

Boosted Photoactivity of Titania Nanotube Layers Doped with a Suspension of Gold Nanoparticles

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Abstract

In the present work, we report on the behavior of synthesized gold nanoparticles suspension, incorporated in titanium dioxide nanotube layers (TiO₂-NT) and fabricated by electrochemical anodization in 0.4 wt% hydrofluoric acid solution and we study its photocatalytic response. Gold nanoparticles were characterized using Transmission electron microscopy and X-ray diffraction. Scanning electron microscopy was used to study the morphology of TiO₂ nanotube layers doped by gold nanoparticles. Boosted photocatalytic performances on the degradation of an azo dye were obtained by using TiO₂ nanotube layers doped by gold nanoparticles (Au/TiO₂-NT), compared to undoped TiO₂ nanotube layer (TiO₂-NT) catalysts. Under UV irradiation, this new nanomaterial, with noble metal-semi conductor heterojunction (Au/TiO₂-NT) exhibits a synergetic effect in accelerating the electron transfert, resulting in an enhanced photoactivity recorded in the kinetics of degradation of Acid Orange 7 (AO7). Chronoamperometry was used to highlight higher photocurrent produced by gold-titania interface submitted to UV irradiation.

Keywords

Gold Nanoparticle, TiO₂-NT, Photocurrent, Photodegradation

1. Introduction

Nowadays, the use of nanostructured surfaces, in conjunction with nanoparticles is of growing interest. Numerous studies actually deal with high quality and miniaturized sensor applications [1] [2] [3], medical applications and biosensors

[4] [5] [6] [7], photocatalytic applications [8] [9] [10] [11] and pollutant removers [12] [13] [14] [15] just to mention the most commonly found in the recent literature. From these applications viewpoints, the development of future nanocomposite requires the investigation of different elaboration pathway in order to promote synergetic effect between the nano compounds and to obtain a device cheaper and easier to be fabricated and handled. Titanium dioxide nanotube arrays are fabricated using anodic oxidation of titanium metal [16] [17] [18]. This electrochemical method is simple, cost effective and scalable. Titanium reacts with atmospheric oxygen, making a passive layer on the outside surface which is very convenient to build titanium dioxide nanotube arrays on a titanium wire electrode [19] [20]. Titanium based nano-array systems establish inherent properties such as high surface area [21], controllable nanotube dimensions [22], tunable geometries and surface chemistry [23] and so forth. These additional and valuable new properties given by the nanostructuration lead to considering titanium dioxide nanotube arrays as suitable photocatalyst material. In addition, recent studies report the use of TiO_2 material for the degradation of organic compound such as methyl orange [24] [25], acid Orange 7 [26] [27], rhodamine [28] [29] and many others. Numerous studies are dedicated to the improvement of the dye degradation by TiO_2 composite using inorganic doping [30] [31], chemical activation [32] [33] and development of synergetic effect [34] [35] [36].

Titanium dioxide nanotube arrays are of great interest in order to develop synergetic effect with noble metal nanoparticles, and nanostructured TiO_2 can act as efficient nanoparticle carriers. Supported gold nanoparticles exhibit outstanding catalytic properties when deposited on titanium dioxide [37] [38]. The major problem with a gold based catalyst is the deactivation during the catalytic reaction; this phenomenon is often attributed to the agglomeration of gold nanoparticle [39] and sintering [40]. A larger surface area of the support is needed to maintain a better dispersion of the gold nanoparticles and avoid deactivation. Titanium dioxide nanotube arrays provide higher surface area than flat titanium dioxide and numerous defects on the surface make them an interesting alternative as a support for gold nanoparticles.

Our study is dedicated to the improvement of photocatalytic properties of TiO_2 nanotube layers (TiO_2 -NT), researching the establishment of a synergetic effect between the gold nanoparticles and the TiO_2 nanotube layers. In the present work, we report on the behavior of gold nanoparticles embedded on titanium dioxide nanotube arrays using a simple immersion technique. Photocurrent measurements and photodegradation experiments are used as a probe of the enhanced photocatalytic performance of the elaborated new nanocomposite.

2. Experimental

2.1. Fabrication of TiO_2 Nanotube Layers (TiO_2 -NT)

Titanium foils supplied by Good Fellow, 99.6% in purity are used as starting material. The samples were degreased by successive sonication in trichloroethylene, acetone, methanol and water, followed by rinsing in deionized water and

blown dried with nitrogen. In order to purify the degreased surface, each foil was immersed in a HF: HNO₃: HCl: H₂O solution. The samples were dried in an oven at 100°C and stored in a desiccator before use.

Anodization of the cleaned titanium sample was carried out at room temperature (20°C) in 0.4 wt% HF aqueous solution with the anodizing voltage maintained at 20 Volts for 20 minutes using a home-made anodization apparatus.

Annealing treatments of the samples were performed using a Nabertherm tubular furnace under air, for 2 hours at 500°C to obtain crystallized anatase (TiO₂-NT).

2.2. Fabrication of Gold Nanoparticles in Suspension

Before the experiment, all the glassware was carefully cleaned using Aqua regia (mixture of concentrated Nitric and Hydrochloric acid in 1: 3 volume ratio) and rinsed with ultra-pure water (18 MΩ·cm), blown dried with nitrogen and stored at 110°C. The starting solution is made by dissolving the designated amount of hydrogen tetrachloroaurate (III) hydrate (HAuCl₄ · xH₂O) purchased from Sigma Aldrich to reach 1mM solution. Gold nanoparticles in suspension were synthesized using a two steps method. First, reduction of hydrogen tetrachloroaurate (III) precursor species was carried out using sodium borohydride under vigorous stirring for 15 minutes at room temperature, gold metal nano core were obtained. To avoid undesired aggregation, nanoparticles surface were post functionalized using PEG (Mw 5000). A red ruby, stable gold nanoparticles suspension is obtained.

2.3. Fabrication of Gold Doped TiO₂ Nanotube Layers (Au/TiO₂-NT)

Annealed TiO₂ nanotube layers (TiO₂-NT) were immersed in the previously prepared gold nanoparticles in suspension for one minute. Samples were rinsed with ultra-pure water and submitted to UV post treatment at $\lambda = 254$ nm for 15 minutes. Gold doped TiO₂ nanotube layers (Au/TiO₂-NT) were obtained.

2.4. Gold Nanoparticles Characterization

2.4.1. Transmission Electron Microscopy (TEM)

The TEM of each samples were obtained using a Hitachi transmission electron microscope (H7650). The accelerating voltage used was 80 kV. Micrographs were made using a Hamamatsu AMT CCD camera placed in a side position. To perform the characterization, 10 µL sample solution was deposited on a 300 mesh carbon-coated copper grid and dried at room temperature. The average gold core diameter and standard deviations were calculated for each nanoparticle sample by averaging 50 nanoparticles from the TEM images using AMT Image Analysis Software (Advanced Microscopy Techniques Corp.).

2.4.2. X Ray Diffraction Analysis

The gold nanoparticles in suspension were dried on a wax pellet and the XRD measurements were performed using a Bruker D2 Phaser diffractometer. The

pattern was recorded by CuK α 1 radiation with λ of 1.5406 Å and nickel monochromator filtering wave at tube voltage of 30 kV and tube current of 10 mA. The scanning was done in the region of 2θ from 30° to 80° at 0.24°/min and the time constant was 5 s. The spinning speed of the sample holder was 15°/min. The size of the nanoparticles was calculated using the Scherrer's formula.

2.4.3. UV-Visible Spectroscopy

The UV-vis spectra of the suspensions of gold nanoparticles were recorded using a Perkin Elmer Lambda 35 double beam UV-vis spectrophotometer from 300 to 700 nm. Ultra-pure water was taken as reference. The maximum absorbance of the surface plasmon resonance (SPR) band was recorded for each sample.

2.5. TiO₂ and (Au/TiO₂) Nanotube Layers Characterization

2.5.1. Field Emission Scanning Electron Microscopy (FESEM)

The samples were characterized by a Field Emission Scanning Electron Microscope (FESEM) from ZEISS ultra- + (ZEISS-Germany) used with a low voltage (900 V and 3 kV) to limit charge effects. The working distance was from 2.5 mm to 6 mm. Samples were observed on a flat view or with a various tilt angle (33° and 40°). Pictures were obtained with the secondary electron detector in-lens. These characterizations were carried out at the Neel Institute of Grenoble.

2.5.2. Chronoamperometry

Photocurrent density of the different TiO₂ layers were measured using a Voltalab PGZ 301 apparatus in a three-electrode configuration with the TiO₂ layer as the working electrode, a Pt wire as the counter electrode and a Ag/AgCl (in KCl saturated solution) as the reference electrode. Photocurrent measurements were performed in Na₂SO₄ 5%wt electrolyte with a constant voltage of 1200 mV.

2.6. Photocatalytic Degradation of Acid Orange 7

Photo-catalytic experiments were conducted in 3 mL of AO7 solution (from Acros Organics, also called Orange II, CAS# 633-96-5) with a concentration of 5.0×10^{-5} mol/L, placed in a cylindrical quartz glass reactor (1.5 cm diameter by 16 cm tall). The surface area of the immersed samples was approximately 3 cm². The quartz glass reactor was irradiated with polychromatic fluorescent UV lamps (Philips TDL 8W (total optical power, 1.3 W), 300 mm long, wavelength range 350 - 400 nm) in a configuration providing about 2 mW/cm² at the sample surface. The photo-catalytic decomposition of AO7 solution was monitored by the decrease of the solution's absorbance at 468 nm using a Perkin Elmer Lambda 35 double beam UV-vis spectrophotometer from 300 to 700 nm. Deionized water was taken as the reference.

3. Results and Discussion

3.1. TiO₂ Nanotube Layers (TiO₂-NT)

Titanium dioxide nanotube arrays were obtained by anodizing a Ti Foil. **Figure 1** shows FESEM top down images of the TiO₂ nanotube layers (TiO₂-NT). We

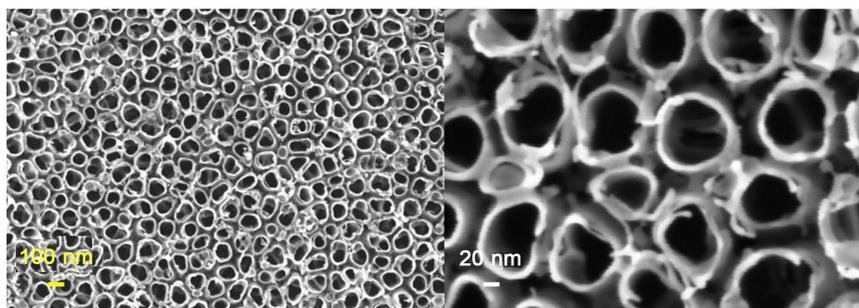


Figure 1. Top down FESEM images of TiO₂ nanotube arrays obtained by anodization at 20 V.

observe the growth of ordered nanotube arrays grown on top of the Ti measuring approximately 84 nm in diameter.

3.2. Gold Nanoparticles in Suspension

The TEM images of the PEG stabilized gold nanoparticles in suspension are shown in **Figure 2**. All the pictures exhibited a quasi-spherical shape, with some aggregates.

Taking this spherical geometry, the average core diameter of the nanoparticles is 7 nm and with a standard deviation of 1 nm.

3.3. UV-Visible Spectroscopy

All spectra were characterized by the Surface Plasmon Resonance (SPR) band (see **Figure 3**). Considering the PEG-stabilized gold nanoparticles in suspension with initial concentration of 1.25×10^{-4} M in tetrachloroaurate (III) hydrate as the starting point, four dilutions are made (volume ratio): 1/2, 1/4, 1/6 and 1/10. The maximum of the SPR band was recorded for each dilution and used to calculate the extinction coefficient ϵ . The maximum absorbance around 520 nm was plotted versus the auric molar concentration, (see **Figure 4**). After checking the required parameters of validity ($A < 3$, no turbidity, monochromatic light), the extinction coefficient ϵ can be obtained from the slope of the linear region of the maximum absorbance-concentration graph, according to Beer Lambert law (Equation (1)):

$$A = \epsilon lc \quad (1)$$

where A is the maximum absorbance of the SPR band, ϵ is the extinction coefficient; l is the length of the optical pathway and c the concentration. Considering the linearity, the extinction coefficient ϵ is found to be 1.8×10^3 L·mol⁻¹·cm⁻¹.

3.4. Calculation of Nanoparticles Concentration in the Suspension

First, the average number of gold atoms per nanoparticle was determined using the average core diameters measured with HRTEM images analysis. Supposing unaltered spherical shape and a uniform Face-centered cubic (fcc) structure, the average number of gold atoms per nanoparticle N was calculated using Equation (2) (all units converted into IS):

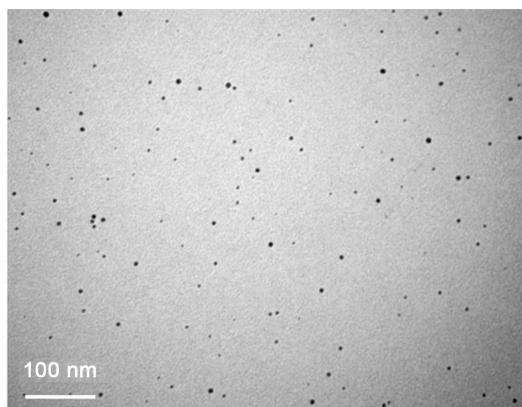


Figure 2. High transmission electron microscopy images of PEG-stabilized gold nanoparticles in suspension with initial precursor concentration of 1.25×10^{-4} M in tetrachloroaurate (III) hydrate.

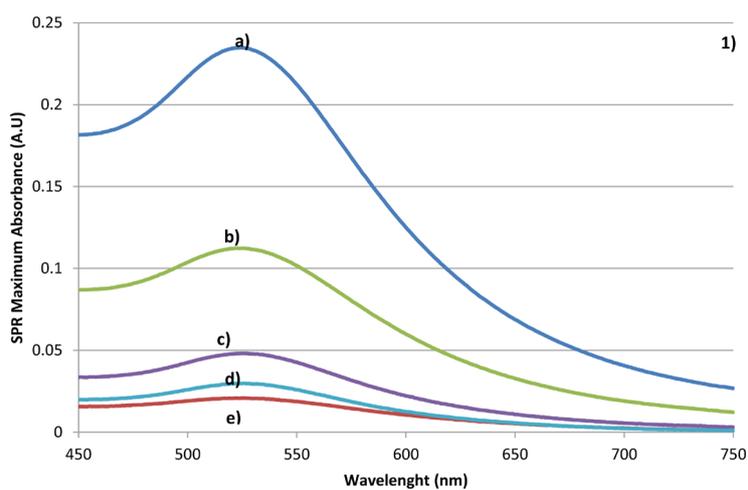


Figure 3. Absorbance spectra of diluted PEG grafted gold nanoparticles suspensions from the suspension with initial concentration of 1.25×10^{-4} M with volume dilution ratio: (a) 1/1 (b) 1/2 (c) 1/4 (d) 1/8 (e) 1/10.

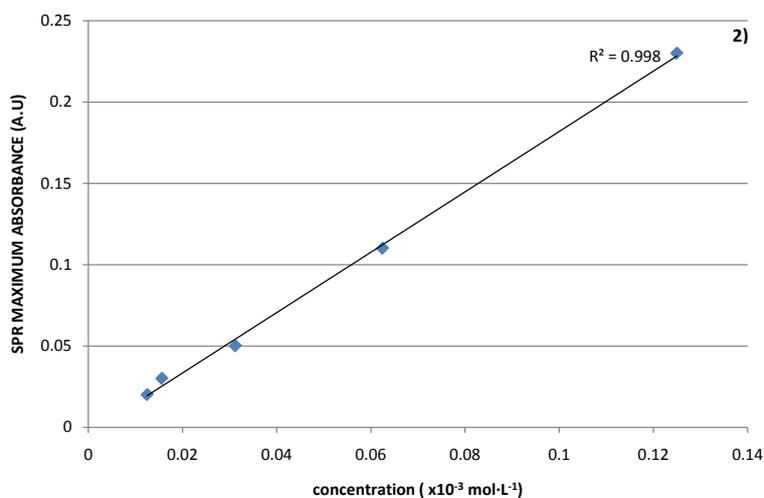


Figure 4. Plot of SPR maximum absorbance versus corresponding gold concentration and linear fitting for diluted PEG grafted gold nanoparticles suspensions.

$$N = \frac{\pi \rho N_A}{6M} D^3 = 3.09 \times 10^{28} D^3 \quad (2)$$

where ρ is the density for fcc gold: $19.3 \text{ g}\cdot\text{cm}^{-3}$, N_A is the Avogadro's constant and M is the molar weight of gold ($197 \text{ g}\cdot\text{mol}^{-1}$).

The nanoparticles concentration for the sample is then calculated by dividing the total number of gold atoms equivalent to the initial amount of chloroauric species introduced (*i.e.* 3.461×10^{19} gold atoms) over the calculated average number of gold atoms per nanoparticle considering a volume of one liter and a total reduction from gold (III) ionic species to gold atoms. The average number of gold atoms per nanoparticle is 10,600 atoms/nanoparticle so the nanoparticle concentration of the sample was 3.26×10^{15} nanoparticles $\cdot\text{L}^{-1}$ with an average core diameter of 7 nm.

3.5. X-Ray Diffraction Analysis

XRD pattern of the analyzed gold nanopowder is presented in **Figure 5**. Intense peaks, relative to Bragg's reflections, could be indexed considering the face-centered cubic gold structure. No diffractions due to crystallographic impurities were found. A strong diffraction peak located at 38.71° was ascribed to the (111) facets of the gold face-centered cubic structure. The other diffraction peaks relative to (200) (220) and (311) facets were much weaker. The broadening of Bragg's peaks indicates the formation of nanoparticles. The mean size of gold nanoparticles was calculated using the Debye Scherrer's equation by determining the width of the (111) Bragg reflection. The crystallite size, D , is evaluated by measuring the width of the curves produced and using the Scherrer Formula (3):

$$D = \frac{0.9 * \lambda}{\beta \cos \theta} \quad (3)$$

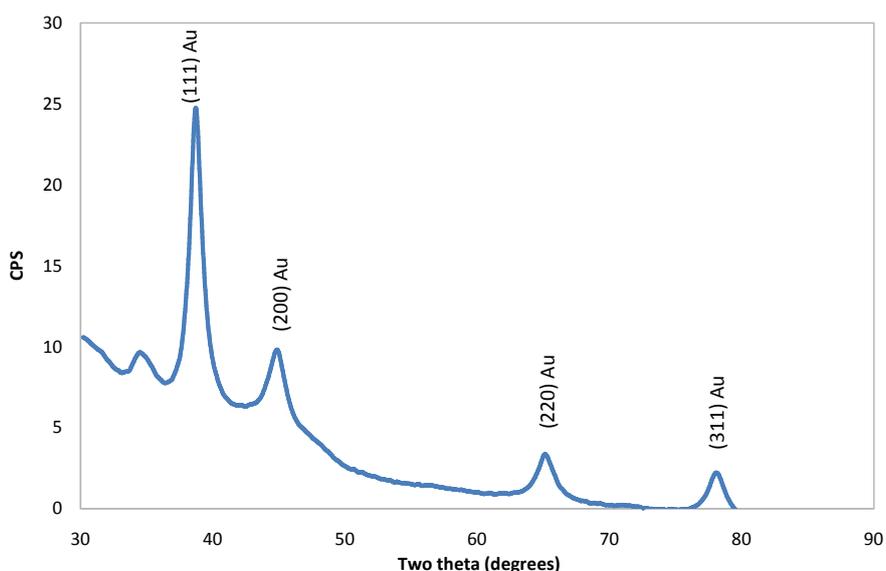


Figure 5. XRD spectrum of the gold nanopowder, the principal Bragg reflections are indexed.

where λ is the wavelength of the Cu K $\alpha 1$ line in \AA , β is the full-width at half max (FWHM) of the peak. The size of the crystallites was found and is approximately 8 nm for gold nanopowder. This result is in good agreement with the average core diameter of 7 nm previously found.

3.6. Gold Doped TiO₂ Nanotubes Layer (Au/TiO₂-NT)

(Au/TiO₂-NT) nanocomposite layers were obtained by immersion of the TiO₂ nanotubes layer into the as-prepared gold nanoparticle suspension. **Figure 6** shows FESEM top down images of TiO₂ nanotube layers doped by gold nanoparticles (Au/TiO₂-NT). We observe inclusion of some gold nanoparticles into the titanium dioxide nanotube layers, creating interfaces between gold and TiO₂ at the nanometric scale.

Chronoamperometry was used to assay the photocurrent generated by the different samples. As seen in **Figure 7**, prompt generation of photocurrent was detected for all annealed TiO₂ samples. The anatase crystallization of the anneal-

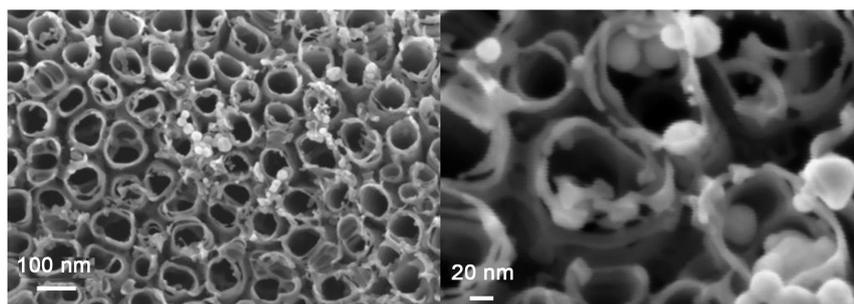


Figure 6. Top down FESEM images of (Au/TiO₂-NT) nanotube arrays: gold nanoparticles entrapped in the TiO₂ matrix.

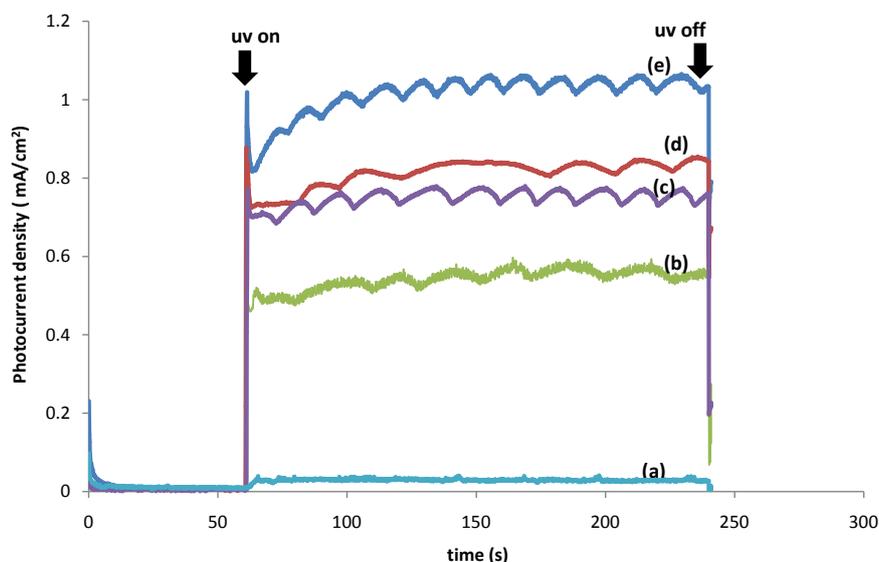


Figure 7. Photocurrent generated under illumination (a) Unannealed TiO₂-NT; (b) Flat TiO₂ annealed at 500°C for 2 hours undoped; (c) gold doped Flat TiO₂ annealed at 500°C for 2 hours; (d) undoped TiO₂-NT annealed at 500°C for 2 hours; (e) Gold doped Au/TiO₂-NT annealed at 500°C for 2 hours.

ed samples at 500°C in air for 2 hours is responsible for this activation because no photocurrent is measured when a pure titanium sample or unannealed TiO₂-NT with the same macroscopic area is assayed. Compared to the flat TiO₂ sample, TiO₂-NT layers exhibited a greater photocurrent density, the amount of TiO₂ anatase species and the rise of the specific area generated by the anodization were responsible for the difference. The photocurrent response for the gold doped samples Au/TiO₂-NT is about 25% higher than for the undoped TiO₂-NT sample. This boosted photocurrent is attributed to Au/TiO₂ interface and specifically to the enhanced potential difference at the noble metal-semiconductor junction. This enhanced potential difference accelerated the photo induced electron-hole separation, resulting in an enhanced photocurrent for the Au/TiO₂-NT samples.

3.7. Kinetics of Photodegradation

These photodegradation tests are used to investigate the influence of the gold nanoparticles on the photocatalytic behaviors of TiO₂ samples. We recorded the decrease of the AO7 absorbance at $\lambda = 468$ nm versus the UV irradiation time. A typical overlay of the spectra is presented in **Figure 8** for an undoped TiO₂-NT layer annealed at 500°C for 2 hours.

Figure 9 summarizes the kinetics curves obtained for the different samples tested. The curve (a) shows the photodegradation of AO7 without TiO₂; AO7 is not degraded by the UV radiation alone. We did not find significant adsorption of the AO7 during the first step of the experiments, i.e., immersion of the different samples in the AO7 under stirring in the dark. Curve (b) illustrates the photodegradation kinetics of AO7 with an undoped TiO₂-NT layer annealed at 500°C. Curve (c) shows the result for AO7 with a gold doped TiO₂-NT layer annealed at 500°C.

The mechanistic pathway of the photodegradation of the dye is based on the presence of oxidative species, such as singlet oxygen (¹O₂) and hydroperoxy

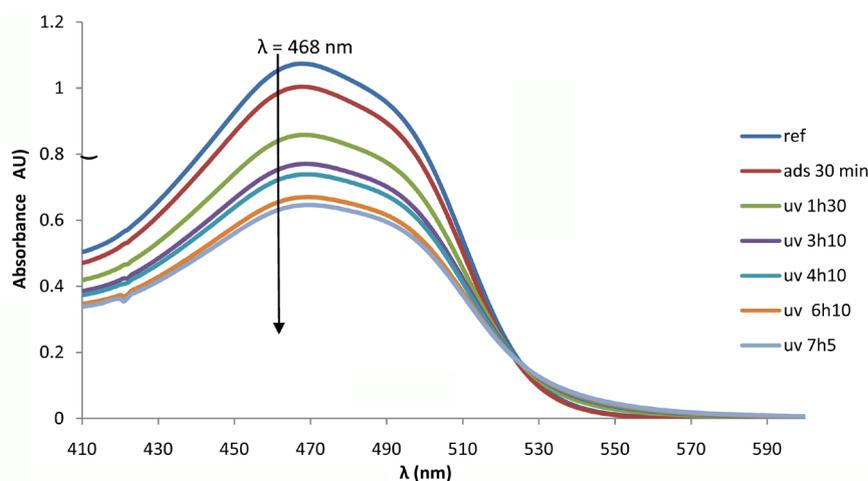


Figure 8. Absorbance versus irradiation time for the Acid Orange Seven (AO7) under irradiation at $\lambda = 365$ nm, in the presence of an undoped TiO₂-NT layer annealed at 500°C for 2 hours.

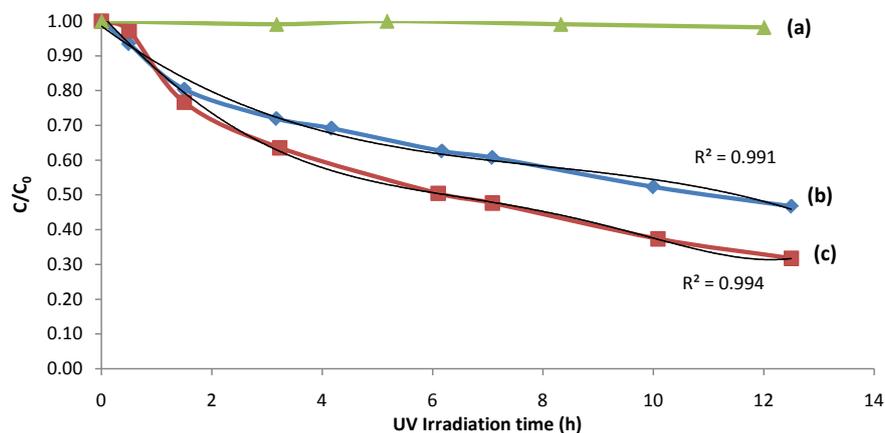


Figure 9. Photodegradation kinetics of Acid Orange 7 (AO7) under UV irradiation at $\lambda = 365$ nm: (a) AO7 without TiO_2 sample; (b) AO7 with undoped TiO_2 -NT layer annealed at 500°C ; (c) Gold doped TiO_2 -NT layer annealed at 500°C .

(HO_2^*) radicals [41]. Primary oxidation products are naphthoquinone derivatives, such as coumarin, 2-naphthol, 1,2-naphthoquinone and 2-hydroxy-1,4-naphthoquinone [42]. Further oxidation can lead to smaller aromatic compounds and eventually, to CO_2 and inorganic ions in the case of a complete mineralization.

Gold nanoparticles embedded in TiO_2 -NT nanotube layers have a significant boosting effect on the photodegradation kinetics of the Acid Orange 7, compared to the result of AO7 with undoped TiO_2 -NT layers. A synergic effect at the interface between the gold nanoparticles and the titanium nanotube layers host gives a rise of the photodegradation rate for the gold doped sample. This increased photocatalytic effect is relative to a better hole-electron separation so higher amounts of Reactive Oxygen Species (ROS) were produced and in that case photodegradation kinetic of AO7 was accelerated.

4. Conclusion

We examined the doping effect of gold nanoparticles in suspension, embedded in TiO_2 -NT layers using a simple immersion technique. Enhanced photocurrent density was measured in relation to the gold doped TiO_2 -NT samples. Noble metal semiconductor interface promoted higher photocurrent density than the undoped TiO_2 -NT samples. A local potential difference at the noble metal-semiconductor junction promoted the hole-electron separation. This enhanced photoactivity of the gold doped samples was brought to light using Acid Orange 7 photodegradation tests. A higher rate of photodegradation of the organic dye was obtained for the gold doped samples. A small amount of gold nanoparticles, well dispersed in the nanotubes was sufficient to ensure a boosted photoactivity of anatase crystallized TiO_2 -NT layers. This doping effect can open the way to new nanomaterial with improved photoactivity.

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