

# Characterization and Evaluation of Standard Enthalpy of Vaporization and Kinetic Studies of Volatile bis(*N*-ethyl-5-methyl-salicylaldimine)nickel (II) Complex

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## Abstract

The bis(*N*-ethyl-5-methylsalicylaldimine)nickel(II) [Ni(5-me-saletN)<sub>2</sub>] complex was synthesized and characterized by elemental analyses, FT-IR, TG-DTA, mass spectrometry and vapour pressure measurement studies. The TG curve of the complex showed a single-step weight loss commencing from 490 K to nil residue at 600 K, without competing fragmentation step. The non-isothermal vaporization activation energy value determined by Coats-Redfern method yielded the value of  $93.5 \pm 7 \text{ kJmol}^{-1}$ . The dynamic TG run proved the complex to be completely volatile. And the equilibrium vapour pressure  $(p_e)_T$  of the complex over the temperature range of 421 - 524 K, determined by the TG-based transpiration technique yielded the value of  $94.2 \pm 1.2 \text{ kJmol}^{-1}$  for its standard enthalpy of vaporization  $(\Delta_{\text{sub}}H^\circ)$ . The entropy of vaporization  $(\Delta_{\text{sub}}S^\circ)$  was calculated from the intercept and found to be  $249.4 \pm 2.6 \text{ Jmol}^{-1}\cdot\text{K}^{-1}$ .

## Keywords

Nickel Complexes, Enthalpy of Vaporization, Thermogravimetry, Vapour Pressure, Kinetic Studies

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## 1. Introduction

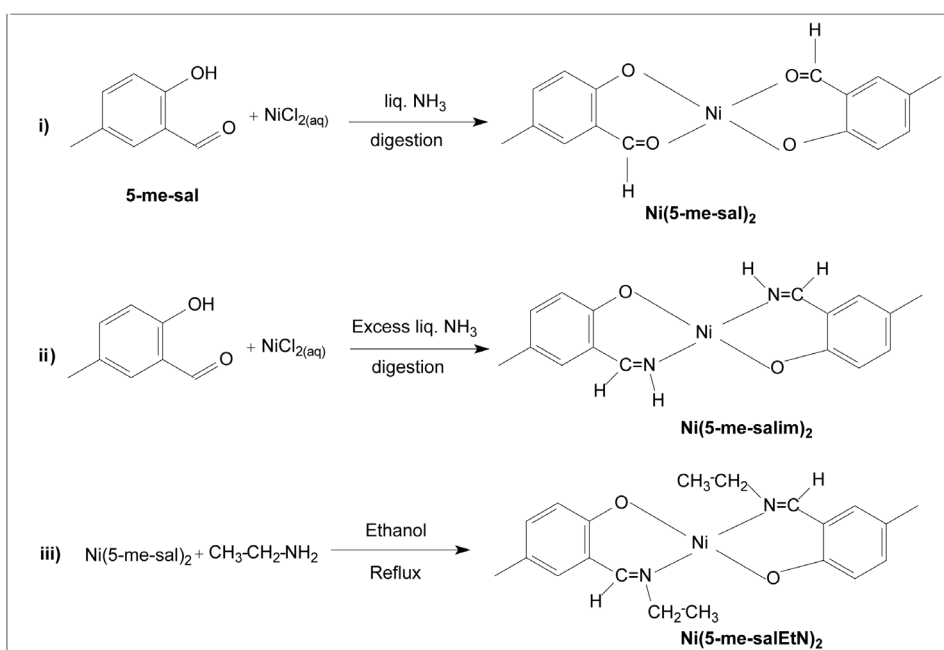
Nickel films are widely used in the metallization of ferrites and as decorative and corrosion-resistant coatings. Ni and its alloy (NiSi) are also promising materials for microelectronic devices because the electrical resistivity is low and the oxidation resistance is high [1] [2]. Therefore, chemical vapour deposition (CVD) for Ni or its alloy film has been widely investigated by researchers of microelectronics and nanotechnology, because CVD is the method to deposit film with good-quality step coverage on patterned structures [3]-[5]. Ni precursor is tetracarbonylnickel(0) Ni(CO)<sub>4</sub> [6], whose boiling point is 43°C and volatility is very useful as a precursor. However, it has the critical problem that it is too highly toxic. Therefore, an alternative precursor is strongly required. Until now, several Ni precursors such as a glyoximate [7], β-diketonato [8], or cyclopentadienyl nickel [9] [10] were utilized as an alternative. All of them are solids below 150°C, and their volatilities are far lower than that of Ni(CO)<sub>4</sub>. In CVD, using one of them as a precursor, a high deposition temperature was needed. This paper presents the development and characterization of bis(*N*-ethyl-5-methyl-salicylalimine)nickel(II) volatile precursor of nickel. The original result on this complex is presented here in order to assess the enthalpy of vaporization. This work focuses on the rapidity with which a reliable equilibrium vapour pressure could be determined with the horizontal TG/DTA thermoanalyser in a time frame which is nearly one-tenth of that required for a similar measurement with Knudsen cell mass spectrometry (KCMS) or conventional transpiration with suitable effluent gas analyser. The thermal stability and vaporization kinetics of bis(*N*-ethyl-5-methyl-salicylalimine)nickel(II) under non-isothermal and isothermal TG methods are also presented.

## 2. Experimental Section

### 2.1. Materials and Development of Precursors

The chemicals liquid ammonia (Qualigens, India), NiCl<sub>2</sub>·6H<sub>2</sub>O (purity—97%), ethylamine (purity—70%, Loba Chemie, India), ethanol (purity—99.9%, S.D. Fine, India), 5-methyl-salicylaldehyde (purity—99%, Merck, Germany) were used as such in the syntheses of nickel complex.

The volatility of the nickel precursor was improved by altering the ligand chemistry of the complex. The polarity of the metal-oxygen bond responsible for the ionicity of the complex could be reduced by substituting nitrogen for oxygen in the carbonyl group (C=O). Based on these considerations, the second nickel complex bis(5-methyl-salicylalimine)nickel(II), Ni(5-me-salim)<sub>2</sub> was prepared and was subsequently characterized. At this juncture it was decided to substitute H of =NH as shown in (III) of **Scheme 1** by the ethyl group, thereby



**Scheme 1.** Graphical representation of the Ni(5-me-salEtN)<sub>2</sub>.

tuning the volatilization behaviour. This complex was considered as an ideal precursor for the CVD of metallic nickel and further characterization as well as the vapour pressure measurements was deemed to be essential, as it is imperative to know the vapour pressure of the precursor for an accurate control of the feed rate to the CVD reactor.

## 2.2. Characterization

The FT-IR spectrum was recorded using a Perkin-Elmer FT-IR spectrometer (RX1, FT-IR) in the region of 4000 - 400  $\text{cm}^{-1}$ . The C, H and O analyses were carried out by a CARBO-ERBA-11008 CHNS-O rapid elemental analyzer. The electron spray ionization mass spectra of the complex was recorded on a MICRO MASS QUATRO II triple quadrupole spectrometer having a JASCO PU-980 HPLC pump connected to it.

## 2.3. Non-Isothermal TG

The thermogram of the complex was carried out with Perkin-Elmer Pyris-Diamond TG-DTA. The thermogravimetric analysis was performed at a linear heating rate of 10  $\text{Kmin}^{-1}$  over the temperature range of 420 - 520 K. The temperature calibration was done by the method of fixed melting points by using International Practical Temperature Scale 1968 (IPTS-68; amended in 1975) recommended standards for indium, tin, and aluminum. Approximately 3 mg of the complex was taken for the experiment. A sintered high-density alumina crucible was used as sample and reference holders, and  $\alpha$ -alumina powder was used as the reference material. The purge gas was high purity nitrogen dried by passing through refrigerated molecular sieves (Linde 4A) at a flow rate of 12  $\text{dm}^3\cdot\text{h}^{-1}$ .

## 2.4. Isothermal Vapour Pressure Measurements

A horizontal thermal analyser was adopted as transpiration set up for vapour pressure measurements. The block diagram of the thermal analyser, modified for its functioning in the transpiration mode is given elsewhere [11]. The details of the precise flow calibration for the carrier gas, using a capillary glass flow meter and corrections for apparent weight losses in isothermal mode, are described elsewhere [11] [12]. The sample, finely powdered using a mortar and pestle was spread on an alumina crucible and was carefully flushed with nitrogen gas at the rate of 6  $\text{dm}^3\cdot\text{h}^{-1}$  at ambient temperature. Vapour pressure measurements were carried out at 10  $\text{Kmin}^{-1}$ . After the temperature stabilization, subsequent changes in isothermal steps were done at a heating rate of 2  $\text{Kmin}^{-1}$ .

## 2.5. Vaporization Kinetics

The conventional non-isothermal thermogravimetric run was carried out at the heating rate of 10  $\text{Kmin}^{-1}$ . Among the several methods available for the kinetics evaluation of TG weight loss data Coats-Redfern was followed in the present paper to study the vaporization kinetics.

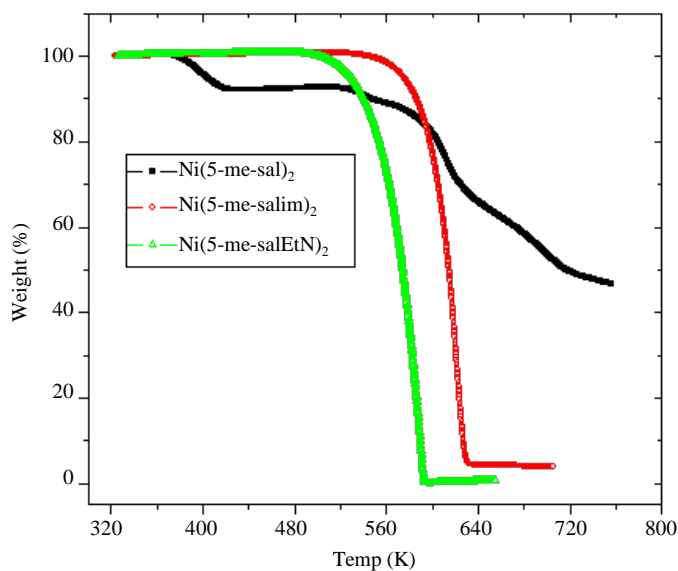
# 3. Results and Discussion

## 3.1. Characterization

The C and H analyses of the complex confirmed the assigned composition of the calculated values. The C, H and Ni analyses are provided. Anal. Calcd. for  $\text{Ni}(\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2)$  [ $\text{Ni}(\text{5-me-salEtN})_2$ ]: C, 62.66 (62.53); H, 6.26 (6.23) and Ni, 15.14 (15.18). The FT-IR absorption spectral bands of [ $\text{Ni}(\text{5-me-salEtN})_2$ ] complex observed at 1612, 1543, 1328, 524, 484  $\text{cm}^{-1}$  are assigned to  $\nu_{(\text{C}=\text{N})}$ ,  $\nu_{(\text{C}=\text{C})}$ ,  $\nu_{(\text{phenolic C-O})}$ ,  $\nu_{(\text{Ni-N})}$  and  $\nu_{(\text{Ni-O})}$  respectively.

## 3.2. Thermal Properties of bis(5-methyl-salicylaldehydato)nickel(II) [ $\text{Ni}(\text{5-me-sal})_2$ ]

The bis(5-methyl-salicylaldehydato)nickel(II) [ $\text{Ni}(\text{5-me-sal})_2$ ] Complex was characterized by thermogravimetry (TG) for its melting, vaporization, decomposition temperatures and volatility. However, the thermogram as shown in **Figure 1** indicates the loss of two water molecules after the first step at about 417 K, and 35% residue after the second weight loss step at about 620 K; subsequent decomposition of the complex gives about 46.68% residue at the temperature of about 756 K, making it unsuitable for CVD applications. This clearly shows that



**Figure 1.** Non-isothermal TG curves of bis(5-methyl-salicylaldehyde)nickel(II); Ni(5-me-sal)<sub>2</sub>, bis(5-methyl-salicylaldimine)nickel(II); Ni(5-me-salim)<sub>2</sub> and bis(N-ethyl-5-methyl-salicylaldimine)nickel(II); Ni(5-me-salEtN)<sub>2</sub>.

the nickel atom has been retained in the residue and not transferred into the vapour phase as evidenced by the black end residue analysis.

### 3.3. Thermal Properties of bis(5-methyl-salicylaldimine)nickel(II) [Ni(5-me-salim)<sub>2</sub>]

The TG trace of this complex shown in **Figure 1** indicates a residue of 3.8% at 705 K after a smooth single step weight loss. The interesting inference was the enhanced thermal stability and volatility as well as the expulsion of two H<sub>2</sub>O molecules caused by the coordination of azomethine nitrogen (–CH=NH) with Ni instead of carbonyl oxygen.

### 3.4. Thermal Properties of bis(N-ethyl-5-methyl-salicylaldimine)nickel(II); [Ni(5-me-salEtN)<sub>2</sub>]

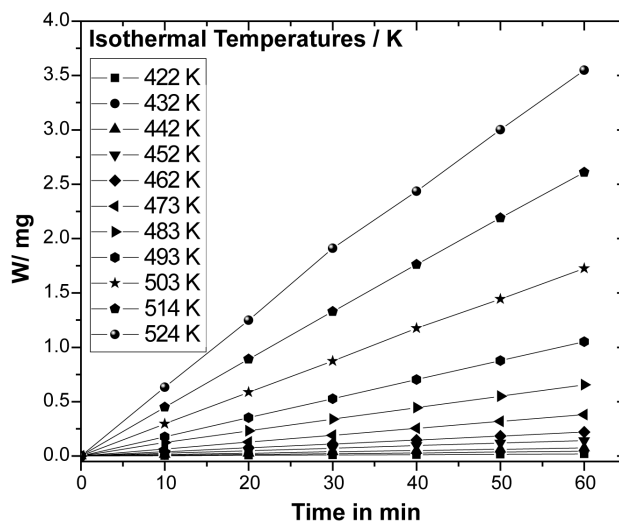
The non-isothermal TG run of the complex (**Figure 1**) revealed that the complex was thermally stable up to 490 K. The weight loss, commencing from about 490 K, occurred in a single step and ended up with total weight loss around 580 K. The fact that the complex showed a single step weight loss ending up with nil residues helped to identify the complex as a potential precursor for CVD of nickel films. As the objective of this investigation was to evaluate the equilibrium vapour pressure by a TG-based transpiration technique, *a priori* knowledge of the molecular weight of the complex was required. For this purpose, the Ni(5-me-salEtN)<sub>2</sub> was subjected to ESI-MS analysis and its molecular mass was determined.

### 3.5. Determination of Standard Enthalpy of Vaporization ( $\Delta_{\text{sub}}H^\circ$ ) of Ni(5-me-salEtN)<sub>2</sub>

The equilibrium Vapour pressure ( $(p_e)_T$ ) could be calculated using the Equation (1),

$$(p_e)_T = WRT/MV_c \quad (1)$$

Which is derived from Dalton's Law of partial pressures for a mixture of ideal gases, where  $W$  is the mass loss (mg) at an isothermal temperature caused by the flow of  $V_c$  (dm<sup>3</sup>) of the carrier gas (measured at  $T_c = 298$  K),  $R$  is the universal gas constant (8.314 Jmol<sup>-1</sup>·K<sup>-1</sup>) and  $M$  (383.11024 g·mol<sup>-1</sup>) is the molecular mass of the monomeric precursor vapour species. Attainment of equilibrium condition was also ascertained by the plot of isothermal time against mass loss for all the isothermal steps as shown in **Figure 2** and the lines passing through origin show that equal amount vaporized at equal intervals of time. The values of observed



**Figure 2.** Mass loss against time of 1 h isothermal holding.

mass loss, isothermal temperatures and the calculated equilibrium vapour pressure ( $p_e$ ) are listed in **Table 1**. The plot of  $\log(p_e/mP_a)$  against  $1000/T$  (K) is shown in **Figure 3**. The vapour pressure ( $p_e$ ) could be represented by the least square expression (2),

$$\log(p_e/mP_a) = 13.0263 \pm 0.1370 - 4.9206 \pm 0.0643/T \text{ (K)} \quad (422 - 524 \text{ K}) \quad R^2 = 0.9991 \quad (2)$$

Multiplying the slope of the expression by  $-2.303 R$ , a value of  $94.2 \pm 1.2 \text{ kJmol}^{-1}$  was derived for the standard enthalpy of vaporization ( $\Delta_{\text{sub}}H^\circ$ ). The entropy of vaporization ( $\Delta_{\text{sub}}S^\circ$ ) calculated from the intercept was found to be  $249.4 \pm 2.6 \text{ Jmol}^{-1}\cdot\text{K}^{-1}$ .

### 3.6. Determination of Activation Energy ( $E_a$ ) of Ni(5-me-salEtN)<sub>2</sub>

The rate constant  $k$  for the vaporization enthalpy of the complex was determined in the temperature range of 480 – 580 K for every 10% weight loss of the complex. The expression for  $k$  is given by

$$k = d\alpha/dt \quad (3)$$

where  $d\alpha/dt$  is the derivative of the fraction vaporized with respect to time and  $k$  is the rate constant of vaporization. By using this equation,  $k$  was calculated for every 10% weight loss,  $\alpha$  is defined by the expression (4) as

$$\alpha = \frac{\%W_t - \%W_i}{\%W_t - \%W_f} \quad (4)$$

where  $\%W_t$  is the percent weight at any time  $t$  and  $\%W_i$  and  $\%W_f$  respectively, are the initial and final per cent sample weights [13].

From the mass loss and the TG/DTA experiments,  $\alpha$ , the fraction reacted was evaluated as a function of temperature. A typical plot of  $\alpha$  versus temperature is shown in **Figure 4**.

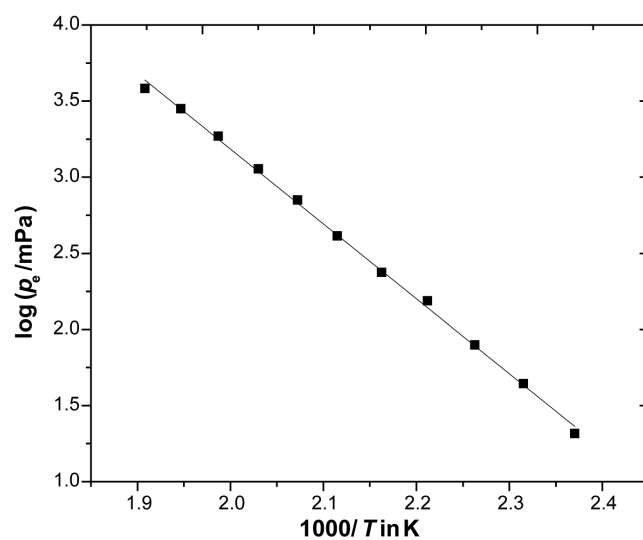
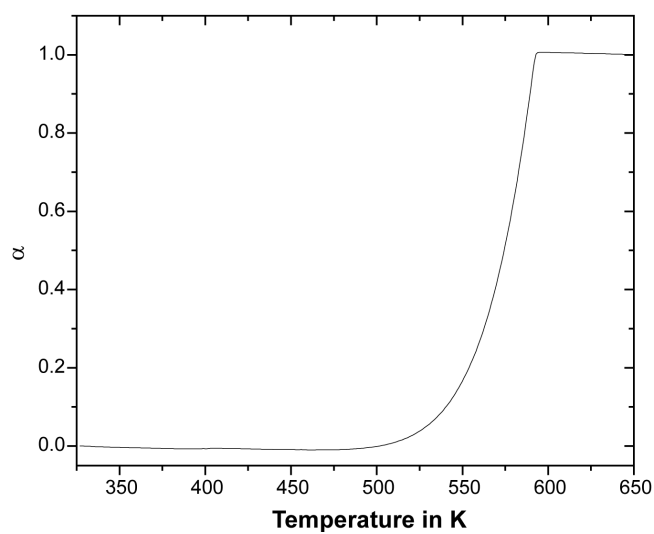
The kinetics of the complex was followed by employing the Coats-Redfern approximation which gives the expression (5)

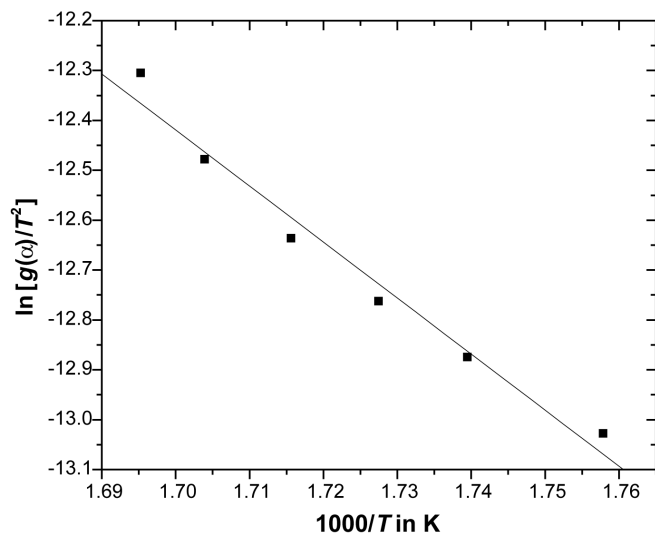
$$\ln \left[ \frac{g(\alpha)}{T^2} \right] = \left[ \frac{AR}{\beta E} \left( 1 - \frac{2RT}{E} \right) \right] \frac{E}{RT} \quad (5)$$

The  $g(\alpha)$  function describes the mechanism [14] of the reaction. Straight line with high-correlation coefficient and low standard deviation were selected to represent the possible mechanism. Therefore, the best fit was obtained with A2 (Avrami-Erofe'ev model) which corresponds to nucleation and growth mechanism. A plot of  $\ln [g(\alpha)/T^2]$  versus  $1000/T$  (K) gives a straight line when the correct  $g(\alpha)$  function is used in the equation

**Table 1.** Isothermal vapour pressure of Ni(5-me-salEtN)<sub>2</sub>.

Isothermal Vapour Pressure		
<i>T</i> /K	Mass Loss (mg)	<i>p</i> <sub>v</sub> (MPa)
421.91	0.0192	20.7065
431.95	0.0408	44.0014
441.93	0.0732	78.94369
452.11	0.1427	153.8971
462.42	0.2193	236.5076
472.84	0.3808	410.6797
482.55	0.6564	707.9049
492.57	1.0528	1135.409
503.29	1.725	1860.353
513.65	2.6093	2814.041
524.09	3.5478	3826.181

**Figure 3.** Clausius-Clapeyron plot of Ni(5-me-salEtN)<sub>2</sub>.**Figure 4.** Plot of fraction reacted  $g(\alpha)$  versus temperature.



**Figure 5.** Plot of  $\ln[g(\alpha)/T^2]$  versus  $1000/T$  in K; where  $g(\alpha) = [-\ln(1-\alpha)]^{1/2}$ .

and the plots are shown in **Figure 5**. From the slope, the activation energy ( $E_a$ ) for the vaporization of the complex yielded the value of  $93.5 \pm 7 \text{ kJmol}^{-1}$ .

#### 4. Conclusion

The thermodynamic and vaporization kinetics of  $\text{Ni}(5\text{-me-salEtN})_2$  complex was carried out. The TG-based transpiration technique was used to evaluate the vapour pressure data of  $\text{Ni}(5\text{-me-salEtN})_2$ . The standard enthalpy of vaporization and entropy of vaporization of the complex have been evaluated. The activation energy obtained for the Coats-Redfern method was found to agree with the value of enthalpy of vaporization. The thermal degradation mechanism for  $\text{Ni}(5\text{-me-salEtN})_2$  system is a A2 (Avrami-Erofe'ev model) which corresponds to nucleation and growth mechanism.

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#### References

- [1] Jervis, T.R. (1985) Metal Film Deposition by Gas-Phase Laser Pyrolysis of Nickel Tetracarbonyl. *Journal of Applied Physics*, **58**, 1400-1401. <http://dx.doi.org/10.1063/1.336114>
- [2] Lee, P.S. (2002) Enhanced Stability of Ni Monosilicide on MOSFETs Poly-Si Gate Stack. *Microelectronic Engineering*, **60**, 171-181. [http://dx.doi.org/10.1016/S0167-9317\(01\)00592-5](http://dx.doi.org/10.1016/S0167-9317(01)00592-5)
- [3] Stauf, G.T., Driscoll, D.C. and Dowben, P.A. (1987) Enhanced Stability of Ni Monosilicide on MOSFETs Poly-Si Gate Stack. *Thin Solid Films*, **153**, 421-430. [http://dx.doi.org/10.1016/0040-6090\(87\)90202-1](http://dx.doi.org/10.1016/0040-6090(87)90202-1)
- [4] Ishikawa, M. (2004) Ni Precursor for Chemical Vapor Deposition of NiSi. *Journal of Applied Physics*, **43**, L1833-L1836. <http://dx.doi.org/10.1143/JJAP.43.1833>
- [5] Kada, T., Ishikawa, M., Machida, H., Ogura, A., Ohshita, Y. and Soai, K. (2005) Volatile CVD Precursor for Ni Film: Cyclopentadienylallylnickel. *Journal of Crystal Growth*, **275**, e1115-e1119. <http://dx.doi.org/10.1016/j.jcrysgro.2004.11.198>
- [6] Tonneau, D., Auvert, G. and Paueau, Y. (1989) Deposition of Nickel Microstructures by  $\text{CO}_2$  Laser-Assisted Decomposition of Nickel Tetracarbonyl. *Journal of Applied Physics*, **66**, 165-168. <http://dx.doi.org/10.1063/1.343898>
- [7] Becht, M. (1995) Nickel Thin Films Grown by MOCVD Using  $\text{Ni}(\text{dmg})_2$  as Precursor. *Journal de Physique IV*, **C5**,

465, 472.

- [8] Maruyama, T. and Tago, T. (1993) Nickel Thin Films Prepared by Chemical Vapour Deposition from Nickel Acetylacetonate. *Journal of Materials Science*, **28**, 5345-5348. <http://dx.doi.org/10.1007/BF00570088>
- [9] Brissonneau, L. and Vahlas, C. (1999) MOCVD-Processed Ni Films from Nickelocene. Part I: Growth Rate and Morphology. *Chemical Vapor Deposition*, **5**, 135-142. [http://dx.doi.org/10.1002/\(SICI\)1521-3862\(199908\)5:4<135::AID-CVDE135>3.0.CO;2-1](http://dx.doi.org/10.1002/(SICI)1521-3862(199908)5:4<135::AID-CVDE135>3.0.CO;2-1)
- [10] Kang, J.K. and Rhee, S.W. (2000) Metalorganic Chemical Vapor Deposition of Nickel Films from  $\text{Ni}(\text{C}_5\text{H}_5)_2/\text{H}_2$ . *Journal of Materials Research*, **15**, 1828-1833. <http://dx.doi.org/10.1557/JMR.2000.0264>
- [11] Pankajavalli, R., Mallika, C., Sreedharan, O.M., Premila, M. and Gopalan, P. (1998) Vapour Pressure of C60 by a Transpiration Method Using a Horizontal Thermobalance. *Thermochimica Acta*, **316**, 101-108. [http://dx.doi.org/10.1016/S0040-6031\(98\)00304-9](http://dx.doi.org/10.1016/S0040-6031(98)00304-9)
- [12] Arockiasamy, S., Sreetharan, O.M., Mallika, C., Raghunathan, V.S. and Nagaraja, K.S. (2007) Development, Characterisation and Rapid Evaluation of Standard Enthalpies of Vaporisation and Fusion of Volatile Bis(NR-Salicylaldimine) Nickel(II) (nR= Methyl to Pentyl) Complexes for Its MOCVD Applications. *Chemical Engineering Science*, **62**, 1703-1711. <http://dx.doi.org/10.1016/j.ces.2006.12.001>
- [13] Burnham, L., Dollimore, D. and Alexander, K. (2001) Calculation of the Vapor Pressure-Temperature Relationship Using Thermogravimetry for the Drug Allopurinol. *Thermochimica Acta*, **367**, 15-22. [http://dx.doi.org/10.1016/S0040-6031\(00\)00652-3](http://dx.doi.org/10.1016/S0040-6031(00)00652-3)
- [14] Joseph, K., Sridharan, R. and Gnanasekaran, T. (2000) Kinetics of Thermal Decomposition of  $\text{Th}(\text{C}_2\text{O}_4)_2 \cdot 6\text{H}_2\text{O}$ . *Journal of Nuclear Materials*, **281**, 129-139. [http://dx.doi.org/10.1016/S0022-3115\(00\)00241-5](http://dx.doi.org/10.1016/S0022-3115(00)00241-5)