

The Use of Sodium Silicate as a Corrosion Inhibitor in a Saline Drilling Fluid: A Nonaggressive Option to the Environment

Fernando B. Mainier*, Anne A. M. Figueiredo, Alan Eduardo R. de Freitas, André Armando M. de Alencar Junior

Escola de Engenharia, Universidade Federal Fluminense, Niterói, Brazil

Email: *fmainier@uol.com.br

How to cite this paper: Mainier, F.B., Figueiredo, A.A.M., de Freitas, A.E.R. and de Alencar Junior, A.A.M. (2016) The Use of Sodium Silicate as a Corrosion Inhibitor in a Saline Drilling Fluid: A Nonaggressive Option to the Environment. *Journal of Environmental Protection*, 7, 2025-2035.

<http://dx.doi.org/10.4236/jep.2016.713157>

Received: November 8, 2016

Accepted: December 13, 2016

Published: December 16, 2016

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Abstract

The function of a corrosion inhibitor in drilling mud compositions is the corrosion protection of the equipment involved in drilling operations. Many compositions involve environmentally several products such as fatty amines of high molecular weight, polyoxylated amines, amides, imidazolines, nitrogen heterocyclic products, etc. The potential advantages of the use of silicates are the effective protection of carbon steel, especially in aerated saline fluids, low costs and non-aggressive behavior to environment. Gravimetric and electrochemical tests were carried out using an aerated solution of 3.5% NaCl and the addition of sodium silicate ($\text{Na}_2\text{SiO}_3 \cdot 9\text{H}_2\text{O}$) as a corrosion inhibitor at concentrations of 250 to 2000 mg/L. The efficiencies of the corrosion protection of carbon steel using silicate concentrations greater than 1250 mg/L were greater than 92%.

Keywords

Sodium Silicate, Corrosion Inhibitor, Environment, Aerated Saline Fluid, Carbon Steel

1. Introduction

The direct and indirect costs of a failure or fracture in oil well drilling are relatively high, especially when they are associated with the time lost during the recovery process, the placement of a new drill and/or the loss of the oil well [1] [2].

To minimize these costs, facts must be established in the following programs:

- A specific inspection of the drill column and a study of the variables associated with the drilling techniques and the history of the column used, especially taking into

consideration some parameters such as aggressive media knowledge of penetration, contamination (high salinity, oxygen, carbon dioxide, hydrogen sulphide), the thickness limits of the drill pipe, the evaluation of internal and external pitting (**Figure 1**), rotation speed, vibration, bending and coaxial loads;

- An evaluation of specific additives (corrosion inhibitor) to eliminate or reduce the corrosiveness of drilling fluids that are compatible with the environment.

According to Fink [1], drilling muds are mixtures of natural and synthetic chemical compounds used to cool and lubricate the drill string and bit, to remove rock bit cuttings from the bottom of the hole and carry them to the surface, to overcome the fluid pressure of the formation, to avoid damage of the producing formation and to prevent drill pipe corrosion.

There is no doubt that such points are vital in studies of the faced corrosion problems when drilling oil wells and all items justify a careful evaluation.

The function of corrosion inhibitors in saline fluids is to inhibit and/or neutralize the aggressive contaminants present in and/or forming on the surface of carbon steel. They provide a resistant and adherent film that prevents the development of corrosive reactions. The aggressiveness of saline fluids can be increased by incorporating contaminants such as dissolved oxygen (O_2), carbon dioxide (CO_2) and hydrogen sulphide (H_2S).

However, the objective of this article is to show, through laboratory experiments, that corrosion inhibitors based on sodium silicate can be added to aerated saline fluids with an excellent performance for the corrosion protection of carbon steel. In addition, the growing concern for environmental conservation in terms of chemicals is pointed out in this way.



Figure 1. Pitting in a drill pipe.

2. Materials and Methods

2.1. Gravimetric Test

The evaluated material in this work was a carbon steel AISI 1020 (0.20% carbon, 0.34% Mn, 0.35% Si, 0.011% S and 0.009% P) in the confection of the coupons. The coupons used in the gravimetric tests (weight loss) have the following dimensions: 3.17 cm outside diameter, 0.76 cm internal diameter and 0.19 cm thickness. The metal surface was prepared with grade 100 sandpaper until grade 500. They were cleaned with acetone and ethanol and dried with hot air. The test coupons were weighed up to the fourth decimal place using a digital electronic balance.

Gravimetric assays (weight loss) were performed in acrylic containers with a capacity of 20 L possessing a system of continuous circulation through the injection of compressed air in order to maintain a constant oxygenation rate. To ensure the corrosive medium remained under the same experimental conditions, the solution was continuously filtered where the deposits, precipitates and corrosion products of the coupons were constantly removed, as shown in the diagram in **Figure 2**.

The solutions used in the experiments were characterized as saline water containing 3.5% (% mass) of NaCl, 800 mg/L of MgCl₂ and 600 mg/L of CaCl₂. As a corrosion inhibitor, sodium silicate hydrate (Na₂SiO₃·9H₂O) of high purity at concentrations of 250, 750, 1000, 1250, 1500, 1750 and 2000 mg/L was used. The concentrations were referred to as SiO₃²⁻. The exposure time of the coupons was 24, 72 and 96 hours of immersion. At the end of each test, the coupons were photographed.

The final mass of each coupon was obtained after the exposure time referenced. Each coupon was washed with water, pickled with Clark solution for 40 seconds according to ASTM G 31 - 72 [3], cleaned with ethanol and dried with hot air.

The difference in weight for an exposed period of 24 - 96 hours was taken as the total weight loss. Four coupons were used to check the reproducibility of the results. From the average weight loss results, the corrosion rate (CR) and the percentage of corrosion inhibition efficiency (% E) were calculated using equations [3] [4] [5] [6].

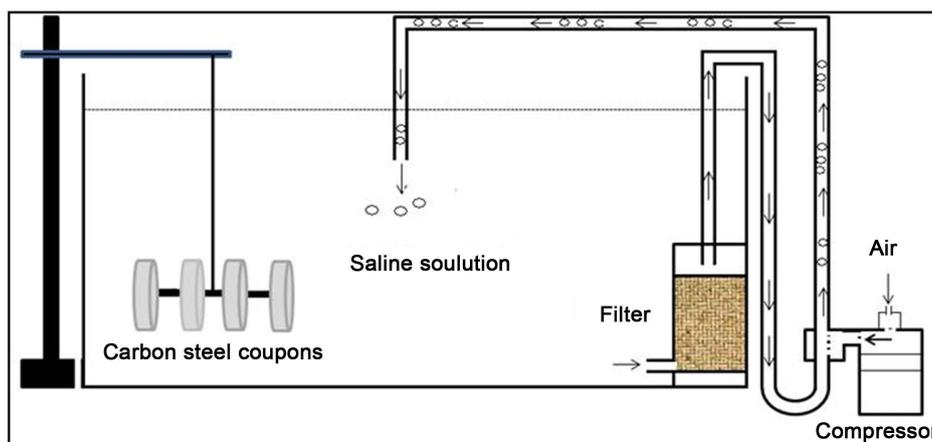


Figure 2. Assay carbon steel immersed in seawater.

$$\text{Corrosion rate} = CR = K(W_o - W_i)/\rho \cdot S \cdot t \text{ (mm/year)}$$

$$\text{Inhibition Efficiency} = \% E = [(W_o - W_i)/W_o] \times 100$$

where:

K is a constant of 8.76×10^4 (value to mm/year);

W_o and W_i are the loss weight in the absence and presence of the inhibitor;

S = area of the coupon (cm^2);

ρ = density in g/cm^3 ; carbon steel is 7.8 g/cm^3 ;

t = exposure time, h.

The degree surface coverage (θ) of the silicate compound obeys the adsorption law of Langmuir and was calculated using equations [7] [8] [9].

$$\text{Degree of surface coverage} = (C_{\text{inhib}}/\theta) = 1/K_a + C_{\text{inhib}}$$

where:

C_{inhib} = inhibitor concentration;

θ = surface coverage (θ);

K_a = equilibrium constant for the adsorption of the inhibitor molecules on the metal surface.

The evaluation of average corrosion rates for carbon steel coupons was based on NACE-RP-07-75 [7], as in **Table 1**.

2.2. Electrochemical Tests

Samples of carbon steel AISI 1020 for the polarization measurements were cut so as to obtain an electrode surface of 0.20 cm^2 . The metal was then embedded in a polyester resin and polished with No. 600 emery paper. Next, the samples were degreased with acetone and washed in distilled water before being inserted in a conventional polarization cell as shown in **Figure 3**.

The polarization curves were recorded using an Autolab potentiostat, Type III by varying the potentials 60 mV/min . All measurements were carried out at a constant temperature of 25°C under aerated conditions and without agitation.

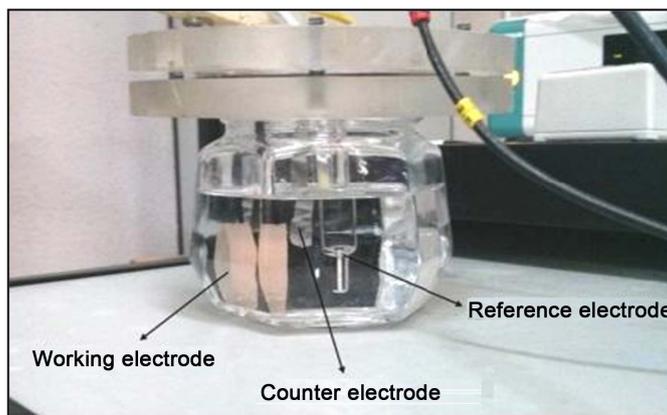


Figure 3. Conventional polarization cell.

Table 1. Average corrosion rate [7].

Average corrosion rate, mm/y	Corrosiveness
<0.025	Low
0.025 - 0.12	Moderate
0.13 - 0.25	High
0.25	Severe

3. Results and Discussion

The visual appearance (**Figure 4**) of the carbon steel coupons immersed in the saline solution containing 3.5% (% mass) of NaCl, 800 mg/L of MgCl₂ and 600 mg/L of CaCl₂ as well as the results of corrosion rates (**Table 2**) with concentrations of 250 and 1500 mg/L of sodium silicate hydrate. The related graphics (**Figures 5-7**) for exposure times of 24 h, 72 h and 96 h can be seen in the following.

To understand the mechanism of corrosion inhibition, the adsorption behavior of silicate adsorbents on the metal surface must be known. The predominant adsorption mode will be dependent on factors such as the properties of the composition of the silicate, the dissociation of the silicate in the saline solution, the temperature, possible chemical changes and the nature of the surface charge on the metal.

The values of the surface coverage (θ), corresponding to different concentrations of silicate in the saline solution, have been used to explain the best isotherm to determine the adsorption isotherm process. The percentage of inhibition efficiency (% E) and the degree of surface coverage (θ) were tabulated in **Table 3** using a gravimetric assay.

Figure 8 confirms that the inhibition processes are due to the adsorption of the silicate compound by the carbon steel surface. This is because a straight line is obtained when $(C_{\text{inhib}}/\theta)$ is plotted against C_{inhib} , and the linear correlation coefficient of the fitted data is close to 1.

The results of the electrolytic tests are presented in **Table 4** and in **Figure 9**.

Based on the graph in **Figure 9** the polarization parameters are shown in **Table 4**, such as: corrosion potential (E_{corr}), resistance polarization (R_p) and corrosion current density (I_{corr}), and the calculated efficiency of sodium silicate in the saline solution. In this figure it is clear the efficiency of addition of sodium silicate on concentration of 1500 mg/L.

It is important to note that, during the gravimetric assays (weight loss), precipitation and the formation of colloidal mixtures of calcium silicates and magnesium hydroxide (considering the high presence of Ca²⁺ and Mg²⁺ ions in the saline solution) occurred. This was removed continuously by the filtering system although such precipitation and colloids do not interfere in the corrosion rate results obtained in gravimetric testing. However, the system was a static electrochemical cell and did not have continuous filtration, allowing the formation of precipitates and colloids on the electrodes. In this way, the random occurrence of deposits interfered in electrochemical results as shown in **Table 4** and graph (**Figure 9**).



Figure 4. The visual appearance of coupons immersed for 96 h.

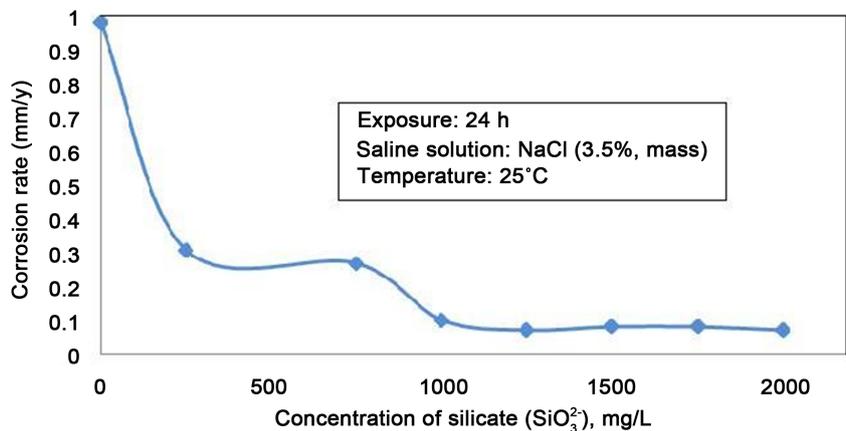


Figure 5. Rate of corrosion of carbon steel coupons in the saline solution with the silicate corrosion inhibitor, 24 h.

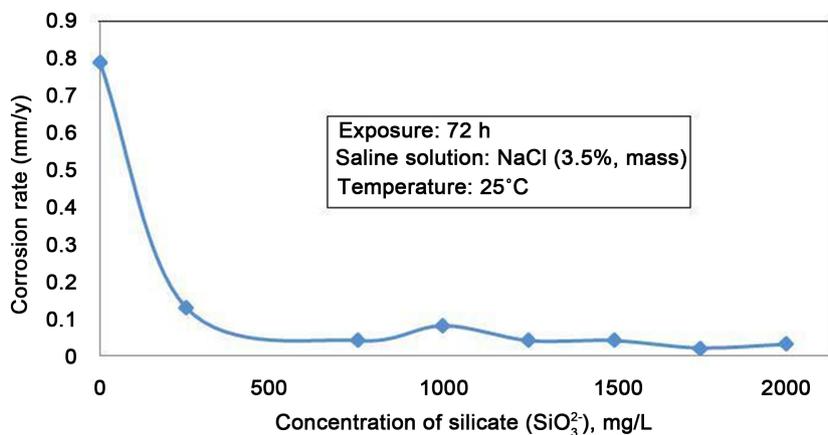


Figure 6. Rate of corrosion of the carbon steel coupon in the saline solution with the silicate corrosion inhibitor, 72 h.

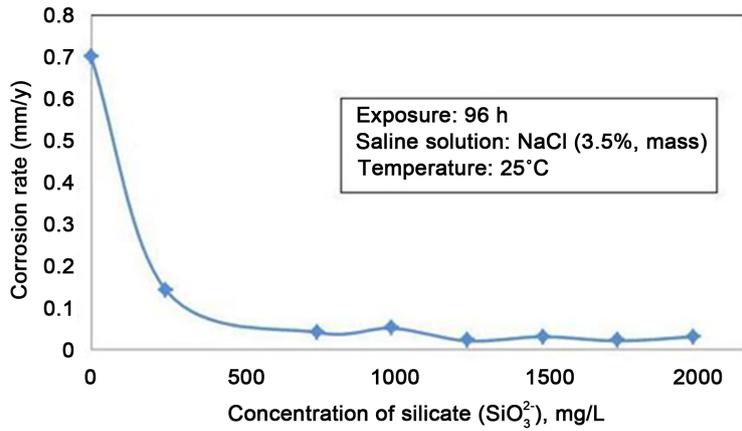


Figure 7. Rate of corrosion of the carbon steel coupon in the saline solution with the silicate corrosion inhibitor, 96 h.

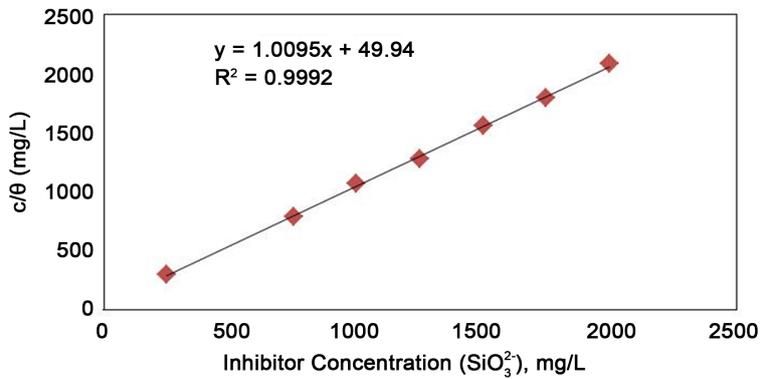


Figure 8. Langmuir adsorption isotherm for the silicate on the carbon steel coupons in the saline solution.

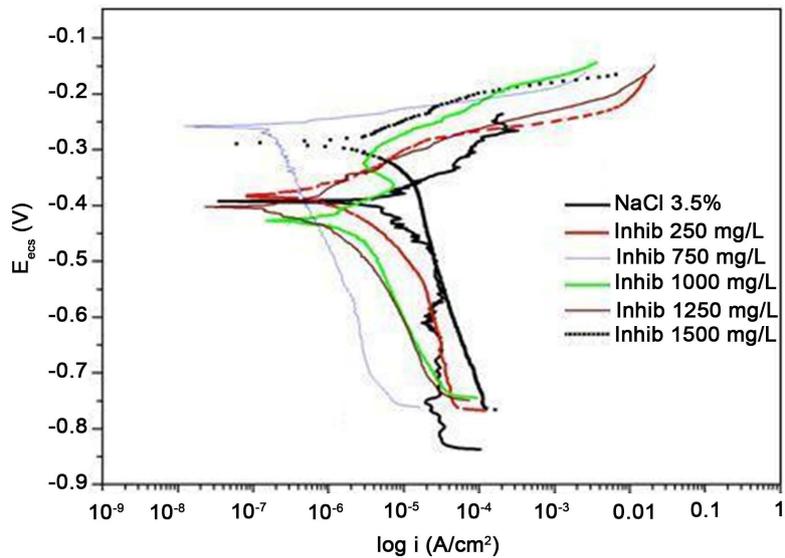


Figure 9. Anodic and cathodic Tafel polarization curves for carbon steel in the absence and presence of various concentrations of the corrosion inhibitor to the base sodium silicate.

Table 2. Laboratory tests of carbon steel coupons in saline water containing 3.5% (% mass) of NaCl, 800 mg/L of MgCl₂ and 600 mg/L of CaCl₂.

Corrosion inhibitor Silicate (SiO ₃ ²⁻), mg/L	Average corrosion rate, mm/y			Efficiency %		
	24 h	72 h	96 h	24 h	72 h	96 h
0	0.99	0.79	0.70	-	-	-
250	0.31	0.13	0.14	68.69	83.54	80.00
750	0.27	0.04	0.04	72.73	94.94	94.28
1000	0.10	0.08	0.05	89.90	89.87	92.86
1250	0.07	0.04	0.02	92.93	94.94	97.14
1500	0.08	0.04	0.03	91.90	94.94	95.71
1750	0.08	0.02	0.02	91.92	97.47	97.14
2000	0.07	0.03	0.03	96.20	96.20	95.71

Table 3. Data from gravimetric testing for 96 h.

Corrosion Inhibitor	SiO ₃ ²⁻ (mg/L)	Efficiency %	Surface coverage (θ)	Concentration/θ (mg/L)
250		80.00	0.8000	312.50
750		94.28	0.9428	795.50
1000		92.85	0.9285	1077.00
1250		97.14	0.9714	1286.80
1500		95.71	0.9571	1567.23
1750		97.14	0.9714	1801.52
2000		95.71	0.9571	2089.64

Table 4. Electrochemical parameters obtained by the Tafel polarization technique for carbon steel in the absence and presence of various concentrations of silicate (SiO₃²⁻).

Conc., mg/L	I _{corr} , A	I _{corr} , A/cm ²	E _{corr} , V	%E
Blank	3.50 × 10 ⁻⁷	1.82 × 10 ⁻⁶	-0.393	-
250	1.48 × 10 ⁻⁷	8.30 × 10 ⁻⁶	-0.384	54.40
750	3.00 × 10 ⁻⁷	1.55 × 10 ⁻⁶	-0.258	91.48
1000	1.19 × 10 ⁻⁷	6.79 × 10 ⁻⁶	-0.427	62.70
1250	5.56 × 10 ⁻⁷	3.64 × 10 ⁻⁶	-0.405	80.00
1500	2.80 × 10 ⁻⁷	1.46 × 10 ⁻⁶	-0.288	19.78

Due to the large amount of deposits and colloids of silicates next to the electrodes, the corrosion rate results and the inhibition efficiency related to 1500 mg/L of silicate (19.78%) are not consistent when compared to gravimetric testing.

Many authors referenced have discussed the mechanism of the formation of and the nature of the protective film formed on a metal surface in solutions of sodium silicate. The difference in the proposals of the mechanism of protection and the conclusions

may be due to the variation in the concentration of the sodium silicate solution, the temperature, the ratio of Na_2O to SiO_2 , the experimental conditions, pH, the contamination of certain ions in the saline solution that can form precipitates and colloids and the presence or absence of dissolved oxygen [10] [11] [12] [13] [14]. The proposals of the mechanisms of the formation of the protective film are based on the research of the authors previously referenced and presented as models in **Figure 10** and **Figure 11**.

Sodium silicate has a strong anodic inhibition in a highly alkaline medium, providing that the SiO_3^{2-} ions migrate to the anode region of the metal surface and react with Fe^{2+} ions, thus forming a protective film of iron silicate (FeSiO_3), as can be seen in **Figure 10**.

The Mg^{2+} and Ca^{2+} ions present in the saline solution may migrate to the cathodic area, find the OH^- ions and form a protective film of $\text{Mg}(\text{OH})_2$ and $\text{Ca}(\text{OH})_2$. Over time, it is possible to obtain a uniform protective film on the metal surface by adsorption, in an alkaline medium, by inhibiting anodic and cathodic reactions, providing an excellent carbon steel protection against corrosion in a saline solution. This film is made up of iron silicate (FeSiO_3), calcium silicate (CaSiO_3), magnesium silicate (MgSiO_3) and silica (SiO_2), these being associated with calcium hydroxide and magnesium hydroxide, as shown in **Figure 10**.

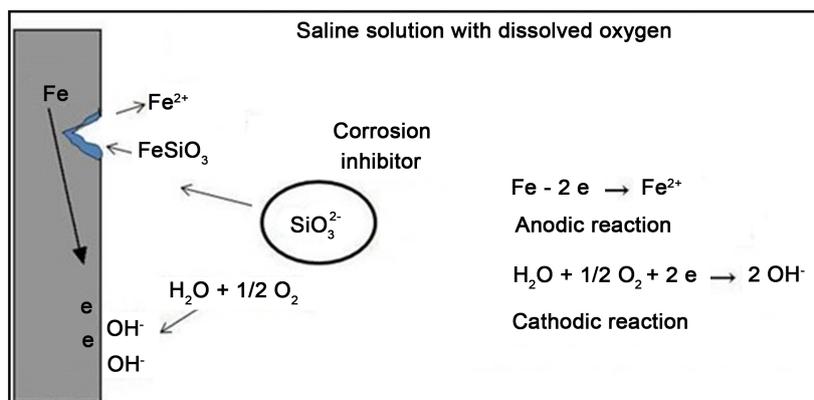


Figure 10. Silicate formation on a metal surface in a saline solution with dissolved oxygen.

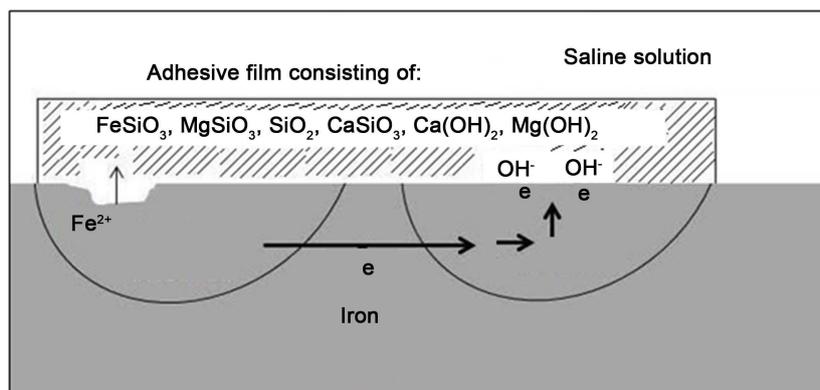


Figure 11. Proposed barrier mechanisms to silicate in the saline solution.

4. Conclusions

On the basis of the bibliographic references and laboratory tests carried out, it is concluded that:

- Through gravimetric testing of the addition of sodium silicate in different concentrations, the formation of a protective film on the surface of the carbon steel, which provided a drastic reduction in the corrosion rate, according to the increase of the concentration of the inhibitor and the immersion time, was observed;
- The concentration of inhibitor that showed better performance was 1250 mg/L, with values of corrosion rates of 0.07 mm/year, 0.04 mm/year, 0.02 mm/year and 0.03 mm/year and 92.93%, 94.94%, 97.14% and 94.00% efficiency for immersion times of 24, 72, 96 and 120 hours respectively;
- For the addition of sodium silicate above 1250 mg/L, significant changes in the rate of corrosion were not observed;
- The electrochemical tests were not conclusive, compared to gravimetric testing, due to the presence of Ca^{2+} and Mg^{2+} ions in the saline that propitiated the formation of precipitates and colloids (calcium silicate, magnesium silicate, calcium hydroxide and magnesium hydroxide) on the electrodes. This disrupted the transmission of the electrochemical parameters;
- The adsorption capacity of the sodium silicate on the carbon steel is based on the Langmuir isotherm;
- The mechanism proposed for the corrosion protection of carbon steel is based on the formation of a uniform and transparent film with a gelatinous aspect on the steel surface constituted of iron silicate, calcium silicate, magnesium silicate and silica, these being associated with calcium hydroxide and magnesium hydroxide;
- Considering the good results of carbon steel corrosion protection obtained with the addition of sodium silicate in a saline fluid, application in the drilling fluids of oil wells, especially for salt water-based fluids, is possible since these fluids present many products that are toxic, corrosive and harmful to the environment.

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