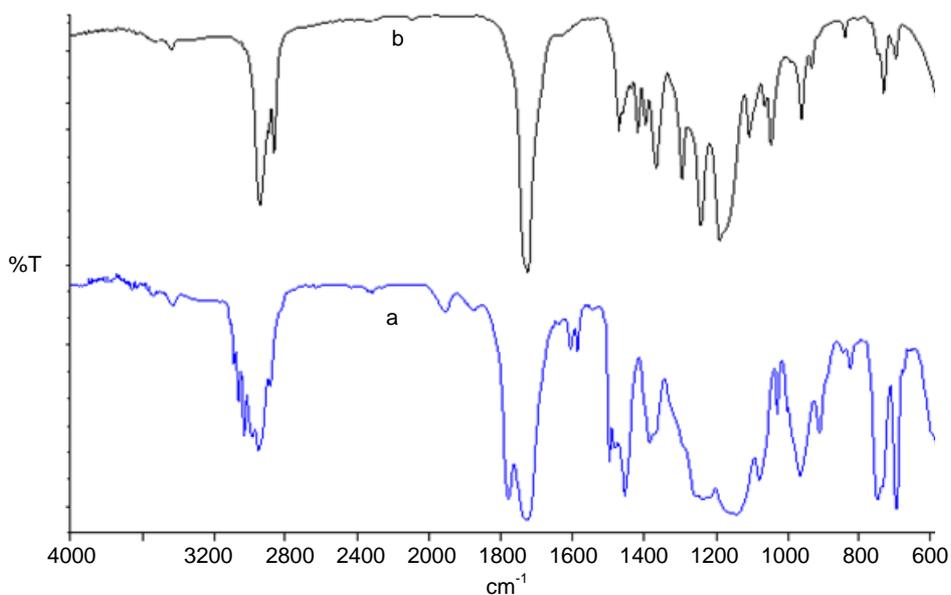


Figure 1. FT-IR spectra of (a) Poly(BMA) ended lacton obtained by ATRP method (b) Miktoarm star shaped P(BMA-co- ϵ -CL) obtained by ring opened polymerization of poly(BMA) ended lacton with ϵ -CL.

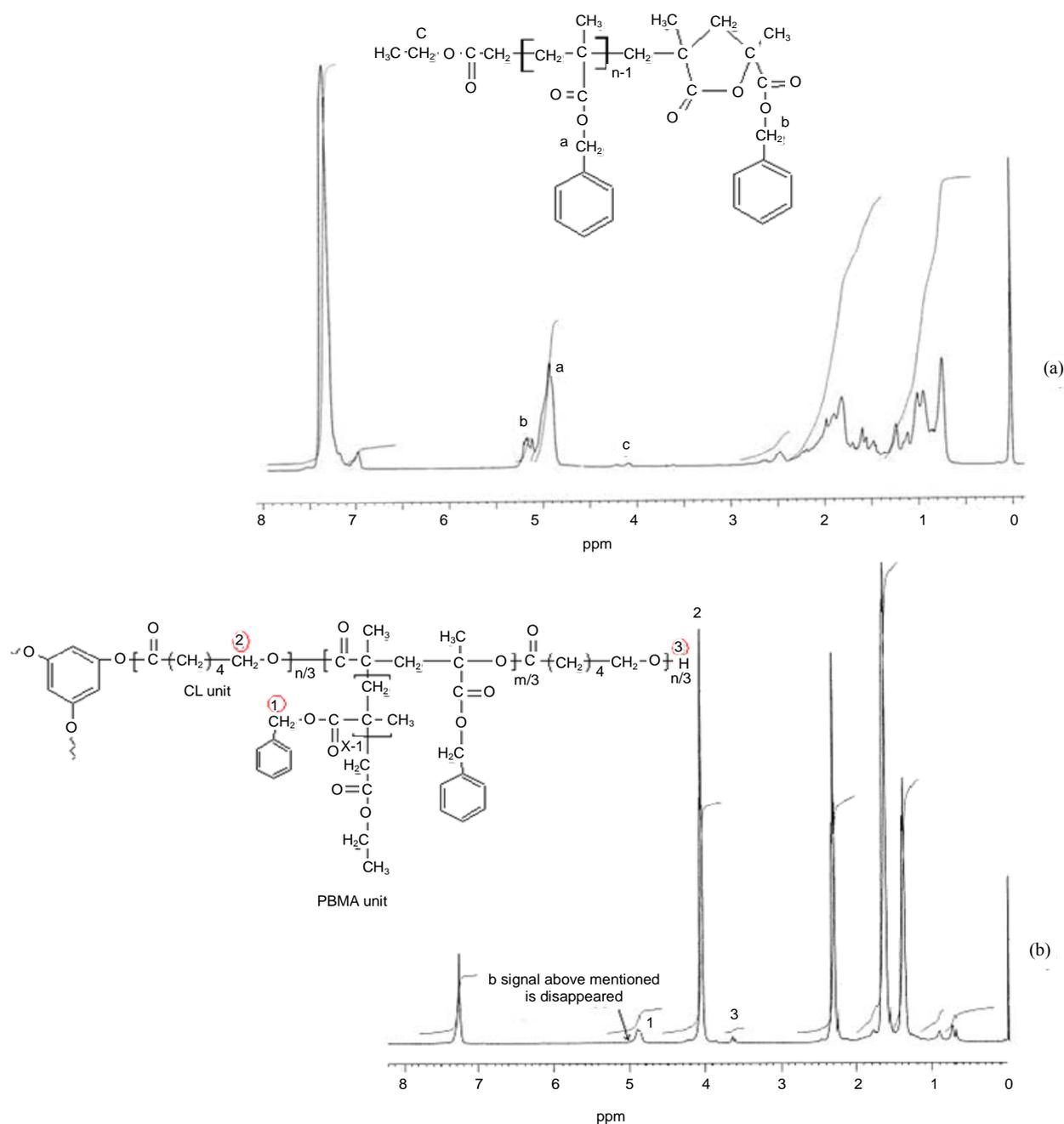


Figure 2. ¹H-NMR spectra of a) Poly(BMA) ended lacton by prepared ATRP method; b) Miktoarm star shaped poly(BMA-co-ε-CL).

$$C = \frac{\int \text{height of CH}_2 \text{ protons in BMA units}}{\int \text{height of CH}_2\text{OC- protons in MMA units}} = \frac{2m_1}{2m_2}$$

$$\text{On simplification; } m_1 = \frac{C}{C+1}$$

where m_1 and m_2 are the molar fractions of BMA and -ε-CL in the copolymer, with lacton end group. Copolymer compositions of BMA and -ε-CL were determined

as 9.5% and 90.5%, respectively.

The copolymer composition of P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer was determined based on the integral values of the separate signals, which are characteristic for MMA and BMA moieties at 3.63 ppm and 4.89 ppm, respectively. The copolymer compositions by using the integration values of the separate signals, can be calculated according to following equation:

$$C = \frac{\int \text{height of CH}_2 \text{ protons in BMA units}}{\int \text{height of CH}_3 \text{ protons in MMA units}} = \frac{2m_1}{3m_2}$$

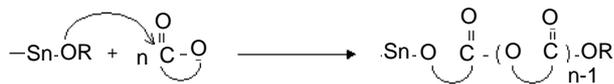
$$\text{On simplification; } m_1 = \frac{3C}{3C+2}$$

where m_1 and m_2 are the molar fractions of BMA and MMA in the copolymer, with lacton end group. Copolymer compositions of MMA and BMA were determined as 68% and 32%, respectively.

The number-average molecular weight ($M_{n,NMR}$) of poly(BMA) ended lacton has been calculated by following equation :

$$\begin{aligned} M_{n,NMR} \text{ (PBMA ended lacton)} \\ = (I_{4.91}/I_{5.18}) \times M_{BMA} = (1.80/0.18) \times 176 = 1780 \end{aligned}$$

The polymerization of ϵ -caprolacton is commonly initiated by alcohols in the presence of a tin catalysts [20,21]. Some groups of authors [26-28] advanced schemes in which cyclic esters activated by $\text{Sn}(\text{Oct})_2$ are attacked by nucleophiles like hydroxyl groups coming from the chain ends and reforming the unit terminated with $-\text{OH}$. Two of groups cited above proposed detailed chemistry in which a secondary oxonium ion appears. The schemes mentioned above strikingly differs from the one in which tin atoms are becoming a part of the active species at the growing polymer chain-end [29].



Poly(MMA) 7-(chloroacetoxyloxy) coumarin ended (CAPMMA) was prepared from reaction of poly(MMA) 7-hydroxy coumarin ended (HAPMMA) with chloroacetylchloride in presence of triethyl amin. **Figure 3** shows a typical $^1\text{H-NMR}$ spectrum of HAPMMA (a in **Figure 3**) and CAPMMA (b in **Figure 3**). It clearly shows that besides the typical signals of coumarin group, which is the end group of PMMA chain at 8.3 ppm (δH (a)), 3.7 ppm (δH (b)) $-\text{OCH}_3$ next to chloride, which is on the other end group of PMMA chains, and 3.6 ppm (δH (e)) is characteristic for CH_3 protons of MMA units repeating. Otherwise, when a in **Figure 3** compared with b in **Figure 3**, it is seen that the terminal $\text{Cl}-\text{CH}_2\text{C}=\text{O}$ methylene (g) signals of the macroinitiator (HAPMMA) at 4.27 ppm. c in **Figure 3** shows $^1\text{H-NMR}$ spectrum of poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton. The signals of 4.95 ppm represents $-\text{OCH}_2-$ (peak h), the signals at 5.18 ppm represents $-\text{OCH}_2-\text{Ar}$ (peak i) on lacton group. The signals at 3,6 ppm represents CH_3 (peak k) in MMA units. The FT-IR spectra of poly(methyl methacrylate) ended 7-hydroxy coumarin, poly(MMA) 7-(chloroacetoxyloxy) coumarin ended (CAPMMA), p(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock

copolymer and miktoarm star shaped poly[(ϵ -caprolacton) (PCL)-co-(BMA-*b*-MMA-*b*-BMA)] were showed in **Figure 4**. The most characteristic bands for HAPMMA (a in **Figure 4**) are at 3450 cm^{-1} (OH stretching) and 1730 cm^{-1} ($\text{C}=\text{O}$). Incorporation to HAPMMA of chloroacetyl groups was further evidenced as a shoulder by 1758 cm^{-1} for $\text{C}=\text{O}$ ($\text{Cl}-\text{CH}_2\text{C}=\text{O}$) b in **Figure 4**. **Figure 4** (c) shows FT-IR spectrum of p(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer. The strong band at 1731 cm^{-1} is due to the stretching vibration of the ester carbonyl group in both MMA and BMA units. The bands at 1162 cm^{-1} and 1253 cm^{-1} are symmetric stretching and asymmetric stretching of the $-\text{C}(\text{O})-\text{O}-\text{C}$ group, respectively. The intensity band at 1775 cm^{-1} showed a lactone ring with five members. It may be suggested that a lactone ring forms by means of removing benzyl bromide at the chain end, as suggested in **Scheme 4**. The FT-IR spectrum of miktoarm star shaped poly[(BMA-*b*-MMA-*b*-BMA)-co- ϵ -CL)] synthesized from ring opened copolymerization of p(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized with ϵ -caprolacton at 130°C not showed the band at 1779 cm^{-1} , which characterizes $\text{C}=\text{O}$ in lacton ring with five members. This phenomenon shows that lacton groups form copolymer with ϵ -CL. The ^1H NMR spectrum of miktoarm star shaped poly[(BMA-*b*-MMA-*b*-BMA)-co- ϵ -CL)] showed peaks at 7.16 - 7.32 ppm reveal the phenyl ring and the signal at 4.92 ppm assigned to methylene adjacent to oxygen in the BMA unit, and not showed the signal at 5.18, which characterizes methylene of benzyl group on lacton ring. This is another important evidence of ring open for lacton group of p(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized during copolymerization with ϵ -CL.

3.1. GPC Measurements

The average molecular weights of the poly(BMA), three armed ϵ -PCL and their copolymers in THF were determined by GPC calibrated using poly(methyl methacrylate) standards, and the single GPC traces of the poly(BMA) and all the star shaped homo and copolymers are showed in **Figures 5** and **6**. **Figure 5** shows the typical GPC curves of poly(BMA) ended lacton (a curve in **Figure 5**), miktoarm star shaped P(BMA-co- ϵ -CL) (b curve in **Figure 5**) were showed as compared with those of original star-shaped PCL. The M_n and polydispersity for P(BMA) ended lacton by ATRP method are 2900 and 1.20, respectively. Those of miktoarm star shaped P(BMA-co- ϵ -CL) are 8600 and 1.24, respectively. In this case, the GPC trace of miktoarm star shaped P(BMA-co- ϵ -CL) shows a very significant shift to higher molecular weights as compared to linear poly(BMA), providing that the new polymer was readily formed. In **Figure 5**, the GPC traces of the starting poly(BMA) is also presented.

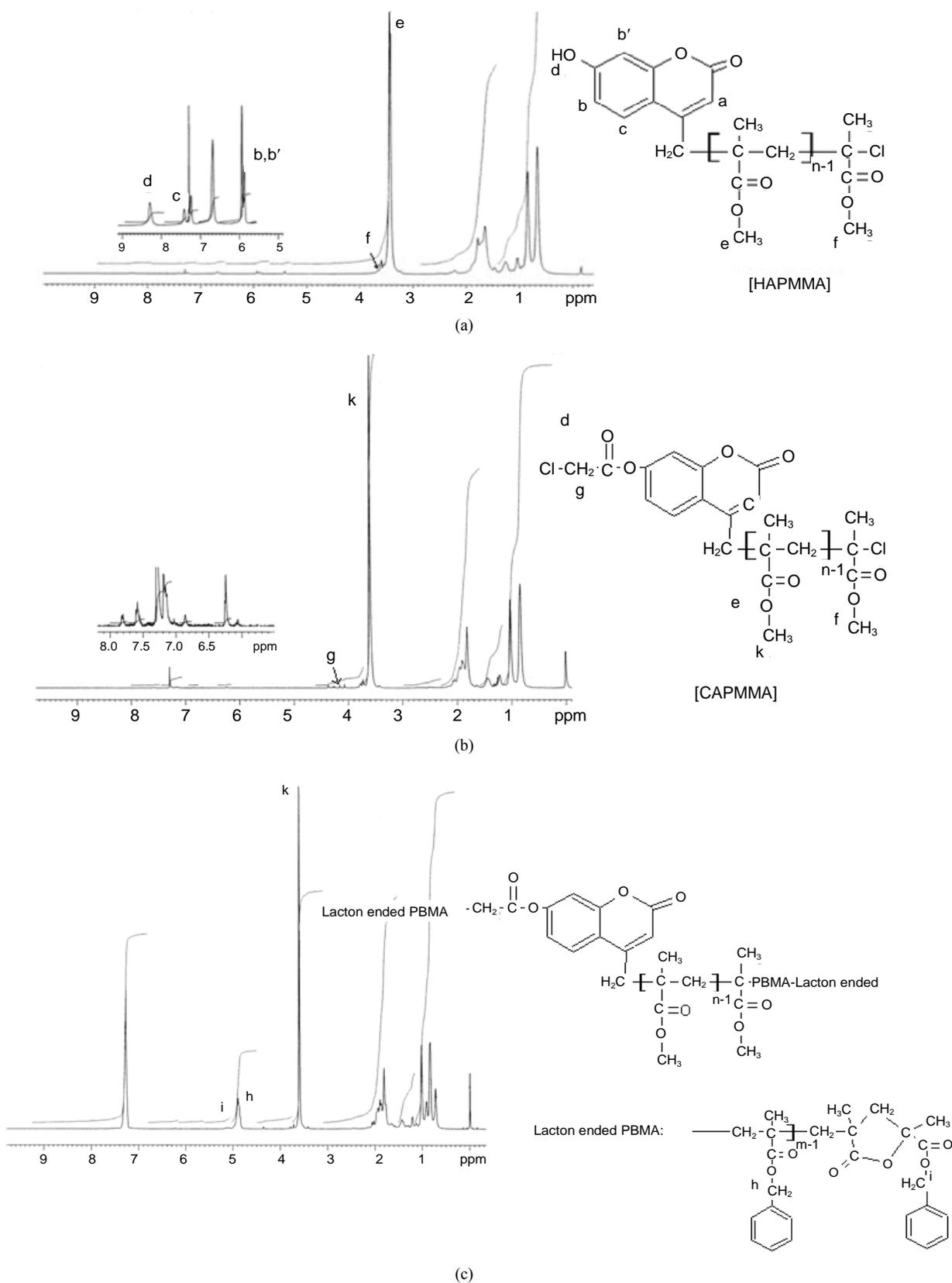


Figure 3. $^1\text{H-NMR}$ spectra of (a) Poly(MMA) 7-hydroxy coumarin ended (HAPMMA); (b) Poly(MMA) 7-(chloroacetoxy) coumarin ended (CAPMMA); (c) P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer.

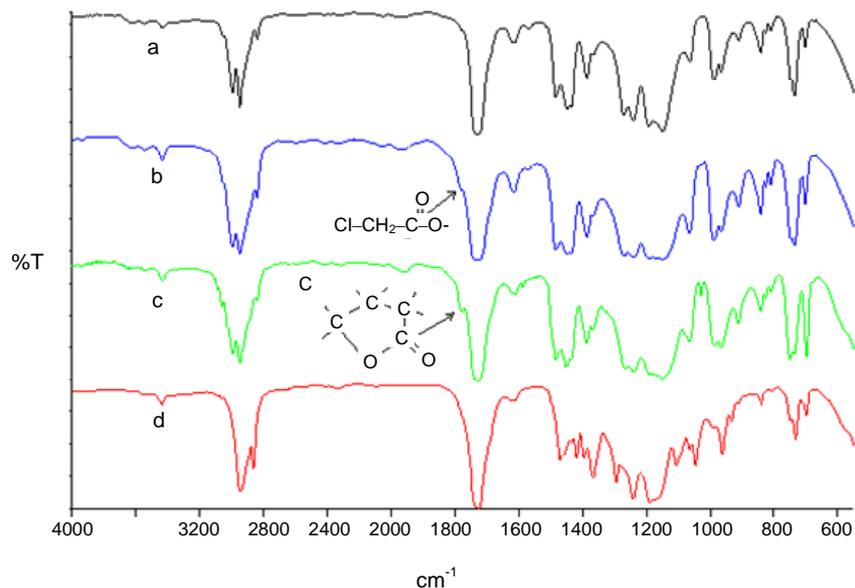
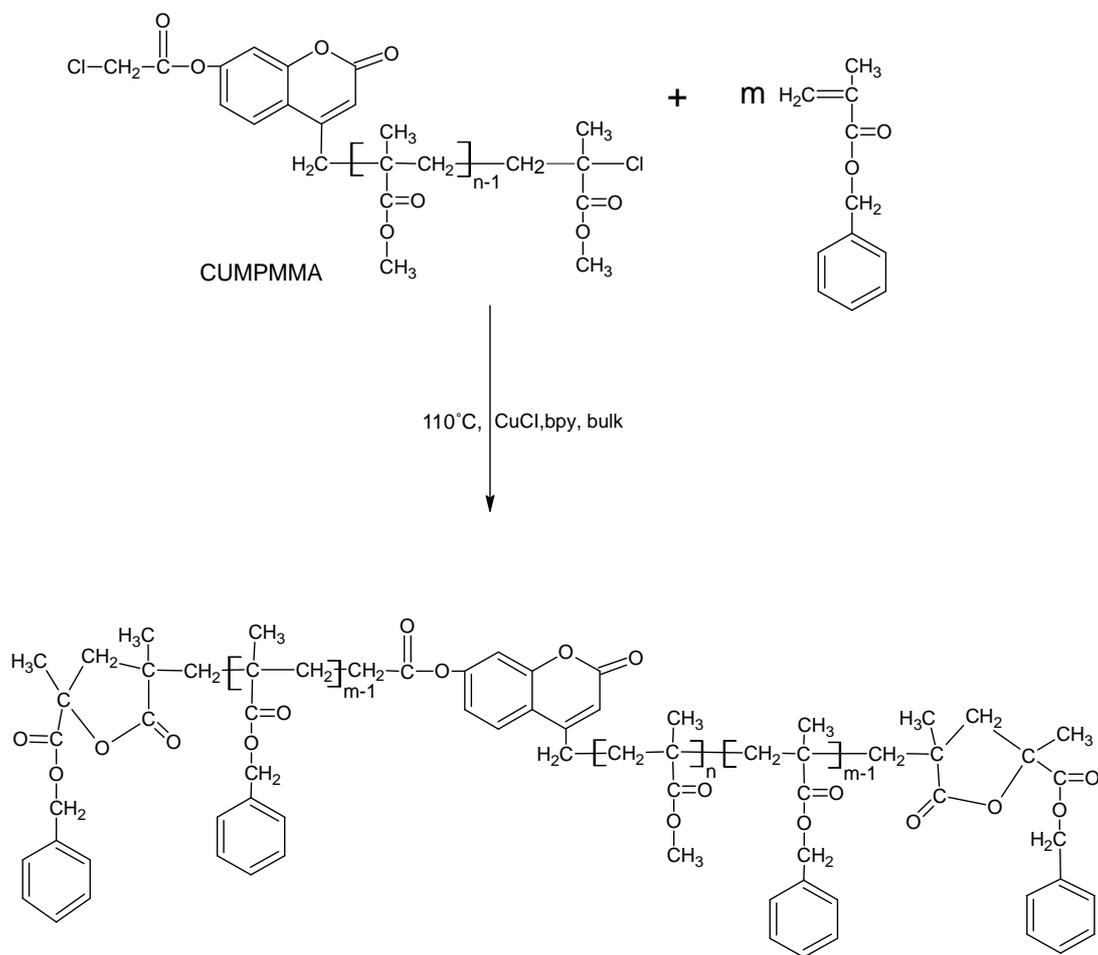


Figure 4. FT-IR spectra of (a) Poly(methyl methacrylate) ended 7-hydroxy coumarin; (b) PMMA 7-chloroacetoxyloxy coumarin ended (CAPMMA); (c) P(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized diblock copolymer; (d) Miktoarm star shaped P[(BMA-*b*-MMA-*b*-BMA)-*co*-(ϵ -CL)].



Scheme 4. The reaction of P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer.

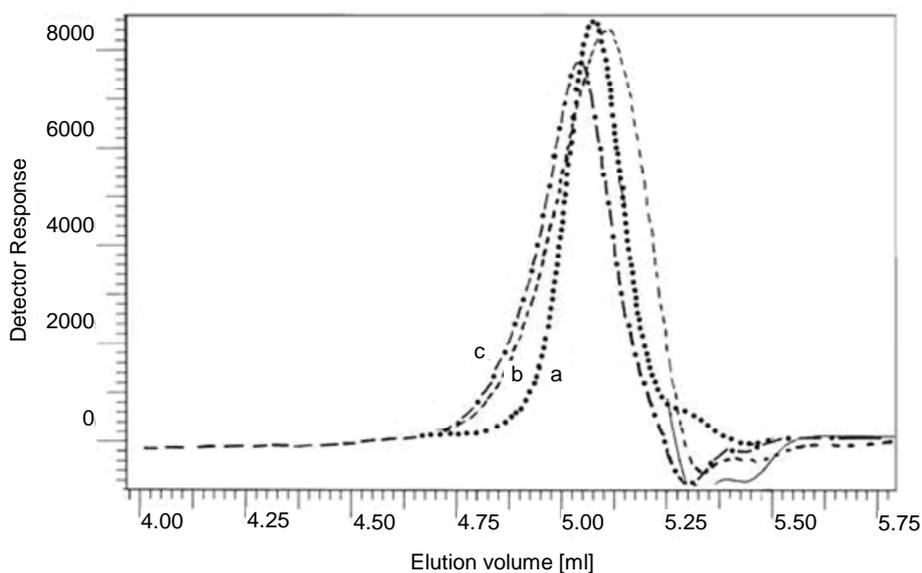


Figure 5. GPC curves of (a) P(BMA) ended lacton; (b) Mikroarm star shaped P(BMA-co- ϵ -CL); (c) Three armed P(ϵ -CL).

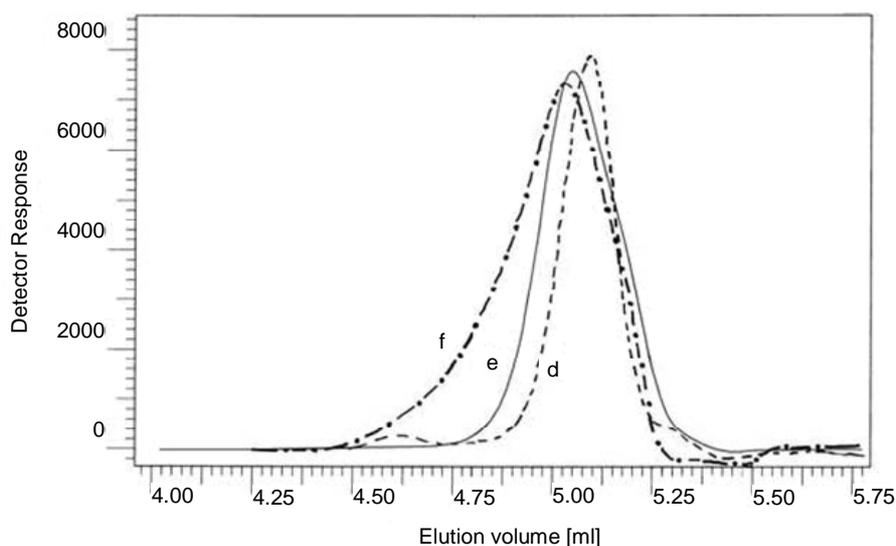


Figure 6. GPC curves of (d) PMMA ended 7-hydroxy coumarin; (e) P(BMA-b-MMA-b-BMA) dual lacton functionalized diblock copolymer; (f) Mikroarm star shaped P[(BMA-b-MMA-b-BMA)-co-(ϵ -CL)].

The $M_{n,NMR}$ value of poly(BMA) ended lacton was estimated from the ratio of the methylene (b) and terminal one (a). The $M_{n,NMR}$ value (2900) was close to $M_{n,GPC}$ (3050) (Table 1). That of poly(BMA-co- ϵ -CL) mictoarm star shaped (f curve in Figure 6) was 8600 and its polydispersity was 1.24. According to 1H -NMR (b in Figure 2), ϵ -PCL composition in the well-defined mictoarm star-shaped P(BMA-co- ϵ -CL) with terminal hydroxyl group was 89.5% (by mol). The $M_{n,NMR}$ value of ϵ -PCL in three-armed P(BMA-co- ϵ -CL) was estimated from the ratio of the methylene proton (c) and terminal one (d). Integration heights of c is 200 mm, that of d is 4. So, $M_{n,NMR}$ value of PCL in copolymer is 5700. The

peaks made it possible to evaluate the polymerization degree of PCL in copolymer. The value $n_{CL} = 16$ means that contain CL unit per arm. The M_n , GPC value of three armed poly(ϵ -CL) (e curve in Figure 6) prepared under synthesis conditions of mikroarm star shaped P(BMA-co- ϵ -CL). This means that contain 29 CL units per arm. Decreasing the polymerization degree of PCL in copolymer is as a result of being difficult in ring open of lacton group.

To prepare poly(BMA-b-MMA-b-BMA) dual lacton functionalized was used poly(MMA-b-BMA) ended 7-chloroacetyloxy coumarin. Figure 6 shows the typical GPC curves of poly(MMA-b-BMA) ended 7-hydroxy

Table 1. GPC data of polymers.

Polymer	M_n	$HI (=M_w/M_n)$
P(BMA) ended lacton	2900	1.20
Miktoarm star shaped P(BMA-co- ϵ -CL)	8600	1.24
Star shaped P(ϵ -CL)	10,000	1.20
PMMA ended 7-hydroxy coumarin	4000	1.10
P[(BMA- <i>b</i> -MMA- <i>b</i> -BMA) dual functionalized	8900	1.15
Miktoarm star shaped P[(BMA- <i>b</i> -MMA- <i>b</i> -BMA)-co-(ϵ -CL)]	11,500	1.55

coumarin (d curve in **Figure 6**), poly(BMA-*b*-MMA-*b*-BMA), macromonomer, dual lacton functionalized (e curve in **Figure 6**) and miktoarm star shaped poly(BMA-*b*-MMA-*b*-BMA)-co- ϵ -CL (f curve in **Figure 6**). The $M_{n, GPC}$ value of poly(MMA-*b*-BMA) ended 7-chloroacetyloxy coumarin was 4000 and M_w/M_n is 1.1. That of poly(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized was 8900 and polydispersity was 1.15. The polymerization degree of BMA, MMA and BMA in macromonomer dual lacton functionalized was 14:100:14, respectively. The GPC results were summarized in **Table 1**. They indicate that in each block copolymer the peak is shifted toward a lower molecular weight region compared with that of its original linear poly(BMA) with a little change in molecular weight distribution.

3.2. DSC Measurements

The glass transition temperatures of poly(ϵ -caprolacton)(PCL) three armed, miktoarm poly[(ϵ -caprolacton) (ϵ -PCL)-co-benzylmethacrylate(BMA)] three armed, miktoarm poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton with coumarin core, poly(MMA) ended 7-hydroxy coumarin and poly(MMA-*b*-BMA) ended 7-chloroacetyloxy coumarin were measured by DSC. The samples were heated from 20°C to 200°C at a rate of 20°C/min. Representative DSC curves of the polymers were shown in **Figure 7**. The melting point of poly(ϵ -caprolacton) (PCL) three armed was measured at 59.5°C (a in **Figure 7**). In case of miktoarm star shaped poly[(ϵ -caprolacton) (PCL)-co-benzyl methacrylate(BMA)] (b in **Figure 7**) is 62.5°C. The increasing in this temperature means that BMA units in copolymer influences the crystallinity of PCL. The T_g of the linear poly(BMA) prepared by ATRP method is 54°C [22] and is lower than the T_m of PCL. But not observing of T_g for the poly(BMA) with amorphous character is due to overlap with the peak area of T_m for poly[(ϵ -caprolacton)(PCL)-co-benzyl methacrylate (BMA)]. The poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton with coumarin core (c in **Figure 7**) bearing MMA, BMA and CL units showed only a T_m at 66.9°C.

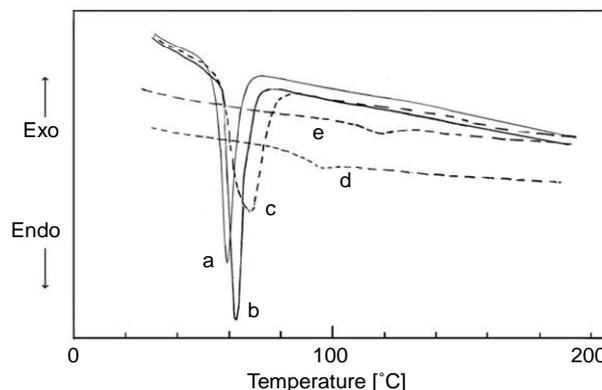


Figure 7. DSC curves of (a) Three armed P(ϵ -CL); (b) Miktoarm star shaped P(BMA-co- ϵ -CL); (c) Miktoarm star shaped P[(BMA-*b*-MMA-*b*-BMA)-co-(ϵ -CL)]; (d) P(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized diblock copolymer; (e) PMMA 7-chloroacetyloxy coumarin ended.

Although the T_g of the PMMA block is higher than the T_m of PCL block, T_g of PMMA is not observed. This is probably due to solubilization of PMMA amorphous blocks within PCL liquid phase melting at 66.9°C. The T_g of PMMA 7-chloroacetyloxy coumarin ended used in synthesis of poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton with coumarin core was measured as 83°C by DSC. This temperature is between that of BMA and MMA homopolymers and it is showing a single phase behavior. Most of the polymers having relatively low molecular weight are generally compatible [30].

3.3. Thermogravimetric Studies of the Polymers

Thermal studies of miktoarm star shaped poly[(BMA- ϵ -CL)] and poly(MMA) ended 7-hydroxy coumarin, poly(MMA) ended 7-chloroacetyloxy coumarin and poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton with coumarin core were investigated by thermogravimetric analysis. The thermogravimetric curves of polymers measured between room temperature to 500°C at a heating rate of 10°C/min under nitrogen flow are shown in **Figure 8** in comparison with those of poly[(ϵ -caprolacton)(PCL), poly(BMA) ended lacton by ATRP method. In general, the polymers showed a thermal degradation with a two stages. For first stage, the miktoarm star shaped P(ϵ -CL-co-BMA) (b curve in **Figure 8**) decomposes at 235°C. The second stage, that is, more rapid weight loss random chain scission and depolymerization to BMA starts at 320°C. While poly(ϵ -caprolacton) three armed (e curve in **Figure 8**) decomposes at 310°C, linear poly(BMA) ended lacton prepared by ATRP method starts at 150°C. Thus, P(ϵ -CL-co-BMA) three armed decomposes (at 230°C) between those of homopolymers with both units. As was seen, b curve in **Figure 8**, while thermal stability of poly((ϵ -CL-co-BMA) observes lower than that of poly(BMA-*b*-MMA-*b*-BMA) ended dual lacton with cou-

marin (a curve in **Figure 8**) core to 317°C, that of mikroarm star shaped P(ϵ -CL-co-BMA) three armed after it temperature increases with a important different. This propably results in depolymerization of MMA block in poly(BMA-*b*-MMA-*b*-BMA). Poly(MMA) prepared by free radical method in this study decomposes in a single stage at about 272°C, but that poly(MMA) ended 7-hydroxy coumarin (c curve in **Figure 8**) decomposes at the first stage between 292 and 420°C by 97% weight loss, and it may be result in that coumarin group retards depolymerization tendency of PMMA ended 7-hydroxy coumarin in comparison that of poly(MMA) prepared by free radical polymerization method. Poly(MMA) ended

7-hydroxy coumarin (c curve in **Figure 8**) decomposes at the first stage between 292°C and 420°C by 97% weight loss. The initial decomposition temperature or stability of mikroarm star shaped P[(BMA-*b*-MMA-*b*-BMA)-co-(ϵ -CL)] (d curve in **Figure 8**) is higher in comparison to the other copolymers. It means that the MMA and ϵ -CL moities delay chain scissions and volatile during degradation of mikroarm star shaped P[(BMA-*b*-MMA-*b*-BMA)-co-(ϵ -CL)] from ambient temperature to 500°C. The thermal stability of this polymer are considerably different than that of poly(MMA) prepared by free radical method [31]. Some thermal degradation characteristics of polymers are summarized in **Table 2**.

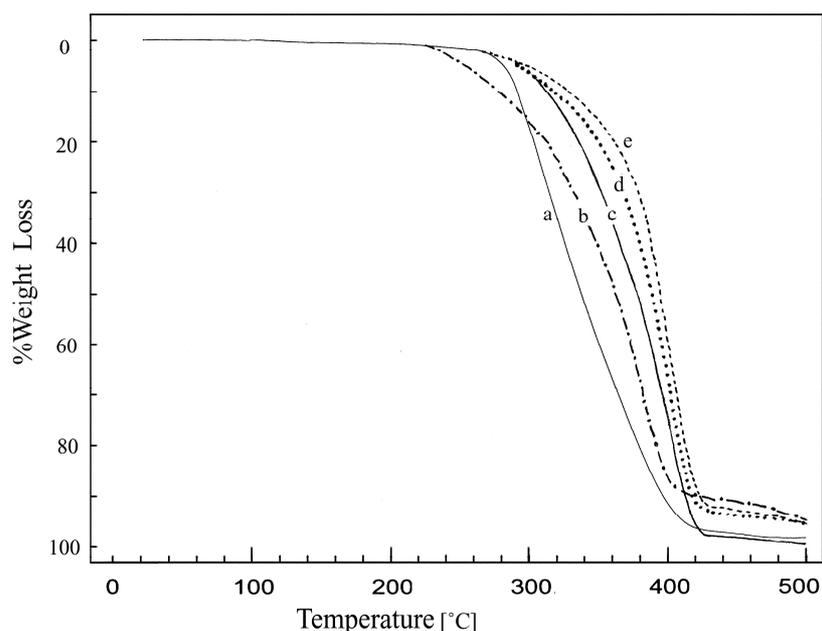


Figure 8. TGA curves of (a) P(BMA-*b*-MMA-*b*-BMA) dual lacton functionalized diblock copolymer; (b) Mikroarm star shaped poly(BMA); (c) Poly(methyl methacrylate) ended 7-hydroxy coumarin; (d) Mikroarm star shaped P(BMA-*b*-MMA-*b*-BMA)-co-(ϵ -CL)]; (e) Three armed P(ϵ -CL).

Table 2. TGA data for some degradation characteristic of mikroarm star poly[ϵ -CL-co-benzyl methacrylate (BMA)], poly(ϵ -caprolacton)(PCL) three armed and their copolymers.

Polymer	^a T _i (°C)	^b T _{%50} (°C)	%Weight Loss at 300°C	%Weight Loss at 350°C	%Residue at 400°C
PBMA ended lacton [22]	150	250	25	61	12.0
Miktoarm star shaped P(ϵ -CL-co-BMA)	230	370	15	36	12.0
P(BMA- <i>b</i> -MMA- <i>b</i> -BMA) dual lacton functionalized diblock copolymer	250	397	3	16	43.0
PMMA ended 7-hydroxy coumarin	292	321	18	59	9.0
Miktoarm star shaped P[(BMA- <i>b</i> -MMA- <i>b</i> -BMA)-co-(ϵ -CL)]	295	400	3	28	23.5
P(ϵ -CL) three armed	310	384	-	22	30.0

^aInitial decomposition temperature; ^bDecomposition temperature at 50% weight loss.

4. Conclusion

In this paper, it was carried out synthesis of miktoarm star copolymers using “core-first” method by ring-opening (ROP) process using appropriate (co)polymers with poly(benzyl methacrylate) lactone functionalized and P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer as macromonomer. Low-molecular-weight poly(benzyl methacrylate) one and P(BMA-*b*-MMA-*b*-BMA) dual functionalized diblock copolymer prepared by atom transfer radical polymerization (ATRP) of benzyl methacrylate at 110°C was used as starting materials for ring open polymerization. Then, the ROP of ϵ -caprolactone with one and dual functionalized macromonomers was carried out in the presence of tin(II)bis(2-ethylhexanoate) ($\text{Sn}(\text{Oct})_2$) at 130°C. Although poly(benzyl methacrylate) ended lactone with ϵ -caprolactone shows ring open polymerization in presence of phloroglucin by $\text{Sn}(\text{Oct})_2$, poly(benzyl methacrylate) lactone ended did not show ring opening reaction under conditions above mentioned.

5. Acknowledgements

The authors wish to thank the Firat University Research Fund for financial support of this Project (FUBAP-1650) and M. Fatih Coskun, who provided chromatograms (GPC) at our university.

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