

Optimization and Characterization of Nano Aluminum Trihydrate-Based Flame-Retardant Materials in the Rotating Packed Bed Reactor

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Abstract

A CO₂/NaAlO₂/H₂O absorption-reaction system was adopted to prepare nano aluminum trihydrate (ATH) particles for cables use in a rotating packed-bed (RPB) reactor (Phase I). The Taguchi experimental design was used to give an orthogonal array, L₉(3⁴). The parameters were the speed of rotation (A), gas-liquid flow rate ratio (B), aging temperature (C) and aging time (D), each having three levels. According to the signal/noise (*S/N*) ratio analysis, optimum condition for BET area (*S_w*) was found to be A3B3C1D1, while the parameter significance was DABC. Empirical formula for *S_w* was also obtained and discussed. Subsequently, the optimum size (10.7 nm) of ATH powders was used to blend with the matrix including plastics and additives (Phase II) for characterization further. The examination of composite materials showed that the limiting oxygen index (LOI) was higher than 28, indicating an excellent fire-retarding property. The results also showed that the LOI for nano ATH is higher than that bought from the market, while the fire-retarding grade examination (UL-94 test) for both ATHs was found to be of a V-0 grade. In addition, the mechanical properties for nano ATH were measured prior to using the ATH from the market.

Keywords

Absorption, Aluminum Trihydrate, Rotating Packed Bed, Taguchi

1. Introduction

Flame retardants are used in many fields for products such as cell phones, carpets and insulating sheets, to prevent them from being burnt or to at least slow down a fire in an emergency [1]. Most fires and casualties in recent years resulted from inflammable polymers and the toxic gas released during a fire [2].

Therefore, the selection of flame retardants is a big problem. Flame retardants can be divided into groups identified by their halogen base, phosphor base, nitrogen base, borium base, silicon base, antimony base, metal oxide or hydroxide [3] [4] [5] [6]. The most familiar flame retardant on the market is the halogen compound because the halogen can react with a free radical quickly, gradually implementing an excellent flame retardant effect. However, a halogen compound damages the atmosphere, reduces the ozonosphere, releases corrosive acidic gases and generates toxic dioxin. Therefore, with the rise in environmental considerations, various countries have made strict regulations governing current flame retardant materials including being halogen-free, low smoke and heavy metal-free [7] [8] [9]. An inorganic metal hydroxide as the flame retardant in polymers does not produce toxic gas during combustion. The product from combustion harms the environment, whereas a metal oxide performs an endothermic reaction in combustion and releases water vapor; the polymer surface temperature is reduced and the metal oxide covering the polymer surface forms a protection layer to obstruct flammable gas, so as to retard the flame [10] [11] [12].

At present, the most familiar inorganic metal hydroxides are magnesium hydroxide (MH) and aluminum trihydrate (ATH) [13] [14]. Both of these have filling, flame retardant and smoke suppression functions. They are the major inorganic flame retardants used in present industrial circles. ATH decomposes at about 180°C, absorbing a considerable amount of heat in the process and giving off water vapor. In addition to behaving as a fire retardant, it is very effective as a smoke suppressant in a wide range of polymers, especially in polyesters, acrylics, ethylene vinyl acetate, epoxies, PVC and rubber [15].

In comparison to other flame retardants, the preparation, use and discarding processes for ATH are free of hazardous material, and they have low smoke generation, low corrosivity and low toxicity [7] [8] [10] [11] [12] [14] [16]. ATH is selected as the flame retardant for this experiment because its maximum dehydration temperature (250.4°C) is lower than the maximum thermal degradation temperature of linear low density polyethylene (LLDPE) (380.1°C), and the flame retardant should release moisture earlier than the substrate, so as to successfully perform the flame retardant effect. The maximum dehydration temperature (396.2°C) of MH is higher than the substrate, so it cannot perform the flame retardant effect [17] [18]. Alternatively, ATH loses water gradually at 210°C - 370°C to form an alumina powder. The dehydration reaction releases 2.5 portions of hydrate water resulting in a great deal of heat being absorbed, so that the ATH reduces the burning temperature of the substrate, *i.e.* the “cooling technique” of flame retardant technology, showing its smoke suppression effect [19].

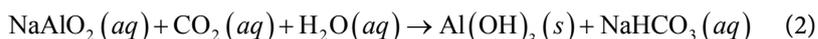
As ATH is a typical polar inorganic material, it has poor affinity with a non-polar organic polymer. The interfacial bonding force is weak, so that the workability and physical mechanical properties of the composite with this flame retar-

dant are degraded. The ATH of nano-particles can enhance the interfacial interaction, and ATH is dispersed in plastic substrates more uniformly, so as to improve the mechanical and flame retardant properties of blend composition [20] [21] [22] [23].

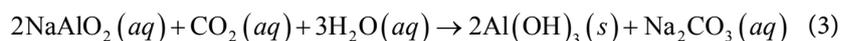
The methods for preparing ATH include the Bayer precipitation method, the Bayer addition method, and aqueous caustic and aluminum chloride syntheses. The sodium aluminate solution and CO₂ synthesis, e.g. a high-gravity rotating packed bed (HGRP), can prepare nano-scale ATH [24] [25] [26] [27]. In addition, synthesis of ATH was explored recently by Chen *et al.* [28] who obtained pseudo-boehmite nanoparticles with a direct precipitation method and the transparent dispersions of boehmite nanoparticles was obtained by a following hydrothermal treatment. The formation mechanism of ATH was also discussed in this paper. Jiao *et al.* [29] found that Al³⁺/AlO₂⁻ ratio can affect the initial pH and hence the ATH morphologies. As we know, sodium aluminate is an important commercial inorganic chemical. It works as an effective source of aluminum hydroxide, which is also known as hydrated alumina, aluminum hydrate or alumina tri-hydrate (ATH), for many industrial and technical applications. The reaction can be expressed as:



when carbon dioxide is excess:



when carbon dioxide is less:



It was found that factors affecting the formation of ATH using HGRP were carbonation of sodium aluminate (SA), temperature, speed, SA concentration, G/L ratio and terminal pH [25]. Other factors affecting the formation of ATH were precipitation time, soda content, additives, amount of seed and surface area of the seed [30]. The gravity field generated by the high speed centrifugal force of HGRP can tear and thin the liquid film, however, the gas-liquid mass transfer rate was good, the reaction time was short, and the composite powder size distribution was narrow, so the nano-scale powder could be prepared [26] [27].

As we known, there have not been any extensive studies on the effects of process parameters on the particle size and specific surface area of ATH product for rotating packed bed reactors. In addition, no ATH flame-retardant data was available. Therefore, it would be highly beneficial for the optimization of reaction processes to produce a high quality ATH. Taguchi statistical analysis was performed to get the relative importance of each parameter and to obtain the optimal condition. This study used HGRP to prepare nano-scale ATH, and used the Taguchi method to find the optimum conditions for preparing ATH. Afterwards, the nano-ATH produced under optimum conditions and the commercially available ATH powder were mixed with a linear low density polyethylene (LLDPE), an ethylene ethyl acrylate copolymer (EEA), zinc borate (ZnB) and

nanomontmorillonite (MMT) to obtain cable materials in a fixed proportion, respectively, so comparison of the flame resistance, mechanical properties and thermal analysis of the composites after mixing the nano-scale and commercially available powders was possible. In order to reach the purpose of this work, a framework was proposed, as shown in **Figure 1**. Two phases were adopted: Phase I was designed for the preparation of ATH, while Phase II was designed for the preparation and characterization of the cable materials.

Steps to accomplish the purpose of this work were:

1) The Taguchi method was used to look for optimum conditions for preparing nano-ATH. The particle size, grain size, powder morphology and thermal analysis were analyzed.

2) ATH, prepared under the optimum conditions and the commercially available ATH, were mixed in the same proportions, respectively. The flame retardancy, mechanical properties and thermal analysis of the composites were compared.

2. Experimental Feature

2.1. Experimental Design for the Synthesis of ATH

In order to reduce the experimental cost, an experimental design for the synthesis of ATH was required. Factors adopted here were centrifugal speed of RPB (A), gas liquid flow rate ratio (B), aging temperature (C) and aging time (D), each having three levels, as shown in **Table 1**. According to the Taguchi method, the orthogonal array is listed in **Table 2** for experimental design. This reduces the experimental number from $3^4 = 81$ to $L_9(3^4) = 9$. In addition, parameter significance and optimum condition could be obtained by using signal/noise (S/N) analysis.

2.2. Materials and Method for the Synthesis of ATH

The starting materials used for the synthesis of ATH were NaAlO_2 and CO_2 gas.

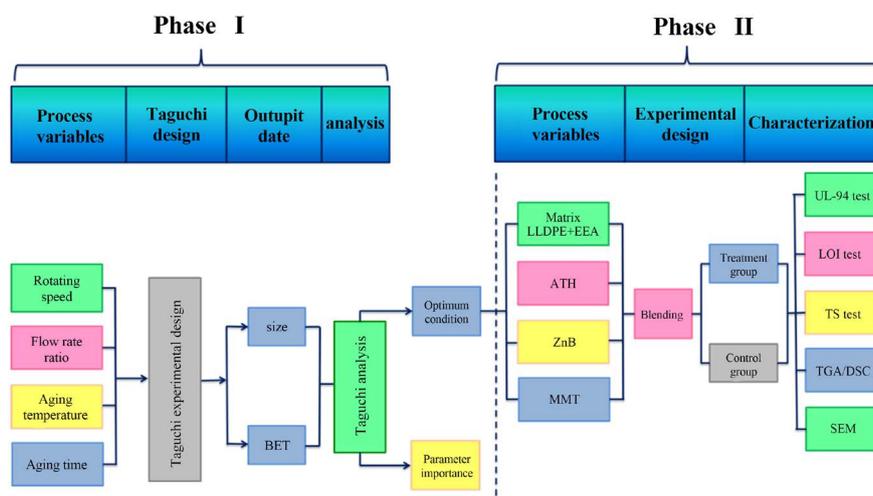


Figure 1. A framework for preparation and characterization of cable materials.

Table 1. Factors and levels for the synthesis of ATH.

Factor	Level		
	1	2	3
A	1000 rpm	1500 rpm	2000 rpm
B	0.6	0.8	1.0
C	50°C	60°C	70°C
D	1 hr	2 hr	3 hr

Table 2. Orthogonal array for the synthesis of ATH.

No.	A	B	C	D
A1	1	1	1	1
A2	1	2	2	2
A3	1	3	3	3
A4	2	1	2	3
A5	2	2	3	1
A6	2	3	1	2
A7	3	1	3	2
A8	3	2	1	3
A9	3	3	2	1

All of the chemicals used were of an analytical grade (Alfa Aesar). The experimental apparatus of the RPB is shown schematically in **Figure 2**, which was similar to that reported in the literature [31] [32]. The apparatus consisted of a rotating packed bed, pumps, flow meters, speed controller, water bath and reservoir. The rotating packed bed contained wired packing with several layers. A known volume and concentration of reactant (2 M NaAlO₂ solution) was taken and fed into the reactor at a desired rotating speed from the top of the reactor through the action of tubing pumps. The reactant temperature was controlled at 25°C. In addition, a desired flow rate of CO₂ gas was fed into the reactor. The reactant was pumped through the distributor into the rotating packed bed and made contact with CO₂ gas, where the primary step of the reaction occurred. Then, the product slurry flowed out of the reacting chamber and was collected in the product reservoir. The time from the beginning of this process to the end is termed the reaction process time when the solution pH reached 12. At the end of operation, the slurry was withdrawn and heated to the desired temperature level for the aging process. After the aging process, the slurry was filtered and rinsed with acetone, and the cake was dried in a hot air oven for one hour. The dried powders were analyzed by XRD to identify the composition of the product. The decomposition temperatures of the dried powders were determined by means of XRD (Rigaku, D/MAX2200) analysis. Then, the ATH powders were characterized further by means of TGA/DSC (TA, SDT-Q600), SEM (JEOL, JSM-6500F) and BET (PMI, BET 201-APCW) analyses.

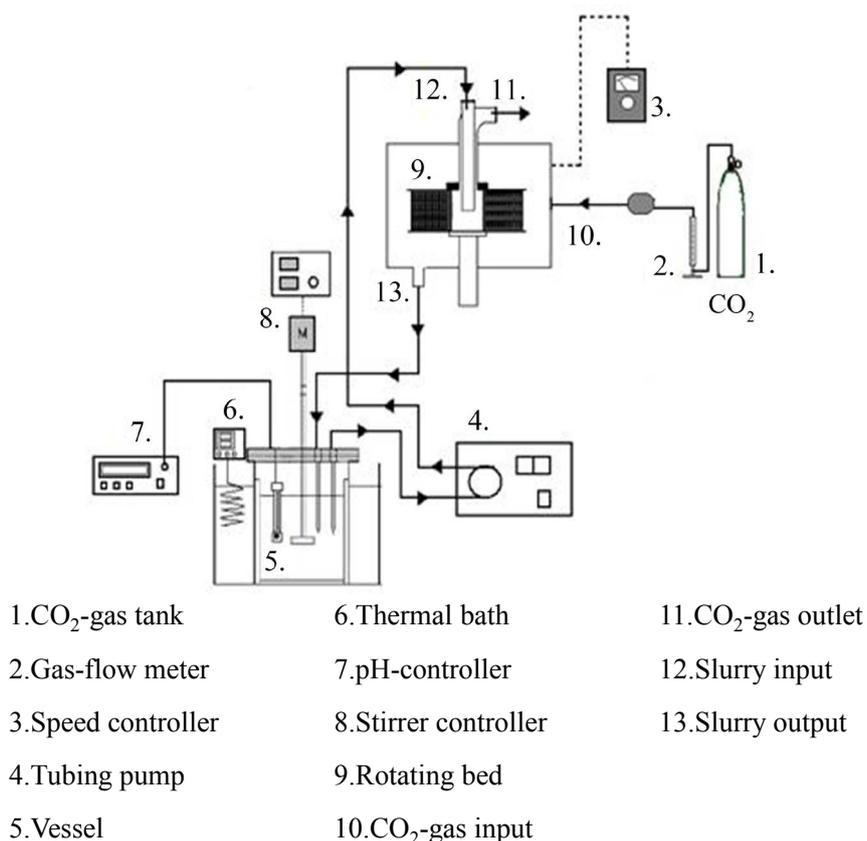


Figure 2. A RPB device for preparation of nano-scale ATH; please see the reference [31] [32].

2.3. Blending Tests for Flame Retardant Cable Materials

Materials for blending tests, ATH, MMT, ZB, EEA and LLDPE, were employed by using a twin screw (C. W. Brabender, Laboratory Double Sigma Blender) at a working temperature of 170°C. In order to carry out the blending test, three series (A, B and C) of tests were explored here, as shown in **Table 3**. Series A was only matrix LLDPE tested for comparison. Series B and C used the same ingredients except ATH; Series B's ingredients were obtained in the market, while Series C's ingredients were obtained in this work. After blending, the composite materials were used in thermal compression to prepare specimens for tests, including mechanical property and flame retardant tests, as shown in **Figure 3**. The former was measured by a tensile strength machine (QC-TECH, CTS-2KN); the latter was measured by the limiting oxygen index (LOI) (Suga Test Instrument Co. Ltd., ON-1) and UL-94 test (UL-1). In addition, the thermal properties and dispersion of composite materials were examined by TGA/DSC and SEM analyses, respectively.

3. Results and Discussion

3.1. Morphology of Nano-ATH Crystals

The SEM photographs shown in **Figure 4(a)** and **Figure 4(b)** indicate irregular

Table 3. Contribution of materials in blending tests at different series. *ATH was gotten in the market.

Series	No	Matrix		Flame retardant agents		
		LLDPE (wt%)	EEA (wt%)	ATH (wt%)	ZnB (wt%)	Clay (wt%)
A	A0	100	0	0	0	0
	B1	40	10	40	10	0
B*	B2	40	10	40	9	1
	B3	40	10	40	7	3
	B4	40	10	40	5	5
	C1	40	10	40	10	0
C	C2	40	10	40	9	1
	C3	40	10	40	7	3
	C4	40	10	40	5	5



Figure 3. Specimens prepared using thermal compression. (a) before compression; (b) after compression.

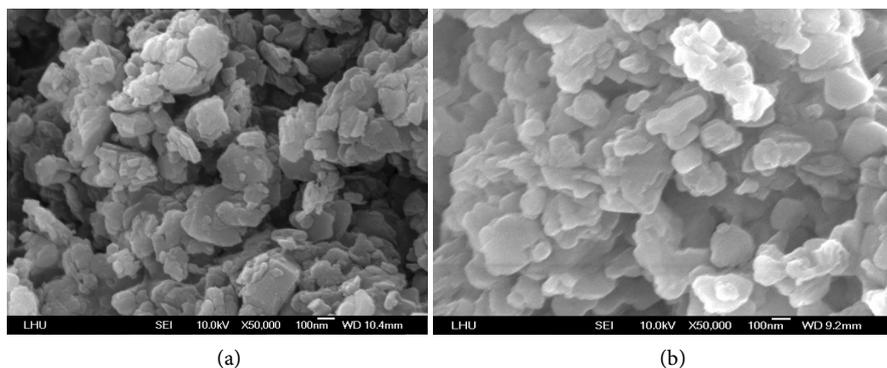


Figure 4. Typical ATH SEM photographs from different runs. (a) Sample A2; (b) Sample A7.

crystal forms with the two forms prepared in this study. The two pictures show similar irregular shapes. The pictures also show aggregates with tiny particles. Therefore, it is hard to determine crystal size from particle size analyzer. Due to

this, the crystal size can be determined by using XRD analysis and BET measurement further.

3.2. BET Specific Surface Area of Nano-ATH

The specific surface areas obtained were in the range of 44.75 - 237.06 m²/g as shown in **Table 4**. The BET specific area can be used to calculate the particle size. The equation is:

$$d_{\text{BET}} = \frac{6}{\rho S_w} \quad (4)$$

The calculated sizes were in the range of 10.82 - 57.30 nm.

In order to effectively determine S_w and understand the factors affecting S_w , the relationship of S_w and variables can be correlated by using linear regression. The result was:

$$S_w = 0.04534 \exp\left(\frac{2544.05}{T}\right) N^{0.1081} R^{0.8046} t^{-0.7235} \quad (5)$$

The regression error is 7.48%. The result showed that S_w increased with an increase in $N(A)$ and $R(B)$, while it decreased with an increase in $T(C)$ and $t(D)$. Moreover, this equation can be used to predict the S_w , and hence d_p , when the conditions are given.

3.3. Confirmation of Nano-ATH Crystals from XRD Analysis

ATHs obtained here can be characterized by using XRD analysis. **Figure 5** shows the XRD patterns for night runs. As compared with a standard card, the patterns showed two ATH forms, *i.e.* Bayerite (β -phase) and Gibbsite (α -phase), mixed together. The Bayerite form includes major peaks at (0, 0, 1), (1, 1, 0), (1, 1, -1) and (1, 3, -1); while the Gibbsite form shows peaks at (0, 0, 2), (0, 2, 1), (3, 1, 1) and (2, 0, 2). This indicates that the two forms provide different crystalline structures. The major peak with an (h k l) value of (0 0 2) at 18.45° is the characteristic of monoclinic ATH as per JCPDS card number 76-1782.

Table 4. BET specific area and crystal size for both methods.

No.	Condition	S_w (m ² /g)	d_{BET} (nm)
1	A1B1C1D1	181.70	14.11
2	A1B2C2D2	119.11	21.53
3	A1B3C3D3	87.05	29.46
4	A2B1C2D3	44.75	57.30
5	A2B2C3D1	95.91	26.73
6	A2B3C1D2	109.25	23.47
7	A3B1C3D2	139.32	18.41
8	A3B2C1D3	85.84	29.87
9	A3B3C2D1	237.06	10.82

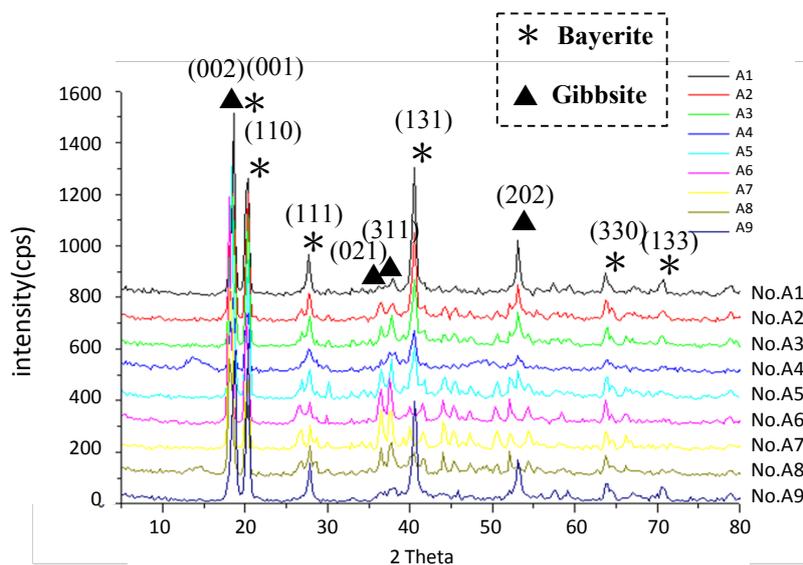


Figure 5. XRD patterns for ATHs prepared in here.

For comparison, mean crystal size was also determined using XRD analysis with the aid of FWHM (full width at half maximum). The size, d_{XRD} , can be estimated using the Scherer equation:

$$d_{\text{XRD}} = \frac{K\lambda}{\beta \cos \theta} \quad (6)$$

where K is a dimensionless shape factor (a typical value of 0.9), β is the line broadening at the full width at half maximum (FWHM), λ is the X-ray wavelength and θ is the Bragg angle (in degrees). **Table 5** lists the estimated sizes of ATH particles, which was close to that determined by BET areas. **Figure 6** was a plot of d_{XRD} vs. d_{BET} for comparing both methods.

3.4. Thermal Property of Nano-ATH Crystals

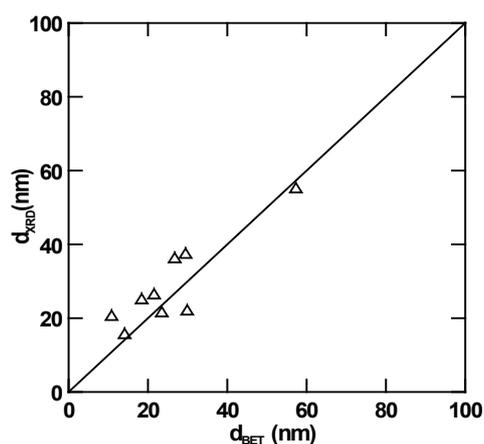
The thermal property of ATH was also examined, as shown in **Figure 7(a)-(c)**. In **Figure 7(a)**, the heat flow of ATH (No. 3) changed with the heating temperature. From DSC analysis, the enthalpy (ΔH) and maximum endothermic temperature (T_{en}^{**}) could be determined simultaneously. **Figure 7(b)** showed the DSC spectra of each ATH obtained. On the other hand, **Figure 7(c)** shows the TGA measurement spectra showing the weight loss during heating and giving the maximum decomposition temperature (T_d^*). All data obtained is listed in **Table 6** showing the effect of particle size. It can be found that the ΔH is higher at a smaller size since a smaller particle gives a higher surface area resulting in a higher ΔH . In addition, the weight loss for a smaller particle is higher than for a larger one due to the smaller particle having a larger surface area, which favors water loss.

3.5. Taguchi Analysis

The S/N ratio is calculated according to the larger-the-better value and the

Table 5. Nano ATH crystal size estimated using Scherer equation.

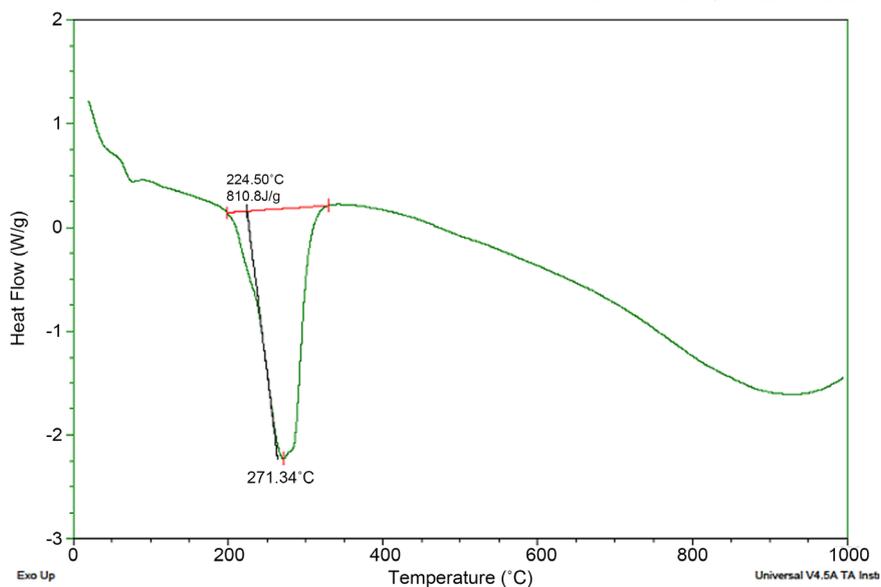
No.	FWHM	2 θ	d_{XRD} (nm)
A1	0.55	18.314	15.89
A2	0.37	18.314	26.64
A3	0.302	18.187	37.64
A4	0.258	18.244	55.39
A5	0.307	18.230	36.45
A6	0.427	18.776	21.80
A7	0.383	18.287	25.32
A8	0.42	18.413	22.27
A9	0.442	18.314	20.81

**Figure 6.** A plot of d_{XRD} versus d_{BET} showing the reliable methods for both.

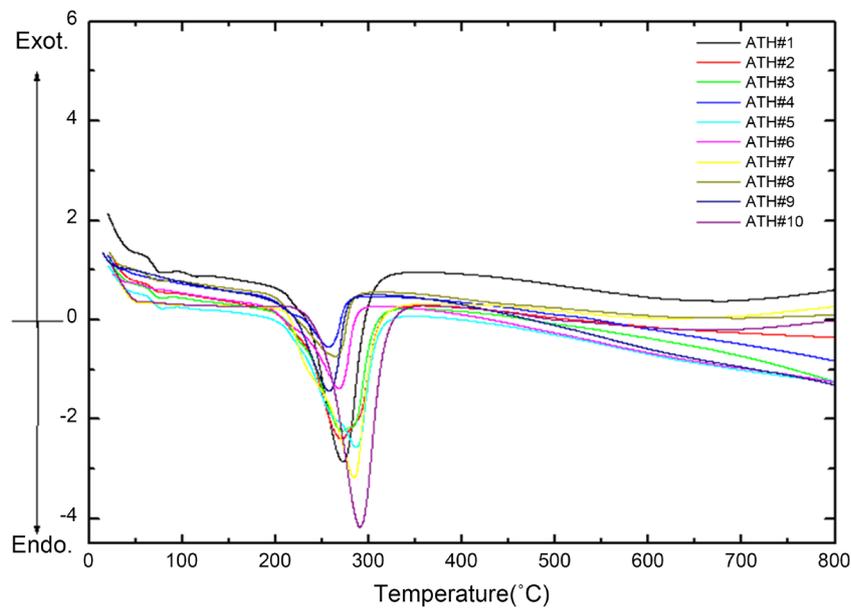
Sample: Al(H)3#3
 Size: 17.2000 mg
 Method: DSC Heatflow

DSC-TGA

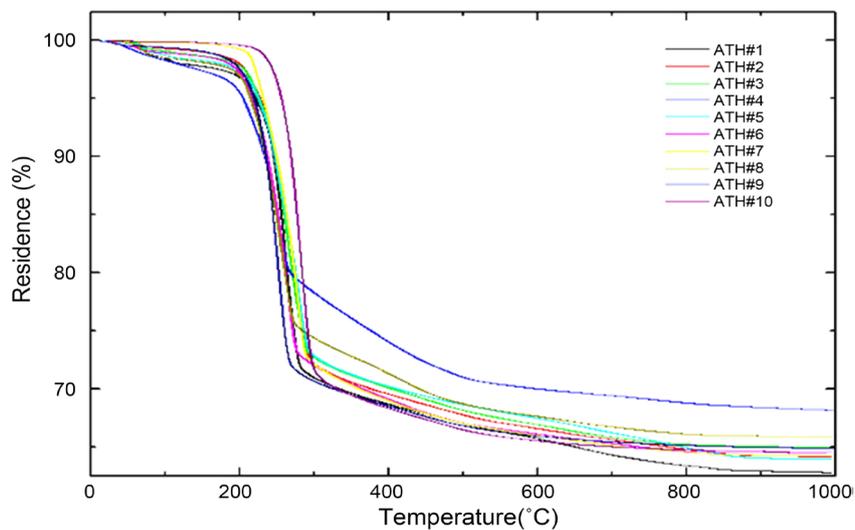
File: C:\TA\Data\SDT\chia\Al(OH)3#3.001
 Operator: summy
 Run Date: 11-Jan-2012 11:04
 Instrument: SDT Q600 V20.9 Build 20



(a)



(b)



(c)

Figure 7. TGA/DSC analyses for ATHs showing thermal property of each ATH. (a) Sample A3; (b) DSC analysis for all samples; (c) TGA analysis for all samples

Table 6. Thermal data of nano-ATH from TGA/DSC analysis.

	T_d^* (°C)	T_{en}^{**} (°C)	ΔH (J/g)	Weight loss (%)	d_{BET} (nm)
A1	285.17	272.17	919.0	64.27	14.11
A2	296.11	270.43	865.3	65.44	21.53
A3	291.37	271.34	810.8	65.83	29.46
A4	267.68	257.11	495.3	69.41	57.30
A5	296.11	286.10	865.1	66.21	26.73
A6	278.34	267.66	848.0	65.26	23.47

Continued

A7	290.19	283.99	895.6	65.61	18.41
A8	274.79	263.82	744.6	66.73	29.87
A9	268.86	257.11	1056.0	65.09	10.82
A10	302.04	290.33	1085.0	64.98	10.70

T_d^* = maximum decomposition temperature; T_m^{**} = maximum endothermic temperature.

smaller-the-better value from the data obtained in the final stage. The sequences of importance and optimum conditions are found by using the S/N ratio calculations where d_{XRD} takes the smaller-the-better value, and S_w takes the larger-the-better value as shown below:

$$\frac{S}{N} = -10 \times \log \left(\frac{1}{n} \sum_1^n z_i^2 \right) \quad (7)$$

and

$$\frac{S}{N} = -10 \times \log \left(\frac{1}{n} \sum_1^n \frac{1}{z_i^2} \right) \quad (8)$$

where n is the number of data evaluated and z_i is the data presented at the i^{th} level.

According to Equation (7) the smaller-the-better S/N ratio analysis for size is shown in **Table 7**. The three levels were used in here, which was the common levels (3 - 4 levels) adopted in the Taguchi experimental design [29]. The results show that the optimum condition is A3B3C1D1, see bracket; and the sequence of parameter importance is CDAB. Alternatively, the S/N ratio analysis for S_w (the larger-the-better) was also explored. According to Equation (8), the estimated values are presented in **Table 8**. It can be found that the optimum condition is A3B3C1D1, which is the same as that in d_{XRD} ; while the sequence of parameter importance is DABC, which is different from that obtained in d_{XRD} . This indicates that D and A are major factors for S_w , while C and D are major factors affecting d_{XRD} .

The optimum condition, A3B3C1D1 (No. 10), for both cases was also carried out for verification. Characterization of No. 10 for the X-ray spectra and SEM photographs was needed and is shown in **Figure 8**. The result for d_{XRD} was 15.25 nm and for S_w was 239.72 m²/g, leading to a 10.70 nm BET size. Due to this, the optimum conditions were confirmed further as compared with the values in **Table 5**. In addition, S_w can be evaluated using Equation (8), giving 271.6 m²/g in S_w at the optimum condition of A3B3C1D1. The value is higher than 239.72 m²/g, indicating that Equation (8) is reliable.

3.6. Characterization of Flame Retardant Materials

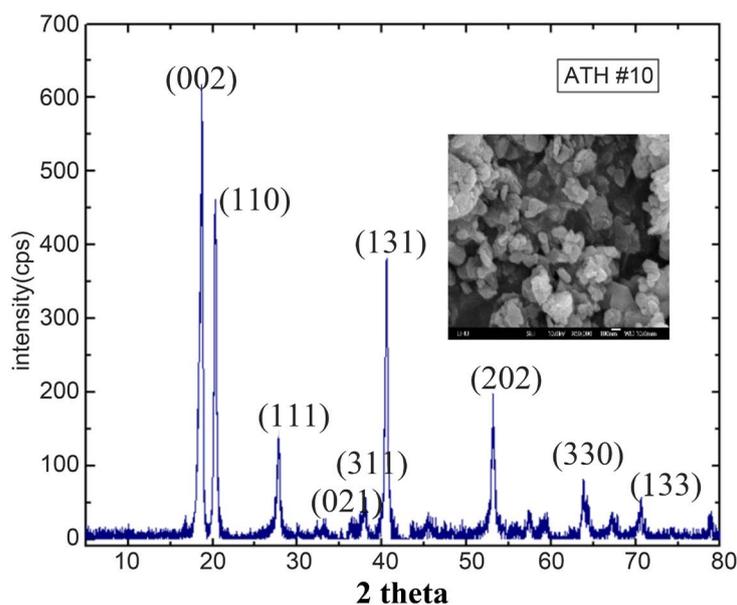
In order to compare the effects of different ATH sources on the blending test, additions of A10 and market ATHs in the blended materials were explored. Sample A10 was prepared under optimum conditions. The particle sizes for both

Table 7. *S/N* analyses for d_{XRD} to find optimum condition and parameter importance.

Level	A	B	C	D
1	-28.02	-28.99	(-25.92)	(-27.21)
2	-30.96	-28.90	-29.91	-27.78
3	(-27.13)	(-28.22)	-30.27	-31.11
Delta	3.83	0.77	4.36	3.90
Rank	3	4	1	2

Table 8. *S/N* analyses for S_w to find optimum condition and parameter importance.

Level	A	B	C	D
1	41.83	40.36	(41.54)	(44.11)
2	37.81	39.94	40.68	41.72
3	(43.02)	(42.35)	40.44	36.83
Delta	5.21	2.41	1.11	7.28
Rank	2	3	4	1

**Figure 8.** XRD spectra and morphology of ATH obtained for No.10.

are 10.70 nm and 29.11 nm for A10 and market ATH, respectively. In addition, the effects of MMT on characterizations for both cases include mechanical properties, LOI, the UL-94 test, TGA/DSC and SEM.

3.6.1. Mechanical Property for Blended Materials

The effects of MMT and a type of ATH on the mechanical properties were investigated. **Figure 9(a)-(c)** presented the effect of both on the tensile strength (TS), elongation at the break (EB) and tensile modulus (E). It was found that mechanical properties among them for ATH prepared in our study were higher

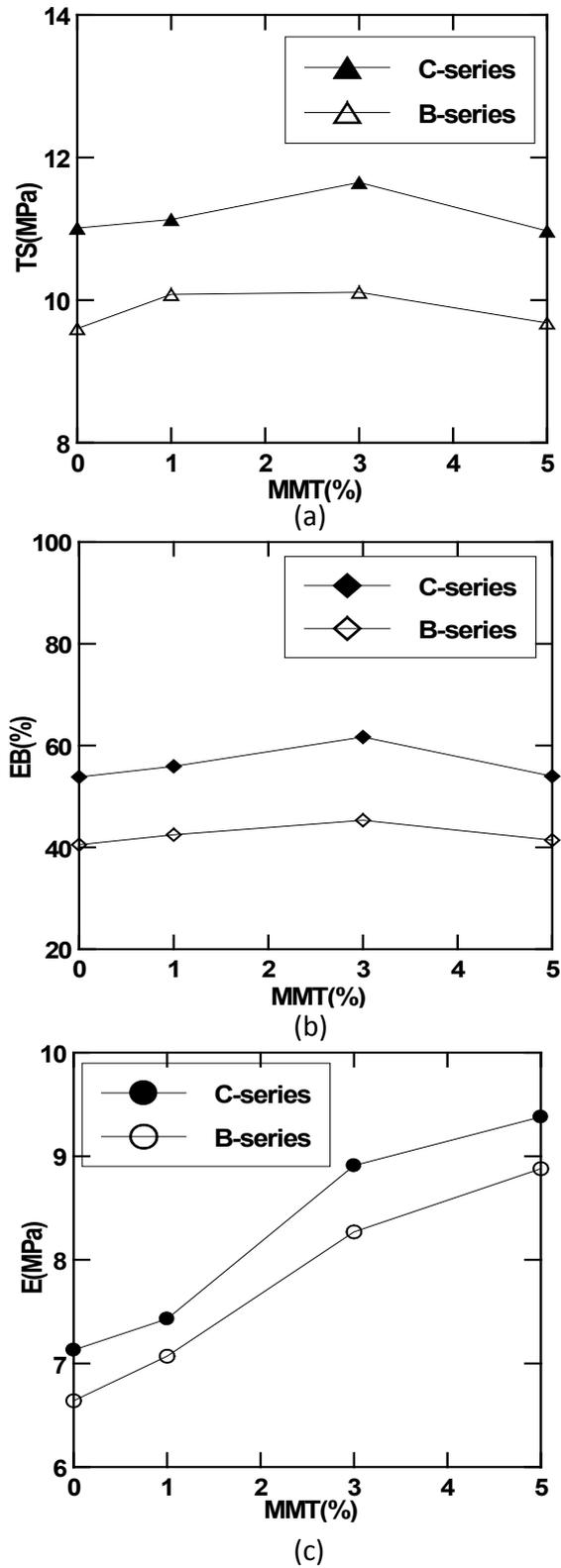


Figure 9. Effect of MMT on the mechanical property of composite materials.

than those for product bought in the market. The possible reason is the effect of size. In addition, the effect of MMT on the TS showed that TS increased with an

increase in MMT at about 3%; then it decreased with MMT, as shown in **Figure 9(a)**. A similar result for EB is shown in **Figure 9(b)**. On the other hand, E increased with an increase in MMT, as shown in **Figure 9(c)**, indicating that the addition of MMT can improve the mechanical strength of E. However, the TS and EB for pure LLDPE were 18.56 Mpa and 551.92%, respectively. The values were much higher than those for blended materials due to the presentation of ATH in the matrix. This indicated that the addition of flame retardant materials resulted in the decrease in mechanical properties. Therefore, how to control the ingredients of composite materials while obtaining a good flame retardant with excellent mechanical properties becomes a significant focus in future applications.

3.6.2. Flame Retardant Test for Blended Materials

There are two methods used to evaluate a flame retardant: one is the LOI method and the other is the UL-94 method. **Figure 10** shows the effect of MMT in both cases. It was found that LOI increased with an increase in MMT, while the LOI in Series C was higher than that in Series B. The values of the LOI showed that the materials were exhibited as sparingly flammable when LOI was greater than 26, which was higher than 22.5 for pure LLDPE, indicating the effect of the addition of retardant materials. From this viewpoint, the flame retardant test showed that ATH becomes an effective material as a flame retardant.

In addition, three major levels in the UL-94 test, *i.e.* V-2, V-1 and V-0, were used to examine, in which V-0 was the highest level among them. The results for different samples are shown in **Table 9**. Comparing the two series showed that the flame time for Series C was always less than for Series B, indicating the effect of ATH size on the flame time because the specific surface area for C is high enough to release more vapor, decreasing the burning temperature and hence, reducing the burning time. All cases passed the minimum UL-94 test except for

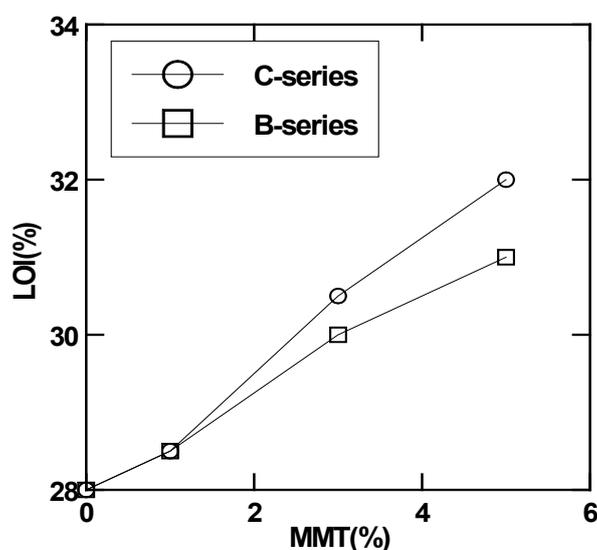


Figure 10. Effect of MMT on LOT.

Table 9. UL-94 test for various samples.

Series	No.	UL-94 level	Flame time for a total of 5 specimens (Sec)	Drip of flaming particles that ignite
B	B1	V-2	183	√
	B2	Fail	472	√
	B3	V-2	103	
	B4	V-0	11	
C	C1	V-2	169	√
	C2	Fail	463	√
	C3	V-2	89	
	C4	V-0	8	

B2 and C2 because the blending materials were not uniformly distributed in the matrix. However, B3 and C3 passed the V-2 test, while B4 and C4 all passed the V-0 test. At 5% MMT, ATH, MMT and ZnB have uniform distribution in the matrix, enhancing more water vapor and absorbing more energy when burning. Therefore, this shows that an excellent flame retardant property was observed as demonstrated with SEM photographs in **Figure 11**.

3.6.3. Thermal Property of Blended Retardant Materials

Thermal data for flame retardant composite materials were measured using a TGA/DSC machine, as shown in **Figure 12(a)-(b)** for a typical example in Series C. Measured data was determined and is listed in **Table 10**, including melting points, decomposition temperatures and weight loss at 600°C. In addition, pure LLDPE and EEA were also measured, as shown in the first two rows. The values in T_d are much higher than for Series B and C. In addition, T_m and T_d for Series C are slightly less than for Series B, showing the effect of size. However, T_m and T_d for both cases decreased with an increase in MMT, supporting the LOI test, as shown in **Figure 10** because the flame retardant in the matrix was decomposed before the matrix, creating an effective flame retardant for the flame retardant composite materials. In general, the decomposition temperatures for ATH were between 200°C - 300°C, resulting in weight loss in the composite materials. In addition, the T_d for Series B and C were around 288°C - 290°C, showing the effect of the formation of Al_2O_3 on the matrix.

4. Conclusion

Nano ATH-based flame-retardant materials for cables were successfully prepared. Synthesis of nano-ATH crystals and the blending of retardant materials in the matrix (LLDPE/EEA) were effectively explored using an RPB reactor and a twinscrew extruder. The obtained BET areas of nano ATH were in the range of 44.75 - 237.06 m²/g, corresponding to the sizes of 15.89 - 55.39 nm, which are close to that determined by the Scherer equation and XRD data. Results showed that the sizes of nano ATH crystals can be controlled by adjusting the parameters.

Table 10. Thermal data obtained for TGA/DSC analyses.

	T_m (°C)	T_d (°C)	Weight loss at 600°C (Wt.%)
Pure LLDPE	123.60	456.16	-0.38
Pure EEA	100.41	452.54	-0.26
B1	122.20	291.58	30.26
B2	122.06	287.80	33.77
B3	121.95	285.24	36.46
B4	121.66	283.91	36.46
C1	122.10	289.85	30.26
C2	121.71	286.53	34.55
C3	121.51	284.21	35.98
C4	121.11	282.73	36.55

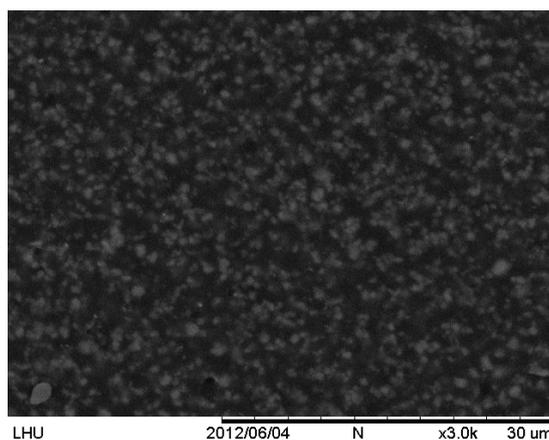
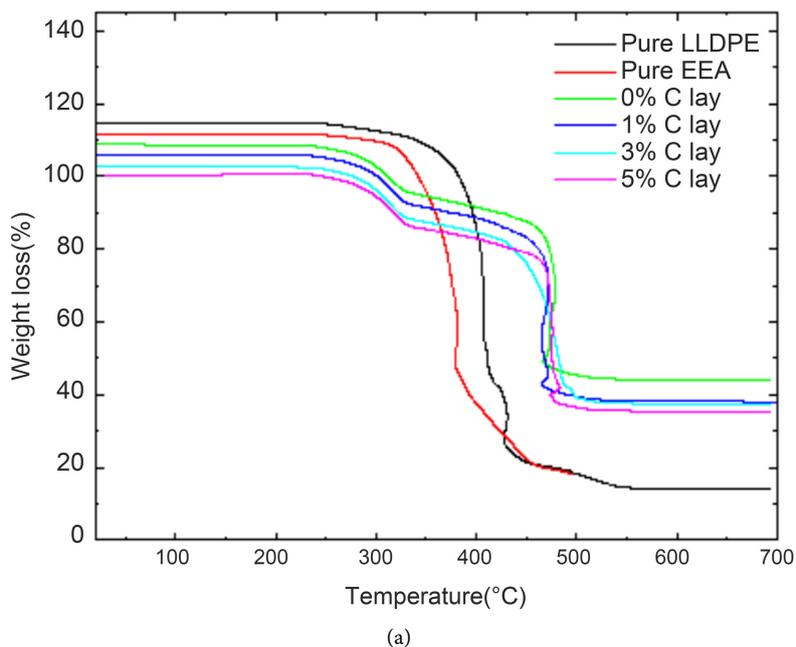


Figure 11. SEM image for No. 10 showing uniform distribution of flame retardant materials in the matrix.



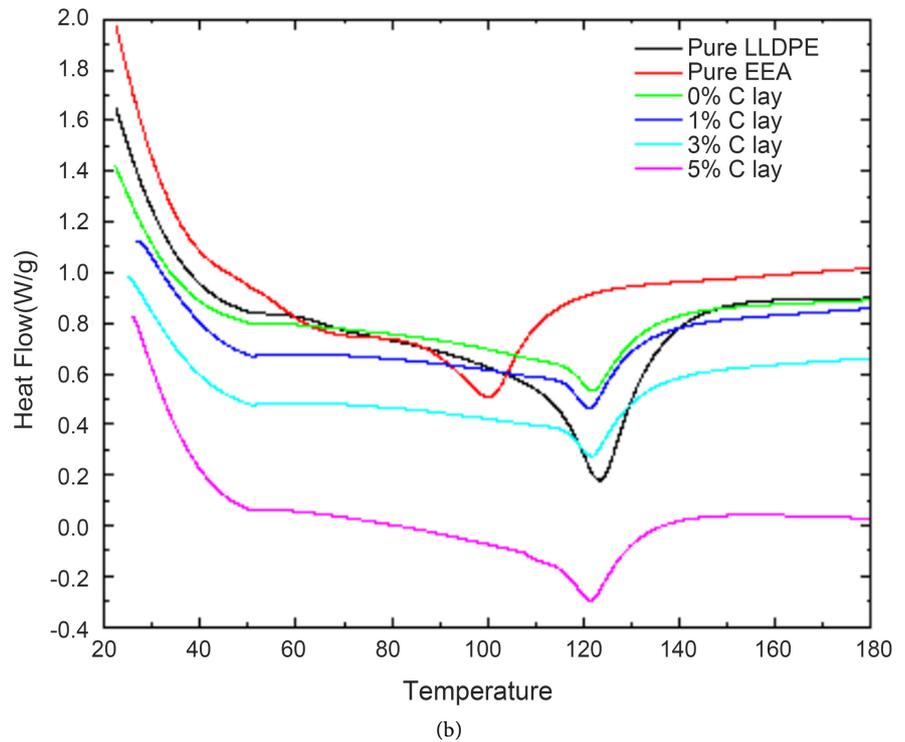


Figure 12. Thermal data analysis for composite materials. (a)TGA analysis; (b)DSC analysis.

The optimum condition A3B3C1D1 in S_w was successfully verified for the “larger-the-better”, obtaining $239.72 \text{ m}^2/\text{g}$ in S_w , which is $271.6 \text{ m}^2/\text{g}$ as predicted by the correlation equation. In addition, the parameter importance shows that D and A are major factors affecting S_w , while B and C are minor factors. All of them supported that the data obtained here is reliable and the S_w expression is useful. Blending materials found that nano-ATH for mechanical and flame-retardant properties tests were superior to those available in the market. SEM photographs show that dispersion of ATH/MMT/ZnB in the polymer matrix is uniform and shows an excellent property. Finally, the process creates an opportunity to capture CO_2 for some industries.

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Conflicts of Interest

The authors declare no conflicts of interest regarding the publication of this paper.

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