

# Change of Spectroscopic Properties of PU/PVC Doped by Single Walled Carbon Nanotubes

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## Abstract

Nanocomposite films based on polyurethane and polyvinyl chloride (PU/PVC) doped by low content ( $\leq 0.045$  wt%) of single wall carbon nanotubes (SWCNTs) were prepared and characterized. The structural, optical, thermal and mechanical properties of the prepared nanocomposites were characterized by various tools. The IR band at  $3437\text{ cm}^{-1}$  in SWCNTs vanishing for doped samples signified that SWCNTs were successfully covalently encapsulated with PU/PVC. The X-ray peaks broaden with reduced intensity which refers to a reduction of the crystallinity for the blend after increase of SWCNTs attributed to changing in the crosslink density. UV/Vis analysis reveals that the optical band gap values were decreased with increased SWCNT attributed to the formation of defects in polymeric matrix with increase in the degree of disorder in the films. These defects produce the localized states in the optical band gap. All results data showed well dispersion of SWCNTs in polymer matrix at all weight percentages and an interaction between SWCNTs and the blend occurred. The thermal analysis showed that the melting point ( $T_m$ ) and degradation temperature ( $T_d$ ) of the obtained films were increased gradually with adding of SWCNTs. Furthermore, the thermal stability was improved after adding SWCNTs into the PU/PVC polymer blend.

## Keywords

Nanocomposites, SWCNTs, X-Ray, TGA, Mechanical Properties

## 1. Introduction

Polyurethane (PU) has attracted an exceptional measure of consideration because of a variety of applications such as thermoplastic, elastomers, scientific device, textiles and smart actuators adhesive [1]. The molecular structure of PU can be easily tailored to meet particular property requirements.

Poly vinyl chloride (PVC) is a polymer utilized as often as possible as a part of uses that require structural features (mechanical strength) abilities. Since PVC has attractive properties, it usually requires the incorporation of different additives [2] [3].

Blending PU with other polymers has been of a considerable interest in recent years. There are many studies on the blends of PVC with PU [4] [5]. The presence of intermolecular bonding of hydrogen between carbonyl groups (C=O) of PU with active  $\alpha$ -hydrogen in PVC chains makes PVC component and PU more compatible in the main blend chain [6]. Adding PVC to PU would be possible to expand the characteristics of two polymers to obtain compatibility of the blends and better mechanical property. Due to well dispersion, highly potential nanoparticles, or tubes in polymeric matrices, high-performance lightweight nanocomposites can be produced [7].

Carbon nanotubes (CNTs) incorporating into the polymeric matrices is an attractive method to combine the optical, electrical and mechanical properties for CNTs with the advantages of the polymer [8] [9]. It has attracted numerous researchers because of their good properties. As a result of the properties with nanometer scale size and aspect ratio, they are seen perfect as a strong reinforcing specialist for high strength polymer nanocomposites.

Many researchers have focused on the preparation of polymer nanocomposites loaded with carbon nanotubes by casting technique, mixed by melt method and electro spinning with enhanced thermal and mechanical structure [10] [11] [12]. The present work seeks for synthesis and investigation of PU/PVC polymer blend doped with low content of single-walled carbon nanotubes ( $\leq 0.045$  wt%) and studies the changes of the structural and spectroscopic properties of the nanocomposites.

## 2. Experimental Work

Polyurethane (PU) have  $M_w \approx 330,600$  from Cargill-Dow Company, South Korea and high molecular weight polyvinyl chloride (PVC) supplied by Fluka. Tetra hydro furan (THF) supplied from Duksan Company (Korea) and it is used as received. Purified single walled carbon nanotubes (SWCNTs) with COOH functional groups (NTX10) were obtained from Nanothinx, Greece. SWCNTs-COOH has a diameter of 8 - 14 nm, length  $\geq 5 \mu\text{m}$  and a purity of 88%.

The PU/PVC (75:25) wt% blend composite was dissolved in Tetrahydrofuran (THF) as a solvent at  $45^\circ\text{C}$  under stirring until 24hour to complete dissolution to form a homogeneous solution. The polymer blend solution was left without stirring at  $30^\circ\text{C} \cong 2$  hour to remove the bubbles during stirring. Also, SWCNT functionalization was mixed with polymer solution under continuous stirring in an ultrasonic to disperse of these CNTs inside the solution. Different concentration of SWCNT (0.015, 0.030 and 0.045 wt%) were added. After that, the films of the (PU/PVC)-SWCNT were obtained. The nanocomposites solution was put in oven at room temperature to evaporate the solvent for about 36 h. The obtained films were kept in vacuum desiccators before their investigation and measurement.

The FT-IR absorption spectra were carried out using the single beam Fourier trans-

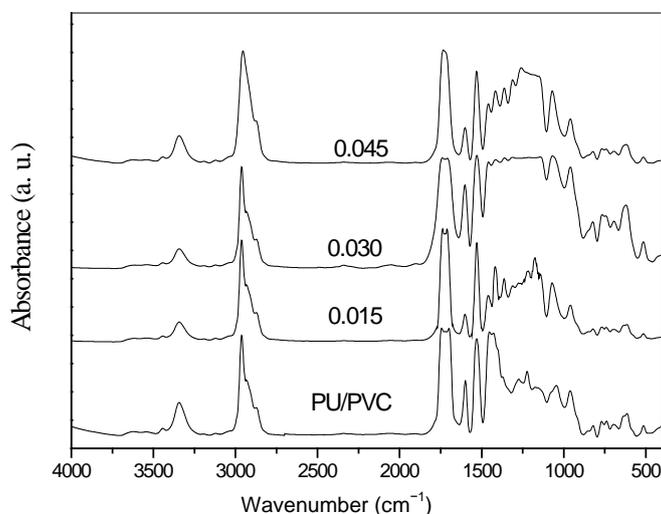
form-infrared spectrometer (FT-IR-430, JASCO, Japan). The FT-IR spectra of the samples were obtained in the spectral range of 2000 - 400  $\text{cm}^{-1}$ . The X-Ray diffraction scans were obtained using DIANO corporation-USA equipped using Cu-K $\alpha$  radiation ( $\lambda = 1.540 \text{ \AA}$ , at 30 kV, the Bragg angle  $2\theta = 5^\circ - 60^\circ$ ). The UV-vis. spectra were measured in wavelength region of 190 - 900 nm by spectrophotometer (V-570 UV-VIS-NIR, JASCO, Japan). Thermogravimetric analysis (TGA) used to characterize the decomposition and thermal stability of prepared samples was carried out using A Perkin-Elmer TGA-7 from 30°C to 550°C with a heating rate of 5°C/min.

### 3. Results and Discussion

#### 3.1. Fourier Transform Infrared (FT-IR)

**Figure 1** displays FT-IR spectrum of both polyurethane (PU) and poly vinyl chloride (PVC). The spectrum of PU shows IR absorption bands at 3340  $\text{cm}^{-1}$  due to NH stretching, at 2950  $\text{cm}^{-1}$  corresponding to CH asymmetric stretching, the band at 2860  $\text{cm}^{-1}$  due to CH symmetric stretching, 1730  $\text{cm}^{-1}$  for C=O stretching of urethane groups, 1602  $\text{cm}^{-1}$  due to NH bonding), 1530  $\text{cm}^{-1}$  attributed to C-C stretching (aromatic)) and at 1065  $\text{cm}^{-1}$  C-O assigned to stretching of urethane groups [13] [14]. The spectrum of pure PVC shows the main characteristic bands at 2960  $\text{cm}^{-1}$  due to CH stretching, 2910  $\text{cm}^{-1}$  assigned to CH<sub>2</sub> asymmetric stretching, 2861  $\text{cm}^{-1}$ , 1434  $\text{cm}^{-1}$  due to CH bending and at 837  $\text{cm}^{-1}$  attributing to C-Cl stretching [15] [16].

The spectrum for PU/PVC blend shows characteristic absorption bands of both PU and. However, the band corresponding to C-N-H bonding vibrations modes at 1602  $\text{cm}^{-1}$  and 837  $\text{cm}^{-1}$  were slightly shift to 1594  $\text{cm}^{-1}$  and 817  $\text{cm}^{-1}$  assigned to hydrogen bonded interactions between C-N-H in PU and C-Cl in PVC groups. The shift of IR absorption band can be used as measure intermolecular interactions between two polymers.



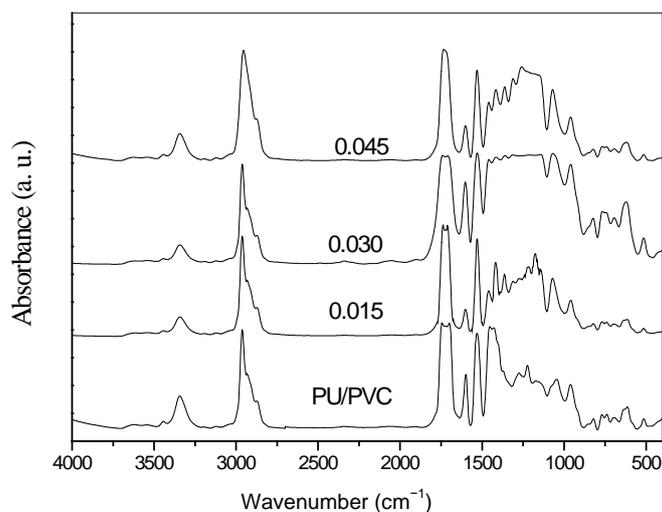
**Figure 1.** FT-IR spectra of pure PU, pure PVC and pure PU/PVC.

It is known that, pure single walled carbon nanotubes have a function group at  $3437\text{ cm}^{-1}$  which refers to OH stretching of the hydroxyl group which can be attributed to the oscillation of carboxyl groups ( $\text{O}=\text{C}-\text{OH}$  and  $\text{C}-\text{OH}$ ). The main band which is observed at  $2922\text{ cm}^{-1}$  is attributed to asymmetric stretching of  $\text{CH}_2$  and at  $2861\text{ cm}^{-1}$  is attributed to symmetric stretching of  $\text{CH}_2$ . The absorption band at  $1735\text{ cm}^{-1}$  is corresponding to  $\text{C}=\text{O}$  stretching of  $\text{COOH}$ . The band at  $1630\text{ cm}^{-1}$  can be associated with the stretching of the CNTs backbone and band at  $1579\text{ cm}^{-1}$  is  $\text{C}=\text{C}$  vibration.

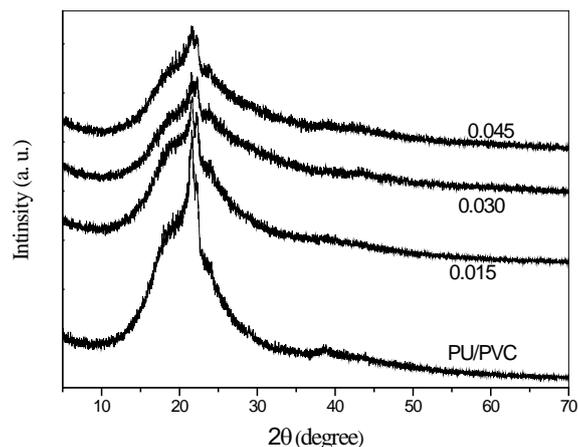
**Figure 2** depicts the FT-IR spectra of PU/PVC doped with low contents of SWCNT. It can be seen that, the intensity of the band at  $3437\text{ cm}^{-1}$  which shows the strongest band in SWCNTs vanishes after addition of SWCNTs, signifying that SWCNTs were successfully covalently encapsulated by two layers of polymeric matrices. The carboxylic acid group may have reacted with OH group in the blend to form ester groups and consist of a broad band at  $3320 - 3400\text{ cm}^{-1}$  which is overlapping of free amine group ( $\text{NH}$ ) in urethane linkage and hydroxyl group ( $\text{OH}$ ) in  $\text{SWCNT}-\text{COOH}$ .

### 3.2. X-Ray Diffraction (XRD)

**Figure 3** displays the X-ray diffraction of PU/PVC doped SWCNTs in the range of  $2\theta = 5^\circ$  to  $70^\circ$ . The spectra show the semicrystalline behavior and are characterized by two peaks at  $2\theta = 21.6^\circ$  and  $22.3^\circ$ . The spectra which contain SWCNTs showed much lower diffraction intensity compared to the blend and these peaks become broadened with reduced intensity, referring to a reduction of the crystallinity of the blend with increase of SWCNTs contents. This change can be attributed to changing in the crosslink density of the polymer blend with increase in SWCNTs. The SWCNTs considerably affect the well short-range microstructural phases of both soft and hard segments of polymeric matrices attributed to the presence of strong interfacial interactions between components. No peaks pertaining to SWCNT appeared in the spectra, indicating the complete/partial dissolution of the SWCNT in the polymer matrix.



**Figure 2.** FT-IR spectra of PU/PVC doped SWCNT.



**Figure 3.** X-ray spectra of PU/PVC doped SWCNT.

### 3.3. Ultraviolet and Visible Study

**Figure 4(a), Figure 4(b)** shows UV-Vis spectra of PU/PVC doped with low content of SWCNTs ( $\leq 0.045$  wt%) in the range of 190 - 600 nm. All spectra exhibit very small absorbance in the ultra-violet range from 200 to 380 nm. The absorption band at 200 nm for the blend was assigned to  $n \rightarrow \pi^*$  transitions and the band at 290 nm was attributed to  $\pi \rightarrow \pi^*$  transition [17]. The main absorption edge for all curves was observed at about 208 nm which shifted toward longer wavelengths by about 20 nm with increasing SWCNT content. This shift in the bands indicates interactions between the polymers chains with the nanotube surface.

The optical energy band gap ( $E_g$ ) for transitions can be determined by using the equations [18]:

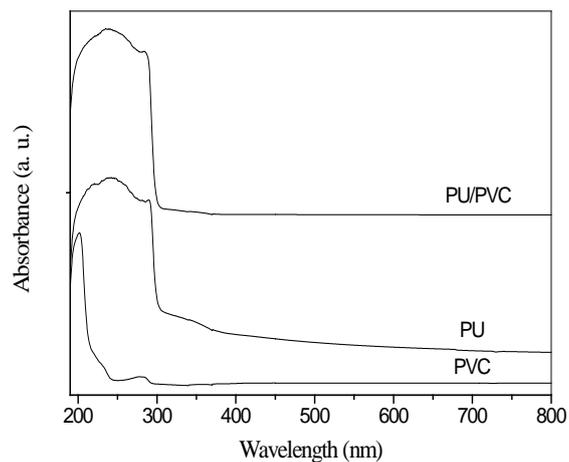
$$(\alpha h\nu) = C(h\nu - E_g)^n \quad (1)$$

where  $C$  is a constant and  $n$  is the order which characterizes the optical absorption processes. It's that can take values as:  $n = 1/2$  for indirect allowed transition,  $n = 1$  for nonmetallic materials,  $n = 3/2$  for direct forbidden transition,  $n = 2$  for direct allowed transition and  $n = 3$  for indirect forbidden transition [19].

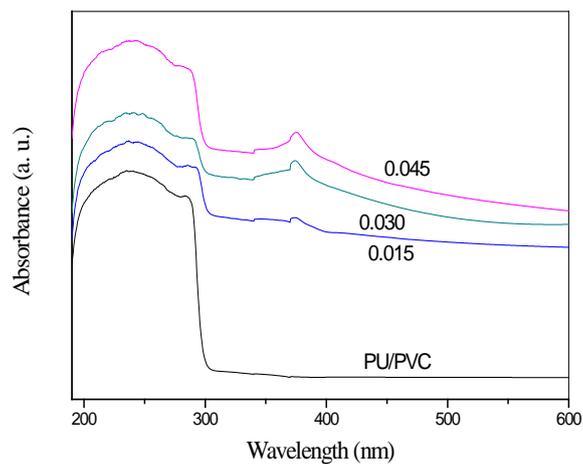
The plot of  $(\alpha h\nu)^{1/2}$  from Equation (1) depends on of photon energy ( $h\nu$ ) as shown in **Figure 5** is used to calculate the indirect optical energy gap. The estimated values of indirect energy gap are recorded in **Table 1**. The indirect band gap values was decreased with increase SWCNT content attributed to the formation of defects in the polymeric matrix with increase in the degree of disorder in the films. These defects produce the localized states in the optical band gap.

### 3.4. Thermogravimetric Analysis (TGA)

**Figure 6** shows the TGA thermograms of PU/PVC doped with single walled carbon nanotubes from 30°C to 650°C. Two major weight losses were observed in all curves. The first weight loss of pure blend was observed from 220°C to 350°C attributed to losses of vapors and gasses from the blend. The initial weight loss were observed for all

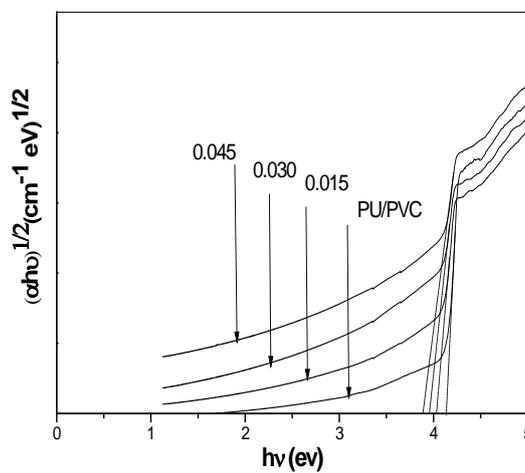


(a)



(b)

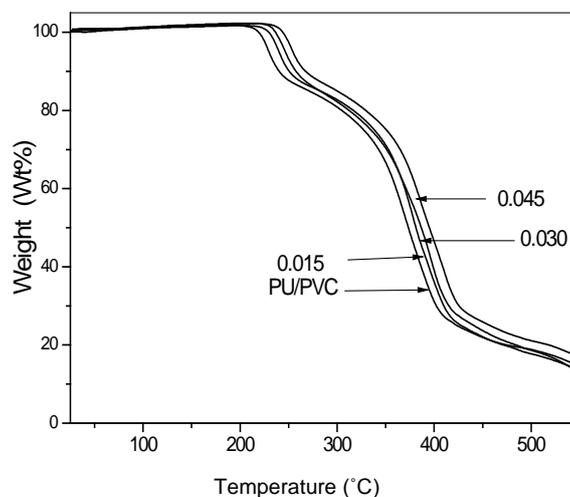
**Figure 4.** (a) UV-Vis spectra of pure PU, pure PVC and pure PU/PVC; (b) UV-Vis spectra of PU/PVC doped SWCNT.



**Figure 5.** The relation between  $(\alpha h\nu)^{1/2}$  versus  $(h\nu)$  of PU/PVC doped SWCNT.

**Table 1.** The calculated values of direct band energy and the activation energy of PU/PVC doped SWCNTs.

SWCNTs (wt%)	$E_{\text{gd}}$ (eV)	Coats-Redfern method (KJ/mol)	Broido method (KJ/mol)
PU/PVC	4.2	242	243
0.015	4.1	220	224
0.030	4.0	217	219
0.045	3.9	212	215

**Figure 6.** TGA thermograms of PU/PVC doped SWCNT.

prepared nanocomposites at lower temperatures may be attributed to the residual solvent evolution. The second weight loss from 360°C to 420°C corresponding to burning of organic phase and the thermal dehydration of inorganic particles. The weight loss at  $t > 420^\circ\text{C}$  ( $\approx 3\%$ ) due to degradation of large polymer chains into small fragments which subsequently undergo further decomposition forming carbonaceous matter at around 510°C and remains constant there after showing the plateau behavior. For doped samples, there is higher stability attributed to the combined effect of the excellent thermal stability of CNTs and their strong interfacial interactions with the polymeric matrix. It is possible that the formation of compact char of carbon nanotubes and polymer matrix during the thermal degradation is beneficial to the improvement of thermal stability of the composites. It is important to point out that the extent of interaction between SWNTs and polymer matrix with good dispersion could be responsible for a higher thermal stability of the composite system.

### 3.5. Calculation of the Activation Energy

The thermodynamics parameters of main decomposition process in the second degradation region of the present samples can be calculated using two different methods:

### 3.5.1. Coats-Redfern Method

Coats-Redfern method is a typical integral method and can be represented in the following form [20]:

$$\log \left[ \frac{-\log(1-\alpha)}{T^2} \right] = \log \frac{R}{\Delta E_a} \left[ 1 - \frac{2RT}{E_a} \right] - 0.434 \frac{E_a}{RT} \quad (2)$$

where,  $T$  is the absolute temperature in Kelvin,  $E_a$  is the activation energy in J/mol,  $R$  is the universal gas constant (8.3136 J/mol K), and the fraction of conversion ( $\alpha$ ) for a weight loss.

The plot between  $\log \left[ \frac{-\log(1-\alpha)}{T^2} \right]$  vs  $\frac{1000}{T}$  of the samples, the relation give a parallel straight lines. So, the activation energies are determined from slopes of these lines as:

$$E_a = 2.303R \times \text{slope} \quad (3)$$

### 3.5.2. Broido Method

Broido introduce a model to evaluate the activation energy associated with second stage of decomposition using equation [21]:

$$\ln \left[ \ln \left( \frac{1}{Y} \right) \right] = \left( \frac{E_a}{R} \right) \frac{1}{T} + C \quad (4)$$

$Y$  is given by :

$$Y = \frac{W_t - W_\infty}{W_i - W_\infty} \quad (5)$$

where,  $Y$  is the fraction of the number of initial molecules not yet decomposed;  $W_t$  is the weight at any time  $t$ ;  $W_\infty$  is the weight at infinite time (=zero) and  $W_i$  is the initial weight. A plot between  $\ln \left[ \ln \left( \frac{1}{Y} \right) \right]$  vs.  $1/T$  gives an excellent approximation to a straight line. The slope is related to the activation energy. The values of the activation energy ( $E_a$ ) for both methods are recorded in **Table 1**. From the **Table 1**, the calculated values of the activation energy were decreased with increasing of SWCNTs addition indicates that SWCNTs intensively affect the polymer.

## 4. Conclusion

Nanocomposites of PU/PVC-doped SWCNTs were prepared and characterized using different techniques. The FT-IR spectra show shift and disappearance of some bands after adding SWCNTs in PU/PVC blend. It indicates a good physical entanglement between the constituents and clearly the interaction between the SWCNT and the blend. The XRD results revealed broad peaks with lower intensity in relation with pure blend, which inferred that the SWCNTs considerably affect the well short range microstructural phases of both the soft and the hard segments of the polymer matrix, resulting in strong interfacial interactions. The UV-Vis shows shift of the bands towards longer wavelength. The TGA study shows that the incorporation of SWCNTs significantly improved the thermal stability caused by the high thermal conductivity. The

calculated values of energies gap were decreased with increasing of SWCNTs.

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