

Room Temperature Synthesis of Aminocaproic Acid-Capped Lead Sulphide Nanoparticles

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ABSTRACT

Aminocaproic acid (ACA) mixed methanolic lead acetate-thiourea (PbAc-TU) complex as precursor for fabrication of lead sulphide (PbS) nanoparticles (NPs) has been explained. The size, structure and morphology of as-prepared ACA-capped PbS NPs were systematically characterized by scanning electron microscopy (SEM), Transmission electron microscopy (TEM), X-ray diffraction (XRD), Uv-vis spectroscopy and Brunauer-Emmett-Teller (BET) techniques. The obtained results show that the synthesized PbS NPs are nanocrystalline, size quantized and their agglomeration shows a mesoporous network of 8.7 nm in pore size. The binding nature of ACA molecules on PbS surface was studied by thermo gravimetric analysis (TGA), Fourier transform infrared (FTIR) and X-ray photoelectron (XPS) techniques. Results indicate that ACA acts as a soft template that restricts the growth of PbS NPs through its binding to Pb surface via nitrogen lone pair.

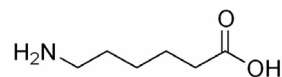
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1. Introduction

In recent years, semiconducting nanoparticles have been studied frequently because of their highly tunable properties obtained by controlling their size, shape and surface morphology [1-4]. Being an important IV-VI group of semiconductors, PbS has attracted an increasing interest and has been extensively investigated due to its unique size quantization properties. Bulk PbS having cubic crystal structure and narrow direct energy band gap (0.41 eV at 300 K) has a large excitonic Bohr radius (around 18 nm). Hence, PbS NPs exhibit strong quantum size effects for relatively large sizes. Also, their absorption edge can be tuned to anywhere between near infrared to violet (0.4 μ m) covering the entire visible spectrum [5]. The combination of such properties makes PbS suitable for efficient electroluminescent devices, such as inorganic-organic bulk hybrid solar cells [6,7], tunable near-infrared detectors [8], and solid state lasers [9]. Further, at a given particle size, the third-order nonlinear optical response of PbS NPs is 30 times greater than that of gallium arsenide (GaAs) and 1000 times greater than that of cadmium sulphide (CdSe), which enables them to be a potential candidate for photonic and optical switching applications [10]. So far, a variety of PbS architectures have been syn-

thesized by various methods. These structures include nanotubes, nanowires, cubes, octahedrons, flower structures, sheet-like shape, dendrite and macrostar-like hierarchical structures. Synthesis techniques generally include hydrothermal, solvothermal, chemical or thermal decomposition routes with conventional surfactants or capping molecules that act as a directing factor to regulate the size and the shape of the synthesized NPs [11-16].

Recently, biomolecule-assisted synthesis has been highly promising due to its novelty and its environmentally friendly character for a large variety of nanomaterials and also for its strong utility in morphology control [17]. To the best of our knowledge, only few references on the synthesis of PbS nanostructures using biomolecules are available in literature [17-20]. Aminocaproic acid (ACA, **Scheme 1**), a derivative of the amino acid lysine, is very effective in the treatment of bleeding disorders. It has functional groups, such as -NH₂ and -COOH (**Scheme 1**), which have a strong capability of coordination with metal ions surface. It was successfully used to produce size and shape controlled NPs including their organized



Scheme 1. Structure of aminocaproic acid (ACA).

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self-assembly nanostructures [21].

In the present contribution, PbS NPs were synthesized via an easily and rapid technique using ACA mixed Pb-Ac-TU complex as a precursor. Utilization of this precursor allows a simple route for the synthesis of well-dispersed PbS NPs at room temperature compared to previously reported synthetic route for the synthesis of PbS NPs using PbAc-TU complex by microwave and sonochemical techniques [22]. Physical, optical and surface properties of the as-prepared PbS NPs show that these NPs are nanocrystalline and size quantized. ACA molecules work as a soft template and are bound to Pb surface through nitrogen lone pair.

2. Experimental

2.1. Reagents

All chemical used were of analytical grade and used without further purification. ACA, lead acetate trihydrate and thiourea were purchased from Sigma-Aldrich, Canada. Methanol was purchased from Fisher chemicals, Canada.

2.2. Synthesis ACA-Capped PbS NPs

In a typical synthesis, methanolic solution of PbAc-TU complex was prepared by dissolving 0.01 mole (M) (0.189 g) of lead acetate trihydrate and 0.01 M (0.038 g) of thiourea one by one in 50 ml of methanol. The solution was stirred for 5 min and 0.01 M of ACA (0.065 g) was added to it. The final mixture was then stirred at 1000 rpm at room temperature using magnetic stirrer. After 1 h of stirring, ACA was completely dissolved and the solution became transparent. Upon 3 h of additional stirring, the transparent solution started to become brown in color indicating the formation of PbS NPs. After 12 h, these NPs were centrifuged, rinsed many times with distilled water and methanol, and dried in a vacuum oven at 50°C for 2 h. Washed and dried PbS NPs were redispersed ultrasonically in methanol for TEM, SAED and Uv-vis characterization.

2.3. Instruments and Characterizations

SEM images were recorded on Jeol JSM-6360/LV scanning electron microscope. TEM and SAED were done using a JEOL JEM 1230 electron microscope operated at 120 kV. The XRD data was recorded on a Siemens D-5000 X-ray diffractometer, using Cu-K α radiation ($\lambda = 1.54059 \text{ \AA}$). The Uv-vis absorption spectrum of agglomerated NPs was recorded on Varian Cary 500 Scan Uv-vis spectrophotometer and the specific surface area was calculated from the linear part of the BET equation. The pore diameter distribution was obtained from the analysis of the desorption branch of the isotherms using Barrett-Joyner-Halenda (BJH) model. TGA characterisation was carried out from 50°C to 650°C using a TA, Q-5000 IR instrument with a heating rate of 10°C/min. FTIR spectroscopy characterization was obtained using a Nicolet (Thermo Fisher) Model 380 FTIR with an attenuated total reflectance (ATR) sampling device (Smart Performer) with a ZnSe crystal. XPS measurements were carried out on Axis Ultra Kratos X-ray photoelectron spectrometer under a vacuum of 2×10^{-8} - 5×10^{-8} Torr; the binding energy values were charge-corrected to the C 1s signal (285.0 eV).

3. Results and Discussion

Figure 1 shows electron microscopic images, SAED and XRD patterns of ACA-capped PbS NPs. SEM image (**Figure 1(a)**) clearly shows the well-defined porous network of NPs agglomeration. TEM image (**Figure 1(b)**) of methanolic dispersion of ACA-capped PbS NPs shows that these NPs are well dispersed and spherical in shape with an average diameter of around 10 nm. The SAED pattern (inset of **Figure 1(b)**) shows concentric rings with bright spots, which indicates the nanocrystalline nature of PbS NPs. All diffraction peaks shown in **Figure 1(c)** match well with the standard XRD lines of face-centered cubic PbS (JCPDS Card No. 05-0592) and are due to reflections from (111), (200), (220), (311), (222), (400), (331), (420), (422) and (511) planes. Furthermore, no other peaks rather than those corresponding to PbS are observed,

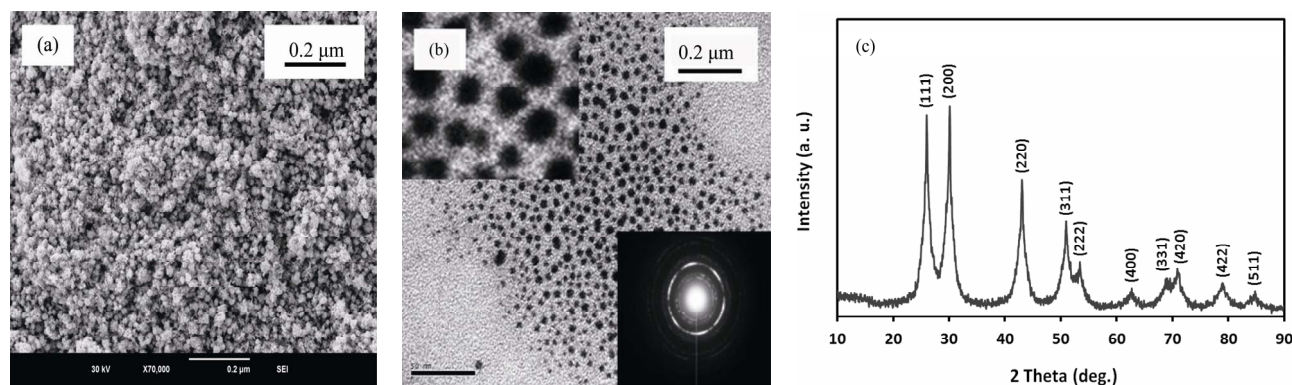


Figure 1. (a) SEM image; (b) TEM image with SAED; and (c) powder XRD pattern of ACA-capped PbS NPs.

which indicates the high purity of the synthesized PbS NPs. The broadening peaks indicate that crystals are in the nanoscale size, as confirmed by the TEM observation. Crystal size was estimated from the broadening peak width of PbS (111) X-ray spectral peaks using the Scherrer formula [23]. It was around 9.2 nm, which is also in close accordance with the size observed by TEM.

The optical absorption spectrum of methanolic ACA-capped PbS NPs is shown in **Figure 2**. As shown, these NPs absorb strongly in visible region due to their strong quantization, compared to bulk PbS.

To examine the specific surface area and pore size distribution, N_2 adsorption-desorption isotherm measurements were performed on ACA-capped PbS NPs agglomerations in powder form. **Figure 3** shows N_2 adsorption-desorption isotherm with a hysteric loop in the range of 0.6 P/P_0 - 1.0 P/P_0 , which indicates that they have a mesoporous structure. Pore size distribution obtained from BJH model is shown in the inset of **Figure 3**. The corresponding specific surface area is 101 m^2/g .

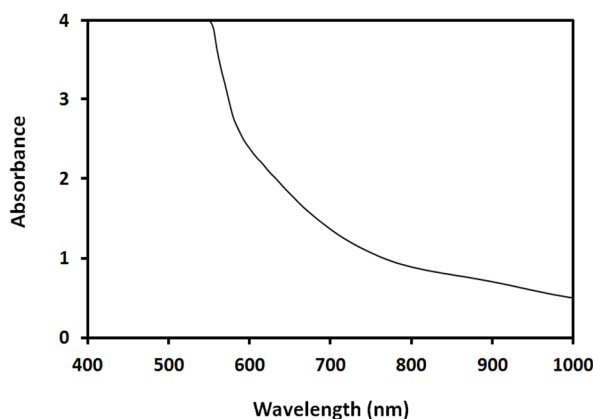


Figure 2. Optical absorption spectrum of ACA-capped PbS NPs.

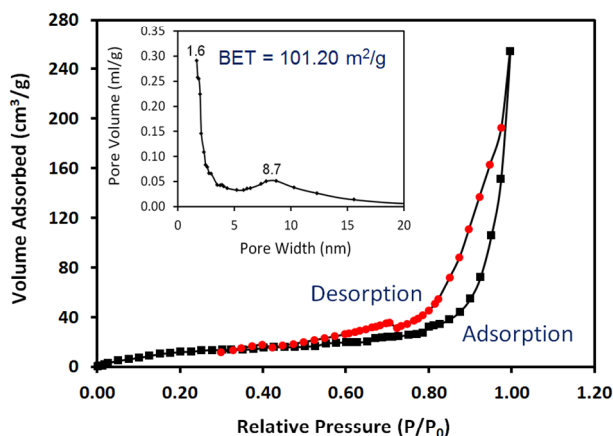


Figure 3. Nitrogen adsorption and desorption isotherms and inset BJH pore radius distribution of ACA-capped PbS NPs powder.

Thermal characterisation from 50°C to 650°C of synthesized ACA-capped PbS NPs powder is summarized in the TGA plot of **Figure 4**, obtained at a heating rate of 10°C/min under air atmosphere. The initial weight-loss of around 4 wt% starting from about 80°C corresponds to the water absorbed on the surface of the NPs. The weight-loss (~23 wt%) observed in the range of 150°C to 450°C corresponds to the combustion and elimination of ACA molecules. This means that the NPs are composed of about 23 wt% of ACA and 73 wt% of PbS NPs. The relatively rapid weight loss in the TGA curve is an indication of a weak binding of ACA molecules to the Pb atom surface.

FTIR spectra of pure ACA and ACA-capped PbS NPs powder were measured to investigate the interaction between ACA and PbS NPs. The corresponding results are respectively shown by curves (a) and (b) in **Figure 5**.

Bands at 2850 cm^{-1} - 2950 cm^{-1} are due to C-H stretching vibrations of methylene groups of the ACA carbon chain for both pure ACA and ACA-capped PbS NPs.

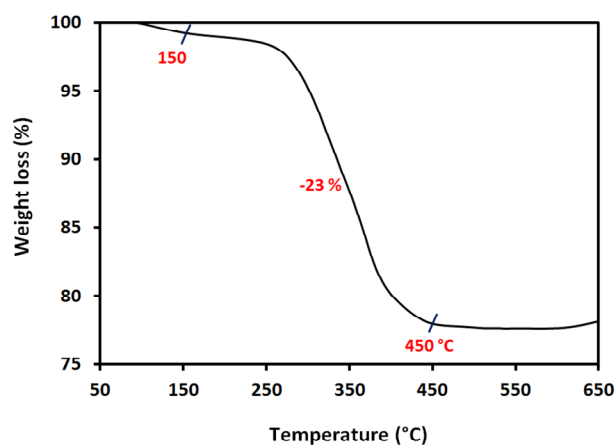


Figure 4. TGA curve of ACA-capped PbS NPs powder.

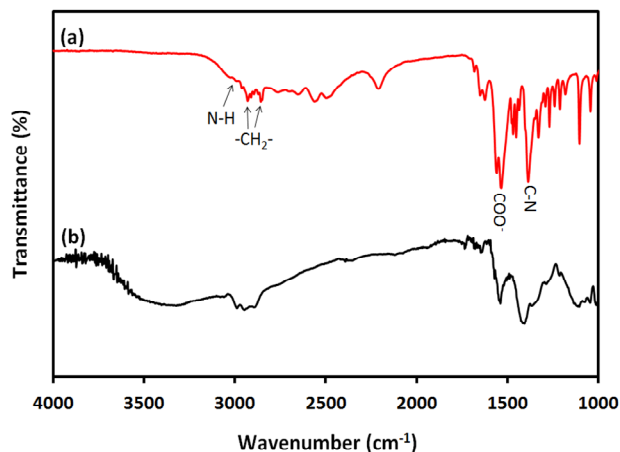


Figure 5. FTIR spectra of (a) ACA and (b) ACA-capped PbS NPs.

The two bands at 1380 cm^{-1} and 3050 cm^{-1} are respectively due to C-N and N-H stretching modes of ACA molecules. These bands also exist in the FTIR spectrum of ACA-capped PbS NPs at same positions, which indicates the binding of amino ($-\text{NH}_2$) groups to the PbS NPs surface through nitrogen lone pair. In FTIR spectrum of ACA, the two bands at 1530 cm^{-1} and 1545 cm^{-1} are attributed to the symmetric and asymmetric stretching vibrations of uncoordinated $-\text{COO}-$ terminal groups of ACA. The spectrum of ACA-capped PbS NPs also shows two bands at the same wavenumbers, which proves that only amino ($-\text{NH}_2$) groups of ACA molecules were bounded on the PbS surface and the free carboxylic ($-\text{COOH}$) groups were oriented outward. Furthermore, the fact that no band appears at 1630 cm^{-1} in the FTIR spectrum of ACA-capped PbS NPs indicates the absence of peptide bond between C=O and C-N moieties, which was usually observed in literature for amino acid capped NPs produced using amino acids [21]. Also, the broad band around 3300 cm^{-1} to 3600 cm^{-1} on curve (b) is due to the presence of adsorbed water in ACA-capped PbS NPs powder sample. Since no free ACA molecules were left on the surface of the NPs, there is no possibility for peptide formation. It is well known that polypeptide chains are always formed through the interaction between uncoordinated amino acid molecules in solution phase. Therefore, the concentration level of amino acid has an important effect on the final morphology of the synthesized nanomaterials. It is known that high concentration of amino acid in bulk precursor affects the self-assembly of the peptide structure and subsequently affects the morphology

of the nanomaterials. From FTIR results, it can be concluded that only amino ($-\text{NH}_2$) groups of ACA molecules bind on the surface of PbS NPs via nitrogen lone pair, while carboxylic ($-\text{COOH}$) groups are located at the other side of the ACA chain.

For further examination of the chemical structure of ACS-capped PbS NPs, XPS spectra of Pb 4f, S 2p, C 1s, O 1s, and N 1s were obtained (**Figure 6**). The XPS spectrum presented in **Figure 6(a)** shows that two Pb peaks appeared at 138 eV (Pb 4f_{7/2}) and 142.9 eV (Pb 4f_{5/2}). **Figure 6(b)** shows that two peaks corresponding to S 2p_{3/2} and S 2p_{1/2} appeared respectively at 161.1 and 162 eV. All binding energy values for Pb 4f and S 2p are slightly higher than what was mentioned in literature for bulk PbS [24]. The difference could be attributed to the coordination action of Pb moiety and ACA molecules. The C 1s region (**Figure 6(c)**) consists of three XPS peaks at 284.9, 285.8, 288.3 eV respectively assigned to C-C, C-N, C-C=O groups of ACA acid molecules [25]. The single and symmetric peak of O 1s at 531.8 eV (**Figure 6(d)**) indicates the presence of two symmetric oxygen atoms in of the carboxylic acid moiety. The XPS of N 1s (**Figure 6(e)**) shows broad peak around 399.8 eV, which is assigned to the nitrogen atom of $-\text{NH}_2$ moiety.

It was reported in literature that PbAc-TU complex is metastable in methanol and the decomposition of simple methanolic lead acetate-thiourea complex takes place by itself at room temperature leading to PbS particles [22]. At an initial stage, $-\text{NH}_2$ moiety of ACA interacts with Pb surface through lone pair in methanolic medium. The interaction of ACA on Pb surface probably decreases the

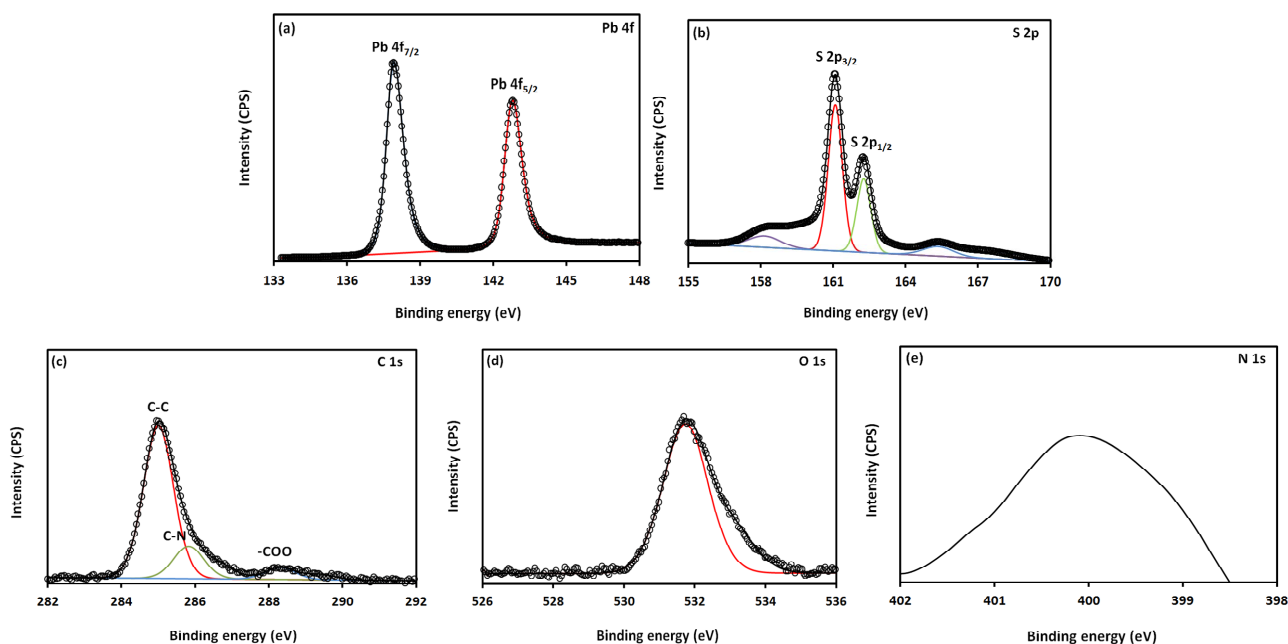


Figure 6. XPS spectra obtained from the ACA-capped PbS NPs: (a) Pb 4f core level; (b) S 2p core level; (c) C 1s core level; (d) O 1s core level; and (e) N 1s core level.

rate of decomposition of PbAc-TU complex. However, this decomposition is further accelerated by continuous stirring. In the reaction mechanism, stirring is a driving force of the reaction. It provides an exposure for the decomposition of PbAc-TU complex. Apart from that, stirring also provides a favorable condition for heterogeneous nucleation, which is ultimately responsible for obtaining regular size and shape of NPs.

4. Conclusion

In summary, a simple soft-chemical route for room temperature synthesis of ACA-capped PbS NPs using ACA mixed PbAc-TU complex as a precursor has been developed. SEM, TEM, SAED, UV-vis and BET studies show that PbS NPs developed via this technique are nanocrystalline, size quantized and present a mesoporous structure. TGA, FTIR and XPS studies of these NPs indicate that ACA acts as a soft template to restrict the growth of PbS NPs through its soft binding via nitrogen lone pair.

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