

Analysis of Spectroscopic, Optical and Magnetic Behaviour of PVDF/PMMA Blend Embedded by Magnetite (Fe_3O_4) Nanoparticles

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Abstract

In the present work, magnetite (Fe_3O_4) nanoparticles have been prepared by a simple chemical method. Polymer nanocomposites based on the blend between poly vinylamine fluoride (PVDF) and (methyl methacrylate) (PMMA) doped with different concentrations of Fe_3O_4 nanoparticles have been prepared. The structural, optical, and magnetization properties of the nanocomposite samples were studied using suitable techniques. The X-ray study reflected that the cubic spinal structure of pure Fe_3O_4 crystal. No small peaks or ripples were found in the X-ray spectra, conforming to good dispersion of Fe_3O_4 within PVDF/PMMA matrices. The FT-IR analysis demonstrated the miscibility between the PVDF and PMMA blend with the interaction between the polymer blend and Fe_3O_4 . The values of the band gap from UV-Vis study were decreased up to 4.21 eV, 3.01 eV for direct and indirect measurements, respectively. The magnetization was measured as a function of the applied magnetic field in the range of $-2000 - 2000$ Oersted. The curves of the magnetization indicated a paramagnetic behavior of pure Fe_3O_4 nanoparticles and PVDF/PMMA- Fe_3O_4 nanocomposites. The values of saturation magnetization for pure Fe_3O_4 are nearly 75 emu/g, exhibiting a paramagnetic behavior, and it is decreased with the increase of Fe_3O_4 content.

Keywords

PVDF/PMMA Blend, Magnetite (Fe_3O_4) Nanoparticles, XRD, FT-IR, UV-Vis Spectroscopy, Magnetization Properties

1. Introduction

Polyvinylidene fluoride (PVDF) is known as one of semi crystalline polymer [1] [2]. PVDF is characterized by its wide applications because of its excellent and

extraordinary properties like smart mechanical strength, and high thermal stability [3] [4]. PVDF is a good response to the piezoelectric and pyroelectric properties. PVDF polymer has at least four polar phases known as (α) alpha, (β) beta, (γ) gamma, and (δ) delta phases. Due to the presence of these phases, PVDF is used to enhance and develop the electronic devices and sensors applications [5] [6]. Different pairs of polymer blends between PVDF and other polymers are examined by several authors, like PVDF/polyvinylpyrrolidone (PVP) [7] [8], PVDF/polyethylene glycol (PEG) [9] [10], PVDF/poly vinyl acetate [11], PVDF/polymethyl methacrylate (PMMA) [12] [13]. As an example of those blends, PMMA is considered one of the most interesting polymers because of its good compatibility with PVDF.

The changes in the structural and morphological PVDF/PMMA polymer blend prepared by different methods are studied by Kim *et al.* [14]. When PVDF/PMMA is prepared by melting technique, the reducing rate of crystallization after adding PMMA prefers the crystal phase formation and the phase formation is reduced. However, when PVDF/PMMA is prepared by the casting method, the addition of PMMA had little effect on the crystalline phases. However, Zhang *et al.* showed that the crystallization behaviour of the PVDF/PMMA blend is highly dependent on the components of mixed solutions [15]. They observed that the phases of PVDF are clear when the weight ratio of PVDF is higher than 30 wt% and the adding of 10 wt% of PMMA could help the growth of PVDF crystallization [16]. The PVDF/PMMA blend is thought to result from the interaction between the oxygen atom of the carbonyl groups in PMMA and the hydrogen atom in PVDF [17] [18].

Magnetite nanoparticles (Fe_3O_4) have attracted increasing interest within the fields of applied nanoscience and technology attributed to their unique and new physicochemical properties that are achieved according to their particle size, shape morphology, and shape of geometric films. Various methods of preparation of magnetite nanoparticles were performed through several techniques, including co-precipitation showing that the addition of nanoparticles to the polymer and/or polymer blend may enhance compatibility between the polymers. Magnetite nanoparticles (Fe_3O_4) are one of most nanoparticles to improve and enhance the magnetic properties for polymer nanocomposites in the industry [19] [20].

Pure blend without nanofiller between PVDF and PMMA blends have been prepared and investigated by FT-IR spectroscopy, X-ray, and thermal analysis [21] [22]. The nanocomposite sample PVDF/ Fe_3O_4 nanoparticles are synthesis by a simple ultrasonication method and investigated by XRD, SEM, DSC and VSM techniques [23]. Lan *et al.* [24] prepared a paramagnetic PMMA/ Fe_3O_4 nanocomposite and they observed high thermal stability and high saturation magnetization ($39 \text{ emu}\cdot\text{g}^{-1}$). The limits of our knowledge are that there is no PVDF/PMMA- Fe_3O_4 in literature, so, the target of the present work is to synthesise novel magnetite (Fe_3O_4) nanoparticles and to evolve the nanocomposite films based on PVDF/PMMA loaded by different concentrations of Fe_3O_4 . The spectroscopic, optical and magnetic properties of PVDF/PMMA- Fe_3O_4 nano-

composites are characterized by different techniques. The nanocomposite films can be exploited for further magnetic applications.

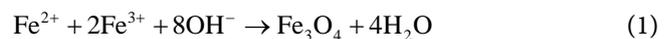
2. Experimental Work

2.1. Materials

Polyvinylidene fluoride (PVDF) pellet has $M_w = 1.8 \times 10^4$ with linear chemical structure $(-\text{CH}_2\text{CF}_2-)_n$ and the polymethyl methacrylate (PMMA) with $M_w = 1.2 \times 10^5$ were purchased from Sigma-Aldrich. All chemicals used in the synthesis of magnetite nanoparticles (Fe_3O_4) such as Ferric chloride hexahydrate ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$), ferrous chloride tetrahydrate ($\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$), and ammonium hydroxide (NH_4OH) were obtained from Sigma-Aldrich.

2.2. Synthesis of Magnetite Nanoparticles (Fe_3O_4)

The magnetite nanoparticles (Fe_3O_4) were synthesized using the promising methods because of its simplicity and ease of implementation with less hazardous and fewer procedures. Ferric chloride ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$) and ferrous chloride ($\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$) were mixed using 2:1 molar ratio. The ferric solutions of both Fe^{2+} and Fe^{3+} were prepared by making their aqueous solutions in distilled water. Then the solutions were heated at 50°C for about 10 min. After heating the solutions, it will be deposited in the presence of an ammonia solution (NH_4OH) with constant stirring by the magnetic stirrer at 50°C . It is also known that the nanocomposite polymeric solution is more stable when ammonium was added. Black color particles of magnetite nanoparticles are deposited. These nanoparticles are then separated from the solution using a strong magnet and washed several times with distilled water. Black magnetite deposits remain. The powder was then dried in a hot air oven at 100°C overnight. General reaction can be written as the following reaction:



2.3. Preparation of PVDF/PMMA- Fe_3O_4 Nanocomposite Films

Polyvinylidene fluoride (PVDF) pellet (4.0 g/50ml) and polymethyl methacrylate (PMMA) powder (4.0 g/50ml) were dissolved individually in tetrahydrofuran (THF) as a common solvent for the polymers at 60°C . Pure solution of PVDF/PMMA was mixed and it stirred using magnetic stirrer about 4 hours at the same temperature. We followed the methods of authors to prepare the polymer blend nanocomposites containing nanoparticles [25] [26]. The different contents (0.0, 0.4, 0.8 and 1.2 wt%) of magnetite nanoparticles (Fe_3O_4) were mixed with the previously polymer solution for 10 min. After that, the nanocomposite solutions (PVDF/PMMA- Fe_3O_4) were immersed in an ultrasonic device to sonicate about 10 min and to ensure that the Fe_3O_4 nanoparticles were dispersed in the polymeric solution. The final solutions were cast in glass dishes and put the dishes in an oven for about 72 hours at 48°C . The nanocomposites (PVDF/PMMA- Fe_3O_4) become a flexible-type freestanding polymeric film with

thickness in the range of 100 μm . **Scheme 1** shows the possible reaction using a chem draw program between PVDF+PMMA embedded by Fe_3O_4 to obtain the PVDF/PMMA- Fe_3O_4 nanocomposite films.

2.4. Characterization Techniques

The X-ray diffraction (XRD) of the nanocomposites films were carried out by PANalytical X'pert Pro MPD diffractometer with Cu-K α radiation ($\lambda = 1.5406 \text{ \AA}$) at operating voltage 15 kV over the range of $2\theta = 5^\circ - 80^\circ$. The complexation between nanocomposites (PVDF/PMMA- Fe_3O_4) was studied using Fourier transform infrared (FT-IR) spectroscopy (FTIR-430, Jascow, Japan) at wavenumber range of $4000 \text{ cm}^{-1} - 400 \text{ cm}^{-1}$. The UV-Vis spectra were recorded by a spectrophotometer (V-570 UV/VIS/NIR, Jasco, Japan) in the frequency range of 190 to 1000 nm. The magnetization curves of the PVDF/PMMA- Fe_3O_4 nanocomposite films were measured using VSM measurement at room temperature (25°C).

3. Results and Discussion

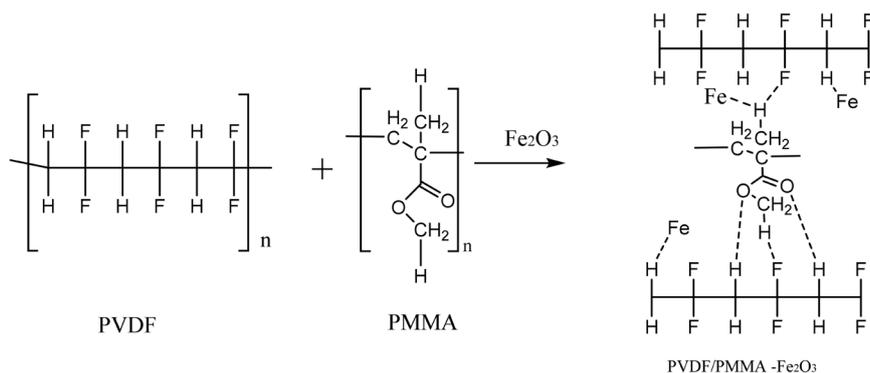
3.1. X-Ray Measurements

The crystalline structure of PVDF/PMMA films doped with different contents (0, 0.4, 0.8 and 1.2 wt%) of magnetite nanoparticles (Fe_3O_4) is recorded using X-ray diffraction as shown in **Figure 1**. The main X-ray peaks of pure Fe_3O_4 powders are founded and inserted in **Figure 1**. The X-ray spectrum exhibited peaks corresponding to (220), (311), (400), (422), (511) and (440) [27] [28] are reflects to Fe_3O_4 crystal with the cubic spinal structure.

The X-ray diffraction measurements of PVDF/PMMA show the semicrystalline structure [29] as the hump and broad peak recorded at $2\theta = 18.38^\circ$ with a peak at $2\theta = 39.07^\circ$. From the spectra of PVDF/PMMA films, no ripples and/or small peaks are found, indicating that the Fe_3O_4 has a good distribution in PVDF/PMMA blend. The movement of the position for Bragg angles from $2\theta = 18.38^\circ$ to $2\theta = 20.04^\circ$ are founded confirms that the crystal structure of Fe_3O_4 is altered by its inserted within PVDF/PMMA matrices. A small decrease in the hump at 18.38° after adding the Fe_3O_4 is seen suggesting that the amorphicity increased. An increase of the hump causes an increase of the amorphous nature inside the nanocomposite films according to Hodge *et al.* method [30]. The interaction between the PVDF/PMMA blend and Fe_3O_4 causing a decrease in the degree of crystallinity of the films. This demonstrating that complexation between Fe_3O_4 and PVDF/PMMA takes place in the amorphous region [31]. The amorphous nature is responsible for the higher conductivity behavior of the prepared samples.

3.2. FT-IR Analysis

The FT-IR spectra of PVDF/PMMA blend doped different concentrations of Fe_3O_4 are shown in **Figure 2**. For pure PVDF, the absorption bands at 1404 and 1066 cm^{-1} are assigned to CH_2 wagging mode. The absorption band at 1167 cm^{-1}



Scheme 1. The possible reaction between PVDF + PMMA embedded by Fe₃O₄ to obtain the PVDF/PMMA-Fe₃O₄ samples.

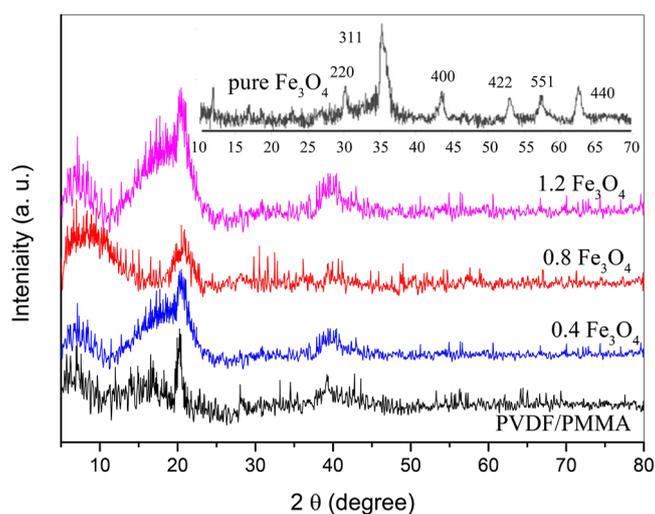


Figure 1. The X-ray diffraction of PVDF/PMMA films doped with different concentrations (0, 0.4, 0.8 and 1.2 wt%) of magnetite nanoparticles (Fe₃O₄).

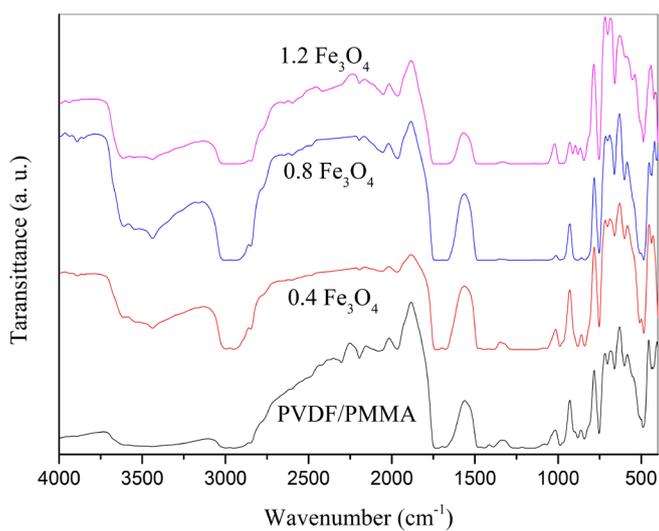


Figure 2. The FT-IR absorption spectra of PVDF/PMMA embedded with different concentration of magnetite nanoparticles (Fe₃O₄).

is assigned to CF_2 asymmetric stretching mode since CF_2 group is absorbed strongly in the region of $1120 - 1350 \text{ cm}^{-1}$. The C-C-C asymmetrical stretching vibration band is observed at 880 cm^{-1} . The absorption band at 796 cm^{-1} corresponds to CF_2 skeletal vibration mode. The absorption band at 478 cm^{-1} corresponds to CF_2 bending vibrational mode [32]. The assignments of FT-IR spectrum of PVDF have been reported as follows: α -phase bands due to CF_2 bending are observed at 482 cm^{-1} , 531 cm^{-1} and 615 cm^{-1} . The main bands ascribed to CH_2 wagging broad mode are observed at 1062 cm^{-1} , whereas the β -phase band due to CF_2 symmetric. The presence of the absorption band at 1729 cm^{-1} is attributable to the stretching of the carbonyl group of PMMA in the blend samples.

The FT-IR spectrum of pure magnetite nanoparticles is inserted inside **Figure 2**. For the magnetite, the band at 580 cm^{-1} corresponds to the vibration of the Fe-O bonds. Additionally, the bands at 1633 and 3400 cm^{-1} can be attributed to the stretching vibration of the hydroxyl groups on the surface of the magnetite nanoparticles [33]. The FT-IR spectra of the nanocomposite are found partially changed after the addition of Fe_3O_4 with a decrease or disappear of some IR bands. These results conform to an interaction between the Fe_3O_4 nanoparticle, and the polymer blend has occurred.

3.3. UV-Vis Analysis

Figure 3 depicts the UV-Vis absorption spectra of pure PVDF/PMMA polymer incorporated with the synthesis Fe_3O_4 nanoparticles. From the spectrum of pure Fe_2O_3 nanoparticles, the sharp absorption in the range of $200 - 400 \text{ nm}$ in the UV region and weak absorption at $400 - 800 \text{ nm}$ of the visible region is observed. The presence of the optical absorption of pure Fe_2O_3 nanoparticles occurs due to two types of electronic transmission mechanisms. The former is due to the contribution of the direct charge transition of $2p$ in (O_2^- ions) $\rightarrow 3d$ (Fe^{3+}) in the UV absorption region, and the other originates from the indirect charge transition of $3d$ (Fe^{3+}) $\rightarrow 3d$ (visible absorption region).

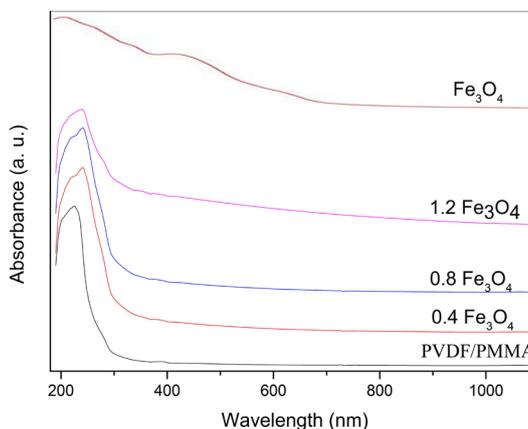


Figure 3. The UV-Vis absorption spectra of PVDF/PMMA embedded with different concentration of magnetite nanoparticles (Fe_3O_4).

The spectrum of pure PVDF/PMMA polymer blend shows an absorbance band at 219 nm attributed to the $n \rightarrow \pi^*$ transition. Moreover, the spectrum displays the sharp edge which towards to red shifted from 219 nm to 241 nm as an increase of absorbance values for doped the films by the Fe_2O_3 nanoparticles, where the color of films is changed from the transparent color to pinkish red. These observations demonstrate that the link between the Fe_2O_3 nanoparticles with the functional groups in the PVDF/PMMA and caused a change of the band gap energy.

The estimated values of E_g for PVDF/PMMA- Fe_2O_3 nanocomposite films can be determined by [34] [35] [36]:

$$\alpha = \frac{B(h\nu - E_g)^n}{h\nu} \quad (2)$$

where, α is the absorption coefficient, $h\nu$ is the photon energy, B is a constant. The values of n are equal to 1/2 for allowed direct and it is equal to 2 for allowed indirect allowed transitions. The absorption coefficient (α) was determined from the equation: $\alpha = \frac{2.303 \times 10^3 * A}{L}$, where A is the absorption constant and L is the sample thickness. **Figure 4** and **Figure 5** show the relation between $(\alpha h\nu)^2$ and $(\alpha h\nu)^{1/2}$ versus $h\nu$ of the PVDF/PMMA- Fe_2O_3 samples. The portion of the straight line (dot lines in the figures) s are an approach to zero absorbance gives values of E_{gs} of direct and indirect transitions, respectively [37]. After doping Fe_2O_3 in pure polymer blend, the values of E_{gs} are gradually decreased with the increase of Fe_2O_3 concentrations due to explained in terms of the formation of charge transfer complexes between the functional groups of PVDF/PMMA and the atoms of Fe_2O_3 [38]. The embedded Fe_3O_4 nanoparticles form an intermediate band among the PVDF/PMMA structures and thus decrease the band energy gap (E_g) of nanocomposite films by absorbing the wavelength of lower energies. The values of the band gap of nanocomposite films decrease to 4.21 eV, 3.01 eV as compared to 4.94 eV and 4.50 from direct and indirect, respectively for pure PVDF/PMMA blend. The reduction of the values of the band gap energy is assumed to increase with a degree of disturbance to generate the localized state in the nanocomposites, as the incorporated of Fe_2O_3 produce energy levels in the band gap in the PVDF/PMMA matrix leading to decrease of the band gap energy. The estimated band-gap energy values of Fe_2O_3 are nearly 2.28 eV and 2.1 eV for allowed direct transition and allowed indirect transition, respectively. The values of E_{gs} of Fe_2O_3 are slightly higher and blue shifted compared to the results of previous reports because of the quantum size confinement in the nanoparticles.

3.4. Magnetic Analysis

Figure 6 shows the magnetization (emu/g) of pure magnetite nanoparticles (Fe_3O_4) and 0.4; 0.8 and 1.2 wt% of Fe_3O_4 embedded in PVDF/PMMA polymer blend was measured as a function of the applied magnetic field (Oe) in the range

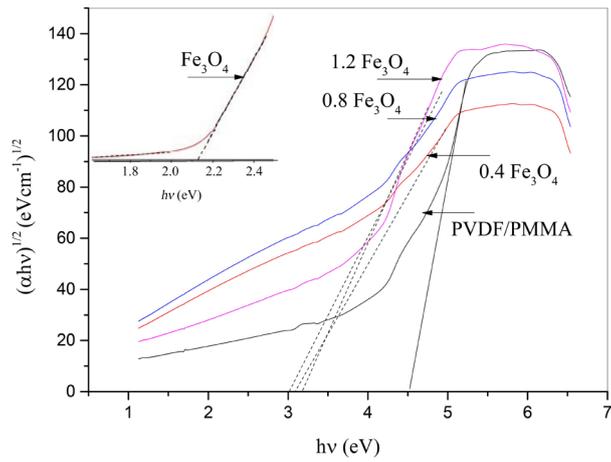


Figure 4. The relation between $(\alpha h\nu)^{1/2}$ and $h\nu$ of PVDF/PMMA embedded with different concentration of magnetite nanoparticles (Fe_3O_4) and the spectrum of pure Fe_3O_4 (insert the figure).

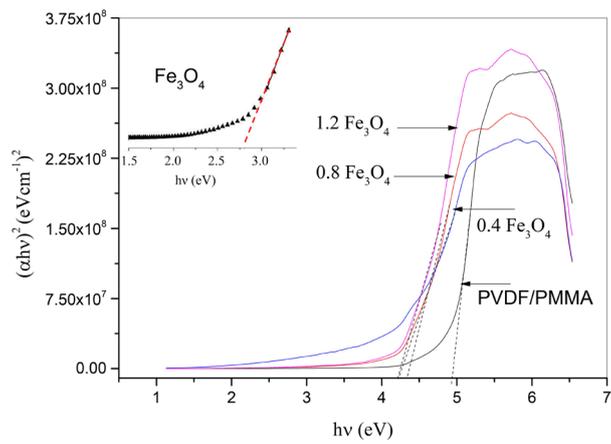


Figure 5. The relation between $(\alpha h\nu)^2$ and $h\nu$ of PVDF/PMMA embedded with different concentration of magnetite nanoparticles (Fe_3O_4) and the spectrum of pure Fe_3O_4 (insert the figure).

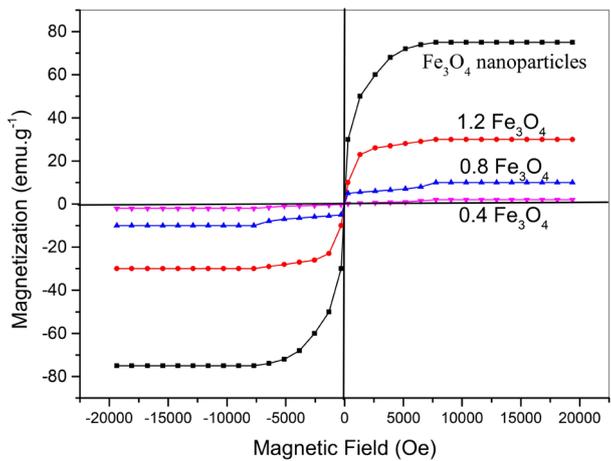


Figure 6. The magnetization curves of Fe_3O_4 and PVDF/PMMA embedded with 1.2, 0.8 and 0.4 wt% of magnetite nanoparticles (Fe_3O_4).

of (−2000 - 2000 Oersted) at room temperature. The curves of the magnetization indicate a paramagnetic behaviour at room temperature for pure Fe₃O₄ nanoparticles and PVDF/PMMA-Fe₃O₄ nanocomposite films. No hysteresis loop from the magnetization curves is observed.

The values of saturation magnetization for pure Fe₃O₄ are nearly 75 emu/g, exhibiting a paramagnetic behavior, and it is decreased with decrease of Fe₃O₄ content embedded in PVDF/PMMA polymeric matrices as the following: ≈30 emu/g for the sample 1.2 Fe₃O₄-PVDF/PMMA; 9.98 emu/g for the sample 0.8 Fe₃O₄-PVDF/PMMA and 4.98 emu/g for the sample 0.4 Fe₃O₄-PVDF/PMMA. As expected, the magnetization of the PVDF/PMMA blend is near to zero value [39] [40] [41]. This is mainly because Fe₃O₄ nanoparticles are embedded into a nonmagnetic polymer matrix. With the decrease in the Fe₃O₄ content, the magnetic saturation value also decreases. These results suggest that the paramagnetic behavior observed in the nanocomposites arose from the magnetic Fe₃O₄ nanoparticles attributed to the decreasing tendency of the nanoparticles to aggregate on decreasing their content. Due to the magnetic properties of the prepared nanocomposites, the magnetic composite films can be further exploited for magnetic applications [42].

4. Conclusion

New nanocomposite films consist of PVDF/PMMA blend loaded by different concentrations of magnetite nanoparticles (Fe₃O₄) using the casting method. The change of the structural, optical, magnetization properties of these nanocomposites has been studied in detail. The structural and chemical complexations between the components were characterized by XRD and FT-IR measurements. The X-ray spectra display the semicrystalline structure of all samples with a decrease of the crystalline degree as an increase of Fe₃O₄ contents. Further, the X-ray spectrum exhibited some peaks that are reflected in the cubic spinel structure of the nanocomposites with good dispersion of Fe₃O₄ nanoparticles within PVDF/PMMA. The FT-IR spectra confirm the miscibility between the PVDF/PMMA blend and show the chemical interaction between the polymer blend and Fe₃O₄. From UV-Vis analysis, the embedded of Fe₃O₄ nanoparticles form an intermediate band inside PVDF/PMMA structures and causing a decrease of the band energy gap (E_g). The magnetization was measured as a function of the applied magnetic field in the range of −2000 to 2000 Oersted at room temperature. The magnetization curves indicate a paramagnetic behavior of the PVDF/PMMA-Fe₃O₄. The values of saturation magnetization for pure Fe₃O₄ are nearly 75 emu/g, exhibiting a paramagnetic behavior, and it is decreased with the decrease of Fe₃O₄ content. Due to the magnetic properties of the prepared nanocomposites, the films can be further exploited for magnetic applications.

Conflicts of Interest

The author declares no conflicts of interest regarding the publication of this paper.

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