

# Environmental Assessment of Acid Mine Drainage Pollution on Surface Water Bodies around Ngwenya Mine, Swaziland

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Received December 7<sup>th</sup>, 2013; revised January 5<sup>th</sup>, 2014; accepted February 2<sup>nd</sup>, 2014

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## ABSTRACT

Twenty-three water samples and three “yellow boy” samples were obtained from different water bodies located at the foot of the Ngwenya Mountain on top of which the old Ngwenya Iron Ore Mine is located. The samples were analysed for pH, electrical conductivity (EC), redox potential (ORP) and temperature (T). The dominant Fe species was determined using a UV-VIS spectrophotometer. Selected anions namely: halogens ( $F^-$ ,  $Br^-$ ,  $Cl^-$ ), the nutrients ( $NO_2^-$ ,  $NO_3^-$ ,  $PO_4^{3-}$ ) and the best indicator for AMD pollution ( $SO_4^{2-}$ ) were analysed using Ion Chromatography (IC) while the selected heavy metals, namely: Cr, Mn, Fe, Ni, Co, Cu, Zn, Pb and Cd were analysed using Flame Atomic Absorption Spectrometry (FAAS). The physico-chemical parameters ranges obtained were pH (6.32 - 8.63), EC (11.00 - 585.33  $\mu S/cm$ ), ORP (-93.67 - 79.33 mV) and T (7.60°C - 18.57°C). The levels of the Fe species (ppm) in the water samples were  $Fe^{2+}$  (0.56 - 3.17) and  $Fe^{3+}$  (0.00 - 0.73). Measured mean anion ranges in ppm were  $F^-$  (0.00 - 0.15),  $Cl^-$  (1.5 - 11.19),  $NO_3^-$  (0 - 13.25) and  $SO_4^{2-}$  (0.05 - 22.29). The mean ranges for the heavy metals in the water samples in mg/l were Cr (0.00 - 2.46), Mn (0.00 - 0.30), Fe (0.00 - 0.34), Ni (0.00 - 0.34), Co (0.00 - 0.29), Cu (0.00 - 0.07), and Cd ((0.01 - 0.11) whilst Zn and Pb were not detectible. The average pH in all the sites was within the acceptable range (6.5 - 8.5) according to WHO standards except for two sites.  $Fe^{2+}$  was the dominant Fe species in the water samples. The water was free from anion pollution because all measured anion levels were below the guideline by WHO. The relative percentages of sites polluted by the respective heavy metals were: Cd (83%), Cr (71%), Ni (46%), Co (42%), Mn (17%), Fe (4%) while no site was polluted by Cu, Zn and Pb. The State needs to set up sound remediation strategies to save the populations around the Ngwenya Mine from the heavy metal pollution they are exposed to by using water polluted by the mining activities.

## KEYWORDS

Acid Mine Drainage; Heavy Metal Pollution; Redox Potential; Physico-Chemical Properties; Iron Speciation

### 1. Introduction

A previous study [1] revealed that the acid mine drainage (AMD) phenomenon occurs in the old Ngwenya Mine, Swaziland. Acid mine drainage is the acidic water that is produced in mining areas from the oxidation of pyrite. The formation of AMD is facilitated biochemically by bacteria called *Thiobacillus ferrooxidans* and *Ferrop-*

*lasma acidarmanus* [2-4]. AMD is rife in abandoned mines and once produced, it can exist for ages thereby polluting the environment concerned in perpetuity [5]. AMD has detrimental effects on surface waters if existent [6,7]. It is therefore worthwhile that the surface water bodies around the mine should be assessed for acid mine drainage pollution. This would enhance the quality of the affected water bodies. Water is life and a safe

drinking water is a human birthright [8]. Use of polluted water upsets the whole ecological system [9-12]. AMD pollution causes the release of heavy metals into the environment [10,13] rendering them available for intake by biological systems. High levels of heavy metals are a nuisance to the environment due to their adverse effects on both plants and animals including human beings [9,14]. They have low biodegradability in the body and accumulate in both soft and hard tissues in the human body [15].

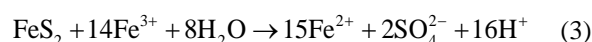
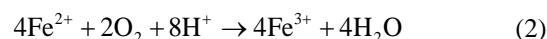
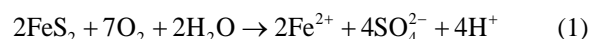
It is hypothesized that acid mine drainage from the Ngwenya Mine ore and dumping sites deteriorates the quality of water in their vicinity [16]. Mobility of heavy metals is enhanced by acidic waters rendering them liable to contaminate the environment [17,18] and causing adverse effects on life generally. As a result, acidic waters are a menace to the environment as well as being incapable to carry out its functions well [2,19]. Such water is unsafe for drinking, irrigation, swimming or cooling purposes in industry. The increased acidity interferes with the characteristics of pure water rendering it a poor coolant in the industry and engines. Low-pH water is inhabitable to aquatic fauna and flora and animals in general [12]. The human population around Ngwenya Mine that can be potentially affected by the AMD stands at 11147 as per 2007 report [20].

AMD can develop throughout the mining process [4], that is, in underground workings, waste rock dumps, open pit mine fractures, tailing deposits, ore stockpiles, etc. Once it occurs, the adverse effects can be costly because of high clean-up costs and irreversibility because of lack of available technology [6,21]. Waters with low pH kill all forms of living organisms in them, save only the highly tolerant ones. This leads to poor water quality and release of bad odour due to the decomposing dead organisms. Moreover, in low pH water, heavy metals become soluble and as such are potential contaminants to soils, sediments and terrestrial plants that use such water for growth [9]. Aquatic organisms can also accumulate the dissolved metals in their bodies which eventually are consumed by higher organisms and sequentially causing pollution by biomagnifications. This continues up to the trophic ladder until these metals get into humans and on accumulating beyond maximum contamination limits (MCL) and cause detrimental effects on human health. Heavy metals are systemic toxins with specific carcinogenic, teratogenic, nephrotoxic, neurotoxic and fetotoxic effects [15]. Reproductivity in crops is hindered by watering with acidic water. Quite often, low pH water contains high concentrations of copper ions that interfere with photosynthesis, and thus resulting in poor yields. This eventually negatively affects food security [10].

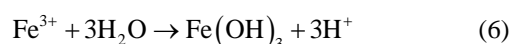
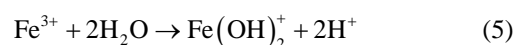
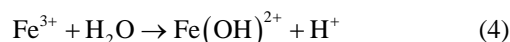
Furthermore, acidic waters are unsuitable for tourism

purposes. The water is corrosive; hence it is not good for swimming. It has bad odour due to the decomposition of dead organic matter which strips off the water of a substantial amount of dissolved oxygen, rendering the waters anoxic. The water is not fit for aesthetics as acidic waters are characterized by an orange-brown colour. This is the colour of ferric hydroxide which is a product of pyrite oxidation. The ferric hydroxide precipitate is called “yellow boy” [22].

Equations (1)-(3) summarize the oxidation of pyrite:



The formation of “yellow boy” removes dissolved oxygen from water [19]. The “yellow boy” is produced in the AMD receiving streams because there is no pyrite to be attacked by the ferric ion. The corresponding step-wise reactions are as follows:



**Figure 1** is a typical picture of “yellow boy” produced in one of the tributaries to Mlondozi River which is very close to the mine.

In summary, the influx of acid mine drainage into streams can severely degrade both habitat and water quality. It often produces an environment devoid of most aquatic life and unfit for desired uses. The severity and extent of damage depend upon a variety of factors including the frequency, volume, and chemical nature of the drainage. The size and the buffering capacity of the



**Figure 1.** The photograph of typical “yellow boy”.

receiving stream are other vital factors. The transport route from the source to the receptor may equally be a determining factor of the extent of damage by the AMD.

## 2. Materials and Methods

### 2.1. Description of the Study Area

The study area is the surface water bodies found around the Ngwenya Iron Ore Mine, Swaziland. The streams studied are used by the communities in the study area for domestic purposes, that is, they are the sole providers of potable water in this area. Some of these streams are feeders for the Hawane Dam which supplies water to the capital city of Swaziland, Mbabane. **Figure 2** below shows the areas in the vicinity of the Mine and the streams that source their water from the Ngwenya Mountain where the AMD producing mine is located [1].

### 2.2. Experimental

Water samples were collected from streams and dams around the foot of the Ngwenya Mountain. The sampling points are shown in **Figure 3** below. Water samples were collected into 1 L prewashed polyethylene bottles as described elsewhere [1]. They were thereafter placed in a cooler box containing ice blocks and transported to the laboratory.

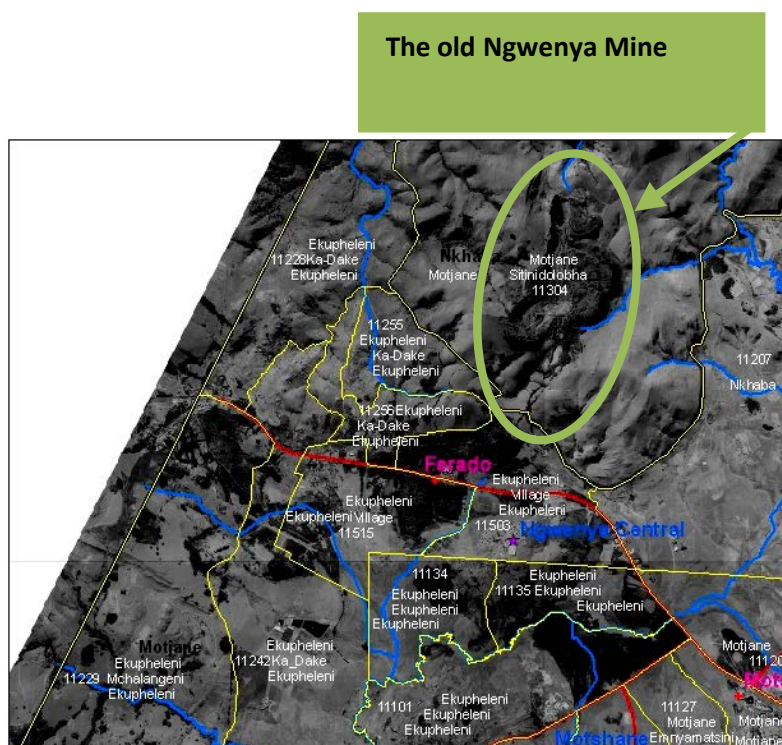
#### 2.2.1. Physico-Chemical Properties

Upon arrival in the laboratory, about 100 ml of each of the water samples were poured into a small beaker and then quickly analysed for the physico-chemical parameters, namely: pH, temperature ( $T$ ), redox potential (ORP) and electrical conductivity (EC). These analyses were carried out using electrochemical methods during which WTW Multi 340i probes were used. The pH and the ORP were measured by one probe, the pH-Electrode Sen Tix 41-3, which can be switched from measuring pH to measuring ORP and vice versa. The EC and  $T$  were measured by another probe, the TetraCon<sup>R</sup> 325. The ORP values were converted to Eh (-log of electron concentration) using Equation (7) below:

$$Eh_{mV} = ORP_{mV} + 1000[0.244 - 0.00066(T - 25^{\circ}\text{C})] \quad (7)$$

This is addition of the temperature-adjusted potential of the saturated calomel electrode, where  $T$  is the measured temperature of the sample in  $^{\circ}\text{C}$  [23].

The water samples were taken out of the freezer, allowed to defrost and then filtered through  $0.45 \mu\text{m}$  HVLP polymer membrane filter using a Nalgene filter funnel with a V-700 model vacuum pump. The filter funnel and the collecting flask were rinsed first with deionized water three times and then with the next sample to be filtered to avoid cross sample contamination.



**Figure 2.** Areas around Ngwenya Mine and streams whose source is the Ngwenya Mountain, Source: Central Statistical Office of Swaziland [20].

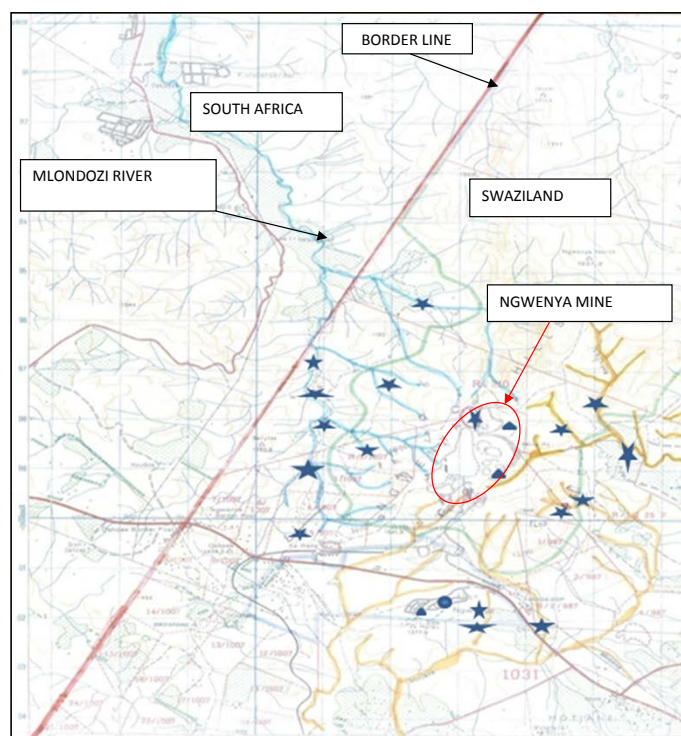


Figure 3. Sampling points (★) in areas around the Ngwenya Mine.

Each of the approximately 200 ml filtered water samples was put in a cleaned bottle, carefully labelled and then kept in the chest freezer for further analysis.

### 2.2.2. Anion Analysis

Filtered water samples were analysed by Ion Chromatography (IC) for seven anions, namely the halogens ( $F^-$ ,  $Br^-$ ,  $Cl^-$ ), the nutrients ( $NO_2-N$ ,  $NO_3-N$ ,  $PO_4-P$ ) and  $SO_4^{2-}$ , the best indicator for AMD pollution [24]. The Metrohm, 861 Advanced Compact IC used was set at 17 minutes analysis time for each sample by the ASUPP4-250. mtw method. The IC offers high sensitivity and multiple analyte determination in a single run. A 1000 ppm multi-standard stock solution containing the selected anions was prepared following the instrument manual's procedure. Five working standards of 2, 4, 6, 8 and 10 ppm were freshly prepared daily from the stock solution. 1.7 M  $NaHCO_3$ /1.8 M  $Na_2CO_3$  eluent (mobile phase) was prepared and used at a flow rate of 1 ml/min. Fresh suppressant solution of 50 mM of sulfuric acid and deionized water were used from the beginning of the analysis. The water samples were manually injected into the IC using a 10 ml syringe through a 0.45  $\mu m$  filter membrane. The results were recorded electronically as the instrument is connected to a computer.

### 2.2.3. Fe Speciation

A SHIMADZU UV-VIS Spectrophotometer was used for

the iron speciation with a 1 cm cuvette quartz cell and wavelength range of 190 nm - 2500 nm.

The estimation of the quantity of native  $Fe^{2+}$  and total Fe was achieved by the UV-VIS analysis. A 10 ppm standard ferrous solution was prepared by weighing 0.0702 g of analytical grade ammonium ferrous sulfate hexahydrate into a 1 L volumetric flask and enough deionized water to dissolve the salt was added. The solution was acidified by adding 2.5 ml concentrated sulfuric acid and the solution was made up to the mark with deionized water. The complexing agent, 1, 10-phenanthroline, was prepared by dissolving 0.1 g in 100 ml deionized water. The reducing agent, hydroxylamine hydrochloride, was prepared by dissolving 10 g of it in 100 ml deionized water. 100 ml of 0.1 M of each of sodium acetate and acetic acid were prepared separately. The acetic acid-sodium acetate buffer (pH 4.5) was prepared by mixing 65 ml of the 0.1 M acetic acid and 35 ml of the 0.1 M sodium acetate in a 100 ml volumetric flask. Fresh buffer solution was prepared whenever required.

Five working standards of 0.1, 0.3, 0.5, 1.0 and 1.5 ppm Fe were each prepared in 100 ml volumetric flasks. 10 ml of each of the water samples were measured into small beakers. 20 ml of deionized water was put into another small beaker to serve as a blank. To each of all the flasks (standards, water samples and blank), 1 ml of hydroxylamine hydrochloride and 5 ml of 1, 10-phenanthroline were added. Each solution was buffered with 8

ml of the acetic acid/sodium acetate buffer. All the solutions were allowed to stand for more than 15 minutes for full colour development. Both the standards and samples were measured at 508 nm, the  $\lambda_{max}$  for  $Fe^{2+}$ .

The native  $[Fe^{2+}]$  was determined by adding 5 ml of the complexing agent to 10 ml of each sample using deionized water as blank. These solutions were allowed to stand for 20 minutes for the complex to fully develop its colour. The samples were then analysed for the ferrous ion levels at 508 nm. The results were recorded and from total iron results, the  $[Fe^{3+}]$  in each sample was estimated by subtraction.

### 2.2.4. Heavy Metal Analysis

Heavy metal levels in the water samples were determined by the Flame atomic absorption spectrometry (FAAS) which uses the air-acetylene mixture. The heavy metals investigated were Cr, Mn, Fe, Ni, Co, Cu, Zn, Pb, and Cd. The heavy metal levels were determined using the instrument operating parameters given in Table 1 below. For each of the heavy metals, five calibration standards were prepared by serial dilution of the respective 1000 ppm stock solution obtained from Rochell Chemicals, RSA.

## 3. Results and Discussion

### 3.1. Calibration Curves and Detection Limits

Table 2 below shows the correlation coefficients, regression equations and detection limits of the analysed heavy metals in the water samples.

### 3.2. Physico-Chemical Properties

The pH and EC of the surface water bodies around Ngwenya mine were summarized in Table 3 and shown in Figure 4.

With the exception of Hawane (6.32) and Mbuluzana (6.48), the pH values for the water bodies were within the acceptable range for health risk, being (6.5 - 8.50), set by USEPA, 2008 [8]. Ndlotane, Mlondozi and yellow-boy water had electrical conductivities above maximum admissible limit set at 250  $\mu$ S/cm by WHO, 2008

Table 1. Instrument operating parameters used in the analysis of respective heavy metals.

Heavy metal	Cr	Mn	Fe	Ni	Co	Cu	Zn	Pb	Cd
WL	357.9	279.5	248.3	232.0	240.7	324.8	213.9	217.0	228.8
SW	0.2	0.2	0.2	0.2	0.2	0.5	1.0	1.0	0.5
LC	7.0	5.0	5.0	4.0	7.0	4.0	5.0	10.0	4.0

WL—Wavelength (nm), SW—Slit width (nm), LC—Lamp Current (mA).

Table 2. Wavelengths, correlation coefficients, regression equations and detection limits ( $C_m$ ) for the heavy metal analysis.

Metal	wavelength (nm)	R <sup>2</sup>	Regression equation	C <sub>m</sub> (ppm)
Cr	357.9	0.987	y = 0.004x + 0.002	0.03
Mn	279.5	0.999	y = 0.064x + 0.003	0.11
Fe	248.3	0.997	y = 0.036x - 0.008	0.63
Ni	232.0	0.997	y = 0.026x + 0.003	0.15
Co	240.7	0.993	y = 0.018x - 1 × 10 <sup>-5</sup>	0.10
Cu	324.8	0.998	y = 0.039x - 0.003	0.15
Zn	213.9	0.987	y = 0.081x + 0.027	0.59
Pb	217.0	0.999	y = 0.006x + 0.001	0.11
Cd	228.8	0.996	y = 0.127x - 0.005	0.13

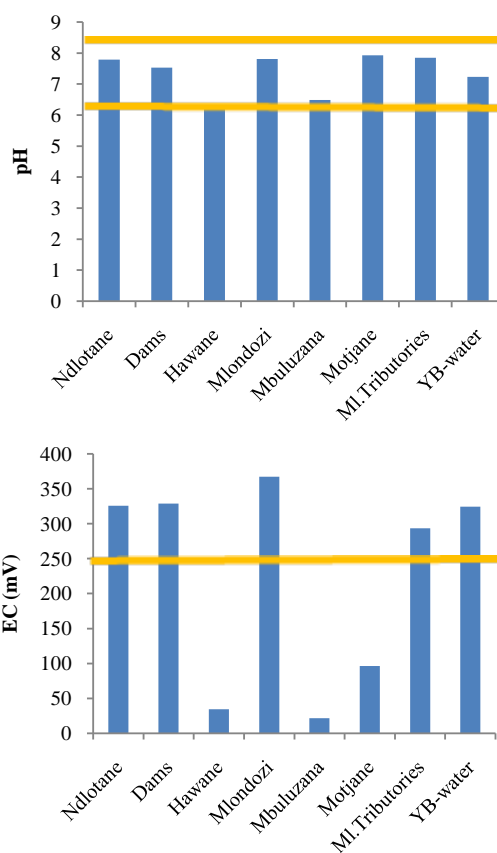


Figure 4. Mean pH and electrical conductivities of the water bodies located around the Ngwenya Mine.

[8]. This implies that the affected water bodies were heavily laden with ionic species which may pose a health hazard to the users. On the other hand, EC values for Hawane, Motjane and Mbuluzana water samples were

**Table 3.** Mean values of the physico-chemical parameters of the water samples.

sample site	pH	EC ( $\mu\text{S}/\text{cm}$ )	ORP (mV)	Temp. ( $^{\circ}\text{C}$ )	Eh(mV)
Quarry Dam	7.76	457.20	-53.60	17.28	195.50
Upper Dam	7.78	201.33	-54.00	18.13	194.53
Down Dam	7.05	327.33	-13.00	18.50	235.29
Hawane Dam	6.32	34.67	53.00	16.93	302.32
Ndlotane upstream	7.70	269.67	-49.33	17.23	199.79
Ndlotane Wetland	7.73	334.00	-51.67	17.60	197.22
Ndlotane Downstream	7.93	374.00	-62.33	17.67	186.51
Mlondozi Downstream	7.77	322.00	-54.00	18.30	194.42
Mlondozi 2	7.73	585.33	-51.67	18.57	196.58
Mlondozi 3	7.85	328.00	-58.33	18.57	189.91
Mlondozi 4	7.74	265.00	-52.67	18.47	195.65
Mlondozi Control	7.64	298.67	-46.67	18.13	201.87
Mlondozi Upstream	8.25	379.00	-93.67	15.97	156.30
New Skom Stream	7.82	177.33	-68.67	16.90	180.68
Stini Dolobha Stream	7.47	324.33	-49.00	16.83	200.39
Mbuluzana Upstream	6.33	11.00	53.67	17.10	302.88
Mbuluzana Downstream	6.54	19.00	41.33	17.07	290.57
Ngwenya Mine stream	6.58	35.33	38.67	17.03	287.92
Motjane Stream 1	7.80	143.67	31.33	17.43	280.33
Motjane Stream 2	7.62	60.00	21.33	17.57	270.24
Motjane Stream 3	8.02	101.33	44.33	17.77	293.11
Motjane Downstream	8.25	80.33	57.33	17.77	306.11
Ngwenya Stream	8.63	108.00	79.33	17.73	328.13
Mlondozi MS Y-Boy	7.14	399.67	-28.67	7.60	226.82
Ndlotane WL. Y-Boy	7.12	279.00	-27.33	11.07	225.86
Mlondozi-4. Y-Boy	7.43	294.67	-45.33	11.97	207.27

within the acceptable limits for drinking water and aesthetic value [8]. **Table 3** below shows that the environ-

ments in the water bodies of Hawane, Mbuluzana and Motjane were oxidising as shown by the positive ORP while reducing environments prevailed in all the other studied water bodies.

**Figure 5** shows the relationship of pH to Eh in the water bodies grouped in **Table 3**. The scatter diagram shows that generally when pH increases Eh decreases, alkaline water has low Eh. In other words, these parameters are inversely related to one another.

### 3.3. Anions in the Surface Water Samples

Generally, anions are less toxic than heavy metals. In this study, the toxicity of the anions detected in the surface water is in the order:  $\text{F}^- > \text{NO}_3\text{-N} > \text{Cl}^- > \text{SO}_4^{2-}$  [25]. This order corresponds to their respective maximum acceptable limits which are as follows: 1.5 ppm, 10 ppm, 250 ppm and 500 ppm [25]. **Table 4** gives the anion levels in surface water bodies around the old Ngwenya Mine. The sulfates are indicator ions for AMD occurrence; they are a product of pyrite oxidation together with the ferrous and the hydrogen ions. It must be noted that phosphates were missing because they are highly adhesive to soil thus the reason why they may not be detected in the water. The Down dam and the Upper dam are surrounded by mine dumps, fields and homesteads. The high nitrate levels must have been from the fertilizers that were used for agricultural purposes and the chlorides from common salt used for domestic purposes. The sulfates must have been the leachate from the mine dumps. The Hawane dam recorded the least of these ions. It is located out of the mining area so the water is fed to it by a stream originating from the mine area which may have undergone river self-purification hence the decreased ion levels.

The high chloride level in Ndlotane river may be attributed to domestic wastes from homesteads along the stream. Chlorides cause a brackish taste in water and are indicator ions for sewage pollution [25]. Nitrates leach from the fields, where artificial fertilizers were applied, down gradient into the streams. Elevated nitrate levels were also observed in sites surrounded by forests of wattle trees, that is the Quarry dam and NG-(M). This result was also observed in [1]. Nitrates are carcinogenic, cause chronic effects ( $\text{NO}_3\text{-N} > 10$  ppm cause methemoglobinemia) and algal bloom which have adverse effects on the environment and human health [26]. The sulfates must have percolated with the acid mine drainage polluted water from the Quarry Dam upgradient [1]. This assertion is reinforced by the high  $[\text{SO}_4]$  recorded in SD-MS which originates from just below the dam. Ingestion of excess sulfates gives water an unpalatable taste and causes laxative effects [25]. The flouride ion recorded the least levels of all the anions such that it is

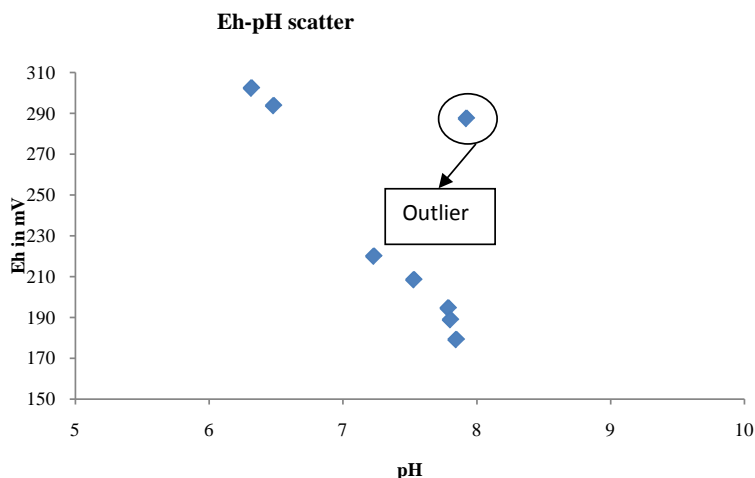


Figure 5. The  $E_h$  – pH diagram for water samples.

Table 4. Anion levels (ppm) in water bodies around the old Ngwenya Mine.

sample ID	F <sup>-</sup>	Cl <sup>-</sup>	NO <sub>3</sub> -N	SO <sub>4</sub> <sup>2-</sup>
Quarry dam	ND	1.57	12.41	22.29
Down Down	ND	11.19	13.25	5.34
Up-Dam	0.06	3.88	3.17	0.73
Hawane Dam	ND	2.65	0.76	0.83
ND-U	ND	2.56	0.70	0.37
ND-WL	ND	11.37	ND	1.86
ND-D	ND	9.11	0.44	1.31
ML-(1)	0.15	2.71	ND	2.09
ML-(2)	0.05	2.70	ND	2.04
ML-(3)	0.03	2.96	1.44	2.04
ML-(4)	0.09	2.67	1.17	1.84
ML-(C)	0.04	2.59	ND	1.35
NS-ML	0.02	4.36	ND	2.33
SD-M	0.15	3.75	ND	6.53
MS-ML	0.00	3.81	ND	5.23
MB-U	ND	1.50	ND	0.05
MB-D	ND	2.82	2.96	0.41
NG-(M)	0.02	2.21	12.41	0.72
Motj-(1)	0.02	1.85	0.94	0.20
Motj-(2)	0.05	2.52	ND	0.71
Motj-(3)	0.10	1.58	ND	0.75
MOTJ-D	0.10	3.32	0.40	0.46
NG-Str.	0.16	2.24	ND	1.61

ND—not detectible.

not detectible in most of the sites. Flourides are the most toxic of the anions detected in this study. They cause dental flourosis when their level is greater than 1.5 ppm in drinking water [25]. It is inferred that the users of the water found around Ngwenya are safe from anion toxicity because the levels of all the anions studied are less than their maximum contamination limits. The water in the three sites that had nitrate levels above the MCL are all not used for human consumption.

### 3.4. Speciation of Iron

The Fe<sup>2+</sup> and total iron in selected water samples were quantified by complexation and UV-VIS spectrometry. The Fe<sup>3+</sup> levels were obtained by subtraction. Table 5 below shows that the dominant Fe ion in the water samples was the Fe<sup>2+</sup>. This result is evidence that the waters were affected by AMD pollution because the Fe<sup>2+</sup> is one of the immediate products of pyrite oxidation. In the “yellow boy” water, the Fe<sup>3+</sup> is dominant because it is gradually produced from the oxidation of the Fe<sup>2+</sup> and becomes hydrolysed to form the ferric hydroxide precipitate.

### 3.5. Heavy Metals in Water

AMD favours heavy metal dissolution in water. The drinking water sites and “yellow boy” water found around Ngwenya Mine were investigated for heavy metal pollution. Heavy metal levels in surface water change due to dilution by rainwater which leads to increase in pH hence precipitation of the dissolved heavy metals out of solution accumulating in the sediments. Decreased water turbulence and evaporation also favour metal precipitation from waters to sediments [28].

The “yellow-boy” water was filtered off from the

orange-brown precipitate (an indication of AMD pollution), collected from the banks of acid effluent receiving streams and wetland. Generally, Cr, Mn, Fe and Cu recorded the highest levels while Ni and Cd had the lowest levels in “yellow-boy” water. This suggests that there is some interesting chemistry that must be investigated on this manifestation of AMD pollution, “yellow-boy”, especially because relatively extreme levels of the heavy metals were found in it. The information on heavy metal levels in the studied sites was collected and summarised in **Table 6** below.

Most of the sites studied were polluted by Cd (83%). Ingestion of water with high Cd content increases salivation and severe stomach irritation leading to diarrhoea [9]. 71% of the drinking water supply sites had Cr mean levels above the maximum admissible level of 0.10 ppm for drinking water and aesthetic value [8]. This implies

**Table 5. Contributions of the [Fe<sup>2+</sup>] and [Fe<sup>3+</sup>] in the total iron concentration (ppm) in the water bodies.**

Sample site	Fe(tot)	Fe(II)	Fe(III)
Quarry Dam	0.56	0.56	0.00
Down Dam	1.18	1.18	0.00
Ngwenya Mine stream	0.76	0.76	0.00
Mbuluzana down stream	0.97	0.66	0.31
Ndlotane Upstream	1.08	1.08	0.00
Ndlotane wetland	1.81	1.71	0.10
Mlondozi mine stream	0.66	0.66	0.00
ML.minestream Y-boy	1.39	0.66	0.73

**Table 6. Summary of the state of each heavy metal in all the sites under investigation.**

Heavy metal	Mean (ppm)	SD	Range (ppm)	%>MCL	Guideline
Cr	0.93	0.77	0.00 - 2.46	71.00	0.1
Mn	0.03	0.04	0.00 - 0.30	17.00	0.05
Fe	0.11	0.09	0.00 - 0.34	4.00	0.3
Ni	0.10	0.09	0.00 - 0.34	46.00	0.1
Co	0.10	0.09	0.00 - 0.29	42.00	0.1
Cu	0.03	0.02	0.00 - 0.07	0.00	1.3
Zn	0.00	0.00	0.00 - 0.00	0.00	5
Pb	0.00	0.00	0.00 - 0.00	0.00	0.015
Cd	0.06	0.03	0.01 - 0.11	83.00	0.005

%—percentage of sites whose heavy metal level is above maximum admissible level.

that the communities that use these water supplies bear the health risk of Cr poisoning; Cr is carcinogenic [9]. Underdevelopment of roots in plants can be caused by Cr exposure which leads to poor plant growth [9]. The average Ni content for all the drinking water samples exceeded the MCL set at 0.10 ppm by WHO [8]. This result implies that all the dams and streams investigated need attention on Ni pollution. Ni is carcinogenic and long-term exposure can result in loss of body weight, heart and liver damage, and skin irritation [9]. 42% of the water bodies under discussion need attention on Co pollution. Ingestion of excess Co causes nausea, vomiting and severe heart effects. Mn mean level was also above the MCL (0.05 ppm) with 17% of the investigated sites polluted with Mn. Only 4% of the sites were polluted with Fe. Excess Fe in the human body may cause toxemia of pregnancy, tumors and may damage heart muscle tissues [29]. In cattle, excess dissolved Fe discourages adequate water intake and consequent adverse effect on milk production [19]. It is worth noting that the yellow-boy water from the wetland recorded high levels of some metals (Mn and Fe). This suggests that wetlands can be used to naturally attenuate dissolved heavy metals in the environment [27]. All the water bodies were free from Cu, Pb and Zn pollution. Finally, the quality of surface water in the vicinity of the Ngwenya Iron Ore Mine is deteriorated due to the activities in the old mine, and it is highly probable that some current human activities exacerbate the situation. Hence, the water pollution by heavy metals is mainly anthropogenic.

#### 4. Conclusion

Safe drinking water is a birthright for all humans while all biota has a right to clean habitable water. Water bodies around the Ngwenya Mine receive contaminants from the old abandoned mine driven by AMD and also by runoff. The presence of AMD pollution in the receiving streams is evidenced by the occurrence of ferric hydroxide (yellow boy) at their banks. There is some interesting Chemistry that should be investigated in the yellow boy because relatively high heavy metal levels were found in its water. The AMD mobilizes the heavy metals into the aquatic environment exposing biota to heavy metal pollution through the use of such contaminated water. Heavy metals pollute the environment through chronic effects. Even if they are available in the water in trace amounts they accumulate in water bodies on a long time exposure until pollution levels are reached. Thus, the population around the mine bears the risk of heavy metal pollution. The State needs to set up sound remediation strategies to save the population around the Ngwenya Mine from the heavy metal pollution they are currently being exposed to by using water polluted by



some previous mining activities. This study shows that there is no health risk from anion pollution in waters around the mine.

## Acknowledgements

We would like to thank the Swaziland National Trust Commission (SNTC) for granting us the permission of entry from time to time into the study area and an anonymous reviewer for his comments and constructive suggestions.

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