

# Activated Carbon and Clay Minerals for the Sorptive Removal of Denatonium Ions from Denatonium Benzoate Solutions

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## Abstract

This study assessed the feasibility of utilizing activated carbon and clay minerals for treating water impacted with the bittering agent denatonium benzoate (DB). Our specific study objectives were to 1) evaluate denatonium ion sorption to smectite clay minerals (bentonite and hectorite) and activated carbon (powdered and granular) at constant pH and ionic strength and 2) examine the impact of pH on denatonium ion sorption to each solid material. The experimental results indicated that high doses (33,000 mg/L) of as-received granular activated carbon and as-received clay minerals completely removed denatonium from aqueous solutions containing 100 - 1000 mg/L denatonium benzoate. Powdered activated carbon at doses of 5 - 100 mg/L exhibited favorable monolayer sorption of denatonium ions from a pH 6.95, 70 mg/L aqueous denatonium benzoate solution with a Langmuir separation factor ( $r$ ) of 0.481, a maximum sorption capacity ( $S_m$ ) of 74 mg/g, and a Langmuir constant of 15.3 L/g. A maximum removal of 23% of denatonium was achieved at the highest powdered activated carbon dosage employed. Denatonium ion removal with peroxide treated bentonite and peroxide treated hectorite did not result in complete removal of the ion and exhibited favorable sorption as evidenced by Freundlich  $1/n$  values ranging from 0.803 to 1.194; Freundlich constants ( $K_f$ ) ranged from 8 ng/L to 575 ng/L. Denatonium ion sorption to peroxide treated bentonite appeared to depend on pH while hectorite sorption of denatonium ions was independent of hydrogen ion concentrations. For powdered activated carbon adsorption, as pH increased denatonium ion sorption decreased. Overall, the work demonstrates that denatonium can be effectively removed from water via activated carbon or clay mineral sorption.

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## Keywords

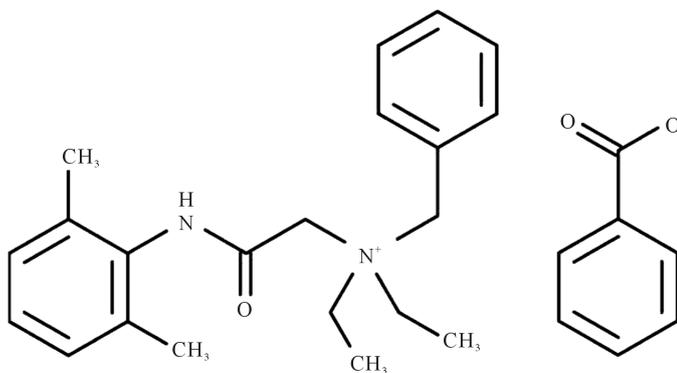
**Denatonium Benzoate, Activated Carbon, Clay Minerals, Sorption, Water Treatment, Wastewater Treatment**

### 1. Introduction

The protection of drinking water supplies from industrial and household contaminants along with the treatment of waters impacted by hazardous and undesirable constituents often involves chemical partitioning at the solid-solution interface [1]. Denatonium benzoate, shown in **Figure 1**, is a water soluble, bittering agent legislatively mandated to be added to antifreeze formulations in at least seventeen American states [2]; the denatonium ion is responsible for imparting the bitter taste. The principal motivation for the legislation is to prevent deaths due to accidental (or intentional) consumption of products containing ethylene glycol at concentrations exceeding 10% v/v [3]. As a component of antifreeze formulations, denatonium benzoate is expected to reach the soil (and possibly groundwater) through antifreeze discharges to the environment [4]. Given the large water solubility of denatonium benzoate (42,400 mg/L at 20°C), irreversible sorption to soils and sediments may not occur resulting in denatonium benzoate seepage into ground waters or runoff into surface waters used as sources for drinking water treatment facilities. As highlighted in a meeting before a congressional subcommittee considering federal legislation mandating the addition of denatonium benzoate to antifreeze formulations, a dearth of knowledge exists related to 1) the environmental fate of denatonium benzoate and 2) the feasibility of current water and wastewater treatment practices for removing denatonium benzoate from impacted waters [5].

The environmental fate of quaternary ammonium compounds (QACs), such as denatonium benzoate, has been studied with the principal focus being surfactant molecules [6]-[9]. To date, data overwhelmingly support a cation exchange mechanism for sorption to pure aluminosilicate minerals and to sediments containing layered aluminosilicates [7]. The majority of aluminosilicate sorption results from a partitioning of QAC molecules between the aqueous phase and interlayer surface(s) of 2:1 layered aluminosilicates whereby exchangeable, inorganic cations are displaced in a manner related to their hydration enthalpy [10].

Several physical, chemical, and biological methods exist for treating water (and wastewater) streams. Common treatment protocols include solvent extraction, ion-exchange, precipitation, biodegradation, and adsorption. Adsorption of dissolved species usually involves activated carbon [11] and/or layered aluminosilicate sorbents [12]. Activated carbon is a common solid sorbent used to remove odorous compounds (e.g. vanillin, methyl-naphthalenes, 2-pyrrolidone, 2-methylisoborneol) [13], explosive compounds such as 1,3,5-trinitro-1,3,5-triazacyclohexane (RDX) and 1,3,5,7-tetranitro-1,3,5,7-tetrazocane (HMX) [11], arsenic [14] and quaternary ammonium surfactants [15] [16] from impacted water and wastewater at treatment plants. Additionally, activated carbon was recently used *in-situ* to treat aquatic sediments containing hydrophobic organic compounds [17] [18].



**Figure 1.** The structure of denatonium benzoate with the quaternary ammonium denatonium cation and the benzoate anion.

Despite undergoing activated carbon treatment at some wastewater treatment plants (WWTP), estuarine sediments still contained detectable levels of QACs [19] [20]. Additionally, alkyl dimethylbenzylammonium chloride and alkyltrimethylammonium chloride QACs were detected in water streams downstream of WWTPs. The detected QACs can be grouped into three classes: dialkyldimethylammonium compounds, benzylalkylammonium compounds, and alkyltrimethylammonium compounds. Generally, extremely hydrophobic compounds within the dialkyldimethylammonium class were detected in large quantities relative to less hydrophobic and less persistent compounds within the same class and in the benzylalkylammonium and alkyltrimethylammonium classes. This observation suggests that super hydrophobic compounds are not extremely biodegradable as biodegradation is extensively employed at many WWTPs [19]. Collectively, these observations highlight the need to experimentally evaluate the efficacy of using traditional sorbents (*i.e.* activated carbon or clay minerals) to remove aqueous denatonium benzoate (DB) prior to DB's widespread use and likely introduction into water and wastewater. Recently, our team evaluated the kinetics of denatonium ion sorption to smectite minerals and determined that 1) sorption was rapid (and near complete) when using 200 ppm initial concentrations of DB and 2) sorption likely proceeded via a cation exchange process [21].

In this work, we studied the constant temperature sorption of denatonium ions to common sorbents utilized in water treatment processes (*i.e.* bentonite clay and activated carbon). The study objectives are to: 1) evaluate denatonium ion sorption to smectites and granular activated carbon at constant pH and ionic strength and 2) examine the impact of pH on denatonium ion sorption to each solid material.

## 2. Experimental

### 2.1. Materials

Denatonium benzoate was obtained from MP Biomedicals (USA) and used without further purification. The mineral sorbents were hectorite (SHCa-1) and bentonite (SWy-2). Each mineral was obtained from the Clay Minerals Society. The reported compositions of hectorite and bentonite clay minerals are shown in **Table 1**.

Granular activated carbon (GAC) produced via steam activation of bituminous coal and acid washing was obtained from NORIT Company<sup>®</sup> in virgin form with grain sizes between 0.4 and 1.68 mm. The carbon had a reported (BET) surface area of 1175 m<sup>2</sup>/g, apparent density of 480 g/L, and moisture content of 2%. The reported GAC pore volume distribution is 0.35 cm<sup>3</sup>/g, 0.22 cm<sup>3</sup>/g, and 0.26 cm<sup>3</sup>/g for micropore, mesopore, and macropore volumes, respectively. This commercially manufactured carbon is widely used in chemical, food, pharmaceutical, and water applications to remove impurities. Sorption experiments used as-received activated carbon in granular form and activated carbon in powdered form. To obtain powdered activated carbon (PAC), as-received GAC was ground and sieved using a no. 325 mesh sieve to obtain <0.045 mm sized particles. The PAC was then rinsed with ultrapure water and oven-dried for 30 minutes at 100°C prior to use. A 1500 mg/L PAC carbon slurry was prepared in ultrapure water for use in experiments. Phosphoric acid and HPLC-grade acetonitrile was purchased from Fisher Scientific and used as received.

**Table 1.** Composition of clay minerals.

Composition	Hectorite	Bentonite
SiO <sub>2</sub>	46.66%	61.46%
Al <sub>2</sub> O <sub>3</sub>	0.86%	22.05%
TiO <sub>2</sub>	0.04%	0.09%
Fe <sub>2</sub> O <sub>3</sub>	0.32%	4.37%
MgO	22.05%	2.94%
CaO	1.18%	14.01%
Na <sub>2</sub> O	1.35%	1.47%
K <sub>2</sub> O	0.14%	0.20%
LOI: 550°C - 1000°C	5.15%	15.47%

## 2.2. Sorption Experiments

Constant sorbent batch sorption tests with clay minerals (in an as-received form and in a peroxide-treated form) utilized initial denatonium benzoate concentrations ranging from 100 mg/L to 1000 mg/L in 10 mM CaCl<sub>2</sub> and were completed in duplicate. All suspensions were prepared by suspending approximately 0.99 g clay in 30 mL aqueous denatonium benzoate (0.033 g/mL) in 50-mL polypropylene conical centrifuge tubes. Prior to each experiment, the clay and denatonium benzoate were mechanically dispersed in background electrolyte by vortexing for at least fifteen seconds. The sorbate and sorbent were mixed end-over-end for 24 hours at 25 revolutions per minute (RPM) in a temperature stabilized hybridization oven (Boekel Scientific, Feasterville, PA). The aqueous concentration of denatonium ions remaining in solution was determined via high-performance liquid chromatography (described below) after collecting 2-mL aliquots from tubes centrifuged for 30 min at 4150 g.

Constant adsorbent batch sorption tests with GAC were also conducted in triplicate to evaluate the sorption of denatonium ions. GAC doses in each sample were 0.033 g/mL (0.8 g/24 mL). A DB stock solution was prepared by mixing 0.2 grams of solid crushed DB into an Erlenmeyer flask filled with 100 mL of ultrapure, organic-free Millipore™ water. Denatonium benzoate doses of 100 to 1000 mg/L were tested. A control experiment was run for each test to ensure that DB losses from solution resulted exclusively via sorption to activated carbon. Amber vials containing the appropriate DB concentrations and activated carbon doses were placed onto a Boekel rotary shaker (Boekel Scientific, Feasterville, PA) for continuous mixing at 25°C at a speed of 60 rpm for 24 hours. Following mixing, sample aliquots were collected in syringes and filtered with Millipore Millex 0.22 μm syringe filters (Bedford, MA) to filter out any remaining activated carbon. All glassware used in experiments was thoroughly washed with Versa-Clean liquid concentrated soap and placed in a bath of 10% HCl for acid-washing after each use. The aqueous concentration of denatonium ions remaining in solution was determined via HPLC as was the case for sorption experiments which used clay mineral sorbents.

Constant adsorbate experiments were conducted with PAC doses ranging from 5 - 100 mg/L and initial DB concentrations of 5 or 70 mg/L. These tests were conducted for 3 hours or 24 hours, but the remaining experimental protocol (mixing speed and method, number of replicates, filtering, glassware cleaning, etc.) was the same as reported for the granular activated carbon experiments.

## 2.3. Chemical Analysis

Aqueous denatonium ion concentrations were determined via high-performance liquid chromatography (HPLC) (Perkin Elmer, Series 275 HRes UV/VIS Liquid Chromatography System with autosampler) equipped with a C-18 analytical column (LiChrospher, 100 × 4.6, 5 μm) and a photo-diode array detector. Denatonium ions were colorimetrically detected at 205 nm in an acetonitrile-phosphoric acid (prepared with Milli-Q water) mobile phase. The column flow rate was 1.0 mL·min<sup>-1</sup> with a sample analysis time of 25 minutes. The HPLC column was equilibrated prior to each measurement by maintaining a mixture composition of 60% phosphoric acid/40% acetonitrile for 5 min prior to sample injection. Following each 50 μL sample injection, the mixture composition was held at 60% phosphoric acid/40% acetonitrile for 8 minutes and then linearly ramped over 2 minutes to 100% acetonitrile and held for 10 minutes. The DB retention time was 6.2 min in standards but varied from 6.4 - 7.1 min in supernatant solutions suggesting that dissolved solutes from the clay mineral might be present. Sorbed concentrations were determined via mass balance. Control samples consisted of denatonium benzoate solutions in centrifuge tubes or amber vials absent solid sorbents; an analysis of a control sample containing the lowest concentration of denatonium benzoate did not show detectable changes in the initial concentration suggesting that sorption to centrifuge walls, glassware, or syringe filters was not a factor. Blank samples consisting of each solid sorbent dispersed in background electrolyte were run with each experiment, and analyses of supernatant solutions for each control did not show any detectable amount of denatonium ions.

## 2.4. Data Analysis

Frequently, the partitioning of a chemical between the aqueous and solid phases is constant such that linear sorption isotherms result, and many isotherm models exist for assessing sorption data. The Freundlich model is an empirical model used to evaluate sorption onto heterogeneous surfaces and multilayer sorption onto environmental surfaces while the Langmuir model is useful for evaluating monolayer sorption processes and estimating the monolayer sorption capacity. In this work, we attempted to model all sorption data with the Freun-

and Langmuir models. The linearized Freundlich model is

$$\log S = \log K_f + 1/n \log C \quad (1)$$

where  $S$  has units of  $\mu\text{g/g}$ ,  $K_f$  is the Freundlich constant ( $(\mu\text{g/L})^{1-1/n}$ ),  $1/n$  is a dimensionless value which provides information on sorption favorability, and  $C$  ( $\mu\text{g/l}$ ) which represents the aqueous equilibrium concentration of denatonium ions. Generally,  $1/n$  values less than unity are interpreted as reflecting a favorable sorption process while values in excess of 2 imply an unfavorable sorption process [22]. The linearized Langmuir model's four types are:

$$\text{Type 1: } \frac{C}{S} = \frac{1}{S_m} C + \frac{1}{K_L S_m} \quad (2)$$

$$\text{Type 2: } \frac{1}{S} = \left( \frac{1}{K_L S_m} \right) \frac{1}{C} + \frac{1}{S_m} \quad (3)$$

$$\text{Type 3: } S = S_m - \left( \frac{1}{K_L} \right) \frac{S}{C} \quad (4)$$

$$\text{Type 4: } \frac{S}{C} = K_L S_m - K_L S \quad (5)$$

where  $S$  is mass loading ( $\text{mg/g}$ ),  $C$  is the aqueous equilibrium concentration of denatonium benzoate ( $\text{mg/L}$ ),  $S_m$  is maximum adsorbate amount per unit weight of adsorbent when surface sites are saturated ( $\text{mg/g}$ ), and  $K_L$  is the Langmuir adsorption constant related to the affinity of binding sites ( $\text{L/g}$ ).

Isotherm parameters were calculated from linear trendline equations fit to experimental data and are reported when appropriate. Correlation coefficients ( $R^2$  values) for these linear trendlines were reported as necessary to describe an isotherm model's ability to accurately describe the experimental data. For the Langmuir isotherm, a dimensionless constant separation factor or equilibrium parameter,  $r$ , was also calculated using the following equation [23]:

$$r = \frac{1}{1 + K_L C_o} \quad (6)$$

where  $C_o$  is the initial concentration of denatonium benzoate ( $\text{mg/L}$ ). This factor numerically expresses the Langmuir isotherm shape or key features and thereby predicts the favorability of an adsorption process as described in Table 2.

Percent removal of denatonium ions from solution was also calculated to quantify activated carbon sorption of the cation. The percent removal was determined using the following equation:

$$\% \text{DBremoval} = \left[ (C_o - C) / C_o \right] \times 100\% \quad (7)$$

### 3. Results

Figure 2 shows the impact of powdered activated carbon dose on sorption from 70  $\text{mg/L}$  aqueous solutions of denatonium benzoate at pH 6.95 for a 3-hour contact time. Generally, the data indicate that denatonium ion removal from solution is proportional to the powdered activated carbon dose. After 3 hours of continuous contact

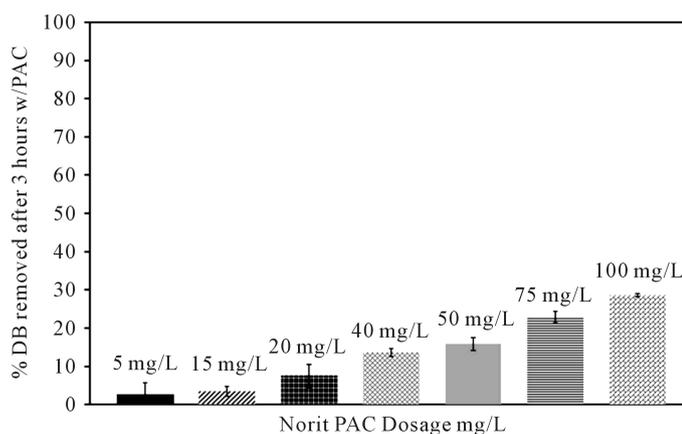
**Table 2.** Definition of  $r$ -factors used to characterize sorption processes based on isotherm shape.

$r$ -Value	Isotherm Type/Description of Adsorption Process
$r > 1$	Unfavorable
$r = 1$	Linear
$0 < r < 1$	Favorable
$r = 0$	Irreversible

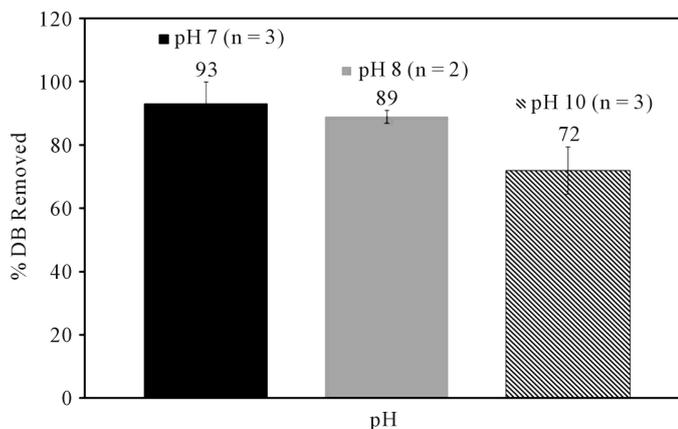
between PAC and aqueous denatonium benzoate, a 27% percent decrease in denatonium ion concentration was observed at the highest PAC dose utilized, while decreases of 16% and 23% occurred for 50 mg/L and 75 mg/L PAC doses, respectively. When granular activated carbon was used at a much higher dose (33,000 mg/L), 100% removal of denatonium ions was achieved when initial denatonium benzoate concentrations of 100 to 1000 mg/L were employed (data not shown). Given the variable aqueous conditions of treatment plant influents, the pH-dependent removal of denatonium ions via powdered activated carbon at 5 mg/L initial DB concentrations and 20 mg/L PAC dosages was investigated as shown in **Figure 3**. As pH increased from 7 to 10, the percent removal of denatonium ions decreased from 93% to 72%, respectively.

The Freundlich and Langmuir (Types 1 - 4) isotherm models were used to model denatonium ion sorption to powdered activated carbon from 70 mg/L denatonium benzoate solutions at 25°C at a 3-hour contact time. **Figure 4** shows the linearized Langmuir Type 3 model's fit with the experimental data. The Type 3 Langmuir model yielded the best fit ( $R^2 = 0.9976$ ) of the isotherm models evaluated based on a comparison of correlation coefficients ( $R^2$  values), and the calculated isotherm parameters were as follows:  $S_m$  equals mg/g and  $K_L$  equals 0.0153 L/g. For our data, the Langmuir separation factor  $r$ -value was 0.481 suggesting that the adsorption of denatonium benzoate onto the powdered activated carbon was favorable (see **Table 2**). The linearized Freundlich isotherm model yielded a weak correlation coefficient ( $R^2$ ) of 0.69, indicating that the data were not well-described by the Freundlich model.

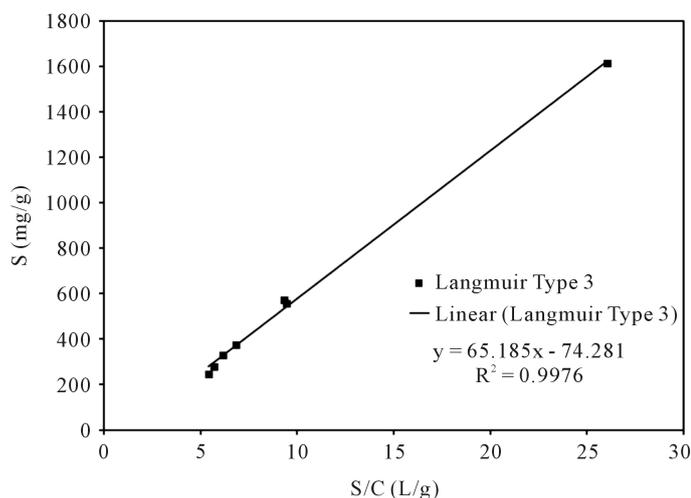
The removal of denatonium ions from a 200 mg/L solution of denatonium benzoate in 10 mM calcium chloride at pH 4 to 10 using as-received bentonite and hectorite resulted in 100% removal (data not shown). Organic



**Figure 2.** Norit PAC dosage impact on denatonium benzoate (DB) removal from ultrapure water ( $[DB]_{\text{initial}} = 70$  mg/L, pH 6.95).



**Figure 3.** Impact of pH on aqueous denatonium benzoate (DB) sorption using powdered activated carbon ( $[DB]_{\text{initial}} = 5$  mg/L,  $[PAC]_{\text{initial}} = 20$  mg/L).



**Figure 4.** Langmuir isotherm of denatonium sorption by powdered activated carbon at  $[DB]_{\text{initial}} = 70 \text{ mg/L}$  at  $\text{pH} = 6.95$  for at a 3-hour contact time.

matter, known to be present on mineral surfaces and dissolved within source waters, could enhance the ability of clay minerals to sequester dissolved aqueous denatonium ions. To eliminate organic matter facilitated sorption of denatonium ions, clay minerals treated with hydrogen peroxide were used as sorbents and are referred to as p-hectorite (peroxide treated hectorite) and p-bentonite (peroxide treated bentonite). The concentration dependent sorption of denatonium ions to p-hectorite and p-bentonite at various initial pH values are displayed in figure 5 with Freundlich model parameters listed in [Table 3](#).

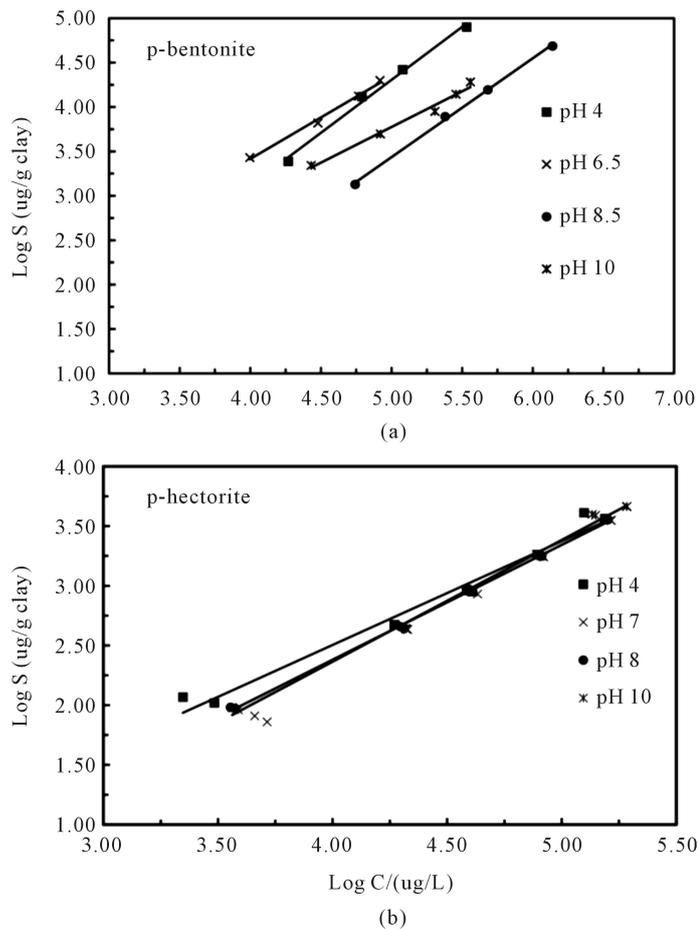
Regarding denatonium ion sorption to p-bentonite, initial hydrogen-ion concentrations largely impacted the sorptive behavior of denatonium ions based on a visual inspection of data shown in [Figure 5\(a\)](#); unfortunately, no clear trend is evident. The  $1/n$  values ranged from 1.194 to 0.803 while the Freundlich constant varied from 0.022  $\mu\text{g/L}$  to 0.575  $\mu\text{g/L}$  for denatonium sorption to p-bentonite. The sorption of denatonium by p-hectorite as a function of initial pH and denatonium benzoate concentration did not show a significant dependence on initial hydrogen ion concentrations as shown in [Figure 5\(b\)](#). However, at the lowest concentrations studied, there does appear to be a weak pH dependence of denatonium ion sorption to hectorite although a clear trend is not apparent. The  $1/n$  values ranged from 0.870 to 1.019 while the Freundlich constant varied from 0.870  $\mu\text{g/L}$  to 1.019  $\mu\text{g/L}$  for denatonium sorption to p-bentonite.

#### 4. Discussion

The increased removal of denatonium ions as a function of activated carbon dosage is consistent with many studies which evaluated the utility of activated carbon removal of contaminants [24] [25]. For example, in a study of the sorptive capacity of granular activated carbon using the cationic dye methylene blue, the authors observed that sorption was directly proportional to the dosage of size-fractionated carbon [24]. Additionally, methylene blue and several other dyes exhibit Langmuir sorption characteristics similar to our observed sorption of denatonium ions onto powdered activated carbon ([Figure 4](#)) [26]-[28]. Sorption of organic molecules or ions to high-surface area sorbents can be influenced by number of solution conditions including pH, ionic strength, temperature, nature of background electrolyte, etcetera [29] [30]. Furthermore, for large organic cations such as denatonium, electrostatic and dispersion interactions are possible given the inherent non-polar nature of the molecular ion and the ion charge [31]-[33]. Thus, one can reasonably conclude that electrostatic and dispersive interactions influenced activated carbon sorption of denatonium ions under the conditions studied since denatonium is a permanently charged cation and the activated carbon sorption sites tend to be negatively charged. Consequently, observed increases in denatonium ion sorption are likely related to an increased quantity of available cation sorption sites with increased powder activated carbon dose. The role of pore structure (size and volume) on denatonium ion adsorption is yet to be examined, but pore structure could play a key role in adsorption. The sorptive behavior of denatonium ions to peroxide treated clay minerals can be explained by considering the impact

**Table 3.** Freundlich sorption parameters for p-bentonite and p-hectorite clays.

p-bentonite				p-hectorite			
pH	$K_f$ (ug/L)	(1/n)	$R^2$	pH	$K_f$ (ug/L)	(1/n)	$R^2$
4	0.022	1.194	0.994	4	0.106	0.870	0.979
6.5	0.478	0.934	0.993	7	0.008	1.086	0.996
8.5	0.007	1.113	0.999	8	0.034	0.962	0.999
10	0.575	0.803	0.983	10	0.019	1.019	0.995

**Figure 5.** Freundlich sorption isotherms of denatonium sorption by (a) p-bentonite and (b) p-hectorite in 10 mM CaCl<sub>2</sub> at the indicated pH values.

of peroxide on the clay composition. Peroxide treatment is a well-established method for reducing the organic matter content [34]. Loss-on-ignition data listed in **Table 1** for the bentonite and hectorite clay minerals indicate 5.47% and 15.15%, respectively, by mass of organic carbon. In our study, the observed decrease in sorption affinity of denatonium ions for bentonite and hectorite clay minerals following peroxide treatment suggests that clay mineral sorption sites and sorption domains within organic matter contribute to the attenuation of denatonium ions. Enhanced sorption of large molecular weight organic cations to mineral surfaces is not unique to our study and has been attributed to enhanced van der Waals interactions between the solute and organic matter in some cases and electrostatic interactions between charged solutes and deprotonated acids in organic matter which can vary with pH [35].

The observation of pH-independent sorption of denatonium ions to p-hectorite coupled with pH-dependent sorption to p-bentonite suggests that sorption does not exclusively occur at the same type of sorption site. Smectite minerals are known to possess internal pH-independent cation sorption sites and external pH-dependent sorption sites. Thus, we conclude that internal cation-exchange sorption sites within p-hectorite are most likely responsible for reduced aqueous denatonium ion concentrations although sorption to external, pH-independent sites cannot be ruled out. Others have observed quaternary ammonium cation sorption to hectorite and concluded that the process likely involved interlayer exchanges of cations [36]. In the case of p-bentonite, a plausible explanation for the observed pH-dependent sorption is that both internal and external sorption sites contributed to the attenuation of denatonium ions. Given the interlayer accessibility of denatonium ions, sorption to internal sites is expected, however the relative number of external sites versus internal sites could favor pH-dependent sorption resulting in hydrogen-ion concentration dependent isotherms. Since  $1/n$  values generally fell between 0.803 and 1.194, the sorption of denatonium ions to each clay mineral could be classified as moderate to highly favorable under the conditions of our study [22].

## 5. Conclusion

The removal of aqueous denatonium ions from water by common solid sorbents used in water treatment processes was studied. Generally, denatonium ion concentrations are significantly decreased when smectite or activated carbon sorbents are utilized. In the case of smectite removal, multiple sorption sites were inferred whereby sorption occurred to pH-independent sorption sites within the interlayer and pH-dependent sites on the external clay surface. Peroxide treated clay minerals showed decreased sorption relative to as-received minerals from which we conclude that oxidized organic matter facilitated denatonium ion sorption. The efficiency of activated carbon removal of denatonium ions increases with decreasing pH and increasing activated carbon dosage which suggest that electrostatics and increased sorption site availability impact denatonium ion sorption. Thus, the use of clay mineral sorbents or activated carbon for the removal of denatonium ions from solution by treatment facilities is expected to be feasible and of great utility under various environmental conditions.

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