

Coacervation Microencapsulation of CaCO₃ Particles with a Fluoropolymer by Pressure-Induced Phase Separation of Supercritical Carbon Dioxide Solutions

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ABSTRACT

We report a method for the coacervation micro-encapsulation of several forms of CaCO₃ microparticles with the fluoropolymer poly(heptadecafluorodecyl acrylate) (poly (HDFDA)) by pressure-induced phase separation of a supercritical CO₂ solution. A suspension of CaCO₃ in CO₂ and dissolved poly(HDFDA) were mixed in supercritical CO₂. After the system pressure was slowly decreased to atmospheric pressure, the microcapsules were obtained. Coacervation was achieved by the precipitation of poly(HDFDA) during the decrease in the pressure of CO₂; the solubility of poly(HDFDA) in CO₂ decreased with the pressure. The structure and morphology of the microparticles were investigated by using a scanning electron microscope (SEM) and an electron probe micro-analyzer (EPMA) equipped with a wavelength dispersive X-ray spectroscopy (WDX).

Keywords: Component; Supercritical Carbon Dioxide; Microencapsulation; Coacervation; Fluoropolymer; Calcium Carbonate

1. Introduction

Polymer microcapsules containing inorganic materials are attracting much attention as the field of supercritical CO₂ (scCO₂) technology. ScCO₂ is the solvent of choice because it is readily available, inexpensive, and environmentally benign. Many investigators have attempted the formation of polymer microcapsules using scCO₂ [1-6]. Rapid expansion from supercritical solutions (RESS) is a well-known process, and a variety of polymer microcapsules have been produced with the help of this process by many investigators [2,3,5-9]. However, the RESS process is limited by the low polymer solubility in CO₂, caused by its low dielectric constant. Relatively few polymers are soluble in CO₂ without a cosolvent. RESS of fluoropolymers such as perfluoropolyether, poly(1,1,2,2-tetrahydroperfluorodecyl acrylate), and poly (heptadeca-fluorodecyl acrylate), which are highly soluble in CO₂ at temperatures near the ambient temperature, produces coating materials [10-12] and submicron to several micron-sized particles and fibers [12,13].

In this work, we try to form microcapsules of CaCO₃ and poly(heptadecafluorodecyl acrylate) (poly (HDFDA)) using scCO₂. In a previous work[9], we proposed a production method for the fluoropolymer microcapsules of talc particles by pressure-induced phase separation of scCO₂. **Figure 1** provides a conceptual framework of our proposed process in comparison with the conventional RESS process.

In RESS, a supercritical fluid solution is expanded across a nozzle, leading to rapid supersaturation and the production of small particles. After a suspension of CaCO₃ in CO₂ containing a dissolved fluoropolymer is sprayed through the nozzle at

atmospheric pressure, microcapsules and small polymer particles are obtained as shown in **Figure 1(a)**. For the industrial applications, we have to restrict the generation of polymer particles not containing CaCO₃ because they degrade the products. Therefore, to prevent the nucleation and the precipitation of polymer particles not containing CaCO₃, the pressure is decreased slowly, and microparticles are collected in the high-pressure cell as shown in **Figure 1(b)**.

The objective of this work is to check the feasibility of the pressure-induced phase separation of the scCO₂ solution to the formation of fluoropolymer microcapsules of several shapes of particles of CaCO₃ and to study the effect of several experimental conditions on particle morphology.

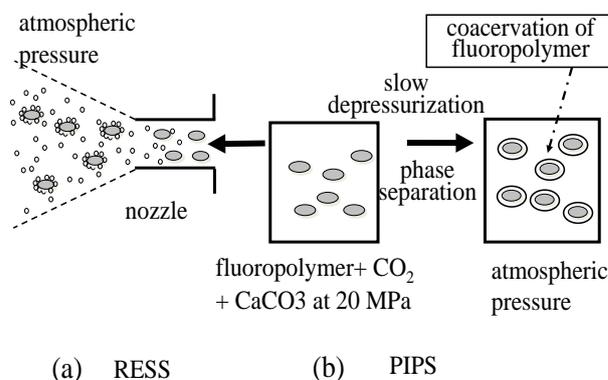


Figure 1. Principles of the formation of polymer microcapsules of CaCO₃ by (a) RESS and (b) pressure-induced phase separation of scCO₂ solutions.

2. Experimental Section

2.1. Materials

CaCO₃ was obtained from Shiraish Calcium.Co., Ltd., and carbon dioxide (CO₂) (99.9% minimum purity) was purchased from Fukuoka Sanso Co., Ltd. The fundamental idea and synthesis of poly(HDFDA) was reported by DeSimone et al. [27], and a similar approach based on their method is employed in the present study. The fluoropolymer poly(HDFDA) was synthesized in a high-pressure cell by the free-radical polymerization of a homogeneous solution of the 3,3,4,4,5,5,6,6,7,7,8,8,9,9,10,10,10-heptadecafluorodecyl acrylate (HDFDA) monomer with an azobis(isobutyronitrile) (AIBN) initiator in CO₂ for 48 h at 333 K and 20 MPa. AIBN and HDFDA were purchased from Aldrich Co. Upon completion of polymerization, the polymer was precipitated from CO₂ directly into a methanol bath. Subsequently, the poly(HDFDA) was washed several times and allowed to dry overnight.

2.2. Experimental Procedure

Known amounts of the fluoropolymer and CaCO₃ were placed in the high-pressure cell (25 cm³) equipped with sapphire windows. The cell was placed in a water bath and the system temperature was maintained at the desired value within +0.1 K. CO₂ was pumped through a preheater to the high-pressure cell. The mixture was stirred by a magnetic agitator for 30 min. The system was slowly depressurized for approximately 30 min at the experimental temperature. Following the decrease in pressure, polymer microcapsules were obtained in the high-pressure cell.

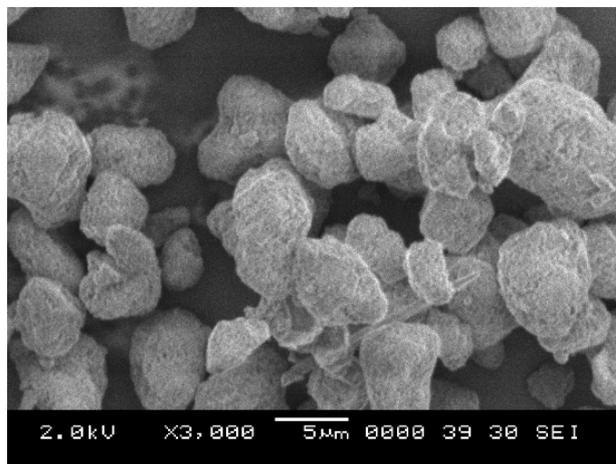
The structure and morphology of the products were analyzed using a scanning electron microscope (SEM, JEOL JSM6060) and an electron probe microanalyzer (EPMA; Shimadzu, EPMA 1610) equipped with a wavelength dispersive X-ray spectrometer (WDX). An EPMA equipped with WDX can identify elements through the use of a crystal monochromator to select X-rays of a particular wavelength. For the SEM sample preparation, polymeric microparticles were mounted on a small glass plate covered with a small piece of double-sided carbon conductive tape. The samples were then sputter-coated with silver palladium and imaged using the SEM and EPMA.

3. Results and Discussion

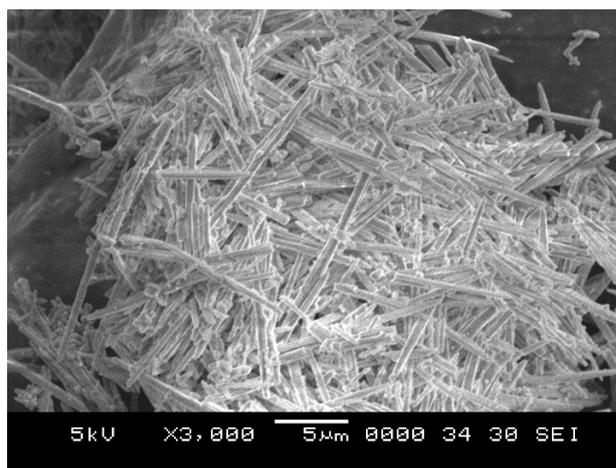
3.1. Evolution of Microencapsulation

Prior to the experiment for microcapsule formation, the phase behavior of the CO₂ + poly(HDFDA) system at 20 MPa and 313 K was confirmed visually by using a high-pressure vessel equipped with sapphire windows. Without the CaCO₃, the mixtures of CO₂ and poly(HDFDA) form a single phase. Details of the phase behavior of the CO₂ + poly(HDFDA) system were reported by Blasig et al. [12] Similar phase behaviors for CO₂ + poly(1,1-dihydroperfluorooctylacrylate) [14] and CO₂ + poly(1,1,2,2-tetrahydroperfluorodecyl acrylate) [13] systems were reported.

SEM photographs of the CaCO₃ and the fluoropolymer microcapsule containing CaCO₃ that was produced by the pressure-induced phase separation of scCO₂ are shown in **Figures 2(a)** and **(b)**.



(a)



(b)

Figure 2. SEM photographs of poly(HDFDA) microcapsules of (a) spheres and (b) whiskers of CaCO₃ particles formed by the pressure-induced phase separation of scCO₂ solutions. Pre-expansion conditions: temperature, 313 K; pressure, 20 MPa; CO₂, 97.9 wt%; poly(HDFDA), 0.20 wt%; CaCO₃, 2.1 wt%.

The system was slowly depressurized from 20 MPa to atmospheric pressure for approximately 30 min at 313 K. The spherical particles of CaCO₃ and CaCO₃ whiskers had a smooth surface. Compared with the SEM photographs of the CaCO₃, the microcapsules of the fluoropolymer containing CaCO₃ have a similar configuration. The surface morphology of the microcapsules reflects the configuration of CaCO₃ in the microcapsules because the coating thickness of CaCO₃ is very small. The primary particle diameter (PPD) and particles size distribution (PSD) of CaCO₃ and microcapsules were determined by a laser diffraction particle size analyzer (SALD-2000, Shimadzu Co. Ltd.).

The PPD and PSD of spherical particles of CaCO₃ are 7.6µm and 0.40, respectively. And the PPD and PSD of microcapsules are 7.7µm and 0.403, respectively. The value of PPD and PSD of the spherical particles of CaCO₃ and microcapsules is almost same. We can not observe the change of particle size.

The CaCO₃ whiskers were also coated by the fluoropolymer. The surface morphology of the microcapsules reflects the

configuration of CaCO_3 whiskers in the microcapsules because the coating thickness of CaCO_3 is very small. But structure of CaCO_3 whiskers coated by the fluoropolymer were more bulky than CaCO_3 whiskers.

Further evidence for the formation of fluoropolymer microcapsules of CaCO_3 can be obtained using EPMA. The peak corresponding to F caused by the fluoropolymer can be observed for the microcapsules, it cannot be detected for CaCO_3 because CaCO_3 does not possess F.

The surface distributions of F, O, and Ca were mapped in an EPMA image. Although the distribution of F in the microcapsules was fairly sharp, it was not detected on the CaCO_3 surface. On the other hand, the distribution of Ca and O on the CaCO_3 surface was sharper and broader. However, the distribution of Ca and O on the microcapsule surface was poorer than that on the CaCO_3 surface. It can be considered that CaCO_3 was completely encapsulated by a thin fluoropolymer film.

It was difficult to check the coating performance for all the collected microcapsules by using EPMA because in the proposed process, an extremely large number of microcapsules were produced. To evaluate the performance of the polymer coating, we examined the stability of the microcapsules in pure water. The CaCO_3 particles or microcapsules were added to pure water (particle concentration: 1 wt%), and the suspended solution was shaken by a mechanical shaker. The stable conditions of the spherical particles of CaCO_3 and microcapsules in water were checked. Although the CaCO_3 was dispersed in pure water for more than 5 min, all the microcapsules floated on water because of the high water repellency of the fluoropolymer. The density of CaCO_3 and microcapsules is almost same (about $2.8 \text{ g} \cdot \text{cm}^{-3}$), because microcapsules contain more than 90 % CaCO_3 . Although the density of microcapsules is higher than that of water, the microcapsules floated on the water. It is inferred that bulk density of microcapsules is lower than that of water. It is difficult to penetrate the water to the void between the microcapsules, because of the repellency of fluoropolymer.

The CaCO_3 was dispersed in water, because the CaCO_3 has hydrophilic surfaces. To check the stability of the microcapsules in pure water, a turbidity measurement was performed using an ultraviolet/visible (UV/VIS) spectrometer at 600 nm wavelength. The turbidity measurement was used to observe the stability of small particle dispersions [29]. We could not observe the dispersed particles through the stability analysis of microcapsules in water because as in the case of pure water, no turbidity was observed. The stability analysis revealed that most of the CaCO_3 particles were coated with the fluoropolymer and were present inside the produced microcapsules.

3.2. Formation Mechanism of Microcapsules

To identify the advantage of the formation mechanism of microcapsules by the pressure-induced phase separation of scCO_2 as compared with RESS, the microcapsules were prepared by RESS. Because RESS is one of the promising methods for the formation of polymer microcapsules and/or composites by using scCO_2 , several investigators have reported the formation of polymer microcapsules and/or composites by RESS [1,2]. The particle formation mechanism by RESS was analyzed thermodynamically [4]. In this work, we attempted the formation of

microcapsules by RESS under the following experimental conditions. The pre-expansion pressure was 20 MPa, and the temperature was 313 K. The feed concentrations of the CaCO_3 and the fluoropolymer were 2.1 wt% and 0.20 wt%, respectively. The feed composition in the RESS experiment was the same as that in the experiment on the formation of microcapsules by the pressure-induced phase separation of scCO_2 . The mixtures of scCO_2 , the fluoropolymer, and the CaCO_3 were expanded across the capillary nozzle ($L = 500 \text{ mm}$, $D = 1.2 \text{ mm}$) to atmospheric pressure. After the expansion, the microparticles were precipitated. SEM photographs of the fluoropolymer microcapsules produced by RESS and containing CaCO_3 were obtained. Compared with the morphology of microcapsules prepared by the pressure-induced phase separation of CO_2 as shown in **Figure 2**, the polymer particles prepared by RESS were observed on the surfaces of the CaCO_3 particles. The polymer does not form a smooth surface at the CaCO_3 particles but is adhered as small particles at the surface of the CaCO_3 .

To examine the coating performance of RESS, the obtained particles were analyzed by EPMA and by performing a stability test in water. F, Ca, and O were detected in the WDX spectrum of the microcapsules. Furthermore, we examined the stability of the microcapsules in pure water to evaluate the performance of the polymer coating. The WDX spectrum and the stability test revealed that most of the CaCO_3 was microencapsulated with the fluoropolymer. However, small polymer particles were precipitated on the surface through the RESS process. The formation mechanism of microcapsules and small polymer particles in the RESS process may be considered as follows. During rapid depressurization both the CaCO_3 and the polymer precipitate from the solutions. And the CaCO_3 particles are formed in the expanding jet. Some polymer coated on the CaCO_3 particles, and some fine polymer particles are generated during the deposition. The evidence for the formation of fine polymer particles by RESS can be obtained by performing the RESS experiment without CaCO_3 . The mean particle diameter

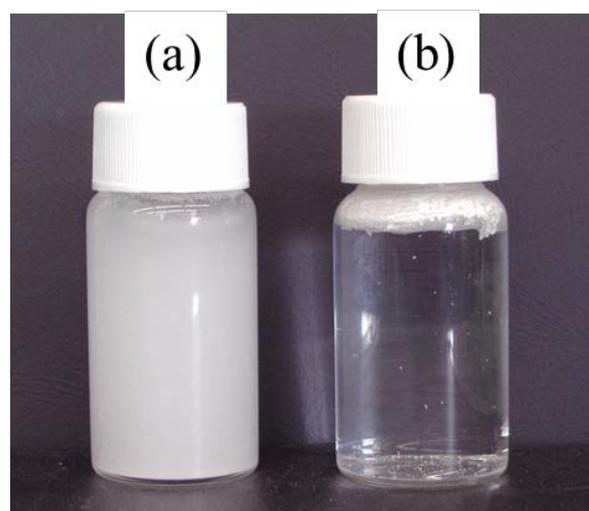


Figure 3. Stability of microcapsules in pure water. (a) CaCO_3 and (b) poly(HDFDA) microcapsules formed by the pressure-induced phase separation of scCO_2 solutions. See Figure 2 for the pre-expansion conditions.

less than 1 μm . With regard to the RESS experiment for the formation of fluoropolymer particles, similar particle morphology was reported by Blasig et al. [12] and Mawson et al. [13]. These fine polymer particles precipitated on and adhered to the CaCO_3 surface by the supersaturation and homogeneous nucleation of the fluoropolymer that was caused by rapid depressurization. To prevent the formation of polymer particles, we have to inhibit the supersaturation of the solute and the homogeneous nucleation caused by the rapid expansion of CO_2 . However, it is impossible to prevent the supersaturation in RESS. We can prevent the formation of polymer particles by the pressure-induced phase separation of CO_2 . Because the depressurizing rate is very slow compared with the conventional RESS process, it is possible to inhibit the large supersaturation of the solute and the homogeneous nucleation of particles. During the slow depressurization, the coacervation was achieved. On the other hand, after the pressure in the high-pressure cell containing no CaCO_3 decreased, polymer foams were obtained. With experimental setup, no pure fluoropolymer were formed. It is inferred that the CaCO_3 suspended in scCO_2 acts as an accelerator for the precipitation of polymer particles and the occurrence of coacervation on the CaCO_3 surface. Furthermore, it is very difficult for the microcapsules to produce forms, because the microcapsules contain about 90 % CaCO_3 . In the conventional coacervation microencapsulation technique, coacervation is induced by a phase separation caused due to a pH change and the addition of a nonsolvent or electrolyte [16]. In contrast, in the present experiment, coacervation was induced by a phase separation caused by a decrease in pressure.

4. Conclusions

The pressure-induced phase separation of scCO_2 has been utilized to produce fluoropolymer microcapsules of several shape particles of CaCO_3 . Prior to depressurization, the polymer and CaCO_3 were mixed in scCO_2 . Fluoropolymer coacervation was achieved during the slow decrease in the pressure. Following the coacervation, we obtained the fluoropolymer microcapsules of CaCO_3 . The products were analyzed by SEM and EPMA equipped with WDX. The CaCO_3 was completely coated with a thin fluoropolymer film. Compared with the microcapsules formed by RESS, the obtained microcapsules had a smooth surface; fine polymer particles on the CaCO_3 surface were not observed.

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