

Spatial Variability and Contamination Levels of Fresh Water Resources by Saline Intrusion in the Coastal Low Lying Areas of the Douala Metropolis-Cameroon

Mary Lum Fonteh^{1,2*}, Theophile Fonkou³, Mathias Fru Fonteh²,
Estella Buleng Tamungang Njoyim⁴, Cornelius Mbifung Lambi⁵

¹Department of Geography, Higher Teacher Training College, The University of Bamenda, Bamibli, Cameroon

²Department of Agricultural Engineering, Faculty of Agronomy and Agricultural Sciences, University of Dschang, Dschang, Cameroon

³Department of Plant Biology, Faculty of Science, University of Dschang, Dschang, Cameroon

⁴Department of Chemistry, Higher Teacher Training College, The University of Bamenda, Bamibli, Cameroon

⁵Department of Geography, Faculty of Social and Management Sciences, University of Buea, Buea, Cameroon

Email: *mariefontehniba3@gmail.com

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Abstract

A study based on the spatial variability and contamination levels of fresh water resources by saline intrusion was conducted in the Douala coastal area. The study was aimed at highlighting the associations between major ions in ground water from which cause-effect relationships could be inferred. Water samples were collected from 19 stations. 3 stations were selected from the mangrove area and 16 stations were selected from the rest of the area partitioned into four transects (coastal transect, inner transect 1, inner transect 2 and inner transect 3). Sampling was done repeatedly during the dry season and these samples were analysed for physico-chemical parameters. Results show that the samples were acidic (pH: 4.7 - 6.7). Total dissolved solids (TDS) and electrical conductivity (EC) values ranged between 70.3 - 3703 mg/L and 136.4 - 7333 $\mu\text{S}/\text{cm}$ respectively indicating medium salt enrichment and brackish waters. High temperatures ($T^{\circ}\text{C}$) and low dissolved oxygen (DO) values of 26°C - 30.3°C and 0.96 - 3.9 mg/L respectively were observed and this could be interpreted as the acceleration of biological and chemical processes of ground water resources. Major ions were within the WHO limits except for Ca^{2+} (20.3 - 85 mg/L) and Mg^{2+} (6.6 - 49.6 mg/L) respectively whose concentrations were slightly higher. The leading cations were $\text{Ca}^{2+} > \text{Na}^{+} > \text{Mg}^{2+} > \text{K}^{+}$ while $\text{SO}_4^{2-} > \text{Cl}^{-} > \text{NO}_3^{-} > \text{HCO}_3^{-} > \text{PO}_4^{3-}$ were the leading anions. The multivariate analysis approach (MAA) values obtained for water variables showed that F1, F2 and F3 accounted for 61.6%, 21.5% and 9.9% respectively

of total variance with strong loadings and these were considered to account for the ground and surface water quality of the area. The main water types were 63.2% $\text{Ca}^{2+}\text{-SO}_4^{2-}$, 26.3% $\text{Ca}^{2+}\text{-Cl}^-$ and 10.5% $\text{Na}^+\text{-Cl}^-$. 89.5% of the water types had secondary salinity implying that anthropogenic sources were the source of salinity. The water was not fit for drinking. The values calculated for percentage sodium and sodium absorption ratio to determine its suitability for agricultural purposes revealed that 63.15%, 15.5%, 5.3% and 15.8% of sampled water were excellent, good, doubtful and unsuitable respectively for irrigation purposes. The need for water resource monitoring and forecasting regarding deterioration in quality is imperative before hazards of the rise in sea level arise.

Keywords

Spatial Variability, Fresh Water Resources, Salinity Intrusion, Douala IV-Cameroon

1. Introduction

The Douala metropolis is the economic hub of the Cameroonian nation with more than 70% of its industries that have given it a considerable importance in the economy and development of the country. In such an area, utilities including water are quite important for human and industrial growth.

Reference [1] stipulates that water is a basic necessity, underlying the reason why the Millennium Development Goals (MDG) advocated for access to clean water for all by the year 2015. Reference [2] indicated that about 50% of the Cameroonian population does not have access to safe drinking water. The situation of the Douala coastal area is particularly worrying because water can be found “everywhere although safe portable water is scarce”. This area has abundant natural water resources, but the water quality parameters in the area are a dodgy issue. Furthermore, water supplies by the National Water Company fall short of the demand for domestic and industrial use in terms of quantity and quality. Big industries (breweries, textile, chemical, soap factories, agrochemical plants and others) in the area have therefore turned to boreholes for water to meet their needs.

The protection of surface and ground water resources from seawater encroachment is a real issue of concern for coastal communities and this area is similarly affected. The quality of ground water resources in the area could be further complicated by the impact associated with an increasing population, urban development and industrial expansion. In addition, the predictions of increasing rise in the sea level in the area as in [3] [4] and associated hazards in the region may exacerbate the problem in the long term. The salinisation of the water resources could also be due to the formation of large piezometric draw-down cones resulting from the increasing extraction of ground water and this could accelerate salt water intrusion (SWI) by reversing the hydraulic gradients into the aquifers as in [5]. Reference [6] [7] [8] indicated that SWI into coastal aquifer

fers, rivers and estuaries is one of the impacts that coastal residents will face because of global warming as a result of thermal expansion of oceans and the consequent rise in sea levels. For instance, in North Africa and Iran, communities have been observed to drink water containing as much as 3000 mg/L of chlorides and 1500 mg/L of sulphates as in [9]. Reference [10] indicated that, salt water (water containing dissolved salts) is not suitable for drinking because it dehydrates the human body systems and should be desalinated before consumption. As indicated, these ions can act as laxative to those who are not used to drinking such water. Even for industrial and agricultural use, salt water also requires desalinisation and this could imply extra costs.

Understanding the quality patterns of ground water resources is a complex issue because inputs can come from a variety of sources: SWI, ascending deep ground water, precipitation, rapid urbanization and improper waste disposal and high agrochemicals use. It is common knowledge that the spatial variation in chemical composition of ground water comes from different types of input and this is extremely difficult to understand but has important practical applications for water resource management and evaluation as in [5] [7] [11]. Therefore, in order to ensure a rational exploitation of these water bodies, there is a need to understand the interaction of the factors underpinning their quality.

A regular monitoring of the quality of ground water cannot therefore be over-emphasized. The on-going study is designed to determine the extent and severity of SWI into water resources of the coastal low lying area of Douala.

2. Materials and Methods

2.1. Study Area

Douala IV is located within the Douala sedimentary basin with coordinates at latitudes 04°03' - 04°07' North of the equator and longitudes 09°37' - 09°41' East of the Greenwich meridian with an average elevation of not more than 10 m above sea level (Figure 1).

The climate is the wet equatorial type and the rainy season extends from April to November, followed by a short dry season typically occurring between December and March. Its annual total rainfall is generally high with values of up to 4000 mm in August and lower values in the month of February. It has, on the average, 180 rainy days a year. Temperatures are high averaging between 24°C and 27°C. The hydro graphical network within the zone of study is very dense and shows a dendritic pattern. The Wouri River constitutes one of the drainage systems of the Douala metropolis and flows all year round through the Akwa, Bonaberi and part of the Bassa zones. This river is fed by many intermittent streams that are normally loaded with most of the solid and liquid wastes channeled from industries, households and waste dumps.

The area is a sedimentary basin of Cretaceous to Tertiary age having a total surface area of 7000 km² and has a maximum width of 60 km. Reference [12] noted that the Douala basin consists of the Cretaceous Mungo River Formation, stratigraphically overlain by the Tertiary Mpundu Formation. The Douala

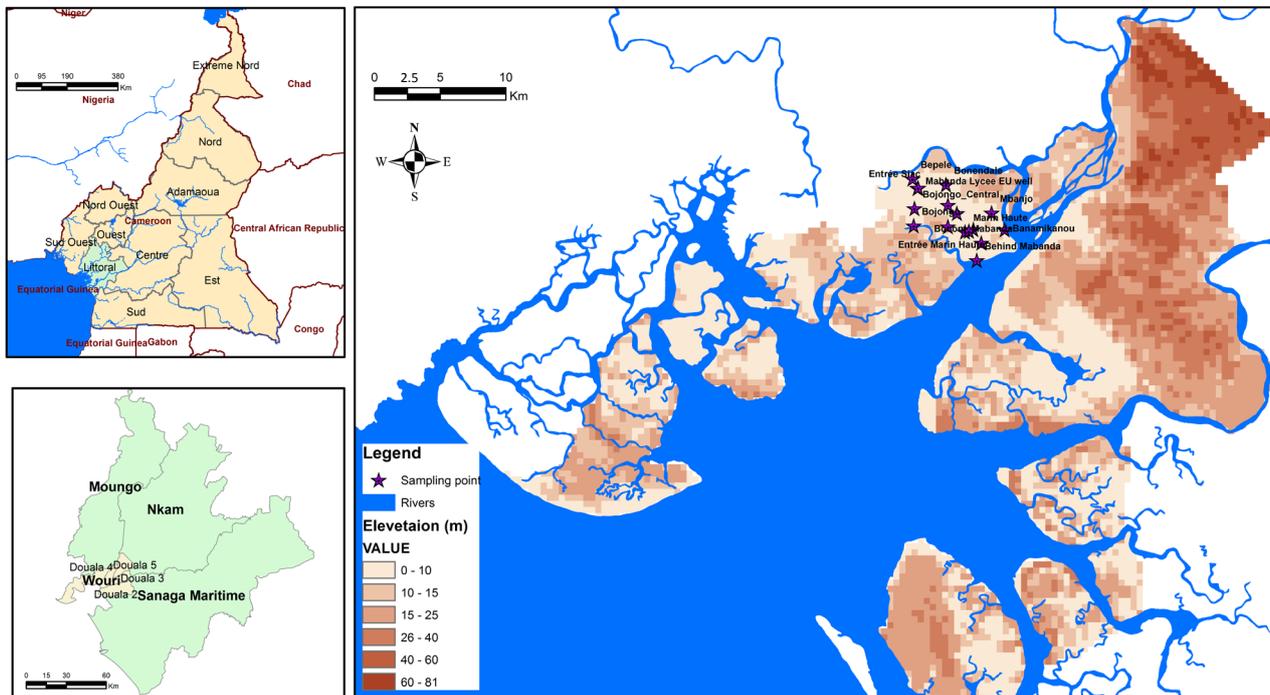


Figure 1. Location of sampling points.

metropolis lies directly on the Mio-Pliocene to Recent alluvial sediments which constitute the Wouri Formation of the Douala basin. It is made up of unconsolidated fine to coarse-grained sand and gravel mixed with silt and clay in various proportions. The alluviums are composed predominantly of quartz and kaolinite with a general thickness that ranges between 50 and 60 m.

Zonal and azonal soils can be distinguished. The sandy ferralitic zonal soils are light coloured, freely drained with top sandy soils and sandy clay sub-soils. These soils are formed in sandy sediments that are rich in quartz with fairly high humus contents. They have a good friability and low water retention capacity; essentially, they result from the decomposition of sedimentary rocks as in [13]. Young azonal (alluvial) soils are found here and have not been sufficiently subjected to the effects of soil forming processes as in [14]. These are immature soils that are good for agriculture.

2.2. Experimental Design

Four coastal zone routes in the Douala IV district of Bonassama were selected to check SWI into open wells between November 2014 to May 2015. The locations and altitudes of selected sample sites were determined on the field using a Garmin GPS 60 CX. This enabled the location of the coordinates to produce sample location maps (Figure 1). The routes and 16 of the sampling stations with their location are given in Table 1.

2.3. Collection and Analysis of Water Samples

Water was drawn from the open wells using drawing buckets tied with ropes.

Table 1. Sampling stations and routes.

	Coastal Transect (CT)	Inner Station 1 (IT1)	Inner Station 2 (IT2)	Inner Station 3 (IT3)
Route 1	Close to EU well	Entre Marin Haute	Bocom (Mabanda)	Bonamikano
Route 2	EU well	Ngwele Entrance	Marche Grand Hangar	Mbanjo
Route 3	Marin Haute	Bojongo	Entree Ndobu	Ndobu Slaughter House
Route 4	End of Ngwele	Mid way Bojongo	Entrée Siac	Bepele

Each water sample was collected into a collector after thorough rinsing with the water to be sampled. Water from the collector was filled into 500 mL capacity plastic bottles and preserved to avoid evaporation. The samples were stored in ice-chest containers at about 4°C prior to laboratory analysis to exclude microbial activity and unwanted chemical activity. Temperature, electrical conductivity (EC), dissolved oxygen (DO), pH salinity and TDS were measured directly on the field using a DO probe and an EC/pH/temperature metre.

Chemical parameters were determined as per standard APHA methods as in [15]. The chemical analyses were performed in the Laboratory for Waste Water Research of the Faculty of Science in the University of Yaounde 1, Cameroon. The total hardness of calcium (Ca^{2+}) and magnesium (Mg^{2+}) were determined by using the EDTA complexometric titration method; chloride (Cl^-) by the argentometric method; nitrate, sulphate and phosphate measurements were done using ion chromatography on a Dionex ICS-900. Reference [16], indicates that the charge balance calculated should be within the acceptable limit of $\pm 5\%$. Sodium (Na^+) and potassium (K^+) were analysed by the flame photometric method; salinity was measured using HACH salinity metre. Bicarbonate (HCO_3^-) was measured through titration by adding 0.02-M HCl solution to the sample while stirring until the endpoint marked by a pH of 4.5, was obtained.

To get a clear grasp of the spatial chemical composition of water resources in the Douala IV area, water samples were plotted on a Piper diagram using AQUACHEM software as in [17]. The Wilcox formula was used to determine the % Na as in [18]. The physico-chemical properties of ground water resources in Douala IV were further evaluated graphically by plotting EC and sodium absorption ration (SAR) on the United States Salinity (USSL) diagram as in [19].

3. Results

3.1. Physical Characteristics of Water Samples

A univariate analysis of the physico-chemical data of ground and surface water in Douala IV is presented in **Table 2**. Water sources characteristics in the study area are not uniform implying that the chemical composition varies over a wide range in physical composition. The average ground water temperature ranged from 27.7°C to 30.1°C which is slightly higher than the present day, mean atmospheric temperature of 27.62°C. Ground water usually maintains a fairly constant temperature which, for superficial aquifers, is normally close to the

Table 2. Statistical summary of physico-chemical data of dry season water resources in the study area.

Parameters	Costal transect (n = 4)				Inner transect 1 (n = 4)				Inner transect 2 (n = 4)				Inner transect 3 (n = 4)				Surface water (n = 4)			
	Min	Max	Mean	Std D	Min	Max	Mean	Std D	Min	Max	Mean	Std D	Min	Max	Mean	Std D	Min	Max	Mean	Std D
Water Temp.	26	29	27.7	0.81	27	30.2	28.14	0.85	27	28.9	27.9	0.69	26.8	29	28.02	0.69	29.9	30.3	30.1	0.21
pH	5.16	6.52	6.07	0.39	4	6.19	4.76	0.725	5.13	6.95	6.27	0.59	3.89	6.21	5.23	0.77	6.49	6.89	6.70	0.20
EC	61	7540	1661.5	2235.53	44.6	594	336.72	221.1	45	1952	629	662	76	226	136.39	61.08	1660	10,500	7388	4977.8
Salinity	0	4.2	0.825	1.28	0	0.2	0.11	0.09	0	1	0.3	0.4	0	0.1	0.01	0.03	0.8	6	4.16	2.90
DO	0.25	2.71	0.955	0.94	0.99	2.47	1.80	0.481	0.18	1.87	1.08	0.65	0.67	5.17	2.68	1.20	3.02	5.63	3.97	1.44
TDS	31	3750	837.75	1112.76	25	300	170.25	113.2	20	965	344	351	32	130	70.33	33.21	810	5300	3703.33	2510.2
NO ₃ ⁻	0	20.14	4.495	6.86	1	57.4	16.46	17.9	0.4	17.2	5.48	5.23	1.2	34	10.91	9.50	0.16	3.33	1.98	1.63
PO ₄ ³⁻	0.04	5.2	1.035	1.41	0.03	6.1	0.83	1.704	0.02	1.2	0.52	0.44	0.12	1.4	0.42	0.36	0.11	0.16	0.14	0.03
HCO ₃ ⁻	0.2	5.5	1.3125	1.49	0.1	5	1.77	1.459	0.1	6.5	2.3	2.31	0.1	15	2.70	4.05	0.11	1.00	0.65	0.50
SO ₄ ²⁻	8	97	44.75	29.17	0	43	11.17	14.58	0	103	47	47.5	0	27	10.50	11.01	59	220	158	86.6
Ca ²⁺	20	80	42.73	18.26	4	84	30.67	22.26	12	256	85	85.4	8	44	20.33	10.98	12	72	36	31.7
Mg ²⁺	2.4	69.6	21.8	21.82	0	19.2	8.20	6.428	0	98.4	20.4	27.6	0	31.2	6.60	9.27	12	110.4	49.6	53.1
Na ⁺	3	73	30.92	23.11	8	53	20.08	11.77	3	63	29.4	19.2	3	123	34.58	34.05	13	23	16.33	5.77
K ⁺	0	1.47	0.45	0.58	0	10.8	1.56	3.358	0	9.6	1.35	3.14	0	5.47	0.64	1.55	0	0.13	0.04	0.075
Cl ⁻	1	223.6	88.58	76.40	0	95.2	19.24	27.73	0.3	61.3	17.8	18.2	0.8	34.6	7.53	9.85	9.7	243	122.9	116.81

mean annual air temperature. In this case, the mean of 27.6°C was slightly higher than the mean atmospheric temperature of 26.6°C. This could accelerate the rate of chemical reactions in such systems.

The dissolved oxygen ranged from 0.95 to 2.68 mg/L for ground water and 3.97 mg/L for surface water indicating a low degree of solubility. All water samples showed DO of below 5 mg/L. The measurement of DO is critical to the scientific understanding of the potential for chemical and biochemical processes in groundwater. Concentrations below 5 mg/L may adversely affect the functioning and survival of biological communities and below 2 mg/L may lead to the death of most fish. The measurement of DO can be used to indicate the degree of pollution by organic matter, the destruction of organic substances and the level of self-purification of the water.

The probable low solubility values could be as a result of higher water temperatures and organic matter pollution. Reference [20] highlighted this fact indicating that low DO could have severe adverse effects on the geochemistry and ecology of different systems.

The entire aquifer system is acidic because all samples (100%) showed pH values less than 7. The pH of ground water ranged from 4.7 to 6.2 (Table 2) indicating that the waters are highly acidic to slightly neutral. Similarly, pH value of 6.7 for surface waters indicate that they were slightly acidic. The presence of azonal and zonal soils and organic litter from mangroves in the area may have

contributed to the acidic nature of the well waters. This could probably result from the decomposition of soil organic matter as in [21]. Additionally, poor sewage disposal methods in this area compounded by the intrusion of brackish water into the area could account for the pH of the water. Reference [22], shows that these processes involve ground water reactions with lateritic soils and humic acids from the decomposition of organic matter in the infiltration zones leading to acidification from dissolved carbon dioxide. Related findings as in [23] attest to this. Reference [24] indicated that acidity in ground water could also result from the reduction in the acid buffering capacity due to soil acidification.

The mean TDS of ground water samples ranged from 70.33 - 837.75 mg/L while that for SW stood at 3703.33 mg/L (Table 2). Most of the ground waters except for two points of the coastal transect, are fresh while the surface water are not fresh following the definitions as in [25] who indicated that TDS values < 1000 mg/L represent fresh water. The EC readings ranged from 136.4 - 1661.5 and 1660 - 10500 for well water and surface water respectively. EC values were distinctly higher at stations close to the coast. The salinity data for this period ranged from 0.008 to 0.825 for IT3 and CT respectively while that for surface water ranged from 0.8 - 6 mg/L.

Generally, the findings have shown that salinity decreased with an increase in distance from the coastal transects to inner transects as expected. The highest salinity values were for the coastal transects at SW1 and SW2.

The huge variation in TDS of surface water and ground waters of 70.33 - 837.7 mg/L and 3703.33 mg/L respectively can be ascribed to the fact that these waters were close to the sea where mangroves could be found, probably indicating an intrusion of salt water. High TDS in water may produce bad taste, odour and colour and may also induce unfavourable physiological reactions in the consumer as in [26]. This could probably account for the bad odour and taste of ground water resources as confirmed by some residents of the study area.

3.2. Chemical Characteristics of Water Samples

The leading cation was Ca^{2+} which had an average percentage concentration of 62.4%, 50.6%, 44.6% and 33.3% as shown in Figure 2 at IT2, IT1, CT, SW and IT3 respectively. This was followed by Na^+ (55.6%, 33.19%, 32.2%, 21.6% and 16%) in IT3, IT1, CT, IT2 and SW respectively. Mg^{2+} was the third and K^+ the last in percentage concentration with values of 48.6% and 2.1% in SW and IT1 respectively in that order (Figure 2).

The major anions were sulphate, chloride, nitrate, bicarbonate and phosphate representing 64.29% (IT2), 63.19% (CT), 34.03% (IT3), 8.41% (IT3) and 1.69% (IT1) respectively (Figure 3). The low K^+ content was probably due to its low content in rocks of the study area.

Nitrate concentrations indicate that 6.3% of sample (one sample) exceeded 10 mg/L (WHO limit for nitrates in drinking water). The probable sources of nitrates could have been domestic detergents (application of fertilizers and domestic waste waters).

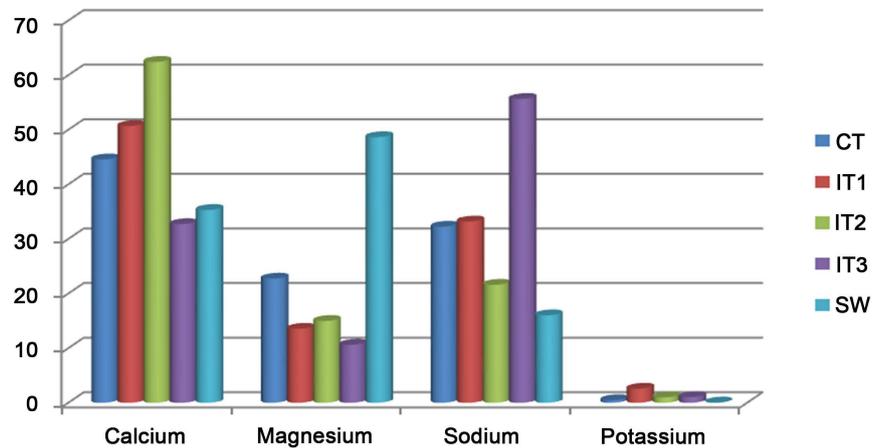


Figure 2. Percentage cations in the different transects.

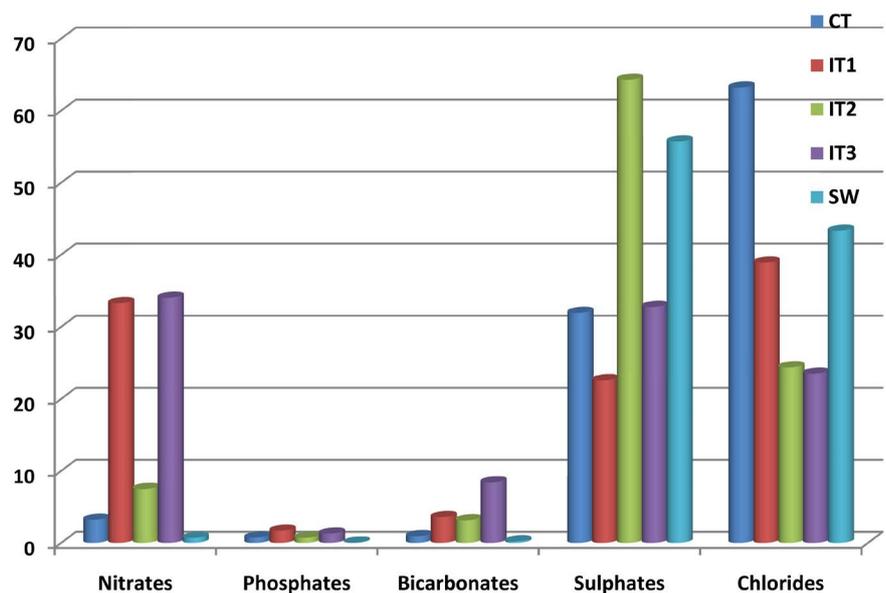


Figure 3. Percentage anions in the different transects.

3.3. Hydrochemistry of Water Samples

The Spearman correlation analysis (**Table 3**) indicates the association between key ions in water resources within the area of study. The results highlight a strong regression statistical analysis with R^2 values of 1 for EC-Salinity, EC-TDS, EC-Mg²⁺, Salinity-TDS, Salinity-Mg²⁺, TDS-Mg²⁺. Strong positive correlations of T°C-DO, pH-SO₄²⁻, EC-SO₄²⁻, EC-Cl⁻, Salinity-SO₄²⁻, salinity-Cl⁻, TDS-SO₄²⁻, TDS-Cl⁻ of 0.9 suggesting that the different aquifer chemistry is controlled by these parameters as in [27]. A strong positive correlation of 0.9 exists between, SO₄²⁻-Mg²⁺, Mg²⁺-Cl⁻. Another minor correlations of Na⁺-HCO₃⁻, K⁺-HCO₃⁻, Ca²⁺-SO₄²⁻ of 0.7, 0.6 and 0.7 respectively which further control the aquifer chemistry of the study area was observed.

The correlations between NO₃⁻-HCO₃⁻ and NO₃⁻-K⁺ of 0.7 and 0.9 respectively could suggest poor sanitation conditions in terms of proximity of toilets to wells and the use of detergents.

Table 3. Spearman's correlation matrix for water samples in Douala IV.

	Water Temp.	pH	EC	Salinity	DO	TDS	NO ₃ ⁻	PO ₄ ³⁻	HCO ₃ ⁻	SO ₄ ²⁻	Ca ²⁺	Mg ²⁺	Na ⁺	K ⁺	Cl ⁻
Water Temp.	1														
pH	0.1	1													
EC	0.1	0.8	1												
Salinity	0.1	0.8	1	1											
DO	0.9	0.2	0	0	1										
TDS	0.1	0.8	1	1	0	1									
NO ₃ ⁻	0	-0.9	-0.9	-0.9	-0.1	-0.9	1								
PO ₄ ³⁻	-0.7	-0.5	-0.1	-0.1	-0.9	-0.1	0.3	1							
HCO ₃ ⁻	-0.3	-0.5	-0.9	-0.9	-0.1	-0.9	0.7	0	1						
SO ₄ ²⁻	0.2	0.9	0.9	0.9	0.1	0.9	-0.8	-0.3	-0.7	1					
Ca ²⁺	-0.5	0.6	0.6	0.6	-0.6	0.6	-0.5	0.3	-0.3	0.7	1				
Mg ²⁺	0.1	0.8	1	1	0	1	-0.9	-0.1	-0.9	0.9	0.6	1			
Na ⁺	-0.7	-0.4	-0.6	-0.6	-0.4	-0.6	0.3	0.3	0.7	-0.7	-0.2	-0.6	1		
K ⁺	-0.1	-0.7	-0.7	-0.7	-0.3	-0.7	0.9	0.4	0.6	-0.5	-0.1	-0.7	0.1	1	
Cl ⁻	0.3	0.5	0.9	0.9	0.1	0.9	-0.7	0	-1	0.7	0.3	0.9	-0.7	-0.6	1

The multivariate analysis approach (MAA) has been used in the last years for the assessment of water quality. This is because water quality assessment often requires that several variables be studied concurrently. The use of MAA to link physical and chemical data will enable an overview of the potential cause-effect associated with concentrations of major ions and physical parameters of multi-variable sets as in [28]. MAA values identify the sources and or controlling factors of SWI into the Douala IV coastal ground water as in [29]. It is a multivariate statistical technique to explore variable distributions. The Factor analysis was performed on the correlation matrix that is the variables were auto-scaled (standardized) so as to be treated with equal importance. All analyses were performed using the PCA option of the factor procedure, followed by the basic set-up for factor analysis procedure from the XLSTAT 2015 statistical software.

The present study chose to interpret a variable or group of variables as those associated with a particular factor where loadings were 0.4 or greater, corresponding to an associated explained variance over 65%. This approximates Comrey's (1973) cut-off of 0.55 for a good association between an original variable and a factor and takes into account discontinuities in the magnitudes of loadings approximating these original variables as in [30].

The results of varimax with the Kaiser normalisation rotated factor loading for water resource chemistry data in Douala IV are shown in **Table 4** and **Figure 4** (biplot) respectively which highlight the controlling factors determining the water chemistry of the study area. These resulting factors are vital because they

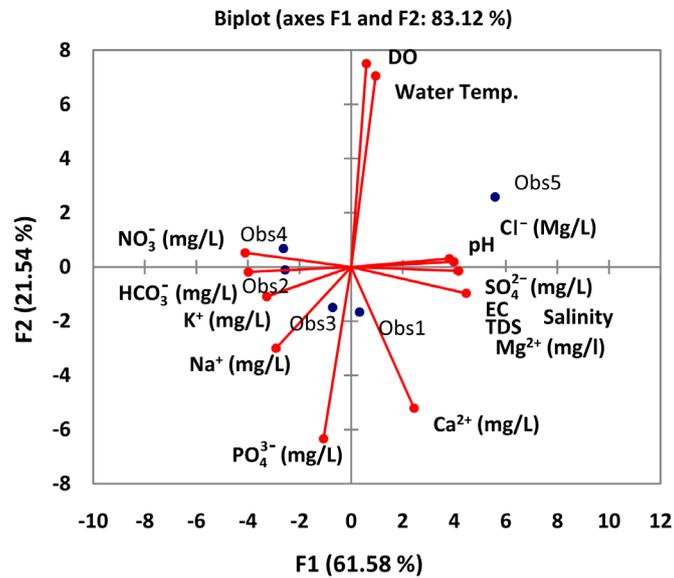


Figure 4. Associations from which cause-effect relationships controlling ground water resources are inferred in the study area.

Table 4. Factor loading, eigenvalues and variances for water resource chemistry data in Douala IV.

Variables	F1	F2	F3	F4
Water temperature	0.211	0.925	-0.307	0.081
pH	0.845	0.040	0.450	0.287
EC	0.990	-0.128	-0.043	-0.043
Salinity	0.990	-0.128	-0.043	-0.043
DO	0.132	0.984	0.119	-0.014
TDS	0.990	-0.128	-0.043	-0.043
NO ₃ ⁻ (mg/L)	-0.911	0.068	-0.383	0.137
PO ₄ ³⁻ (mg/L)	-0.238	-0.832	-0.448	-0.226
HCO ₃ ⁻ (mg/L)	-0.884	-0.024	0.372	0.283
SO ₄ ²⁻ (mg/L)	0.924	-0.020	0.024	0.381
Ca ²⁺ (mg/L)	0.542	-0.684	0.068	0.484
Mg ²⁺ (mg/l)	0.990	-0.128	-0.043	-0.043
Na ⁺ (mg/L)	-0.646	-0.393	0.567	-0.325
K ⁺ (mg/L)	-0.726	-0.143	-0.480	0.471
Cl ⁻ (mg/L)	0.884	0.024	-0.372	-0.283
Eigenvalue	9.237	3.231	1.499	1.033
Variability (%)	61.582	21.540	9.992	6.886
Cumulative (%)	61.582	83.122	93.114	100.00

pinpoint on the controlling processes that lead to the different water facies obtained in this area. Four factor loadings with a 100% variance explained in the dataset are accounted for from **Table 4**. The loading of variables on each factor

and the corresponding percentage of data variance are explained by each factor.

The first principal factor is predominant and accounts for 61.6% of the total variance with high positive loadings of pH, EC, salinity, TDS, SO_4^{2-} , Ca^{2+} , Mg^{2+} and Cl^- and high negative loadings of K^+ , Na^+ , HCO_3^- , NO_3^- . This factor can be termed salinity-dissolved solute factor. The first factor determines the main processes that control ground water chemistry. The association of variables (EC, pH, salinity, TDS, SO_4^{2-} , Ca^{2+} , Mg^{2+} and Cl^-) in factor 1 (61.6%) of total variance indicates that these variables will impact significantly on the resulting water facies that would be obtained. This is in agreement with the observations as in [24] who indicated that the major ions, pH and TDS are sensitive to variations in EC. The fact that SO_4^{2-} , Ca^{2+} and Cl^- ions are under factor one, shows their predominance in the determination of the water types.

Reference [31], noted that a high negative loading of variables in a factor shows that they go contrary to the process that generated the negative high loadings. Positive loading indicates the interrelationship of variables that determine the water types of the study area by linking their different concentrations.

The second factor explains 21.5% variance indicating that the solubility of DO decreases with an increase in temperature; the reason why they could be found under the same factor. MAAs have been carried out as in [32]-[38]. Reference [39] highlighted two factors that were responsible for controlling hydro-chemical processes of ground water in Togo. Their factors were related to the ions with dominant concentrations and therefore the main contributors to groundwater salinity in this study.

Factor 3 was observed to account for 9.9% of the variance while factor 4 accounted for 6.9% of the total variance and it is associated with high pH and Na^+ loading and K^+ , and Ca^{2+} loadings respectively for factors 3 and 4. Their preferences in regulating ground water chemistry could therefore be minimal.

The factor analysis does not produce cause effect relationship but highlights the associations from which such cause-effect relationships can be inferred as in [31]. Factor analysis is not used in isolation but it is used with several graphical techniques to make a meaning from hydro-chemical analysis and to identify the water types in the area under investigation. This is why some equations of bivariate plots of EC and chloride and major ions and chloride were exploited together with the Piper diagram to better understand the underlying controlling processes of ground water resources in the study area as in [17].

3.4. Equations and R^2 Value from Linear Trend Lines for the Determination of Saline Intrusion

The association between key ion chemistry is not enough to deduce the cause-effect relationship of hydro-chemical facies. Reference [40] [41] [42] noted that an early indication of SWI into ground water may be a high EC because sea water has a higher EC value than fresh water.

Additionally, an increase in Cl^- concentration may be the first indication of SWI into an area devoid of other possible sources of saline contamination. Con-

sequently, equations and R^2 value of EC and chloride and major ions and chloride were derived from their corresponding plots to attest to this.

Chloride/EC:

$$Y = 0.0149x + 21.032 \quad (1)$$

$$R^2 = 0.7782$$

Sodium/chloride:

$$Y = -0.0765x + 30.183 \quad (2)$$

$$R^2 = 0.2603$$

Calcium/chloride:

$$Y = -0.0501x + 45.512 \quad (3)$$

$$R^2 = 0.0107$$

Magnesium/chloride:

$$Y = 0.2967x + 6.1289 \quad (4)$$

$$R^2 = 0.7824$$

Sulphate/chloride:

$$Y = 1.0032x + 2.9132 \quad (5)$$

$$R^2 = 0.7252$$

Equation (1) reveals a good correlation of electrical conductivity and the chloride concentration in Douala IV ($R^2 = 0.7782$). This indicates that a mixing between sea water and fresh water are taking place. From Equation (1), the mechanisms of mixing between sea water and fresh water are complex because otherwise, there would have been a perfectly linear correlation between electrical conductivity and the chloride concentration as has similarly been observed as in [43].

From Equation (2), the relationship between sodium and chloride shows a depletion of sodium and a nearly linear relation of calcium and chloride as seen in Equation (3). However, magnesium and sulphate enrichment with chloride, at some points, suggest a strong water aquifer interaction related to direct cation exchange (Equations (4) and (5)). This sample must be from the CT and SW transects as they show a large deviation from the theoretical line of mixing as a result of their high sulphate and magnesium content.

The Piper diagram (Figure 5) indicates the different water types. Reference [44], indicates that the Piper diagram is preferable because it give a better identification of the water samples and the dorminant geo-chemical processes in the water chemistry.

Figure 5, indicates that the water types are of $\text{Ca}^{2+}\text{-}2\text{Cl}^-$, $\text{Na}^+\text{-Cl}^-$, $\text{Ca}^{2+}\text{-SO}_4^{2-}$ types. Five samples (IT3, 2 IT2 and 2 IT1) representing 26.32% of the total samples are of the $\text{Ca}^{2+}\text{-Cl}^-$ type. Two samples (10.53%) of the total samples are of the $\text{Na}^+\text{-Cl}^-$ type and are found in the IT3 transects and 12 samples (one in IT3, two in IT2, two in IT1, four in CT and three in SW) representing 63.2% of the total number of samples are of the $\text{Ca}^{2+}\text{-SO}_4^{2-}$ facies water type.

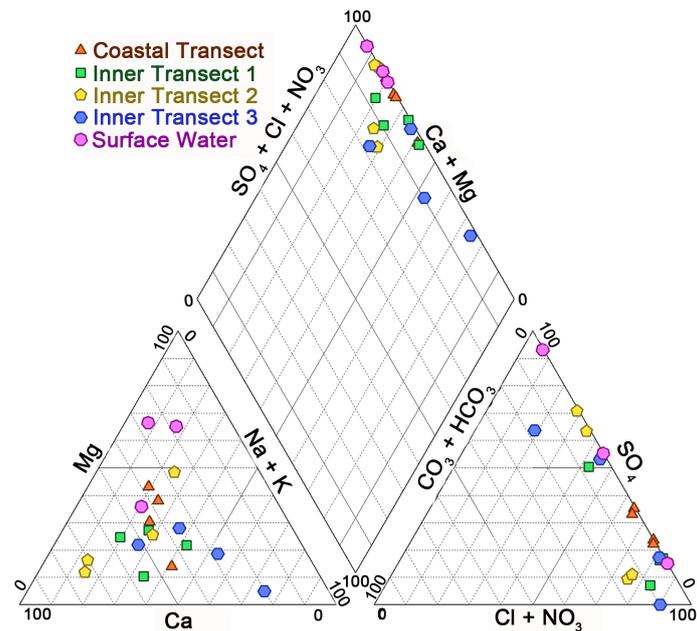
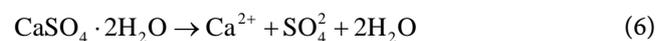


Figure 5. Piper diagram showing the water types and its proportions of Douala IV.

The study shows that the abundance of cations was $\text{Ca}^{2+} > \text{Na}^+ > \text{Mg}^{2+} > \text{K}^+$ indicating that Ca^{2+} is the most abundant followed by Na^+ , Mg^{2+} and K^+ . The relative abundance of anions is in the order of $\text{SO}_4^{2-} > \text{Cl}^- > \text{NO}_3^- > \text{HCO}_3^- > \text{PO}_4^{3-}$. This indicates that the leading anion is SO_4^{2-} followed by Cl^- , NO_3^- , HCO_3^- and PO_4^{3-} . The likely source of Ca^{2+} and SO_4^{2-} could be from the dissolution of gypsum and halite as in [45] which could be present in the catchment area. This could be so because a positive correlation matrix of Ca^{2+} - SO_4^{2-} ($R^2 = 0.7$) exists. Reference [46], shows gypsum to be extensively distributed in the subsurface layer of aquifers and generally its dissolution produces calcium and sulphate as seen in Equation (6).



The correlation between K^+ - NO_3^- ($R^2 = 0.9$) suggests a probably poor sanitation conditions in the area.

The majority of samples are depleted in Na^+ and enriched in Mg^{2+} and SO_4^{2-} (Equations (2), (4) and (5) respectively). This shows that Na^+ usually contained in sea water has been replaced by Ca^{2+} on the one hand, but this could also result from the dissolution of gypsum on the other hand.

4. Water Resource Suitability with Respect to Saline Water Contamination in Douala IV

The temperature values were high. This may increase taste, odour, colour and corrosion problems as already apparent in the water sources from field survey. Ca^{2+} and Mg^{2+} values were much higher than the WHO norms (Table 5). The presence of these alkaline earth metals signifies water hardness. Though hardness has no adverse effect on human health, there is an aesthetic concern because of the unpleasant taste that it imparts. It can prevent the formation of

Table 5. Ground water quality in douala IV compared with WHO (2004) safe drinking water standards.

Parameters	Range in study area	WHO limits
Water temperature (°C)	26 - 30.3	15
pH	4.7 - 6.7	6.5 - 8.5
EC (µS/cm)	136 - 7388	750
Salinity (mg/L)	0.008 - 4.167	-
DO (mg/L)	0.955 - 3.970	5
TDS (mg/L)	70 - 3703	500
NO ₃ ⁻ (mg/L)	1.987 - 16.458	50
PO ₄ ³⁻ (mg/L)	0.143 - 1.035	-
HCO ₃ ⁻ (mg/L)	0.650 - 2.696	200
SO ₄ ²⁻ (mg/L)	10.50 - 158	200
Ca ²⁺ (mg/L)	20 - 85	75
Mg ²⁺ (mg/L)	6 - 49	30
Na ⁺ (mg/L)	16 - 34	200
K ⁺ (mg/L)	0.043 - 1.559	100
Cl ⁻ (mg/L)	7 - 122	200

lather with soap, a problem which the inhabitants of CT are already facing. This may also increase the boiling point of water and cause encrustation in the water supply distribution system.

Fresh water resources in Douala IV are vulnerable to acid contamination as the waters from the area are acidic with a range of 4.7 - 6.7 (Table 5). Additionally, the levels of EC are high as in [47]-[52] which is an early indication of SWI. EC is a good measure of salinity hazard on crops because an excess of salinity reduces the osmotic potential of plants and thus interferes with the absorption of water and nutrients from the soil as in [39]. An evaluation of ground water fitness for agricultural uses, based on standards put in place by international organisations as in [53] [54] is imperative.

The analytical plot of the Wilcox formula (Figure 6) shows that sodium is commonly measured for water used for drinking or agricultural purposes. Elevated sodium contents in certain soil types can degrade soil structure thereby restricting water movement and affecting plant growth.

The analysis showed that 12 water samples (representing 63.2%) were excellent for irrigation; 3 samples (CTc, IT2b, SWa), representing 15.8% of total number of samples, were good; sample CTb (5.3%) was doubtful while 3 samples (CTa, SWb, SWc), representing 15.8%, were unsuitable for irrigation purposes in term of % Na (Table 6).

The sodium adsorption ratio (SAR) was further used to evaluate and confirm the suitability of water for irrigation. The ratio estimates the degree to which sodium will be adsorbed by the soil. High values of SAR, which are related to the

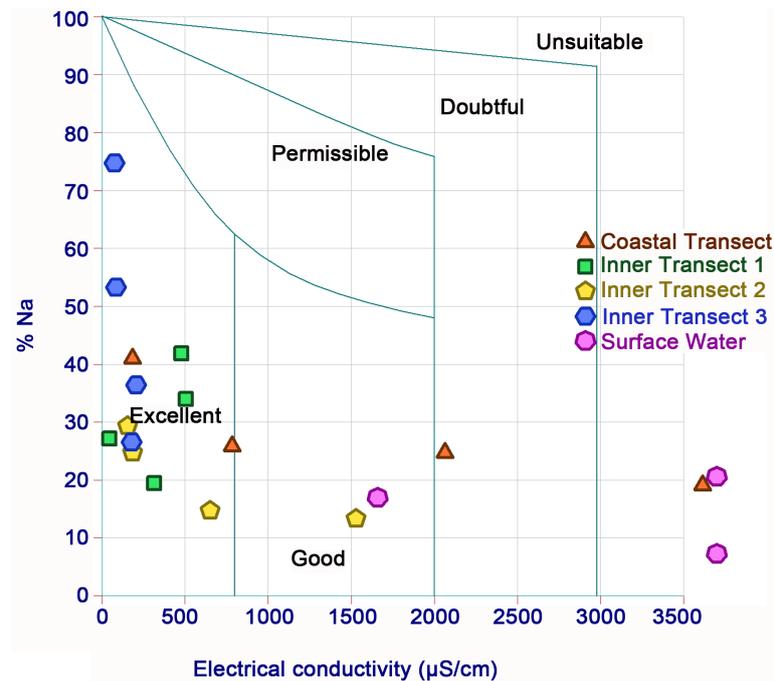


Figure 6. The suitability of water resources in Douala IV for irrigation purposes based on % Na.

Table 6. Sampled points and corresponding quality of water based on % Na.

% Na	Quality of water	Samples
<20%	Excellent	12 (CTd, IT1a, IT1b, IT1c, IT1d, IT2a, IT2c, IT2d, IT3a, IT3b, IT3c and IT3d)
20% - 40%	Good	3 (CTc, IT2b, SWa)
40% - 60%	Permissible	None
60% - 80%	Doubtful	1 (CTb)
80%	Unsuitable	3 (CTa, SWb, SWc)

EC, imply that sodium in the irrigation water may replace calcium and magnesium ions in the soil, potentially causing damage to the soil structure as in [39]. Such studies have been carried out elsewhere with promising implications for policy measures as in [55] [56] [57] [58] [59].

The analytical data for ground water chemistry that was plotted on the US salinity diagram (Figure 7) illustrates that 8 samples, representing 42%, are found in the low salinity and low alkalinity (C1S1) area and is good for irrigation purposes.

Four samples (21.1%) found in the C2S1 type which represents medium salinity and low alkalinity in water qualities showed that the irrigation quality of water was fair in this transects of Douala IV. Type C3S1 had four samples (21.1%) indicating a high salinity and low alkalinity in water, which can be used for irrigation on almost all types of soils with little danger of exchangeable sodium as in [55]. Three samples however, which represented 15.8% of total number of samples were found in type C4S1 indicating a very high salinity and low alkalinity and were thus not suitable for agricultural activities as shown on Table 7.

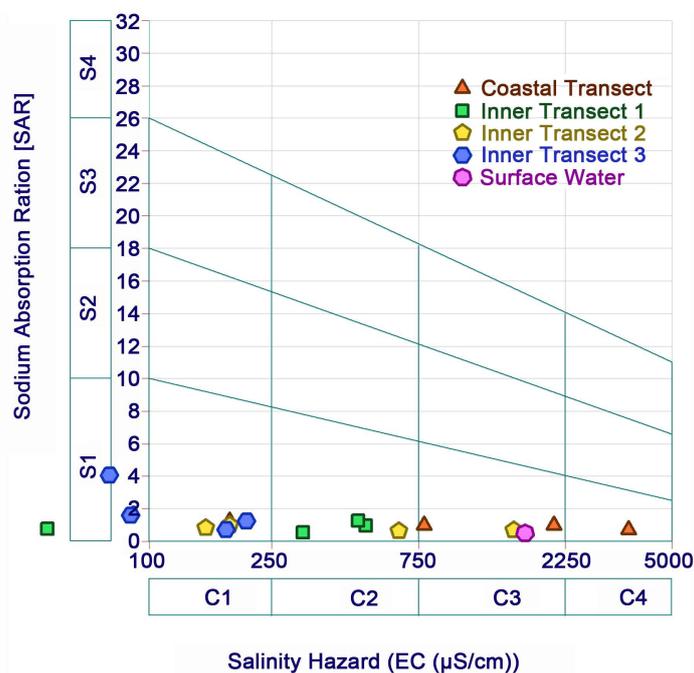


Figure 7. A classification of irrigation waters using the US Salinity diagram.

Table 7. Quality of irrigation water based on electrical conductivity.

Salinity hazard class	Specific conductance ($\mu\text{S}/\text{cm}$)	Samples
Low	0 - 250	8 (CTd, IT1d, IT2c, IT2d, IT3a, IT3b, IT3c and IT3d)
Medium	250 - 750	4 (IT1a, IT1b, IT1c and IT2a)
High	70 - 2250	4 (CTb, CTc, IT2b and SWa)
Very high	>2250	3 (CTa, SWb, SWc)

The spatial variation (CT, IT1, IT2, IT3 and SW) and contamination of ground and surface water resources in the low lying area of the Douala IV metropolis were assessed using both physical and chemical parameters (EC, pH, salinity, TDS, concentrations of Na^+ , SO_4^{2-} , Mg^{2+} , Ca^{2+} , Cl^-). Two points of ground water of IT3 were located close to the Wouri River and the other two were far inland from the coast. These waters are rich in Na^+ , Ca^{2+} , Cl^- and exhibit all the water types ($\text{Ca}^{2+}\text{-Cl}^-$, $\text{Ca}^{2+}\text{-SO}_4^{2-}$ and $\text{Na}^+\text{-Cl}^-$). $\text{Na}^+\text{-Cl}^-$ water types indicates that the water type undergoes primary salinity while $\text{Ca}^{2+}\text{-2Cl}^-$ and $\text{Ca}^{2+}\text{-SO}_4^{2-}$ water types suggest a possible mixing with sea water.

Two points each from the IT2 and IT1 are of the $\text{Ca}^{2+}\text{-SO}_4^{2-}$ water types and the other two points are of the $\text{Ca}^{2+}\text{-2Cl}^-$ water type. All sample points from CT and SW are of the $\text{Ca}^{2+}\text{-SO}_4^{2-}$ water types. These suggest that secondary salinity is taking place at these points indicating pollution from domestic and leaking sewage and that the processes controlling ground water chemistry in the area are varied.

5. Discussion

The interpretation of hydro-chemical analysis in the present study showed that NO_3^- , PO_4^{3-} , HCO_3^- , SO_4^{2-} , Na^+ , K^+ , and Cl^- met the guideline limits while T°C , pH, EC, DO, TDS, Ca^{2+} and Mg^{2+} were above WHO limits.

MAA values obtained for water variables showed that F1, F2 and F3 accounted for 61.6%, 21.5% and 9.9% respectively of total variance with strong loadings and these were considered to determine the ground and surface water qualities of the area. F4 was not considered due to a much lower variance (6.8%). The study indicated calcium followed by sodium to be the most abundant cations. Sulphate was the most abundant anions. Furthermore, most samples were depleted in sodium and enriched in magnesium and sulphate from Equations (2), (4) and (5) respectively.

Reference [5] [31] [47] [48] [49] [50] indicated that depletion in Na^+ and a gain in Mg^{2+} and Ca^{2+} highlight a direct cation exchange observed in situations where sea water is replacing fresh water. When SWI occurs, there is the favoured release of Mg^{2+} and Ca^{2+} from exposed bedrock due to weathering of clay minerals. Consequently, free Cl^- anions are released and there is the adsorption of Na^+ onto mineral surfaces. The fact that Ca^{2+} and Mg^{2+} ions are released and Na^+ are engrossed points to salinization of water resources in the study area. Ca^{2+} is however one of the major cations and this suggests the existence of other sources contributing to the enrichment of groundwater. This could be substantiated by the dissolution of gypsum and associated minerals in the aquifer.

Piper diagrams illustrated that water types are of $\text{Ca}^{2+}\text{-SO}_4^{2-}$ (63.2%), $\text{Na}^+\text{-Cl}^-$ (10.53%) and $\text{Ca}^{2+}\text{-2Cl}^-$ (26.32%) type. This shows that all water types are made up of strong acid ions (SO_4^{2-} , Cl^-) which exceed weak acids (HCO_3^- and CO_3^{2-}). Alkaline earths (Ca^{2+} and Mg^{2+}) exceed alkalis (Na^+ and K^+) in all samples (89.5%) except for two from IT3. **Figure 5** illustrates that these samples are of non-carbonate hardness, indicating that secondary salinity exceeds 50% and the main properties are dominated by alkaline earths and strong acid ions. This explains why the water types are of $\text{Ca}^{2+}\text{-SO}_4^{2-}$ and $\text{Ca}^{2+}\text{-2Cl}^-$ type and implies that the water types are from formations that could be composed of dolomite and limestone or from active recharge zones with short residence times as in [36]. Two samples from IT3 are of alkaline carbonate (primary salinity exceeds 50%) and its chemical properties are dominated by alkalis and weak acids, hence this water type is of the $\text{Na}^+\text{-Cl}^-$ type. These water types could be originating from halite dissolution (saline) or alkaline carbonate enrichment.

Primary salinity implies that the acquisition of salts is a result of natural dissolution of the underlying rocks in the study area and this is seen in two samples from IT3 representing 10.53% of total water types. The other two points which are far from the coastal transects must have started witnessing secondary salinity. The SW, CT, IT1 and IT2 are all experiencing secondary salinity, implying that either salt water from the ocean is gradually intruding into fresh water aquifers or the high rate of urbanisation is influencing secondary salinity. This could pose a fundamental problem to ground water resources in the future if nothing is

done.

The rest, 89.5% of water types is a result of anthropogenic activities, termed secondary salinization. 26.3% of this water type is of the $\text{Ca}^{2+}-2\text{Cl}^{-}$ type. Reference [21] indicated that the occurrence of $\text{Ca}^{2+}-2\text{Cl}^{-}$ water type in coastal aquifers signifies an active process of sea water mixing where Na^{+} from sea water is exchanged for Ca^{2+} adsorbed on clays. $\text{Ca}^{2+}-2\text{Cl}^{-}$ water types have developed in this area and the $\text{Na}^{+}-\text{Cl}^{-}$ water type occur close to the River Wouri. However, Reference [51] indicated that $\text{Ca}^{2+}-2\text{Cl}^{-}$ water types do not develop in the study area. Contrarily, this study underlines the fact that the area was experiencing salinisation. These differences could be attributed to the spatial differences in the studies. Additionally, Reference [36] indicate that typical sea water mixing or base ion exchange is observed in study areas when some water samples belonged to the $\text{Na}-\text{Cl}$ and $\text{Na}-\text{HCO}_3^{-}$ water types. Two samples representing 10.53% of total samples analysed are of the $\text{Na}^{+}-\text{Cl}^{-}$ water type, implying that sea water mixing or base ion exchange occurs in the study area.

Other probable sources of secondary salinity in this area could be urbanization and industrialisation activities. However, inhabitants of Douala IV do not carry out agricultural activities intensively and consequently; the most probable sources of secondary salinity could be urbanization and sea water intrusion into coastal aquifers.

Several authors have noted saline water intrusion into coastal aquifers as a major concern with hazards of probable rise in sea level aggravating the situation as in [60] [61] [62]. This constitutes the most common of all the pollutants of freshwater. Therefore, an understanding of saline intrusion is essential for the management of coastal water resources as in [56].

Varied studies on water quality for different end uses have been carried out in Cameroon. Reference [34], in their study centered on the suitability of ground water for drinking, domestic and agricultural uses in the banana plantation plains of Njombe, concluded that nitrate contamination is taking place. Reference [51] concluded that ground water in a highly urbanised coastal area is vulnerable to acid deposition. Reference [63], in their study, indicated that the continuous use of River Wouri and Dibamba for domestic purposes is unsafe and needs attention. Reference [22] in their different studies for the Gulf of Guinea and of Ndop plain respectively indicated that more than 50% of the water sources studied were suitable only for drinking and irrigation purposes. Their studies did not focus on water quality fitness with regards to hazards of a rise in sea level. Such studies as in Douala IV which is the economic capital of the nation, are thus vital in Cameroon. This study has highlighted a very contemporary issue that needs to be addressed before it gets out of hand. Raising awareness as in [64] to the inhabitants of this area to the effects of SLR hazards is important. Additionally the use of spatial data sets as in [65] as an initial requirement in depicting most vulnerable areas should be envisaged.

6. Conclusions

The interpretation of hydro-chemical analysis in the present study showed that

NO_3^- , PO_4^{3-} , HCO_3^- , SO_4^{2-} , Na^+ , K^+ , and Cl^- met the guideline limits while T°C , pH, EC, DO, TDS, Ca^{2+} and Mg^{2+} were above WHO limits. MAA values obtained for water variables showed that F1, F2 and F3 accounted for 61.6%, 21.5% and 9.9% respectively of total the variance with strong loadings and these were considered to determine the ground and surface water qualities of the area. F4 was not considered due to a much lower variance (6.9%). The piper diagrams illustrated that water types are of Ca^{2+} - SO_4^{2-} (63.2%), Na^+ - Cl^- (10.5%) and Ca^{2+} -2Cl (26.3%) type. All samples except two from IT3 (89.5%) were dominated by more than 50% of alkaline earth ions (Ca^{2+} and Mg^{2+}) and strong acid anions. These samples exhibited non-carbonate hardness indicating that the secondary salinity of ground water dominates over primary salinity. 10.5% (Na^+ - Cl^-) of total number of samples indicated seawater with a Na^+ dominant type or Cl^- dominant type waters indicating typical salinity through seawater mixing .

The main processes that determine the water types of the study area were: 1) reversed ion exchange which is a situation that occurs when alkaline earths metal ions ($\text{Ca}^{2+} + \text{Mg}^{2+}$) are in excess of alkali ions ($\text{Na}^+ + \text{K}^+$) and strong acid ions exceed weak acid ions and; 2) basic ion exchange processes which entails residual sodium carbonate disposition in the area, a situation in which Na^+ - Cl^- water types were identified in the study.

10.5% of water facies of IT3 were of primary salinity while 89.5% of water types were of secondary salinity. This could be attributable to gradual salt water intrusion from the sea into the aquifers as observed in the order of $\text{SW} > \text{CT} > \text{IT1} > \text{IT2} > \text{IT3}$. However, it could also have resulted from urbanization, although it was not clear which urban activities could have provoked the salinization of ground water resources. The fact that two sample points of IT3 experienced secondary salinity did not mean that they constitute a major health hazard in the future.

According to the Wilcox formula and the USSL diagram (**Figure 7**), 15.8% of the ground water samples were not suitable for irrigation purposes. The demand for water has increased resulting in an associated acute shortage of water to meet the daily water consumption needs of the people. Most of the residents of Douala depend on borehole wells for domestic, industrial and agricultural use. Present and future palatability of ground water resources for domestic purposes in Douala IV will go a long way to ensure public health and serve for important practical application of water resource management in the area.

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