

# Physical Properties of Lithium Containing Cadmium Phosphate Glasses

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Received May 1, 2010; revised August 15, 2010; accepted May 14, 2010

## Abstract

Lithium cadmium phosphate glasses were prepared by melt quench technique. These glasses contain a mole % composition of x% Li<sub>2</sub>O-(50-x)% CdO-50%P<sub>2</sub>O<sub>5</sub>. The quantity x varies from 0-40 mole%. The physical properties reported in this paper are mass density  $\rho$ , modulus of rigidity  $\eta$ , coefficient of linear expansion  $\alpha$ , transition temperature  $T_g$ , Softening temperature  $T_s$ , Oxygen packing density, Molar volume and lithium ion concentration. The mass density, oxygen packing density, modulus rigidity, transition temperature and softening temperature show decreasing trend with increasing concentration of lithium ions in these glasses, where molar volume and coefficient of linear expansion increases with increasing concentration of Li<sub>2</sub>O.

**Keywords:** Phosphate Glasse, Mass Density, Modulus of Rigidity, Coefficient of Linear Expansion, Transition Temperature, Molar Volume

## 1. Introduction

The competition of speed, cost and reliability of electronic devices in various applications has led to the research for new materials which can meet the specific requirement. The phosphate glasses are among those materials which have a number of technical and biological applications [1-3]. Phosphate glasses have a variety of technological application due to several unique properties [4]. To investigate the nature of these glasses for their appropriate applications, study of optical, electrical and physical properties are necessary. Most of the researchers studied optical and electrical properties to characterize the phosphate glasses [5-8], whereas the knowledge of physical properties is equally very helpful to study the nature of materials.

A general observation of the thermal expansion curve tells whether the materials under investigation are amorphous or crystalline in nature. Experimental results on physical properties in glassy material have been reported in literature [9-11].

In the present work physical properties such as mass density, oxygen packing density, molar volume, modulus of rigidity, transition temperature, softening temperature and coefficient of linear expansion were studied to examine the effect of alkali metal oxide *i.e.* Li<sub>2</sub>O on the

cadmium phosphate glasses. According to our survey, no data has been reported on the physical properties of cadmium phosphate glass system containing lithium oxide.

## 2. Experimental

Chemicals Li<sub>2</sub>CO<sub>3</sub>, CdO and P<sub>2</sub>O<sub>5</sub> of purity (99.99%) were used to prepare Lithium-cadmium-phosphate glasses. The glass samples were prepared by using 15 g ingredients mixture of composition x%Li<sub>2</sub>O-(50-x) %CdO-50%P<sub>2</sub>O<sub>5</sub> in a platinum crucible. The details of preparation of samples and method of density measurements and modulus of rigidity have already been described elsewhere [12,13]. Oxygen packing density, molar volume and lithium ion concentration were estimated by using following equations respectively,

$$\text{Oxygen packing density} = \{1000 \times \rho \times (O)\} / M \quad (1)$$

$$\text{Molar volume} = M / \rho \quad (2)$$

$$\text{Ion concentration} = \{\rho N_{\text{avo}} P\} / M \quad (3)$$

where  $\rho$  = mass density

M = Molecular weight of glass composition

O = number of oxygen atoms in the composition

$N_{\text{avo}}$  = Avogadro number

and P = nx

where 'x' is the mole fraction in glass composition and

'n' is the number of atoms of element ions in a given oxide, *i.e.*  $n = 1$  for oxides like CdO, ZnO etc and  $n = 2$  for oxides like  $\text{Li}_2\text{O}$ ,  $\text{Na}_2\text{O}$ , etc.

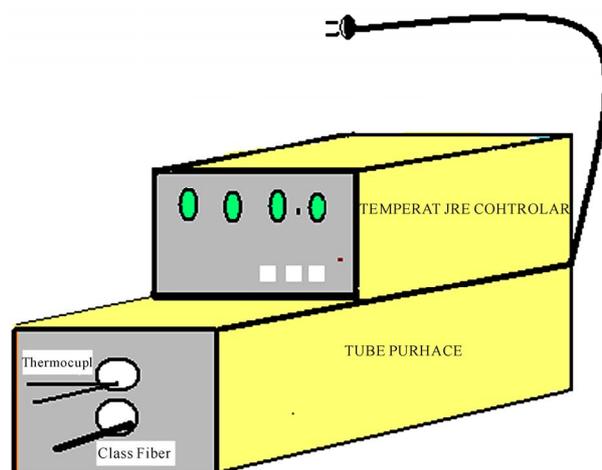
Thermal expansion and coefficient of linear expansion of fibers of lithium phosphate glasses have been studied up to  $600^\circ\text{C}$ . These parameters were measured by using a horizontal tube heating system as shown in **Figure 1**. The heating system consists of two equal bore Pyrex tubes each of 10 cm length. The tubes are fitted together in a single heating unit so that both the tubes can have the same temperature. A Cromel-Alumel thermocouple connected to a Fenwal controller is fitted in one tube and the glass fiber 'F' of length  $L_0 = 10.5$  cm is placed in the other tube. The temperature of the system is controlled by a temperature controller while the heating rate is controlled through a voltage regulator. Each temperature was maintained for five minutes to achieve the equilibrium.

The increase in the length of the fiber for every  $50^\circ\text{C}$  rise in temperature (with constant rate of heating) was measured with a traveling microscope whose least count is  $10\ \mu\text{m}$ . The linear expansion was plotted against temperature as shown in **Figure 2**. This curve was used to estimate transition temperature  $T_g$  and softening temperature  $T_s$ . The coefficient of linear expansion ' $\alpha$ ' is then calculated using equation

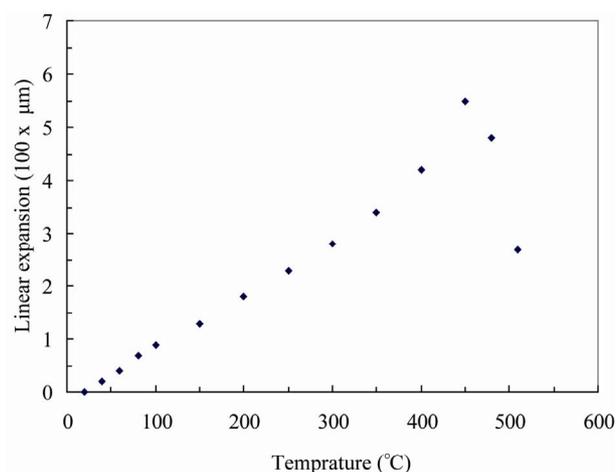
$$\alpha = (\Delta L)/(L_0 \Delta t) \text{ } (^\circ\text{C})^{-1} \quad (4)$$

### 3. Results and Discussion

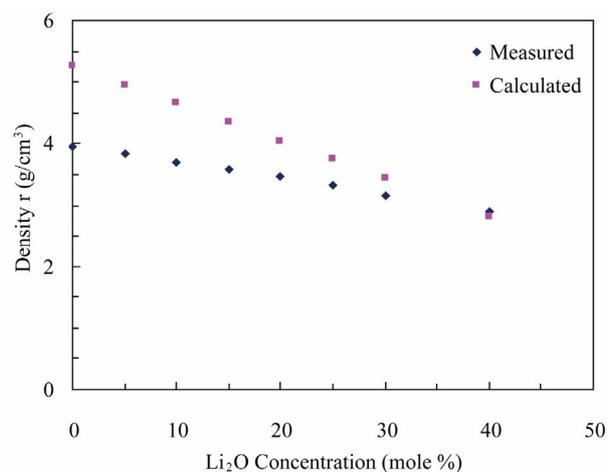
The measured results and estimated values of mass density are listed in **Table 1**. These results indicate a decreasing trend in the density of  $x\% \text{Li}_2\text{O}-(50-x)\% \text{CdO}-50\% \text{P}_2\text{O}_5$  glass system with increasing concentration of  $\text{Li}_2\text{O}$ . Theoretical values of the density were estimated by using the relation  $\rho = \sum \rho_i x_i$ , where  $\rho_i$  and  $x_i$  are the density and fraction of the free oxides respectively. The estimated and the measured values of density of these glasses are depicted in **Figure 3** as a function of  $\text{Li}_2\text{O}$  concentration. It can be seen that estimated values are higher than measured values. This difference in values of density may be due to the variation in atomic arrangement between the structure of glass and molecules of the free oxides. The decrease in the measured density with the increasing concentration of  $\text{Li}_2\text{O}$  agrees qualitatively with that predicted by the composition relation, and it may be due to the replacement of high density CdO ( $8.15\ \text{g/cm}^3$ ) as compared with  $\text{Li}_2\text{O}$  ( $2.02\ \text{g/cm}^3$ ). The estimated oxygen packing density and molar volume and lithium ion concentration are listed in **Table 1** and are shown in **Figures 4, 5, and 6** respectively. These results show that oxygen packing density decreases where as molar volume and lithium ion concentration increase



**Figure 1.** Tube furnace along with temperature controller for heating the fiber to measure its linear expansion with the variation temperature.



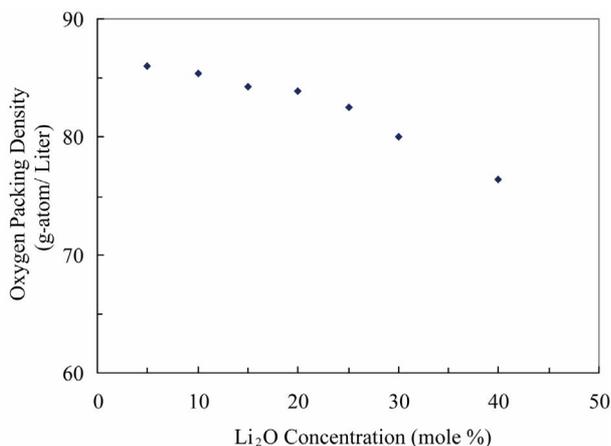
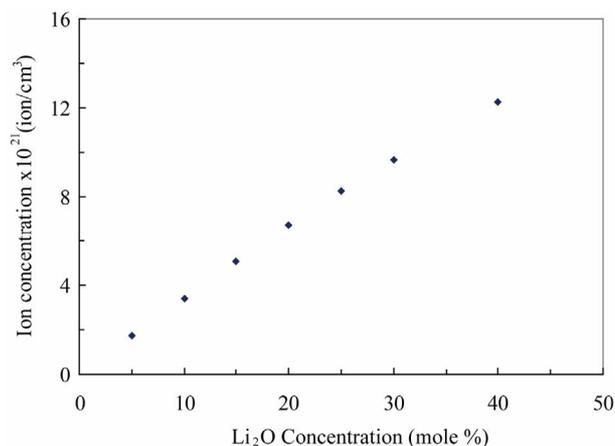
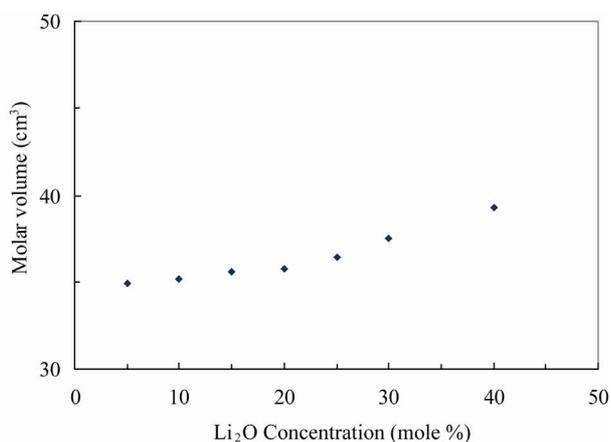
**Figure 2.** A plot of linear expansion  $\Delta l$  versus temperature  $^\circ\text{C}$ .



**Figure 3.** Variation of mass density relative to lithium oxide concentration (mole%).

**Table 1. Variation of mass density, oxygen packing density, molar volume and lithium ion concentration with respect to glass composition in mole%.**

Composition Mole % Li <sub>2</sub> O-CdO-P <sub>2</sub> O <sub>5</sub>	Density $\rho$ (g/cm <sup>3</sup> )		Oxygen Paking Density (g-atom/liter)	Molar Volume V <sub>m</sub> (cm <sup>3</sup> )	Li <sub>2</sub> O ion Concentration (ion/cm <sup>3</sup> ) $\times 10^{21}$
	Calculated	Measured			
0%-50%-50%	5.27	3.95	87.57	34.26	00.0
05%-45%-50%	4.96	3.83	86.05	34.87	1.73
10%-40%-50%	4.66	3.69	85.40	35.13	3.43
15%-35%-50%	4.35	3.57	84.33	35.58	5.08
20%-30%-50%	4.04	3.48	83.92	35.75	6.74
25%-25%-50%	3.74	3.34	82.50	36.36	8.28
30%-20%-50%	3.43	3.17	80.06	37.47	9.64
40%-10%-50%	2.82	2.89	76.42	39.26	12.27

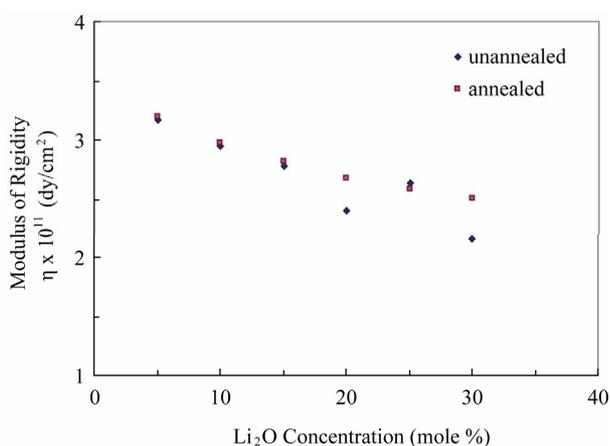
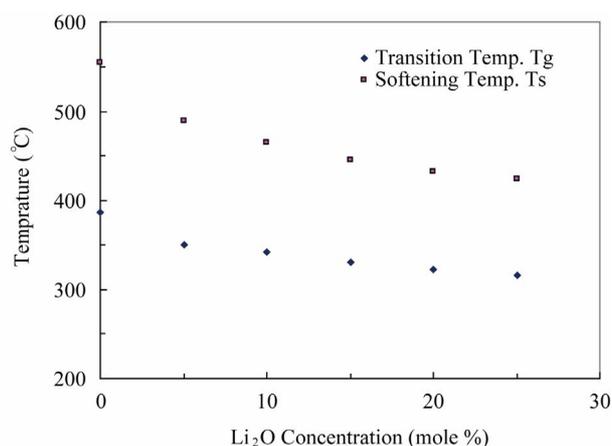
**Figure 4. A plot of oxygen packing density as a function of lithium oxide content.****Figure 6. A graph between ion concentration of lithium and Lithium oxide content.****Figure 5. Variation of molar volume with lithium oxide concentration.**

with increasing concentration of the lithium oxide. The replacement of an intermediate/modifier CdO with Li<sub>2</sub>O which is a modifier only, develops more non-bridging oxygen than bridging oxygen in the glass network [15,16]. The development of non-bridging oxygen may flayer-up the glass system and thus increases its molar volume. The increase in molar volume may cause decrease in oxygen packing density and mass density.

The results of modulus of rigidity ' $\eta$ ' of the lithium cadmium phosphate glasses are listed in **Table 2** and are depicted graphically in **Figure 7** as a function of Lithium Oxide concentration. The results show that ' $\eta$ ' decreases with increasing amount of Li<sub>2</sub>O. The decrease in modulus of rigidity may be due to decrease in oxygen packing density [10,12], which is a consequence of increasing number of non-bridging oxygen due to which

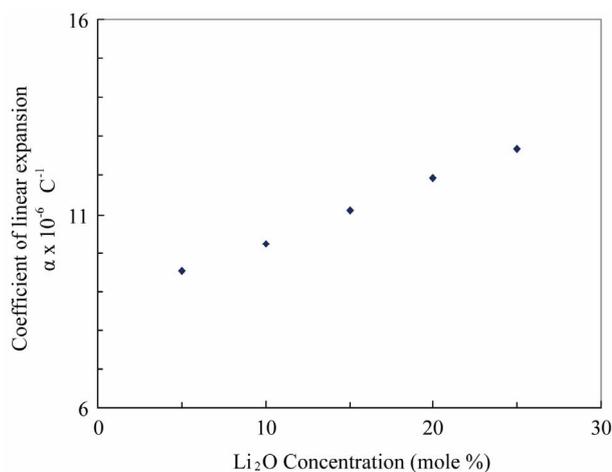
**Table 2. Variation of Modulus of rigidity, transition temperature softening temperature and coefficient of linear thermal expansion with respect to glass composition in mole%.**

Composition Mole % Li <sub>2</sub> O-CdO-P <sub>2</sub> O <sub>5</sub>	Modulus Rigidity $\eta$ (dy/cm <sup>2</sup> ) $\times 10^{11}$		Transition Temperature T <sub>g</sub> (°C)	Softening Temperature T <sub>s</sub> (°C)	Coefficient Of linear Expansion $\alpha$ (°C) <sup>-1</sup>
	Unannealed samples	Annealed samples			
0%-50%-50%	3.29	3.42	387	555	8.99
05%-45%-50%	3.17	3.19	350	490	9.52
10%-40%-50%	2.95	2.97	342	465	10.24
15%-35%-50%	2.78	2.82	330	445	11.11
20%-30%-50%	2.40	2.68	322	433	11.93
25%-25%-50%	2.63	2.57	315	425	12.70
30%-20%-50%	2.16	2.51	----	----	----

**Figure 7. Dependence of modulus of rigidity on lithium oxide concentration.****Figure 8. Graph illustrating the variation of Transition and Softening temperatures with the variation of lithium oxide concentration.**

the glass structure has further expanded.

The linear thermal expansion of these glasses is depicted in a representative **Figure 2**. The estimated values of transition temperature T<sub>g</sub>, softening temperature T<sub>s</sub> and calculated values of coefficient of linear thermal expansion are listed in **Table 2** and are depicted in **Figures 8** and **9** respectively. T<sub>g</sub> and T<sub>s</sub> show a decreasing trend while coefficient of linear expansion increases with increasing concentration of Li<sub>2</sub>O. In ternary lithium cadmium phosphate glasses a high coordination number modifier/intermediate is being replaced by a low coordination number modifier Li<sub>2</sub>O [2,14]. The intermediate oxide CdO probably becomes the part of structure along with the glass former P<sub>2</sub>O<sub>5</sub>. The modifiers do not contribute to the network formation but essentially settle as interstices in the glass structure. The addition of alkali metal oxide cleaves the structure and disturbs the bonding between glass forming cations and oxygen anions.

**Figure 9. Variation of coefficient of linear expansion with respect to Li<sub>2</sub>O concentration.**

This increase, the number of non-bridging oxygens and thus develop a more open structure. Consequently the structure expands the molar volume which causes a decrease in the oxygen packing density and hence a decrease in the density of the glass sample. The decrease in oxygen packing density along with the decrease of density and increase of molar volume of these glasses make them less mechanical resistive which may have caused an increase in the coefficient of linear expansion and decrease in the transition temperature and the softening temperature. The results reported here are similar to those reported by other research workers [13,15-17].

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