

Performance Investigation of Membrane Electrode Assemblies for High Temperature Proton Exchange Membrane Fuel Cell

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Received August 2013

ABSTRACT

Different types of ABPBI (poly(2,5-benzimidazole)) membranes and polymer binders were evaluated to investigate the performance of MEAs for high temperature proton exchange membrane fuel cell (HT-PEMFC). The properties of the prepared MEAs were evaluated and analyzed by polarization curve, electrochemistry impedance spectroscopy (EIS), cyclic voltammetry (CV) and durability test. The results showed that MEA with modified ABPBI membrane (AM) has satisfactory performance and durability for fuel cell application. Compare to conventional PBI or Nafion binders, polytetrafluoroethylene (PTFE) and polyvinylidene difluoride (PVDF) are more attractive as binders in the catalyst layer (CL) of gas diffusion electrode (GDE) for HT-PEMFC.

Keywords: High Temperature Proton Exchange Membrane Fuel Cell; ABPBI (Poly(2,5-Benzimidazole)); Polymer Binders; Gas Diffusion Electrode; Membrane Electrode Assembly

1. Introduction

Polybenzimidazole (PBI) based high temperature proton exchange membrane fuel cells (HT-PEMFCs) have attracted more and more attention in these years due to their advantages over low temperature PEMFCs based on perfluorosulphonic acid polymer electrolytes (e.g. Nafion) [1]. However, the sluggish kinetics of the oxygen reduction reaction (ORR) [2] and the transport limitations of protons and reactants in cathode, especially in the presence of phosphoric acid (PA), limit the cell performance of the high temperature PEMFC. Therefore, enhancing the cell performance is one of the most important issues for high temperature PEMFC being more widely considered as an alternative to the low temperature PEMFC systems.

Membrane electrode assembly (MEA) is the most important component in high temperature PEMFC system and it plays a major role in determining cell performance. It consists of a proton exchange membrane sandwiched between two gas diffusion electrodes (GDEs), which possess a porous structure that allows easy transport of reactant gases and water to and from the catalytically active zone. Therefore, the components of MEA, *i.e.* electrolyte membrane and GDEs, have significant influence on the performance of high temperature PEMFC.

Several types of electrolyte membrane, such as poly [2,2'-(*m*-phenylene)-5,5'-bibenzimidazole (PBI or *m*PBI), poly(2,5-benzimidazole) (ABPBI) and their derivatives [3]; several types of polymer ionomers, such as PBI [4-6], polytetrafluoroethylene (PTFE) [7,8], polyvinylidene difluoride (PVDF) [9,10] and Nafion [11,12], can be used in the MEA as proton exchange membranes and catalyst layer (CL) binders for high temperature PEMFC. The physicochemical properties of these membranes and binders are different from each other; consequently the resultant MEAs have their own advantages and shortcomings. In this work, the MEAs with four types of ABPBI membranes and four types of binders were investigated to evaluate their fuel cell performances. The properties of these MEA were characterized by single cell polarization, electrochemical analysis and durability test.

2. Experimental

2.1. Preparation of GDEs

Hispec 4000 Pt/C catalyst (40 wt% Pt, Johnson Matthey) was used in this study. All GDEs were prepared by our newly developed spraying method [13]. The catalyst powders were deposited onto the microporous layer of commercially available GDL (Freudenberg, Germany). The

catalyst loadings were calculated by weighing the GDEs before applying the catalyst inks, and then after application and oven drying for overnight. The platinum loadings of all GDEs (both anode and cathode) used for this study are $0.5 \text{ mg}\cdot\text{cm}^{-2}$, unless otherwise stated.

2.2. Preparation of MEAs

The membranes used in this study are ABPBI (denoted as AM, AM 55, AMcl and APcl), which were supplied by FuMA-Tech. The properties (thickness, composition, process method etc.) of these membranes vary from each other. For doping with PA, the membranes were immersed in 85% acid solution for certain time at 100°C until their acid doping level of about 3.7 molecules of H_3PO_4 per polymer repeating unit (PRU) were obtained. Before being used, the membrane was taken from the PA bath, and the superficial acid onto the membrane was thoroughly wiped off with lab tissue. Together with gaskets made of fluorinated polymer, the MEA was assembled by sandwiching the doped membrane between two GDEs impregnated with PA in a single cell fixture (BalticFuel-Cells GmbH, Germany) without a preceding hot-pressing step. The active areas of all MEAs are 5 cm^2 , unless otherwise stated.

2.3. Single Cell Test and Electrochemical Characterization

The cells were operated at 150°C (unless otherwise stated) and $2 \text{ N}\cdot\text{mm}^{-2}$ piston pressure in a FuelCon Evaluator C test station (FuelCon, Germany). Pure hydrogen was fed to the anode and air to the cathode respectively, with flow rates (unless otherwise stated) of $100 \text{ ml}\cdot\text{min}^{-1}$ (hydrogen) and $250 \text{ ml}\cdot\text{min}^{-1}$ (air), at ambient pressure. Both hydrogen and air were used as dry gases, directly from the compressed bottles without external humidification.

Electrochemical impedance spectroscopy (EIS) and cyclic voltammetry (CV) were performed using an Autolab PGSTAT 30 Potentiostat/Galvanostat (Metrohm). EIS measurements were carried out at a cell voltage of 0.6 V with amplitude of 5 mV, and in the frequency range of 100 mHz to 20 kHz. The impedance data were obtained by calculation and simulation with Autolab Nova software. Voltammetric measurements, undertaken to study the electrochemical active surface area (EASA), were conducted using dry N_2 at the cathode (working electrode) and dry H_2 at the anode (counter electrode and reference electrode) at cell working temperature. Cyclic voltammograms were recorded from 1.2 V to 0.05 V at a scan rate of $0.05 \text{ V}\cdot\text{s}^{-1}$.

3. Results and Discussion

3.1. Effect of Electrolyte Membrane

Figure 1 shows the FT-IR spectra of the four ABPBI

membranes. The characteristic peaks of PBI are shown in all membranes; hydrogen-bonded N-H stretching at 3184 cm^{-1} , the free non-hydrogen-bonded N-H stretching at 3415 cm^{-1} , and C=C and C=N stretching bands for benzimidazole group at 1612, 1590, and 1443 cm^{-1} can be observed from all the membrane samples [14], indicating that no significant structural changes took place in the bulk polymer of the above ABPBI membranes. However, it should be noted that sharp new peaks in the zone ranged between 800 and 1300 cm^{-1} are observed for the three modified ABPBI membranes (AM, AM 55 and AMcl), which are indicative of interactions between the additives and ABPBI groups [11]. Interactions between the additives and ABPBI groups should increase the stability of the modified ABPBI membranes, as demonstrated in the durability test later in current study (see Section 3.3).

The single cell performance comparison of the MEA with AM, AM 55, AMcl and APcl are shown in Figure 2. It should be noted that the polymer binders used in these MEAs are PVDF (the binder content in CL has been pre-optimized for this comparison), which is the preferred

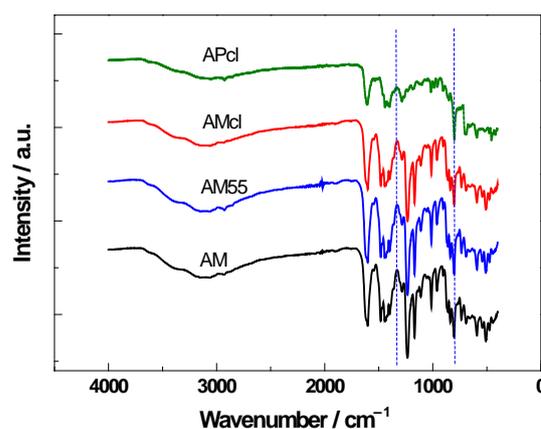


Figure 1. FT-IR spectra of various ABPBI membranes.

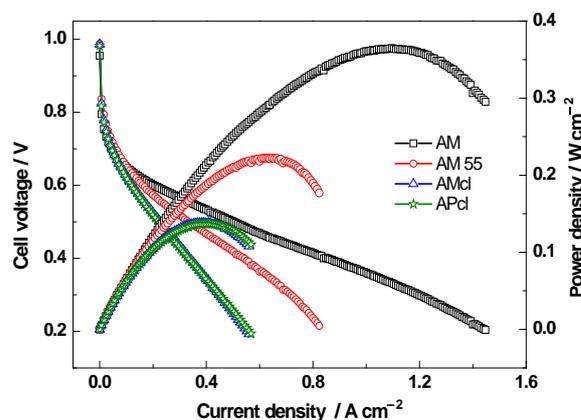


Figure 2. Polarization curves and power density curves of HTPEMFC with MEAs using different ABPBI membranes, operated at atmosphere pressure and 150°C .

binder pre-optimized in Section 3.2. It is clear that the MEA prepared by AM membrane yields much better performance than the MEA prepared with other membranes. At a working voltage of 0.6 V, the current density of the MEA with AM membrane reaches $0.208 \text{ A}\cdot\text{cm}^{-2}$, the maximum power density of the MEA with GDE-1 can reach $0.363 \text{ W}\cdot\text{cm}^{-2}$ at 0.313 V.

To understand the performance difference of the MEAs prepared with these membranes, overall analysis on the polarization curves and electrochemical measurements on both GDEs are performed. Generally, the polarization curve of a PEMFC could be divided into three segments (corresponding to different electrochemical processes) according to its different voltage drop rates. The initial drop of the curve at a very low current density is due to the sluggish kinetics of oxygen reduction at the cathode, determined by the nature of the electrodes. It can be seen in **Figure 2** that all MEAs show similar voltage drop in this region ($0 - 0.1 \text{ A cm}^{-2}$), which is reasonable considering the GDEs used in all MEAs were the same. To prove this point, CV measurements are performed to study the EASAs of the all MEAs, as shown in **Figure 3** (only H_2 desorption peaks were showed for clarity).

The corresponding EASAs were calculated from the H_2 desorption peak of each voltammogram and the results are summarized in **Table 1**. It can be seen that the EASA values of all MEAs are between 17.4 and $22.3 \text{ m}^2\cdot\text{g}^{-1}$, only small differences were detected among these MEAs. The EASA results are certainly consistent with their performances (activation polarization zone) presented in **Figure 2**.

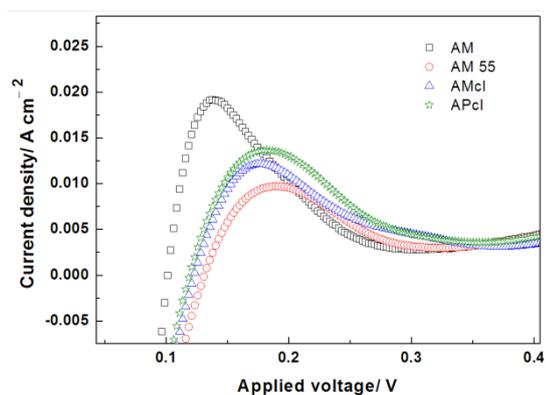


Figure 3. Cyclic voltammograms of the MEAs with different ABPBI membranes.

Table 1. Electrochemical characteristics of the MEAs with different ABPBI membranes.

MEA type	R_{Ω} (Ω)	R_{ct} (Ω)	EASA ($\text{m}^2\cdot\text{g}^{-1}$)
AM	0.05	0.07	22.3
AM 55	0.12	0.13	17.4
AMcl	0.13	0.11	18.7
APcl	0.12	0.13	19.1

The subsequent drop in the polarization curve is ascribed to ohmic loss, which originates from ionic flow through the electrolyte membrane, and from electron flow through the electrode layers, current collectors and flow field plates. As shown in **Figure 2**, the four MEAs present much different decreasing slopes in the linear region, implying that they had different ohmic resistances. To verify the resistances of the single cells with these MEAs, *in situ* impedance measurements are performed at the cell voltage of 0.6 V, as shown in **Figure 4**. Only one semicircular loop can be observed in the Nyquist plot as the electrode process is dominated by ORR [15]. Through simulation with a simple RC equivalent circuit, their corresponding cell resistances (R_{Ω}) and charge transfer resistances (R_{ct}) can be calculated, which are also presented in **Table 1**. It can be seen that there is no significant difference in cell ohmic resistance for the MEAs with the last three membranes (*i.e.* AM 55, AMcl and APcl), which is consistent with the similar decreasing slopes in the linear regions of their polarization curves presented in **Figure 2**. However the ohmic resistance of the MEA with AM membrane is only 0.05Ω , less than half of the other membranes. It suggests that AM membrane has higher proton conductivity than other membranes because all other parts in the test fixture during testing were identical. Moreover, the charge transfer resistance of the AM MEA is much smaller than those of the other three, which suggests that AM MEA yielded a more efficient electrochemical active layer, which means that the interactions between the GDE and the AM membrane are more efficient.

The last voltage drop at high current density is due to mass transport limitations occurring in the electrodes and the membrane. However, from **Figure 2** it can be seen that, for all MEAs, the voltage drop rates in low cell voltage region ($<0.4 \text{ V}$) of their polarization curves are almost the same with that in their linear regions, which means that no obvious mass transfer limitations in these

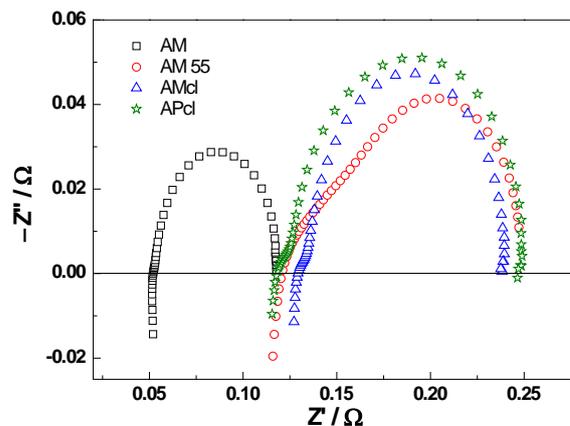


Figure 4. *In situ* impedance curves of the MEAs with different membranes, at a cell voltage of 0.6 V.

MEAs even at the high current densities. It is understandable when considering the high operating temperature (150°C, only water vapor existed in the GDEs) and the high stoichiometries reactants (ca. 7/7.5 for H₂/Air at 0.4 A·cm⁻²) due to the small active area (~5 cm²) even at low gases flow rates [16].

From all these analyses and electrochemical results, it can be concluded that the good performance of the MEA with AM membrane is primarily attributable to the higher proton conductivity and good interactivity of the electrolyte membrane and the GDEs, which makes a more efficient electrochemical active layer, accordingly the minor ohmic resistance and charge transfer resistance. For these reasons, the AM membrane, which performed best at usual working voltage of 0.6 V and maximum power density, is chose for all subsequent studies.

3.2. Effect of Polymer Binders

Figure 5 shows the performance comparison of the MEA with the four different polymer binders. It is clear that the PTFE and PVDF GDEs yield much better performance than the GDEs prepared with PBI and Nafion binders in all regions of the polarization curve. At a working voltage of 0.6 V, the current density of the MEA with PVDF GDEs reaches 0.53 A·cm⁻², 121% higher than that (0.24 A·cm⁻²) of the MEA with PBI GDEs. The maximum power density of the MEA with PTFE GDEs can reach 0.61 W·cm⁻² at 0.35 V. These values are almost the best results yet reported for similar PA-doped PBI fuel cell and operated using air, which are comparable to the performances of the commercial MEAs with high Pt loadings [17]. This is mainly attributable to the properties of PTFE and PVDF binders that they exist in the CLs as a fiber phase, which makes catalyst particles less likely to be encapsulated in the binder, then making more Pt surface available in the CLs. On the contrary, PBI and Nafion polymer ionomers are easily covered on the surface of the catalyst particles, which could impose mass transport limitation in CLs due to the low gas permeability of these films formed on the catalyst sites [15].

3.3. Durability

From above physical characterizations and electrochemical analysis results, it can be concluded that good MEA performance (at usual working voltage of 0.6 V) can be delivered by using PVDF as CL binder and AM membrane as electrolyte membrane. However, the stability or durability of this MEA is also a major concern for the real application and commercialization of HT-PEMFC. The remarkable long term stability of PA-doped PBI MEA is achieved in some research groups' works [9, 10, 18, 19]. To verify the stabilities of the MEA with AM membrane and PVDF binder, a short term durability test was performed, as shown in **Figure 6**.

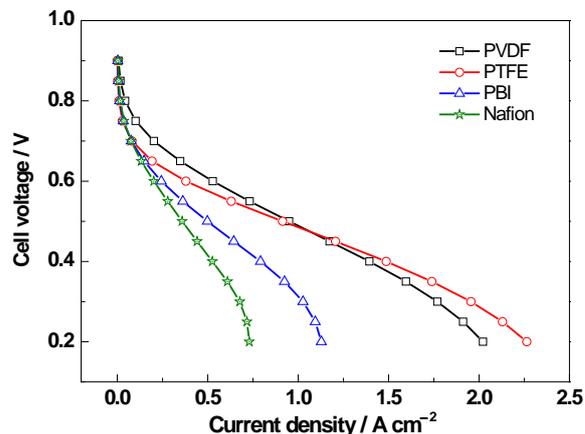


Figure 5. Polarization curves of PA-doped ABPBI (AM membrane) fuel cell using GDEs prepared with different polymer binders, operated with flow rates of 200 ml·min⁻¹ (hydrogen) and 1000 ml·min⁻¹ (air), at ambient pressure and 160°C.

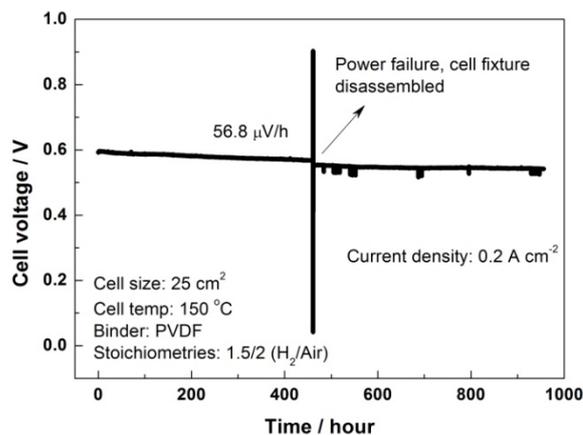


Figure 6. The durability test of the MEA with AM membrane and PVDF binder.

It can be seen that the cell voltage of the MEA remains at ~0.56 V without obvious drop after almost 1000 h operation. The degradation rate calculated by linear fitting of cell voltage data points after the MEA activation is about 56.8 μV·h⁻¹, which are acceptable for most applications [12, 18-20].

4. Conclusion

Four types of ABPBI membranes and four types of polymer binders were evaluated to investigate the performance of MEAs for high temperature proton exchange membrane fuel cell (HT-PEMFC). The results showed that MEA with modified ABPBI membrane (AM) as electrolyte and PVDF as CL binder has satisfactory performance and durability for fuel cell application. At usual working voltage of 0.6 V and cell temperature of 150°C, the peak power density reached 0.363 W·cm⁻², and the current density at usual working voltage 0.6 V was up to

0.208 A·cm⁻², which are comparable to the results yet reported for similar MEAs with Pt loading of ~0.5 mg·cm⁻². The MEA showed good durability in a short term operation: the cell voltage remained at ~0.56 V without obvious drop after almost 1000 h operation.

5. Acknowledgements

This work is supported by Hydrogen and Fuel Cell Technologies RDI Programme (HySA), funded by the Department of Science and Technology in South Africa (project KP1-S01).

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