

A Green Microwave-Assisted Synthesis of New Pyridazinium-Based Ionic Liquids as an Environmentally Friendly Alternative

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Abstract

A green microwave-assisted procedure for the preparation of a series of fifteen new N-alkylpyridazinium ionic liquids with different functionality in the alkyl chain is described. For the first time, target ionic liquids were prepared using standard methodology and under microwave irradiation in short duration of time with quantitative yields. In most cases, the microwave-assisted reaction is an environmentally friendly alternative to traditional methods.

Keywords: Green Procedure, Pyridazinium-Salts, Ionic Liquids, Microwave Irradiation

1. Introduction

Over the past few decades the number of publications concerning room-temperature ionic liquids (RTILs) has increased substantially [1]. (RTILs) provide a new class of solvents where molecules are composed of ions. At normal temperatures, ionic liquids have essentially zero-vapor pressure and are thermally stable over a wide range of temperature. Therefore, they are considered as environmentally friendly alternatives to the classical use of organic solvents.

RTILs have been also widely investigated for a variety of applications: the use as solvents or catalysts for chemical synthesis [2-3], media for electrodeposition of metals [4-6], electrolyte for electrochemical devices such as battery [7], supercapacitors [8,9], as inhibitors of corrosion and as fluids for thermal storage and exchange in solar concentrating power plants [10,11].

Recently, several procedures have been recommended for Green Chemistry involving: solvent-free reactions [12], non-classical modes of activation such as ultrasounds or microwaves. The use of MW irradiation leads to large reductions in reaction times, enhancements in conversions, sometimes in selectivity, with several advantages of the eco-friendly approach [13-16]. The most common benefits of MW irradiation are: 1) very rapid reactions, frequently a few minutes, brought about by high and uniform temperatures and combined with pre-

ssure effects, 2) higher degree of purity achieved due to short residence time at high temperatures, 3) yields often better, obtained within shorter times and with purer products [17].

2. Results and Discussion

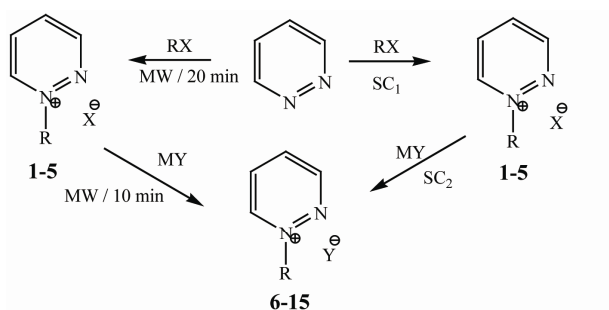
In this work, we have synthesized a variety of a variety of new pyridazinium-based ionic liquids by microwave heating and the protocol is then compared with the similar preparation using conventional heating (oil bath at 80°C). The comparative results are summarized in **Table 1**.

Initially, under standard conditions (SC₁: toluene, 80°C, 18 h), the nucleophilic alkylation of pyridazine with different alkyl halides with different molar ratios (1:1.1; 1:2 and 1:3 eq) afforded the corresponding mono-pyridiazinium halides in 78% - 89% yield as solid powder after solvent removal by filtration (**Scheme 1**) [18]. Alternative anions were subsequently introduced by a metathesis reaction (SC₂: Acetonitrile, 70°C, 3h), with slightly excess of anions namely, tetrafluoroborate or hexafluorophosphate (**Scheme 1**) [19]. The pure metathesis products were obtained after filtration of the salts (metal halides), then followed by filtrate evaporation and washing the residue with dichloromethane followed by further filtration to remove the excess of anion salts (NaBF₄, KPF₆). Finally, evaporation of the filtrate afforded

Table 1. Reaction conditions and reaction yields for the synthesis of pyridazinium-based ionic liquids using conventional preparation (SC) and under microwave irradiation (MW).

S.No.	RX	Yield (%)		Mp (°C)
		SC ₁ ^a	MW ^b	
1	CH ₃ (CH ₂) ₂ Br	82	95	104 - 106
2	CH ₃ (CH ₂) ₃ Br	81	94	94 - 96
3	NC(CH ₂) ₂ Br	87	94	90 - 91*
4	NC(CH ₂) ₃ Cl	78	91	140 - 142
5	EtO ₂ C(CH ₂) ₃ Br	89	95	52 - 55

^a Time (18 h), Temperature (80°C in Toluene); ^b Time (20 min), Temperature (100°C), Power (240 W), Pressure (40 psi); * hygroscopic material.



R = CH₃(CH₂)₂, CH₃(CH₂)₃, NC(CH₂)₂, NC(CH₂)₃, EtO₂C(CH₂)₃
 X = Cl, Br
 MY = NaBF₄, KPF₆

SC₁: toluene, 80°C, 18 h; SC₂: Acetonitrile, 70°C, 3 h

Scheme 1. Synthesis of pyridazinium-based ionic liquids using standard conditions (SC) and under microwave irradiation (MW).

the desired ionic liquids in good yields.

It's noticeable that, the reactivity trend of halides is found to be in the order Br > Cl as expected for an S_N2 reaction, *i.e.* higher temperatures and longer reaction times are required with poorer leaving groups (Cl⁻ < Br⁻)

On the other hand, the microwave-assisted preparation of pyridazinium-based ionic liquids, already synthesized by conventional methods, was explored further with the objective of shortening the reaction time. The reaction conditions were optimized by using either temperature or power control. Use of a small amount of solvent was necessary in these reactions, since the pyridazinium-halides salts are solid. The optimum reaction conditions determined for the synthesis of these unknown ionic liquids was shown in **Table 1**.

Varma reported that an anion exchange metathesis is easily performed under MW activation using a domestic oven [20]. In this way, the preparation of pyridazinium tetrafluoroborate or hexafluorophosphate salts was carried out in the CEM Discover monomode system with a strict

control of power and temperature during the reaction process. The data in **Table 2** indicated that very good yields were obtained within very short reaction times. As observed, the anion nature of exchange agents did not affect the yields.

Interestingly, the alkylation of pyridazine with different molar ratios of alkyl halides afforded only the mono alkylated pyridazinium-salts by using both SC and MW methods without any evidence for formation of dialkylated products. These results were supported on the basis of NMR (see the spectroscopic supplementary materials) and molecular mechanics calculations (MM2). From the MM2 calculations, it's clear that the steric energies of the dialkylated ionic liquids are higher than the monoalkylated products by factor ranged between 3-5 (**Table 3, Figure 1**). These high energies led to the instability of the dialkylated ionic liquid and prevent its formations. Indeed, the instability of the dialkylated ionic liquids can also be attributed to the electrostatic reputations between the two positively charges on the nitrogen atoms of the dialkylated products as well as the reputations of the two negatively charges of the anions outside the ring.

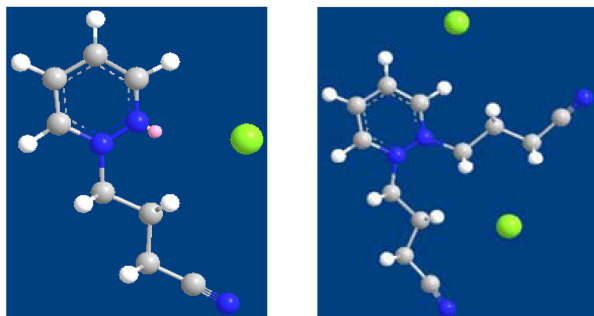
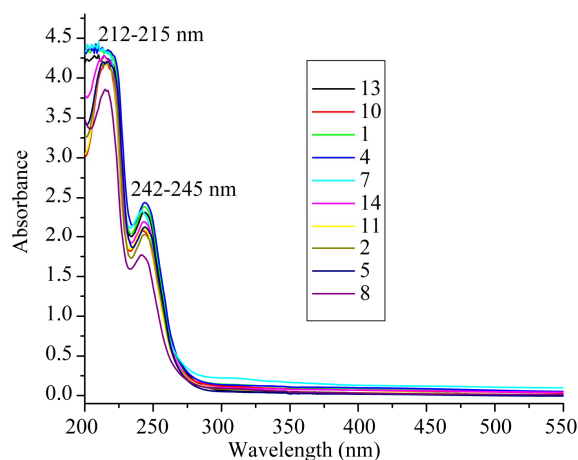
On the other hand, the UV spectra of selected synthesis ionic liquids (**Figure 2**) showed that, all synthesized ionic liquids have a colorless to pale brown color in water solution in concentration of 10⁻³ mol/l solution as well as in solid state with very low absorbance in the visible region. A typical absorption in the UV region at 242 - 245 nm for selected synthesized ionic liquids with shift about 4 nm between them was detected. A broad bathchromic shift to visible region from 525 to 275 nm with low absorbance was monitored when the ionic liquid alkyl chain substituted with terminal chromophoric cyano group (**Figure 2**).

Table 2. Different entries, reaction conditions and reaction yields for the anion metathesis using conventional preparation (CP) and under microwave irradiation (MW).

S.No.	R	X ⁻	M ⁺ Y ⁻	Yield (%)		Mp (°C)
				Anion metathesis (second step) MW ^a	CP ^b	
6	CH ₃ (CH ₂) ₂	Br ⁻	NaBF ₄	95	98	Oil
7	CH ₃ (CH ₂) ₂	Br ⁻	KPF ₆	93	97	Oil
8	CH ₃ (CH ₂) ₃	Br ⁻	NaBF ₄	94	98	Oil
9	CH ₃ (CH ₂) ₃	Br ⁻	KPF ₆	94	98	Oil
10	NC(CH ₂) ₂	Br ⁻	NaBF ₄	92	97	96 - 97
11	NC(CH ₂) ₂	Br ⁻	KPF ₆	93	98	102 - 103
12	NC(CH ₂) ₃	Cl ⁻	NaBF ₄	94	98	Oil
13	NC(CH ₂) ₃	Cl ⁻	KPF ₆	93	97	Oil
14	EtO ₂ C(CH ₂) ₃	Br ⁻	NaBF ₄	92	98	Oil
15	EtO ₂ C(CH ₂) ₃	Br ⁻	KPF ₆	94	98	Oil

Table 3. Steric and optimum energies of both mono- and dialkylated pyridazine based ionic liquids.

Alkyl halide	Energy of monoalkylation	Energy of dialkylation
CH ₃ (CH ₂) ₂ Br	-39.7923	-120.1867
CH ₃ (CH ₂) ₃ Br	-39.3937	-118.4058
NC(CH ₂) ₂ Br	-39.6971	-105.8930
NC(CH ₂) ₃ Cl	-37.2709	-113.8232
EtO ₂ C(CH ₂) ₃ Br	-33.0301	-165.4593

**Figure 1.** An example of the optimization of the energies using MM2 calculations of both mono- alkylated pyridazinium (4) and its corresponding dialkylated pyridazinium - based ionic liquids.**Figure 2.** UV/VIS spectra of some selected ionic liquids ($c = 1 \times 10^{-3}$ mol/l) in water solution.

3. Conclusions

In conclusion, new environmentally friendly pyridazinium-based ionic liquids were prepared by using microwave irradiation. Many advantages for the MW irradiation compared with the standard methods have been recorded. Increasing of the reaction molar ratios between pyridazine and the alkyl halides in both SC and MW did not show any evidence for formation of the dialkylated pyridazine-based ionic liquids. Comparison of SC and

MW methods afforded a lot of advantages and recommendations for the use of the green microwave assisted reactions. These synthesized ionic liquids will be used in the synthesis of novel photochromic materials based ionic liquids as a first attempts for preparation of water soluble photochromic materials based ionic liquid and will be discussed in details in the forthcoming paper.

4. Experimental

4.1. General Comments

All new compounds were synthesized and characterized by ¹H NMR, ¹³C NMR, IR and LCMS.

¹H NMR (400 MHz) and ¹³C NMR (100 MHz) spectra were measured in DMSO at room temperature at 400 MHz. Chemical shifts (δ) were reported in ppm to a scale calibrated for tetramethylsilane (TMS), which is used as an internal standard. The LCMS spectra were measured with a Micromass, LCT mass spectrometer. IR spectra were recorded on a Shimadzu 8201 PC, FTIR spectrophotometer in NaCl and KBr discs (ν_{max} in cm^{-1}). The microwave-assisted reactions were performed using a controllable single-mode microwave reactor, CEM Discovery, designed for synthetic use. The reactor is equipped with a magnetic stirrer as well as a pressure, temperature and power controls. The maximum operating pressure of the reactor is 20 bar. The power and temperature range are 15 - 300 W and 60°C - 250°C, respectively.

4.2. General Procedures for the Synthesis of Pyridazinium Halides (1-5) using Conventional Method

To the solution of pyridazine (1eq) in toluene, was added alkyl halide (1.1 eq) at room temperature, followed by stirring at 80°C for 18 h. The completion of the reaction was marked by the separation of solid from the initially obtained clear and homogenous mixture of pyridazine and alkyl halide in toluene. The product was isolated by filtration to remove the unreacted starting materials and solvent. Subsequently, the quaternary salt was washed with ethyl acetate. In some cases, in each case, the IL/salt was finally dried at a reduced pressure to get rid of all the volatile organic compounds.

4.3. General Procedure for the Synthesis of Pyridazinium Tetrafluoroborates or Hexafluorophosphates (6 - 15) using Conventional Method

The quaternary salt (1 eq) was dissolved in acetonitrile to obtain a clear solution. To this solution of quaternary

halide was added solution of sodium tetrafluoroborate or potassium hexafluorophosphate (1.2 eq), followed by stirring at 70°C for 3 h. The product was isolated by solvent extraction using dichloromethane. The extracts were dried over anhydrous MgSO₄ and evaporated under reduced pressure to yield the desired product.

4.4. General Procedures for the Synthesis of Pyridazinium halides (1-5) under Microwave Irradiation

Pyridazine (1 eq) and different alkyls halides (1 eq) were placed in a microwave reactor vessel and irradiated for 20 minutes at 100°C. the product is collected as described in the conventional method.

4.5. General Procedure for the Synthesis of Pyridazinium Tetrafluoroborate Hexafluorophosphates (6 - 15) under Microwave Irradiation

Pyridazinium-halides salts (1 eq) were placed in a microwave reactor vessel, 1 eq of NaBF₄ or KPF₆ in a small amount of acetonitrile was added. The reaction was carried out for 10 min at 70°C. the product is collected as described in the conventional method.

5. Characterization

5.1. 1-Propylpyridazin-1-ium Bromide (1)

¹H NMR (400 MHz, DMSO) δ : 1.00 (t, $J = 7.6$, 3H), 2.16 (quint, $J = 7.6$, 2H), 4.86 (t, $J = 7.6$, 2H), 8.53 - 8.56 (dd, $J = 1.2$, 1H), 8.61 - 8.65 (dd, $J = 1.2$, 1H), 9.53 - 9.54 (d, $J = 1.2$ 1H), 9.73 - 9.74 (d, $J = 1.2$ 1H); ¹³C NMR (100MHz, DMSO) δ : 12.3 (CH₃), 25.8 (CH₂), 69.7 (CH₂), 138.4 (C-4), 139.0 (C-5), 151.7 (C-6), 157.1 (C-3).

5.2. 1-Butylpyridazin-1-ium Bromide (2)

¹H NMR (400 MHz, DMSO) δ : 0.97 (t, $J = 7.2$, 3H), 1.42 (quint, $J = 7.2$, 2H), 2.10 (quint, $J = 7.2$, 2H), 4.90 (t, $J = 7.2$, 2H), 8.55 - 8.56 (dd, $J = 1.2$, 1H), 8.61 - 8.63 (dd, $J = 1.2$, 1H), 9.53 - 9.54 (d, $J = 1.2$ 1H), 9.74 - 9.76 (d, $J = 1.2$ 1H); ¹³C NMR (100MHz, DMSO) δ : 15.2 (CH₃), 21.3 (CH₂), 34.1 (CH₂), 68.1 (CH₂), 138.4 (C-4), 139.0 (C-5), 151.7 (C-6), 157.1 (C-3).

5.3. 1-(2-Cyanoethyl)Pyridazin-1-ium Bromide (3)

¹H NMR (400 MHz, DMSO) δ : 3.46 (t, $J = 6$, 2H), 5.25 (t, $J = 6$, 2H), 8.62 - 8.66 (dd, $J = 1.2$, 1H), 8.69 - 8.73 (dd, $J = 1.2$, 1H), 9.61 - 9.63 (d, $J = 1.2$ 1H), 9.82 - 9.84 (d, $J = 1.2$ 1H); ¹³C NMR (100MHz, DMSO) δ : 20.4 (CH₂), 62.5 (CH₂), 119.9 (CN), 138.7 (C-4), 140.3 (C-5), 152.9 (C-6), 157.6 (C-3).

¹H NMR (400 MHz, DMSO) δ : 2.48 (quint, $J = 7.2$, 2H), 2.68 (t, $J = 7.2$, 2H), 4.99 (t, $J = 7.2$, 2H), 8.52 - 8.55 (dd, $J = 1.2$, 1H), 8.59 - 8.63 (dd, $J = 1.2$, 1H), 9.51 - 9.52 (d, $J = 1.2$ 1H), 9.76 - 9.77 (d, $J = 1.2$ 1H); ¹³C NMR (100 MHz, DMSO) δ : 13.9 (CH₂), 24.9 (CH₂), 63.8 (CH₂), 120.1 (CN), 136.1 (C-4), 137.2 (C-5), 150.0 (C-6), 154.8 (C-3).

5.4. 1-(3-Cyanopropyl)Pyridazin-1-ium Chloride (4)

¹H NMR (400 MHz, DMSO) δ : 1.15 (t, $J = 7.2$, 3H), 2.26 (quint, $J = 7.2$, 2H), 2.50 (quint, $J = 7.2$, 2H), 4.03 (q, $J = 7.2$, 2H), 4.85 (t, $J = 7.2$, 2H), 8.55 - 8.56 (dd, $J = 1.2$, 1H), 8.61 - 8.63 (dd, $J = 1.2$, 1H), 9.53 - 9.54 (d, $J = 1.2$ 1H), 9.74 - 9.76 (d, $J = 1.2$ 1H); ¹³C NMR (100 MHz, DMSO) δ : 14.2 (CH₃), 24.9 (CH₂), 30.2 (CH₂), 60.4 (CH₂), 64.1 (CH₂), 136.2 (C-4), 136.9 (C-5), 150.2 (C-6), 154.7 (C-3), 172.3 (CO).

5.5. 1-(4-Ethoxy-4-Oxobutyl)Pyridazin-1-ium Bromide (5)

¹H NMR (400 MHz, DMSO) δ : 0.92 (t, $J = 7.6$, 3H), 2.04 (quint, $J = 7.6$, 2H), 4.77 (t, $J = 7.6$, 2H), 8.54 - 8.58 (dd, $J = 1.2$, 1H), 8.66 - 8.70 (dd, $J = 1.2$, 1H), 9.58 - 9.60 (d, $J = 1.2$ 1H), 9.85 - 9.87 (d, $J = 1.2$ 1H); ¹³C NMR (100 MHz, DMSO) δ : 10.3 (CH₃), 23.0 (CH₂), 66.2 (CH₂), 136.0 (C-4), 136.6 (C-5), 149.8 (C-6), 154.6 (C-3).

5.6. 1-Propylpyridazin-1-ium Tetrafluoroborate (6)

¹H NMR (400 MHz, DMSO) δ : 0.94 (t, $J = 7.6$, 3H), 2.05 (quint, $J = 7.6$, 2H), 4.79 (t, $J = 7.6$, 2H), 8.59 - 8.63 (dd, $J = 1.2$, 1H), 8.72 - 8.75 (dd, $J = 1.2$, 1H), 9.63 - 9.64 (d, $J = 1.2$ 1H), 9.88 - 9.90 (d, $J = 1.2$ 1H); ¹³C NMR (100 MHz, DMSO) δ : 10.2 (CH₃), 22.9 (CH₂), 65.9 (CH₂), 135.9 (C-4), 136.5 (C-5), 149.8 (C-6), 154.5 (C-3).

5.7. 1-Propylpyridazin-1-ium Hexafluorophosphate (7)

¹H NMR (400 MHz, DMSO) δ : 0.92 (t, $J = 7.2$, 3H), 1.35 (quint, $J = 7.2$, 2H), 1.99 (quint, $J = 7.2$, 2H), 4.81 (t, $J = 7.2$, 2H), 8.52 - 8.58 (dd, $J = 1.2$, 1H), 8.67 - 8.71 (dd, $J = 1.2$, 1H); ¹³C NMR (100 MHz, DMSO) δ : 10.2 (CH₃), 22.9 (CH₂), 65.9 (CH₂), 135.9 (C-4), 136.5 (C-5), 149.8 (C-6), 154.5 (C-3).

5.8. 1-Butylpyridazin-1-ium Tetrafluoroborate (8)

¹H NMR (400 MHz, DMSO) δ : 0.92 (t, $J = 7.2$, 3H), 1.35 (quint, $J = 7.2$, 2H), 1.99 (quint, $J = 7.2$, 2H), 4.81 (t, $J = 7.2$, 2H), 8.52 - 8.58 (dd, $J = 1.2$, 1H), 8.67 - 8.71 (dd, $J = 1.2$, 1H); ¹³C NMR (100 MHz, DMSO) δ : 10.2 (CH₃), 22.9 (CH₂), 65.9 (CH₂), 135.9 (C-4), 136.5 (C-5), 149.8 (C-6), 154.5 (C-3).

= 1.2, 1H), 9.59 - 9.61 (d, $J = 1.2$ 1H), 9.88 - 9.89 (d, $J = 1.2$ 1H); ^{13}C NMR (100MHz, DMSO) δ : 13.3 (CH₃), 18.8 (CH₂), 34.4 (CH₂), 64.3 (CH₂), 136.0 (C-4), 136.6 (C-5), 149.8 (C-6), 154.5 (C-3).

5.9. 1-Butylpyridazin-1-ium Hexafluorophosphate (9)

^1H NMR (400MHz, DMSO) δ : 0.93 (t, $J = 7.2$, 3H), 1.34 (quint, $J = 7.2$, 2H), 1.99 (quint, $J = 7.2$, 2H), 4.81 (t, $J = 7.2$, 2H), 8.56-8.60 (dd, $J = 1.2$, 1H), 8.70-8.72 (dd, $J = 1.2$, 1H), 9.60-9.61 (d, $J = 1.2$ 1H), 9.90-9.92 (d, $J = 1.2$ 1H); ^{13}C NMR (100MHz, DMSO) δ : 13.2 (CH₃), 18.7 (CH₂), 31.4 (CH₂), 64.4 (CH₂), 135.9 (C-4), 136.5 (C-5), 149.8 (C-6), 154.5 (C-3).

5.10. 1-(2-Cyanoethyl)Pyridazin-1-ium Tetrafluoroborate (10)

^1H NMR (400 MHz, DMSO) δ : 3.37 (t, $J = 6$, 2H), 5.13 (t, $J = 6$, 2H), 8.64 - 8.68 (dd, $J = 1.2$, 1H), 8.77 - 8.80 (dd, $J = 1.2$, 1H), 9.67 - 9.68 (d, $J = 1.2$ 1H), 9.88 - 9.89 (d, $J = 1.2$ 1H); ^{13}C NMR (100MHz, DMSO) δ : 17.7 (CH₂), 59.3 (CH₂), 117.5 (CN), 136.1 (C-4), 137.4 (C-5), 150.8 (C-6), 154.6 (C-3).

5.11. 1-(2-Cyanoethyl)Pyridazin-1-ium Hexafluorophosphate (11)

^1H NMR (400 MHz, DMSO) δ : 3.38 (t, $J = 6$, 2H), 5.13 (t, $J = 6$, 2H), 8.65 - 8.68 (dd, $J = 1.2$, 1H), 8.77 - 8.81 (dd, $J = 1.2$, 1H), 9.68 - 9.69 (d, $J = 1.2$ 1H), 9.89 - 9.90 (d, $J = 1.2$ 1H); ^{13}C NMR (100MHz, DMSO) δ : 17.7 (CH₂), 59.3 (CH₂), 117.5 (CN), 136.0 (C-4), 137.4 (C-5), 150.8 (C-6), 154.6 (C-3).

5.12. 1-(3-Cyanopropyl)Pyridazin-1-ium Tetrafluoroborate (12)

^1H NMR (400 MHz, DMSO) δ : 2.36 (quint, $J = 7.2$, 2H), 2.68 (t, $J = 7.2$, 2H), 4.88 (t, $J = 7.2$, 2H), 8.59 - 8.62 (dd, $J = 1.2$, 1H), 8.71 - 8.74 (dd, $J = 1.2$, 1H), 9.61 - 9.62 (d, $J = 1.2$ 1H), 9.86 - 9.87 (d, $J = 1.2$ 1H); ^{13}C NMR (100 MHz, DMSO) δ : 13.6 (CH₂), 25.1 (CH₂), 63.2 (CH₂), 119.7 (CN), 136.1 (C-4), 136.9 (C-5), 150.5 (C-6), 154.5 (C-3).

5.13. 1-(3-Cyanopropyl)Pyridazin-1-ium Hexafluorophosphate (13)

^1H NMR (400 MHz, DMSO) δ : 2.34 (quint, $J = 7.2$, 2H), 2.69 (t, $J = 7.2$, 2H), 4.91 (t, $J = 7.2$, 2H), 8.61 - 8.64 (dd, $J = 1.2$, 1H), 8.73 - 8.77 (dd, $J = 1.2$, 1H), 9.64 - 9.65 (d,

$J = 1.2$ 1H), 10.0 - 10.1 (d, $J = 1.2$ 1H); ^{13}C NMR (100 MHz, DMSO) δ : 13.5 (CH₂), 25.1 (CH₂), 63.1 (CH₂), 119.6 (CN), 136.0 (C-4), 136.9 (C-5), 150.5 (C-6), 154.5 (C-3).

5.14. 1-(4-Ethoxy-4-Oxobutyl)Pyridazin-1-ium Bromide (14)

^1H NMR (400MHz, DMSO) δ : 1.26 (t, $J = 7.2$, 3H), 2.44 (quint, $J = 7.2$, 2H), 2.61 (quint, $J = 7.2$, 2H), 4.17 (q, $J = 7.2$, 2H), 4.97 (t, $J = 7.2$, 2H), 8.54 - 8.58 (dd, $J = 1.2$, 1H), 8.62 - 8.65 (dd, $J = 1.2$, 1H), 9.54 - 9.55 (d, $J = 1.2$ 1H), 9.75 - 9.77 (d, $J = 1.2$ 1H); ^{13}C NMR (100 MHz, DMSO) δ : 15.8 (CH₃), 27.1 (CH₂), 32.9 (CH₂), 64.5 (CH₂), 67.2 (CH₂), 138.5 (C-4), 139.3 (C-5), 152.1 (C-6), 157.2 (C-3), 177.3 (CO).

5.15. 1-(4-Ethoxy-4-Oxobutyl)Pyridazin-1-ium Bromide (15)

^1H NMR (400 MHz, DMSO) δ : 1.17 (t, $J = 7.2$, 3H), 2.29 (quint, $J = 7.2$, 2H), 2.49 (quint, $J = 7.2$, 2H), 4.05 (q, $J = 7.2$, 2H), 4.85 (t, $J = 7.2$, 2H), 8.57 - 8.60 (dd, $J = 1.2$, 1H), 8.69 - 8.72 (dd, $J = 1.2$, 1H), 9.60 - 9.62 (d, $J = 1.2$ 1H), 9.89 - 9.91 (d, $J = 1.2$ 1H); ^{13}C NMR (100MHz, DMSO) δ : 14.0 (CH₃), 24.7 (CH₂), 30.0 (CH₂), 60.1 (CH₂), 63.9 (CH₂), 136.0 (C-4), 136.7 (C-5), 150.2 (C-6), 154.5 (C-3), 172.0 (CO).

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7. Supplementary Data

Supplementary data (^1H NMR, ^{13}C NMR, ^{11}B NMR, ^{19}F NMR and ^{31}P NMR spectra) associated with this article are available.

8. References

- [1] R. D. Rogers and K. Seddon, "Ionic Liquids: Industrial Applications for Green Chemistry," Oxford University Press, Oxford, 2002.
- [2] J. F. Liu, G. B. Jiang, Y. G. Chi, Y. Q. Cai, Q. X. Zhou and J. T. Hu, "Use of Ionic Liquids for Liquid-Phase Microextraction of Polycyclic Aromatic Hydrocarbons," *Analytical Chemistry*, Vol. 75, No. 21, 2003, pp. 5870-5876. [doi:10.1021/ac034506m](https://doi.org/10.1021/ac034506m)
- [3] J. H. Wang, D. H. Cheng, X. Y. Chen, Z. Du and Z. L. Fang, "Direct Extraction of Double-Stranded DNA into

- Ionic Liquid 1-Butyl-3-Methylimidazolium Hexafluorophosphate and Its Quantification," *Analytical Chemistry*, Vol. 79, No. 2, 2007, 620. [doi:10.1021/ac061145c](https://doi.org/10.1021/ac061145c)
- [4] F. Endres, "Ionic Liquids, Solvents for the Electro-deposition of Metals and Semiconductors," *ChemPhysChem*, Vol. 3, No. 2, 2002, pp. 144-154. [doi:10.1002/1439-7641\(20020215\)3:2<144::AID-CPHC144>3.0.CO;2-#](https://doi.org/10.1002/1439-7641(20020215)3:2<144::AID-CPHC144>3.0.CO;2-#)
- [5] Y. F. Lin and S. I. Wen, "Electrodeposition of Zinc from a Lewis Acidic Zinc Chloride-1-Ethyl-3-Methylimidazolium Chloride Molten Salt," *Electrochimica Acta*, Vol. 44, No. 16, 1999, pp. 2771-2777. [doi:10.1016/S0013-4686\(99\)00003-1](https://doi.org/10.1016/S0013-4686(99)00003-1)
- [6] S. Takahashi, N. Koura, S. Kohara, M. L. Saboungi and L. A. Curtiss, "Technological and Scientific Issues of Room-Temperature Molten Salts," *Plasma Ions*, Vol. 2, No. 3-4, 1999, pp. 91-105. [doi:10.1016/S1288-3255\(99\)00105-7](https://doi.org/10.1016/S1288-3255(99)00105-7)
- [7] J. F. Brennecke and E. J. Magin, "Ionic Liquids: Innovative Fluids for Chemical Processing," *AIChE Journal*, Vol. 47, No. 11, 2001, pp. 2384-2389. [doi:10.1002/aic.690471102](https://doi.org/10.1002/aic.690471102)
- [8] M. Ue, M. Takeda, A. Toriumi, A. Kominato, R. Hagiwara and Y. Ito, "Application of Low-Viscosity Ionic Liquid to the Electrolyte of Double-Layer Capacitors," *Journal of the Electrochemical Society*, Vol. 150, No. 4, 2003, pp. A499-A502. [doi:10.1149/1.1559069](https://doi.org/10.1149/1.1559069)
- [9] A. Balducci, U. Bardi, S. Caporali, M. Mastragostino and F. Soavi, "Ionic Liquids for Hybrid Supercapacitors," *Electrochemistry Communications*, Vol. 6, No. 6, 2004, pp. 566-570. [doi:10.1016/j.elecom.2004.04.005](https://doi.org/10.1016/j.elecom.2004.04.005)
- [10] M. Messali, "A Green Microwave-Assisted Synthesis, Characterization and Comparative Study of New Pyridazinium-Based Ionic Liquids Derivatives towards Corrosion of Mild Steel in Acidic Environment," *Journal of Materials and Environmental Science*, Vol. 2, No. 2, 2011, pp. 174-185.
- [11] L. Moens, D. M. Blake, D. L. Rudnicki and M. J. Hale, "Advanced Thermal Storage Fluids for Solar Parabolic Trough Systems," *Journal of Solar Energy Engineering*, Vol. 125, No. 1, 2003, pp. 112-116. [doi:10.1115/1.1531644](https://doi.org/10.1115/1.1531644)
- [12] P. T. Anastas and J. C. Warner, "Green Chemistry, Theory and Practice," Oxford University Press, Oxford, UK, 1998.
- [13] A. Loupy, "Solvent-Free Microwave Organic Synthesis as an Efficient Procedure for Green Chemistry," *Comptes Rendus Chimie*, Vol. 7, No. 2, 2004, pp. 103-112.
- [14] A. Aupoix, B. Pegot and G. Vo-Thanh, "Synthesis of Imidazolium and Pyridinium-Based Ionic Liquids and Application of 1-Alkyl-3-Methylimidazolium Salts as Pre-catalysts for the Benzoin Condensation Using Solvent-Free and Microwave Activation," *Tetrahedron*, Vol. 66, No. 6, 2010, pp. 1352-1356. [doi:10.1016/j.tet.2009.11.110](https://doi.org/10.1016/j.tet.2009.11.110)
- [15] F. Yi, Y. Peng and G. Song, "Microwave-Assisted Liquid-Phase Synthesis of Methyl 6-Amino-5-Cyano-4-Aryl-2-Methyl-4H-Pyran-3-Carboxylate Using Functional Ionic Liquid as Soluble Support," *Tetrahedron Letters*, Vol. 46, No. 22, 2005, pp. 3931-3933. [doi:10.1016/j.tetlet.2005.03.197](https://doi.org/10.1016/j.tetlet.2005.03.197)
- [16] V. Singh, S. Kaur, V. Saphehiya, J. Singh and G. L. Kad, "Microwave Accelerated Preparation of [Bmim][HSO₄] Ionic Liquid: An Acid Catalyst for Improved Synthesis of Coumarins," *Catalysis Communications*, Vol. 6, No. 1, 2005, pp. 57-60. [doi:10.1016/j.catcom.2004.10.011](https://doi.org/10.1016/j.catcom.2004.10.011)
- [17] M. Deetlefs and K. R. Seddon, "Improved Preparations of Ionic Liquids Using Microwave Irradiation," *Green Chemistry*, Vol. 5, 2003, No. 2, pp. 181-186. [doi:10.1039/b300071k](https://doi.org/10.1039/b300071k)
- [18] M. Messali, "A Facile and Green Microwave-Assisted Synthesis of New Functionalized Picolinium-Based Ionic Liquids," *Arabian Journal of Chemistry*, In Press, Available Online, July 2011. [doi:10.1016/j.arabjc.2011.06.030](https://doi.org/10.1016/j.arabjc.2011.06.030)
- [19] J. G. Huddleston, A. E. Visser, W. M. Reichert, H. D. Willauer, G. A. Broker and R. D. Rogers, "Characterization and Comparison of Hydrophilic and Hydrophobic Room Temperature Ionic Liquids Incorporating the Imidazolium Cation," *Green Chemistry*, Vol. 3, No. 4, 2001, pp. 156-164. [doi:10.1039/b103275p](https://doi.org/10.1039/b103275p)
- [20] R. S. Varma and V. V. Namboodiri, "Solvent-Free Preparation of Ionic Liquids Using a Household Microwave Oven," *Pure and Applied Chemistry*, Vol. 73, No. 8, 2001, pp. 1309-1313. [doi:10.1351/pac200173081309](https://doi.org/10.1351/pac200173081309)